Lanthanide complexes of DOTA–nitroxide conjugates for redox imaging: spectroelectrochemistry, CEST, relaxivity, and cytotoxicity

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The lanthanide(III) complexes (Gd, Eu, Dy, Yb) of DOTA tris(amide) and bis(amide) derivatives (**L₁** and **L₂**) featuring one redox active TEMPO arm were prepared. Ligand **L₂** harbours an alkyne fragment for further functionalization. The X-Ray crystal structures of the ligand **L₂** in complex with Na^+ was solved. The complexes showed in their CV one oxidation wave (0.26 – 0. 34 V vs Fc_+/Fc) due to oxoammonium/nitroxide redox couple and a broad reduction corresponding to the nitroxide/hydroxylamine system. The Eu complexes demonstrated the presence of one water molecule in their coordination sphere. The nitroxide complexes were characterized by EPR spectroscopy, showing the typical 3-line pattern in the high temperature regime, which is quenched upon addition of ascorbate (reduction into hydroxylamine). Under their nitroxide form the complexes show essentially no CEST peak. Conversely, the reduced complexes demonstrate a 12 % CEST peak at 51 ppm, corresponding to the metal bound water molecule. Fast exchange precluded the CEST activity for the amide protons. All the complexes proved to be essentially non-toxic for M21 cells at concentrations up to 50 μM.

1. Introduction

Medical imaging is an extremely important domain for current research due to the necessity of precocious detection, follow-up, and treatment of dangerous illnesses.[1] In particular molecular imaging can provide details of biological processes in tissues in the human body, giving more information for the diagnosis of diseased tissues.[2] The development of molecular imaging requires the development of novel imaging probes capable of giving an effective response to a stimulus. The redox status of tissue is an important marker of its activity. The redox potential is defined by the sum of the redox couples weighted with the concentration of the respective reduced species. The intracellular medium is buffered by several redox couples of which the glutathione disulfide/sulfhydryl GSSG/GSH is the main one (redox potential of -120 to - 70 mV).[3] The redox environment of the extracellular space is dominated by the cysteine disulfide /sulfhydryl redox couple CySSCy/CysSH, with higher concentrations (0.05–0.3 mM) than the GSSG/GSH couple (0.002-0.02 mM) and with less negative redox potential (- 80 to - 20 mV) than GSSG/2GSH, giving a more oxidized compartment than the cytosol.[4] Conditions of oxidative stress are encountered when the natural metabolism of oxygen is disrupted, leading to the over production of reactive oxygen species (ROS) by incomplete reduction.[5] Changes in redox status (oxidative stress or hypoxia) are believed to contribute to cellular signalling (differentiation, apoptosis),[6] but also the proliferation of diseases such as cardiovascular diseases,[7, 8] cancers[9, 10] or neurodegenerative diseases[11].

Among various non-invasive detection techniques, magnetic resonance imaging (MRI) has proved to be a key technique for clinical imaging. It relies on the differential relaxation of water protons of the bulk water in the human body.[12] Its success lies in its high spatial resolution, lack of hazardous radiation, and the fact it is painless and can be functional in addition to anatomical. Approximately one third of the MRI scans are realized with a relaxation contrast agent (CA), which is often a gadolinium complex. Gd^{III} is the metal ion of choice due to its quenched total electronic orbital angular momentum $L = 0$, the highest possible total electronic spin value $S = 7/2$, and a slow longitudinal electronic relaxation, especially at the rather high magnetic field $B_0 \geq 1$ T of the MRI scanners.[13-15] The CA is injected into a vein before the analysis to improve the details and clarity of the images. The CA effectiveness is gauged by its relaxivity r_1, which is the increase in the longitudinal relaxation rate of the water protons (hydrogen nuclei) per mM of contrast agent in biological tissue.[11,14] Recently, Chemical Exchange Saturation Transfer (CEST), another technique of MRI, based on paramagnetic lanthanide ions such as Eu^{III}, Yb^{III}, and Dy^{III} with very fast electronic relaxation lends itself to more sensitive

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specific functional imaging. It was suggested almost ten years ago that lanthanide-based CAs could be also used to sense the redox environment.[16, 17] Two main strategies emerged, which are based on either a redox-reactive ligand or redox-active metal. In the first case the r_1 value of the Gd^{III}-based CA can be tuned by the hydration state (number of coordinated water molecules q), which is itself modulated by a redox event at the ligand. This first strategy is illustrated by DOTA-based Gd^{3+} complexes (DOTA is 1,4,7,10-Tetraazacyclododecane-1,4,7,10-tetraacetic acid) appended by a merocyanine group. This group cyclizes into spyrooxazine upon reduction by NADH, resulting in decoordination and further increase of q from 1 to 2 with change in r_1.[18-20] Several Gd^{3+} complexes functionalized by a thiol-reactive 2-pyridyldithio group were also prepared and shown to react with GSH with a change in q number from 2 to 1.[21, 22] Alternatively, the slower rotational dynamics upon redox reaction of the ligand by reversible (or irreversible) binding of low molecular weight Gd^{3+} complexes to macromolecules can be used to increase r_1 values at constant q. Some Gd^{3+}-DOTA complexes appended by terminal thiol groups demonstrated such behaviour, whereby they react with the Cys-34 residue of human serum albumin.[23]

The second strategy is exemplified by cryptates of europium, the only lanthanide susceptible to adopt two different oxidation states *in vivo*.[24] The Eu^{2+} ion, isoelectronic to Gd^{3+}, is oxidized *in vivo* into Eu^{3+}, which shows a negligible paramagnetic relaxation because of its very fast electronic relaxation.[25-27]

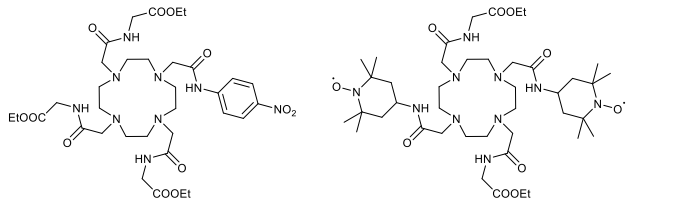


Chart 1. Structures of ligands whose Eu^{3+} complexes were reported as redox-responsive PARACEST agents. [28]

