Vibrational dynamics and structural relaxation in aqueous electrolyte solutions in the liquid, undercooled liquid and glassy states

J. Dupuy, J. Jal, B. Prével, A. Aouizerat-Elarby, P. Chieux, A. J. Dianoux, J. Legrand

To cite this version:


HAL Id: jpa-00251296
https://hal.archives-ouvertes.fr/jpa-00251296
Submitted on 1 Jan 1992

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.
Vibrational dynamics and structural relaxation in aqueous electrolyte solutions in the liquid, undercooled liquid and glassy states

J. DUPUY, J. F. JAL, B. PRÉVEL, A. AOUIZERAT-ELARBY, P. CHIEUX*, A. J. DIANOUX* and J. F. LEGRAND*

Département de Physique des Matériaux, U.A. 172, Université Claude Bernard Lyon I, 43 bd du 11 novembre 1918, 69622 Villeurbanne Cedex, France
*Institut Laue Langevin, BP. 156 X, 38042 Grenoble Cedex, France

Summary

In very concentrated aqueous electrolyte solutions, the competition between dipolar and ionic forces and most of the system properties, are finely monitored by the concentration. For LiCl, RH2O (R being the number of water moles per mole of salt) two domains of concentration exist around even hydration numbers R = 6 and R = 4, which can be reversibly undercooled from the liquid to the glassy state. At other concentrations, for R<12, metastability is achieved only in a restricted temperature range.

The collective behaviour as well as individual atomic motions, have been studied by neutron inelastic scattering in the metastable state of LiCl, 6H2O and LiCl, 4H2O down to the glass transition.

The study of the structural relaxation by neutron spin echo (NSE) allows us to show that the temperature dependence of the response of the system follows a scaling relation with a cross-over at a phase transition present in the liquid phase.

Increasing hydrodynamic interactions in a liquid is certainly the best manner to analyse the intrinsic and microscopic correlations into the system. So the various motions, collective, relaxational, individual and their coupling can be revealed by approaching glass formation conditions.

LiCl, RH2O (or LiCl, RD2O) is one of the most suitable system because by changing R (the number of water molecules per molecule of salt) we change strongly the nature of the medium and by using isotopic substitution on Li+, Cl-, D we can differentiate the dynamics of the various atoms or ions.

Changing R gives the possibility to analyse two types of liquids: the first one, for R>6 is intrinsically inhomogeneous, and described by a regular mixture of pure water and the hydrate LiCl, 6H2O (1). The second one, for R<6 is a mixture in which the competitions between dipolar forces and coulombic forces are finely monitored.

In this paper, we summarize some recent analysis of different kinds of dynamical responses which lead to some questions on their non linear behaviour.
The total dynamical responses of the electrolyte LiCl, 6H2O and LiCl, 4H2O have been analyzed between 0.3 meV and 120 meV, i.e. in a domain of frequencies covering collective vibrational response and individual motion of ions.

1. Dynamical structure factor.

From the dynamical structure factor $S(q,\omega)$ measured with the IN 6 instrument (ILL, Grenoble) at a resolution of 80 μeV, for LiCl, RH2O (with R=6), two types of informations can be derived:

(i) - the frequency distribution $G(E)$ ($E=\hbar\omega$) (extracted as in ref (3), neglecting the incoherent contribution to the scattering), which permits an analysis of the characteristic frequencies (2) which define this low energy "vibrational network" (3).

(ii) - The central mode, i.e. diffusive processes from the quasi elastic behaviour around $E=0$.

A - In the frequency distribution, three fundamental modes of the H-bonded network dominate: bending at 7 meV, restricted translational motion or stretching at 23 meV, and restricted rotational motion or libration at 85 meV (see figure 1). The identification of these modes has been made by comparison with experiment on pure water (2) (10).

Fig. 1. Generalized density of states obtained for LiCl, 6H2O in the glassy and liquid states and assignments of the principal bands for the liquid state (see 2).

The influence of temperature on the different modes has been followed in this desordered system down to the non ergodic situation ($T < 142$ K for LiCl, 6H2O). (Fig. 2)

Fig. 2. Evolution of the three fundamental bands characterizing the bending, stretching, librational motions assigned to H bonded H2O in LiCl, 6H2O versus temperature.
- The broadening of the bending motion which is large in the liquid state is correlated to the existence of different microscopic dynamical configurations. This mode has a significant asymmetry towards high energies which increases with temperature, giving some support to a delocalization of the excitations or a fluctuating coupling between different microscopic configurations (12)

- Among these modes, stretching is the most influenced by the glass transition or equivalently by To i.e the lowest temperature limit of liquid metastability(2) However the limited range of energy that we were able to reach, does not permit us to fully analyze the low temperature behaviour for librational modes. There is another temperature of interest which is characteristic of a third order phase transition in the liquid(quoted by $T_p$ = peritexy at 207 K). It appears that the stretching mode is nearly temperature independent in the low temperature supercooled liquid for $T_o < T < T_p$. This temperature range is characterized by additional microscopic configurations (hydration 5) compared to those of the room temperature liquid.