Paramagnetic Chemical Exchange Saturation Transfer (PARACEST) is an attractive contrast mechanism for MRI, which often employs paramagnetic lanthanide complexes.[29] A pre-saturation pulse is applied to exchangeable protons within the complex, resulting in a transfer of saturated spins to the bulk. As a consequence, the water signal intensity decreases and negative contrast is observed. Very few lanthanide redox-responding complexes were reported to date in PARACEST imaging.[16] All are based on Eu^{3+} complexes of DOTA tetraamide ligands as this lanthanide shows the slowest water exchange rates (Chart 1),[28] and operate according to two distinct mechanisms: The ligand can undergo reduction such as in a nitrobenzene (Chart 1A)[30] or quinolinium group,[31] resulting in an alteration of the water exchange rate. Alternatively, the nitroxide appended Eu^{3+} DOTA complex (Chart 1B) exhibits a T_1-shortening of bulk water protons by the paramagnetic centers, which increases upon reduction.[32] Interestingly a Eu^{3+}/Eu^{2+} switch was recently proposed as a redox PARACEST agent.[33]

We recently reported Ln^{III} complexes, which are sensitive to environmental redox changes. We monitored the change in oxidation state by luminescence, where over 90 % switch could be achieved.[34-36] We have also demonstrated through EPR the spin trap of a major ROS, the hydroxyl radicals, by lanthanide complexes of a DOTA-hydroxylamine derivative.[37]

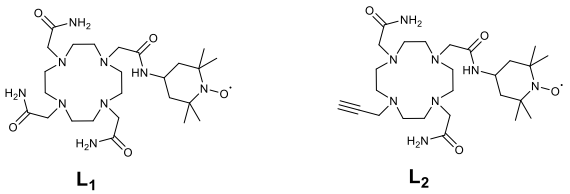


Chart 2. Structures of ligands **L1** and **L2**

Following the elegant strategy developed by Sherry *et al.* to design PARACEST agents responsive to redox changes[32] we herein report a series of lanthanide(III) complexes (Gd, Eu, Dy, Yb) based on a cyclen scaffold containing the redox active TEMPO arm[38] (Chart 2). Two different derivatives $\mathbf{L_1}$ and $\mathbf{L_2}$ have been designed containing three acetamides or with an alkyne. The alkyne was chosen in order to allow future functionalization and bioconjugation of these complexes via click chemistry. We investigate their photophysical, redox, relaxometry and their potential as redox-responsive PARACEST agents.

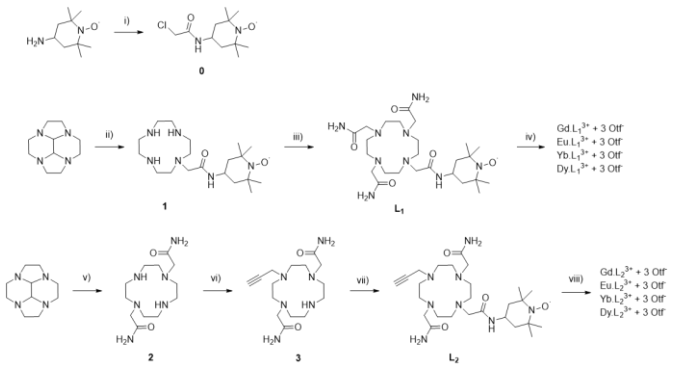
2. Results and discussion

2.1 Synthesis

The synthesis of ligands $\mathbf{L_1}$ and $\mathbf{L_2}$ was performed using the successive functionalization of the cyclen macrocycle (scheme 1). $\mathbf{L_1}$ was synthesised via substitution with the nitroxide acetamide and further functionalization with three bromoacetamide arms. $\mathbf{L_2}$ was synthesised by protection of the cyclen in order to selectively functionalise 2-bromoacetamide arms, deprotection, followed by the addition of a dilute solution of propargyl bromide to perform the mono substitution. Purification and further addition of the nitroxide acetamide afforded the desired product (see experimental).

The Gd^{III}, Eu^{III}, Yb^{III} and Dy^{III} complexes were prepared from $\mathbf{L_1}$ and $\mathbf{L_2}$ by reaction with the corresponding lanthanide salt in H_2O at pH 7.4 and precipitation of the resulting complex.

Ligand $\mathbf{L_2}$ was crystallised from slow evaporation of water solution to give single crystals of quality for X-Ray diffraction. Surprisingly, structural analysis demonstrated the complexation of a Na^+ ion in the cyclen cavity, giving the species $[\mathbf{Na(L_2)}]^+$, 0.5 Br^-, 0.5 Cl^-. The sodium ion is heptacoordinated to four nitrogens of the macrocycle and three oxygens of the acetamide arms (Fig. 1). The Na-N bond distances range between 2.527(10) and 2.643(9) Å, while the Na-O are in between 2.398(3) and 2.436(3) Å. The C2-C3 bond distance in one pending arm is 1.17(2) Å, which confirms its triple bond character. The N-O bond distance in the opposite pending arm is 1.289(4) Å, *i.e.* reminiscent of TEMPO radicals.



Scheme 1. Synthetic procedures towards L_1, L_2 and their lanthanide complexes[a]

[a]Reaction conditions: (i) chloroacetyl chloride, Et_3N, 0°C to rt, CH_2Cl_2, overnight; (ii) 1) 0, THF, 1 day, 2) o-PDA, MeOH, 1 week; (iii) 2-bromoacetamide, K_2CO_3, KI, MeCN, 50°C, 5 days; (iv) Ln salt, H_2O, pH 7, 50°C, overnight; (v) 1) 2-bromoacetamide, MeCN, 1 day, 2) o-PDA, MeOH, overnight, 3) $CsCO_3$, MeOH, Overnight; (vi) 2-propargyl bromide, NaOAc, EtOH, 0 to 50°C, 4 days; (vii) **0**, K_2CO_3, KI, EtOH, 60°C, 5 days; (viii) Ln salt, H_2O, pH 6.5, 50°C, 2 days.

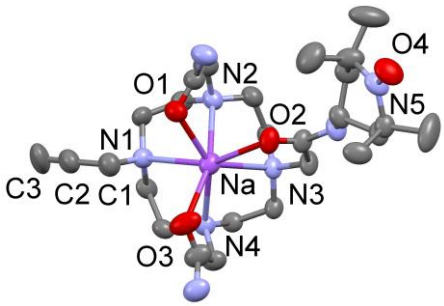


Figure 1. ORTEP plot of $[Na(L_2)]^+$ (the H atoms are omitted for clarity)

2.2 Electrochemistry

Electrochemical investigations were performed by cyclic voltammetry (CV), differential pulse voltammetry (DPV) and rotating disc electrode (RDE) voltammetry in an acetonitrile solution containing 0.1 M tetra-*n*-butylammonium perchlorate (TBAP) as supporting electrolyte. All the potentials are referred to the ferrocenium/ferrocene redox standard.

The ligands demonstrate oxidation potentials of 0.28 V and 0.31 V vs. Fc^+/Fc (0.91 to 0.94 V vs. NHE)[39] respectively for L_1 and L_2. The complexes exhibit a reversible oxidation wave in the range 0.26 – 0.34 V vs. Fc^+/Fc (Table 1, Fig. 2), which is assigned to the oxidation of the TEMPO moiety into the persistent oxoammonium species $TEMPO^+$.[40] These values are slightly more positive than that reported for free TEMPO (0.21 V vs. Fc^+/Fc)[41] due to the presence of the amide function. An irreversible reduction wave corresponding to the reduction of the TEMPO into the hydroxylamine TEMPOH is observable as a broad feature between -1 and -1.5 V vs. Fc^+/Fc (about -0.3 to -0.9 V vs NHE, Fig. 2).[40] The apparent irreversibility of the redox wave is due to a very slow electron transfer for the proton-coupled TEMPOH/TEMPO redox couple and an exceedingly large anodic-to-cathodic peak separation, which shifts the oxidation peak of TEMPOH above the $TEMPO^+$/TEMPO wave.[40, 42]

Each complex, Eu^{III}, Dy^{III}, Gd^{III} or Yb^{III} demonstrates similar features with only slight shift of the oxidation and reduction waves, confirming that the metal ion had very little effect on the redox behaviour of the TEMPO group. It must be stressed that no oxidation or reduction of the Ln^{III} centre was observed in the investigated potential range.

Table 1. Electrochemical data of the ligands and complexes[a]

Complex	$E_{1/2}^{ox}$	Complex	$E_{1/2}^{ox}$
$[Gd(L_1)]^{3+}$	0.26	$[Gd(L_2)]^{3+}$	0.33
$[Eu(L_1)]^{3+}$	0.30	$[Eu(L_2)]^{3+}$	0.34
$[Dy(L_1)]^{3+}$	0.28	$[Dy(L_2)]^{3+}$	0.33
$[Yb(L_1)]^{3+}$	0.32	$[Yb(L_2)]^{3+}$	0.33
L_1	0.25	L_2	0.31

[a] In 0.5 mM CH_3CN solutions containing TBAP as supporting electrolyte. All the potentials values are given in V and referred to the Fc^+/Fc redox couple. *T* = 298 K; scan rate, 0.1 V/sec. Parameters for the ferrocene against the reference used ($AgNO_3$ 0.01 M): $E_{1/2}$ = 0.090 V. In order to convert the potentials from the Fc^+/Fc reference to the NHE 0.630 V must be added to the listed values.[39]

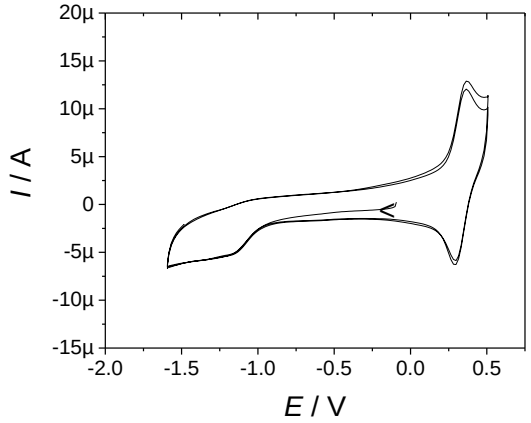


Figure 2 Cyclic voltammetry curve $[Eu(L_2)]^{3+}$ in 0.5 mM CH_3CN solution containing 0.1 M TBAP. Scan rate, 0.1 V s^{-1}, *T* = 298 K, carbon electrode.

2.3 Absorption and luminescence spectroscopy

The UV-Vis absorption spectra of the ligands L_1 and L_2 in CH_3CN show two main bands: 246 nm (ε = 14 200 $M^{-1}cm^{-1}$) and 286 nm (345 $M^{-1}cm^{-1}$) for L_1 and 246 nm (22 530 $M^{-1}cm^{-1}$) and 286 nm (1590 $M^{-1}cm^{-1}$) for L_2.[43] Upon complexation, these bands were slightly red shifted with respect to the ligand, with the $[Gd(L_2)]^{3+}$ complex showing λ of 292 nm (4020 $M^{-1}cm^{-1}$) and 363 nm (ε = 1760 $M^{-1}cm^{-1}$). The complexes also possess very weak epsilon absorption corresponding to the f-f lanthanide transitions.

The luminescence spectra were recorded for the europium complexes since structural information can be extracted from a $^7F_{0\text{-}5}$ band shape analysis. They were recorded via excitation in the ligand at 315 nm providing the expected spectra of the Eu^{III} ion (Fig. 3). A sharp band is observed at 580 nm for $[Eu(L_1)]^{3+}$ that corresponds to the transition $^5D_0 \rightarrow {}^7F_0$. Its sharpness suggests high symmetry in the coordination sphere of the complex and presence of single site symmetry for Eu^{III}. The $^5D_0$

→ $^7F_1$ transition is the magnetic dipole transition independent of the coordination sphere. The hypersensitive electric dipole transition $^5D_0$ → $^7F_2$ is the most influenced by the local symmetry. The distinct ratio of the integrated bands corresponding to the transitions $^5D_0$ → $^7F_2$ over $^5D_0$ → $^7F_1$ in **[Eu(L$_1$)]$^{3+}$** and **[Eu(L$_2$)]$^{3+}$** suggests a different coordination environment. Most particularly the higher ratio in **[Eu(L$_2$)]$^{3+}$** (Fig. 3) supports a lower symmetry around the first coordination sphere. Finally, excitation spectra confirmed the excitation occurred via population of ligand and transfer towards the Eu^{III} metal.