B - Quasi elastic data, i.e. the central mode, permits to obtain informations on the individual motion of H or D in $H_2O$ and $D_2O$. At the composition $R=6$, the system is molecular and although there is no correct diffusion's analysis of the different elements $Li^+$, $Cl^-$ and $H_2O$, the trends
give $D_{Li^+} < D_{Cl^-} < D_{H_2O}$ at 25°C (4)

- The proton delocalization due to vibrations is deduced from the Debye Waller component. It is obtained from the total quasielastic intensity, $I$ and is shown in fig. 3a and fig. 3b for $LiCl$, $6H_2O$ and $LiCl$, $4H_2O$.

Fig. 3a. Debye-Waller factor measured by QENS for $LiCl$, $6H_2O$ in the liquid (●) $223 \text{ K}$ and in the super-cooled liquid : (●) 192 K

Fig. 3b. Debye-Waller factor measured by QENS for $LiCl$, $4H_2O$ in the liquid state (●) $275 \text{ K}$ and in the supercooled liquid state (+) $210\text{ K}$

Insert : Mean square displacement of $H_2O$ in $LiCl$, $6H_2O$ and in $LiCl$, $4H_2O$ in liquid and super-cooled liquid versus ($T$ - $T_o$), $T_o$ being the lowest temperature limit of the liquid metastability.

* Proton delocalization in pure water (11)
The mean square displacement obtained for H₂O in LiCl, 6H₂O and LiCl, 4H₂O and referred to T₀, lowest temperature limit of the metastability (10) is temperature dependent in the liquid state and nearly temperature independent in the supercooled liquid. This effect is induced by a restricted translational motion due to interactions of H₂O with Li⁺ and Cl⁻.

- The diffusion coefficient of H₂O can be obtained from S(q, ω) and compared to the ordinary diffusion law (5). As a matter of fact, the diffusion can be coupled to the hydrodynamics of the medium, the coupling being function of the kinematic viscosity Dₜ, according to δ = D / (D + Dₜ)(5). The q dependence of the half width of the central mode given in fig.4 for LiCl, 6H₂O permits to obtain diffusion coefficient in agreement with literature (4).

\[ \Delta E/q² \ (\text{meV} \ \text{Å}^2) \]

Fig. 4. FWHM of S(q, ω) divided by q² for LiCl, 6H₂O at different temperatures above and below Tₚ.

- The narrowing of the central mode for all q values inferior to qₘₐₓ (= 1.9 Å⁻¹) confirms the coupling of this mode with hydrodynamics (6).

The H₂O diffusion coefficient obtained in this molecular system, for R = 6 and R = 4 permits to confirm the strong influence of the solvation on the dynamics of H₂O when R≤6 (fig 5).
2. Relaxational behaviour.

The relaxational behaviour has been obtained on an "adequate" range of time (10^{-12} sec. to 10^{-9} sec) by NSE experiment performed on IN11 at ILL(9) on LiCl, 6D2O at q = q_{\text{max}} = 1.9 \, \text{Å}^{-1}, q_{\text{max}} being the momentum transfer of the main peak of the structure factor at room temperature.

We deduce from this data (7), the evolution of the characteristic relaxational time $\tau_{\text{NSE}}$ which can be compared to the one obtained from viscosity (8) (fig. 6). Two scaling laws (10) give the evolution of $\tau$, with two characteristic temperatures of the system: lowest limit of the metastability $T_0$ and crystallisation temperature for the first hydrate formed: $R = 5$

![Fig. 6a. Neutron spin echo response at q = q_{\text{max}} = 1.9 \, \text{Å} for LiCl, 6D2O measured and analyzed according to ref. (7).](image)

![Fig. 6b. Relaxational time deduced from NSE (\textbullet) for LiCl, 6D2O compared to $\tau_S$ obtained from kinematic viscosity (\textcircled{o}) according to ref. (8) for LiCl, 5.75 H2O.](image)

This result confirms the connexion between the relaxational process and the third order phase transition, at $T_p$. Slowing down is observed in the supercooled liquid scaled to the lowest limit $T_0$ of the metastability and the non-ergodic character appears down to $T_0$ (7).

In conclusion, in these molecular aqueous solutions, collective vibrational motion of H2O such as stretching, mean square displacement and relaxational response in the hydrodynamic regime and at finite momentum transfer are sensible to the approach of some characteristic temperatures of the system such as $T_0$ and $T_p$; an apparent third order transition in the vicinity of the equilibrium melting point. The existence of diffusive motion confirms the coupling with the hydrodynamics. The extension of these studies to the more concentrated domain, offers the opportunity to study the coupling between ionic conduction and relaxational slowing down.
References:

    A. EL HACHADI - Thesis Lyon July 1991


    55, 3013.

[9] MEZEI, F, ED. Neutron Spin Echo in lecture Note in Physics 128 Springer-Verlag Berlin-

    Complex Systems and related topics Ed by I.A. Campbell, C. Giovannello, Plenum Press N.Y.
    (1990) p. 175.


    " H-Bond characteristic in supercooled and vitreous aqueous solutions"
    " Ice nucleation and growth in Electrolytes : role of ionic forces"