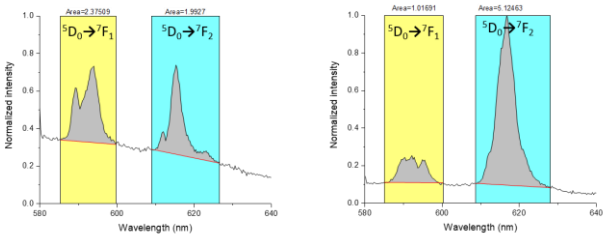


Figure 3 Luminescence spectra of **[Eu(L$_1$)]$^{3+}$ (left)** and **[Eu(L$_2$)]$^{3+}$ (right)** in D_2O upon excitation at 395 nm.

Table 2. UV/Vis data of the nitroxide complexes[a]

Complex	λ (nm) /ε (M^{-1} cm^{-1})
[Gd(L$_1$)]$^{3+}$	245 (30 040), 292 (5770), 363 (3000)
[Eu(L$_1$)]$^{3+}$	245 (26 970), 292 (8110), 363 (4270)
[Dy(L$_1$)]$^{3+}$	245 (34 190), 292 (3620), 363 (1800)
[Yb(L$_1$)]$^{3+}$	246 (22 720), 292 (5220), 363 (2700)
[Gd(L$_2$)]$^{3+}$	292 (4020), 363 (1760)
[Eu(L$_2$)]$^{3+}$	292 (3900), 363(1860)
[Dy(L$_2$)]$^{3+}$	288 (7090), 363 (1840)
[Yb(L$_2$)]$^{3+}$	279 (8780), 363 (918)

[a] In 0.5 mM CH_3CN solutions containing TBAP as supporting electrolyte (0.1 M). *T* = 298 K. In order to ensure full nitroxide radical formation the complexes were subjected to electrolysis at a potential intermediate between the TEMPOH/TEMPO and $TEMPO^+$/TEMPO redox couples.

Luminescence measurements were also used to determine the lifetimes of the complexes via excitation at 396 nm in both H_2O and D_2O. For **[Eu(L$_1$)]$^{3+}$** we determined t=0.43 s in H_2O and t=1.32 in D_2O, representing one water molecule in the coordination sphere of the Ln^{III} ion. This result is mostly as expected due to the potential eight coordinate macrocyclic ligand. Complex **[Eu(L$_2$)]$^{3+}$** demonstrates t=0.44 s and t=1.36 s for H_2O and D_2O, respectively, again calculating one coordinated water molecule. Hence the coordination number evolves upon substitution of one amide with the alkyne, but the number of coordinated water remains constant, q = 1. This water molecule is crucial in the use of these complexes in imaging.

2.4 Electron Paramagnetic Resonance spectroscopy

Both the lanthanide ions and the ligands harbour unpaired electrons and hence can be analysed by electron paramagnetic resonance (EPR).

We first recorded the fluid solution spectra of the free ligands **L$_1$** and **L$_2$** at *T* = 293 K. They show a typical isotropic 3-line pattern centered around g = 2.006 (Fig. 4). The hyperfine coupling constant A_N is 1.7 mT, typical for TEMPO radicals. In the absence of metal ions the broadening of the high field component is solely due to the Brownian rotational motion of the radical. In the fast tumbling regime, the rotational correlation time τ_c can be calculated by using the Kivelson's formula (eq. 1), where C is a constant calculated from the principal values of g and A tensors of the nitroxide radical (6.6 10^{-10} for for di-*tert*-butylnitroxide - this value is not expected to change much between nitroxide radicals),[44, 45] $\Delta H_{(+1)}$ is the peak-to-peak linewidth of the low field line and $h_{(+1)}$ and $h_{(-1)}$ are the height of the low and high field resonances.

$$\tau_c = C\,\Delta H_{(+1)}[(h_{(+1)}/h_{(-1)})^{1/2} - 1] \qquad \text{(Eq. 1)}$$

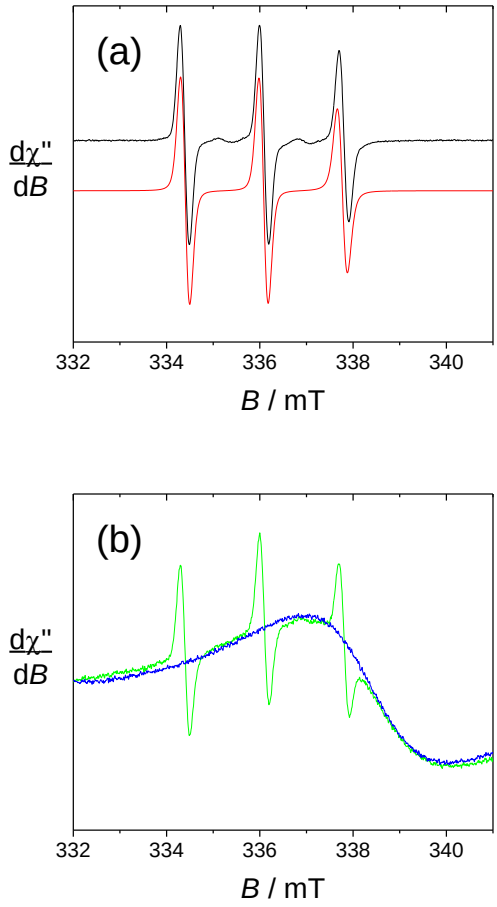


Figure 4 Fluid solution X-Band EPR spectra in water solution. (a) 0.5 mM **L$_1$**; Black lines, experimental spectra; red lines, simulation by using the parameters given in the text. Microwave Freq. 9.42 GHz; power, 3.2 mW, Mod. Amp. 0.2 mT, Freq. 100 KHz. (b) 2 mM **[Gd(L$_1$)]$^{3+}$**; green line. 2 mM **[Gd(L$_1$)]$^{3+}$** in the presence of 2 molar equiv. of ascorbate; blue line. Microwave Freq. 9.42 GHz; power, 1.7 mW, Mod. Amp. 0.2 mT, Freq. 100 KHz.

The rotational correlation times τ_c are 1.6 10^{-10} sec for both **L$_1$** and **L$_2$**, in agreement with their low chemical weight and

comparable hydrodynamic diameters. The spectrum shown in Fig. 4 was simulated accordingly.

The spectra of the gadolinium complexes **[Gd(L$_1$)]$^{3+}$** and **[Gd(L$_2$)]$^{3+}$** at *T* = 293 K show noticeable differences. Both display the sharp 3-line nitroxide pattern (Γ = 0.28 mT) with weaker intensity in comparison to the free ligands. In the case of **[Gd(L$_1$)]$^{3+}$** this 3-line pattern is overlapped with an intense and broad resonance (Γ = 3.24 mT) centered at g = 1.99, which is assigned to the gadolinium signal (Fig. 4). Upon addition of ascorbic acid the TEMPO unit is reduced to TEMPOH, resulting in a quenching of the 3-line pattern without affecting the broad resonance. Strikingly, the gadolinium resonance cannot be observed in the case of **[Gd(L$_2$)]$^{3+}$**. This suggests faster relaxation and hence distinct influence of the nitroxide and coordination environment on the gadolinium relaxation rate.

We also investigated the reversibility of the reduction by adding the oxidant $NaIO_4$ to an ascorbate-reduced sample of **[Gd(L$_2$)]$^{3+}$** (Fig. S15). The initial radical signal of **[Gd(L$_2$)]$^{3+}$** was restored to more than 80 % after 30 min. The ascorbate-reduced sample could be alternatively reoxidized by O_2 bubbling in basic conditions, but the reaction proceeds less efficiently (only 20 % recovery after 30 min).

Low temperature measurements were conducted (4 – 50 K) in order to characterize the Yb, Eu and Dy complexes, wherein fast relaxation prevents observation of the lanthanide at higher temperatures. At *T* = 40 K the complexes exhibit the typical signature of immobilized nitroxide radicals, with broadening and shift of the m_I = +1 and -1 transitions due to anisotropy. Accordingly, similar spectra were obtained for the free ligands in frozen solution (see ESI). Upon cooling down to 5 K, drastic changes were evidenced in the EPR spectra for **[Yb(L$_1$)]$^{3+}$**, **[Dy(L$_1$)]$^{3+}$** and their alkyne counterparts. For a given Ln ion the EPR spectra of the alkyne derivatives **[Ln(L$_2$)]$^{3+}$** show essentially the same temperature dependence than **[Ln(L$_1$)]$^{3+}$** and hence will not be commented on detail.

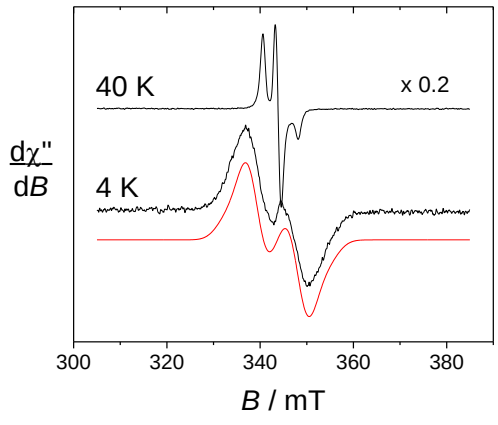


Figure 5 X-Band EPR spectra of **[Yb(L$_1$)]$^{3+}$** in 0.5 mM water solution. Black lines, experimental spectra; red lines, simulation by using the parameters (*S* = 1), |D| = 300 MHz, g_{iso} = 2.00. Microwave Freq. 9.636 GHz; power, 0.2 mW, Mod. Amp. 0.4 mT, Freq. 100 KHz.

The spectrum of **[Yb(L$_1$)]$^{3+}$** taken at 4 K shows an intense doublet at g = 2.00 (Fig. 5), which corresponds to a radical signature but differs markedly from the 3-line pattern of isolated nitroxides. The spectrum of the free ligand **L$_1$** was indeed recorded at this low temperature and shows exclusively the anisotropic 3-line pattern of the nitroxide radical. Hence this behaviour strongly suggests that the paramagnetic lanthanide interacts magnetically with the nitroxide moiety. From the X-Ray crystal structure of **NaL$_1^+$** the N(nitroxide)- metal distance could be estimated at *ca.* 7.8 Å, confirming that a triplet state may be accessible. Considering the feature at 4 K as a spin triplet resonance, and using a simple model with an isotropic g tensor the simulation gives |D| = 300 MHz, which is consistent with the above interspin distance. Note that simulation of this feature by considering an isolated nitroxide anisotropic signal leads to unreasonable ^{14}N hyperfine coupling constants. The addition of ascorbate to **[Yb(L$_1$)]$^{3+}$** results in the full quenching of this feature around g = 2, confirming its assignment. The EPR spectrum of the reduced **[Yb(L$_1$)]$^{3+}$** shows resonances at g = 8.6, 6.41 and 5.2, which are assigned to the parallel features of an isolated Yb^{3+} ion (Fig. 6). The main line at g = 6.41 is assigned to the even isotope, while the satellite ones (approx. 15 % in intensity) correspond to the $^{171}Yb^{3+}$ and $^{173}Yb^{3+}$ isotopes. The $g_{//}$ value of the even isotope compares well with that reported for the $[Yb{\cdot}DTMA{\cdot}OH_2]^{3+}$ complex (6.45),[46] congruent with a similar environment of the metal center.

For **[Dy(L$_1$)]$^{3+}$** the sharp nitroxide lines decrease in intensity upon decreasing the temperature, at the expense of a broader signal centered around g = 2 that we assign to the interacting nitroxide, as observed for **[Yb(L$_1$)]$^{3+}$**. It is distributed over a larger window in comparison to **[Yb(L$_1$)]$^{3+}$**, indicative of a distinct influence of the lanthanide ion on the radical moiety. This is also revealed by the spectrum at *T* = 8K, which still contains both the sharp nitroxide lines and the broader line (see ESI). After addition of ascorbate (reduction of the nitroxide unit) the signal at g = 2 completely disappears, affording a spectrum with an intense resonance at $g_{eff} \approx 11.6$, which is reminiscent of isolated Dy^{3+} complexes (Fig. 6).

The EPR spectrum of **[Eu(L$_1$)]$^{3+}$** at 40 K shows the same nitroxide feature as **[Yb(L$_1$)]$^{3+}$** and **[Dy(L$_1$)]$^{3+}$**. At 4 K the nitroxide signal broadens slightly, but does not show any distinct evolution of its shape, which is likely the consequence of the small magnetic moment of trivalent europium.

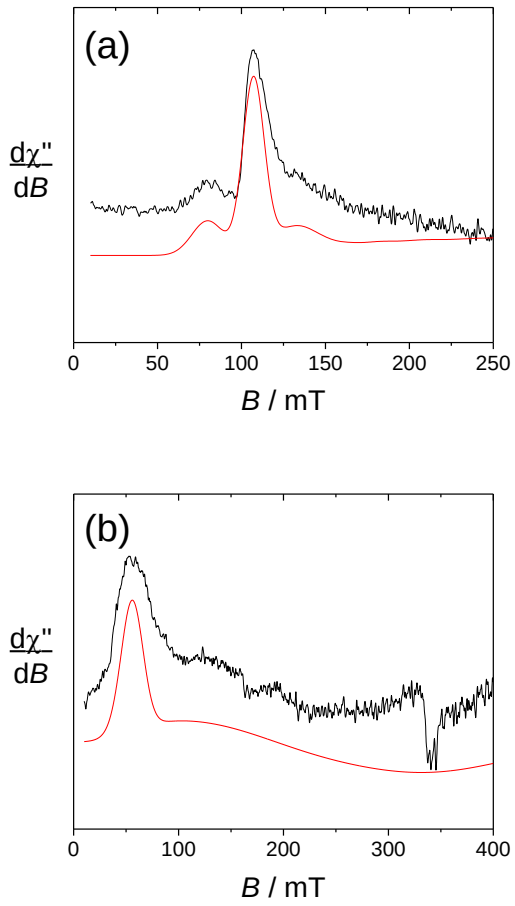


Figure 6 X-Band EPR spectra of (a) **[Yb(L$_1$)]$^{3+}$** and (b) **[Dy(L$_1$)]$^{3+}$** in 4 mM water solution after addition of 2 equiv. of ascorbate. Black lines, experimental spectra; red lines, simulation by using the parameters given in the text. Microwave Freq. 9.636 GHz; power, 2 mW, Mod. Amp. 1 mT, Freq. 100 KHz. T = 5 K.

2.5 DFT calculations on model compounds

The structure of the lanthanide complexes was investigated by DFT calculations. In order to overcome the prohibitive computational cost on complexes of f-metal ions the optimizations were performed on two models wherein the lanthanide ion was replaced by Y^{3+}, namely **YL$_1^{3+}$** and **YL$_2^{3+}$**. The yttrium is indeed a common substitute of Gd^{3+} in imaging. The number of coordinating water molecules was set to one, in agreement with the luminescence data (see above). For **YL$_1^{3+}$** the optimized structure shows a pendent TEMPO arm, with the hydrogens of the water molecule pointing along but opposite to the TEMPO group. For **YL$_2^{3+}$** we first performed a relaxed potential energy surface (PES) scan along the C-N axis of the alkyne arm to determine its preferred orientation. The lowest energy conformation corresponds to the alkyne pointing away from to the metal. It is also interesting to note that the water hydrogens point orthogonal to the alkyne and establish H-bonds with the two acetamide arms. Hence this conformation allows minimization of the steric clash between the coordinating water molecule and the alkyne group, while stabilizing **YL$_2^{3+}$** through H-bonding interactions. The distinct orientation of the water hydrogens in **YL$_1^{3+}$** and **YL$_2^{3+}$** prompted us to perform an additional relaxed PES scan, whereby the water molecule was rotated along the M-O axis (ϕ angle in Fig. 7). For **YL$_1^{3+}$** the variation of energy is monotonic and sinusoidal with alternating minima and maxima every ca. 120° and a total change that does not exceed 0.3 kcal/mol. Model **YL$_2^{3+}$** exhibits a strikingly different behaviour: the total energy change during the scan is one order of magnitude higher, with the lowest energy conformation corresponding to the water hydrogens orthogonal to the alkyne. These results show that the barrier for the rotation of the water molecule is very small for **YL$_1^{3+}$** but significantly higher (immobilized) in **YL$_2^{3+}$**. The functionalization of the macrocycle via an amide linker (TEMPO arm) hence appears to weakly affect the rotational dynamic, in contrast with alkyl linkers (alkyne arm), which do not allow for the formation of stabilizing H-bonds with the coordinated water and prevent rotation. Furthermore the coordination bonds are shorter in **YL$_2^{3+}$**, which is not surprising owing to the lower expected coordination number in this complex.

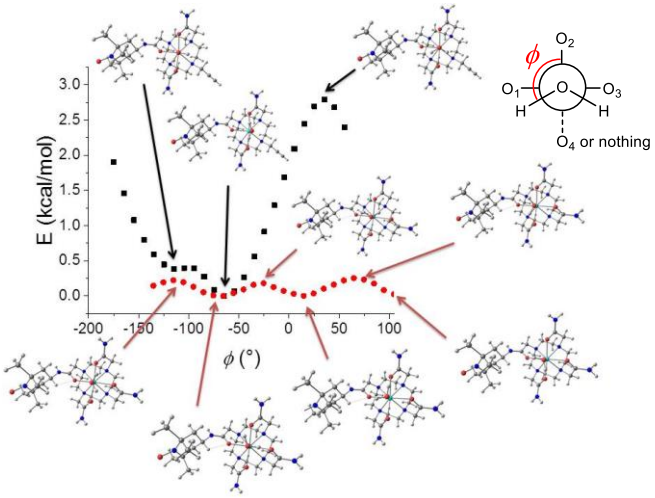


Figure 7. Relaxed PES scan of **YL$_1^{3+}$** and **YL$_2^{3+}$** (B3LYP/6-31g* (C,H,N,O) ECP (Y)). ϕ is the dihedral angle between one water hydrogen, the water oxygen, yttrium and the O2 oxygen of the ligand (amide connecting the TEMPO arm). The O1, O3 and O4 oxygens are from carboxylate groups.

2.6 CEST activity and relaxivity

The ligand design incorporated exchangeable NH_2 protons, which can confer potential CEST activity. In addition, these complexes are 8 or 9-coordinate, thus containing a metal bound water molecule that can confer CEST activity via an exchangeable H_2O on the metal centre. CEST spectra, plotted as a percentage decrease in water proton magnetisation as a function of the presaturation pulse frequency, were recorded in HEPES buffer 0.1 M in H_2O:acetonitrile (2:1) at 500 MHz on both the nitroxide and hydroxylamine forms of complexes **[Ln(L$_1$)]$^{3+}$** and **[Ln(L$_2$)]$^{3+}$** (Ln = Eu^{III}, Dy^{III}, Yb^{III}).

The activity of **[Eu(L$_1$)]$^{3+}$** in its TEMPO form shows a weak CEST spectrum (6% in intensity, Fig. 8) as the relaxation of the water protons neighbouring the TEMPO unit is accelerated by the magnetic moment of the free electron spin of that unit.[32] Upon addition of ascorbic acid the TEMPO unit is reduced, affording a

closed-shell ligand. There is no more acceleration of the relaxation of the vicinal water protons. As a consequence the complex **[Eu(L$_1$)]$^{3+}$** in the reduced form demonstrated a significant CEST peak at + 51 ppm with respect to the bulk water proton resonance. The exchangeable water molecule on the metal centre presents a larger CEST effect of 12 %. The CEST effect demonstrates small pH dependence with values varying from 9 % to more than 13 % between pH 5 and 8. This small pH dependence demonstrates that the metal bound water molecule is little affected by pH changes. The CEST effect in **[Eu(L$_1$)]$^{3+}$** is within the same range, albeit slightly weaker, than for the Eu^{3+} complex of the ligand depicted in Chart 1b in aqueous medium.

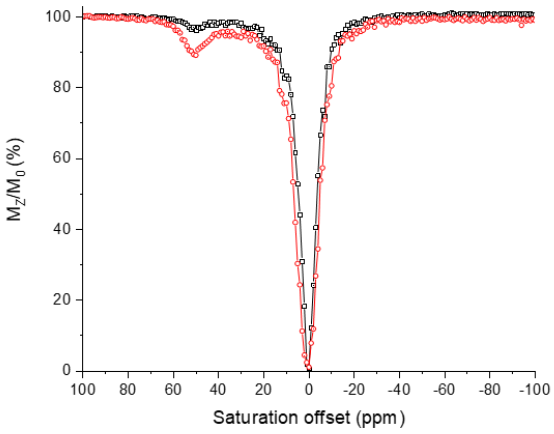


Figure 8. CEST spectrum of 14 mM solution of **[Eu(L$_1$)]$^{3+}$** under its nitroxide (black) and hydroxylamine (red) forms (reduced using sodium ascorbate). Recorded in a water:CD_3CN (2:1) mixture at pH 7.4 (0.1 M HEPES buffer) at 11.7 T. *T* = 298 K, B_1 = 25 µT, irradiation time of 3 s.

Calculation of the rate of exchange via Omega plot calculation and confirmed by RL QUEST plots give approximately 4200 s^{-1} as the k_{ex} for the complex **[Eu(L$_1$)]$^{3+}$**, this corresponds to a proton residence time of 238 µs.

Not surprisingly, both **[Dy(L$_1$)]$^{3+}$** and **[Yb(L$_1$)]$^{3+}$** did not show any significant activity in the PARACEST spectrum, presumably due to faster exchange of the NH_2 protons mediated by these metal ions.

Surprisingly, the PARACEST spectra for **[Eu(L$_2$)]$^{3+}$** under both its TEMPO and TEMPOH form demonstrated no CEST effect for the NH exchangeable protons. This result is in line with the distinct coordination number and conformation of the complex due to the presence of the alkyne. The present data do not yet allow for further rationalizing this behaviour. Not surprisingly, complexes **[Dy(L$_2$)]$^{3+}$** and **[Yb(L$_2$)]$^{3+}$** demonstrated no activity like that of the **[Dy(L$_1$)]$^{3+}$** and **[Yb(L$_1$)]$^{3+}$** (Figures S18, S20).

2.7 Relaxometry

The slow electronic relaxation of the f^7 Gd^{III} ion makes this metal ion inefficient as PARACEST agent, but ensures its importance as relaxation CA. The relaxivity r_1 has been studied by Fast Field Cycling (FFC)-NMR. It shows interesting behaviour. At high field (30MHz), **[Gd(L$_1$)]$^{3+}$** provides a T_1 relaxation time of 0.26s corresponding to a r_1 of 1.92 $mM^{-1}.s^{-1}$ and **[Gd(L$_2$)]$^{3+}$** yields a T_1 relaxation time of 0.16s corresponding to a r_1 of 3.18 $mM^{-1}.s^{-1}$. A very small change between the oxidised and reduced species of 2% and 4% in relaxivity was observed for **[Gd(L$_1$)]$^{3+}$** and **[Gd(L$_2$)]$^{3+}$**, respectively.

At low field (30kHz), **[Gd(L$_1$)]$^{3+}$** provides a T_1 relaxation time of 0.15s corresponding to a r_1 of 3.33 $mM^{-1}.s^{-1}$ and **[Gd(L$_2$)]$^{3+}$** yields a T_1 relaxation time of 0.09s corresponding to a r_1 of 5.28 $mM^{-1}.s^{-1}$. Upon reduction of nitroxide into hydroxylamine, the relaxivity decreases by 13% and 44% for **[Gd(L$_1$)]$^{3+}$** and **[Gd(L$_2$)]$^{3+}$**, respectively.

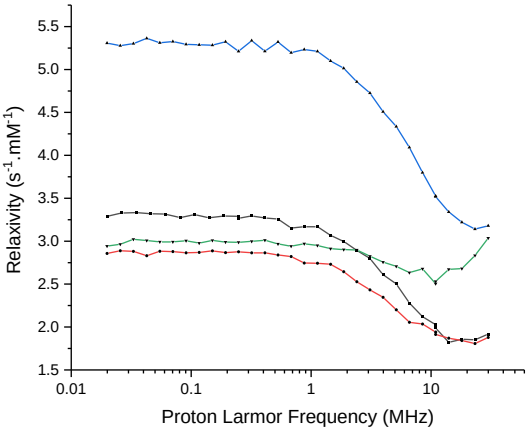


Figure 9. FFC-NMR profiles of 2 mM solution of **[Gd(L$_1$)]$^{3+}$** under its nitroxide (black) and hydroxylamine (red) forms and **[Gd(L$_2$)]$^{3+}$** under its nitroxide (blue) and hydroxylamine (green) forms (reduced using sodium ascorbate). Recorded in water at pH 7.4 (0.1 M HEPES buffer) at 298 K.

In the high-field region, it is known that low exchange rate of the water molecules in the first coordination sphere and fast rotational correlation time of the complex limit the relaxivity of a Gd^{III}-based CA. Besides, in the low-field region, fast electronic relaxation limits the relaxivity under 0.1 T (3-4 MHz).[47,13-15] High relaxivity values in this region are typical of highly symmetric Gd^{III} complexes with long electronic relaxation times.[48] Complex **[Gd(L$_1$)]$^{3+}$** exhibits a higher symmetry than **[Gd(L$_2$)]$^{3+}$** and consequently presents a longer electronic relaxation time, allowing visualization of the Gadolinium signal in the EPR spectrum (0.34 T). Surprisingly, **[Gd(L$_2$)]$^{3+}$** in its reduced form shows similar relaxivity as **[Gd(L$_1$)]$^{3+}$** at low field, but significantly higher value at high field. Much more research beyond the objectives of this article is yet needed to unravel the physical origins of the observed effects.[14] It is worth noting that the influence of the redox status on the relaxivity is weak for **[Gd(L$_1$)]$^{3+}$**, ranging from 20% increase in the low field region to less than 5% in the high field region upon oxidation of the hydroxylamine moiety. Complex **[Gd(L$_2$)]$^{3+}$** shows a better response to the redox stimulus with up to 75% increase in the low field region upon hydroxylamine to nitroxide conversion. At the highest field the difference in relaxivity between both complex oxidation state remains very weak.

The T_1 relaxation time of the water protons is modulated by the redox state without influencing the exchange rate value as demonstrated by D. Sherry and coworkers.[32] This is an interesting trend whereby the formation of a paramagnetic radical center in the vicinity of the Gd^{3+} alters significantly the relaxation of the protons of the coordinated water. Hence complexes **[Gd(L$_1$)]$^{3+}$** and **[Gd(L$_2$)]$^{3+}$** could be considered as gadolinated redox probes though the relaxivity values are still too low to envisage their use as functional MRI probes.

2.8 Biological studies

The toxicity of complexes **[Ln(L$_1$)]$^{3+}$** and **[Ln(L$_2$)]$^{3+}$** has been investigated by MTT assays on M21 cell lines. The range of concentration tested was 1-40 μM, with incubation times up to 72 h. The cell viability upon treatment with **[Dy(L$_1$)]$^{3+}$** and **[Dy(L$_2$)]$^{3+}$** is depicted in Fig. 9 (see ESI for the other compounds). Even after the longest incubation time the viability of the cells treated by up to 40 μM **[Ln(L$_1$)]$^{3+}$** was similar to that of the control sample without complex. Hence all the compounds of this series show negligible cytotoxicity, which is a pre-requisite for their use as imaging agents. Complexes of the **[Ln(L$_2$)]$^{3+}$** series demonstrate slight toxicity above 1 μM, with invariably 60-70% of cell viability for treatments with up to 40 μM of the compounds. The difference of toxicity is clearly inferred to the substitution of one acetamide arm by the alkyne group, but it is yet unclear why this modification has such an effect.

One possible explanation for the absence of toxicity of the **[Ln(L$_1$)]$^{3+}$** series is a lack of cellular penetration, in line with clinically approved MRI CAs, which are injected in the circulating system. We quantified the cellular penetration of one representative complex **[Dy(L$_1$)]$^{3+}$** by using EPR spectroscopy. The complex was incubated with M21 cells at 200 μM, which was a minimal concentration for straightforward detection. The supernatant, cell pellets and cell lysate were next analysed by EPR and the intensity of the nitroxide lines was quantified under non-saturating conditions. No signal was detected within the detection limit of EPR (<0.1% penetration) in either the cell pellets or the cell lysate. The complex was found to be localized in the supernatant, as demonstrated by a similar integration of the nitroxide resonances of the supernatant and a blank containing the complex at similar concentration in a culture medium lacking cells. Both the polar nature of the complex and its tripositive nature most likely prevent cell penetration, which may be an advantage for use as MRI agent. It must be stressed that the alkyne moiety does not affect the cell penetration, but offers a possibility for vectorising the complex by adequate functionalization.

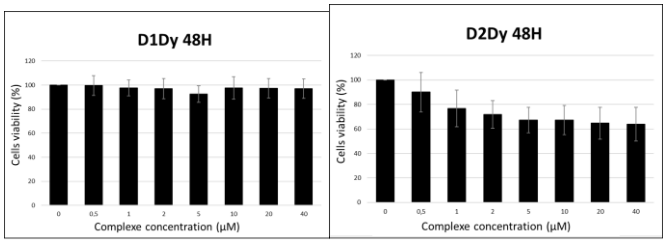


Figure 9. Cell viability after 48h treatment of M21 cells with increasing amounts of (left) **[Dy(L$_1$)]$^{3+}$** and (right) **[Dy(L$_2$)]$^{3+}$**

3. Conclusions

In summary we have prepared two series of lanthanide(III) complexes (Gd, Eu, Dy, Yb) from two DOTA acetamide derivatives appended by one TEMPO arm. One features three acetamide groups (**[Ln(L$_1$)]$^{3+}$**) while the other (**[Ln(L$_2$)]$^{3+}$**) harbours two acetamides and one alkyne group for further functionalization or vectorization through click chemistry. By luminescence measurements we demonstrate that both **[Eu(L$_1$)]$^{3+}$** and **[Eu(L$_2$)]$^{3+}$** possess one water molecule in their first coordination sphere, inferring that the metal ion is 9- or 8-coordinated, respectively. The distinct coordination polyhedron between the series is confirmed by EPR spectroscopy. The lanthanide complexes are redox-active in a physiological window, with the TEMPO being for instance readily reduced into hydroxylamine by sodium ascorbate. Complex **[Eu(L$_1$)]$^{3+}$** is CEST-active under its hydroxylamine form (CEST effect of 12 %). Under its nitroxide form the paramagnetism of the ligand accelerates spin relaxation, resulting in a significant quenching of the CEST effect. Interestingly, **[Eu(L$_2$)]$^{3+}$** does not show a CEST activity, which might be the consequence of different environment and mobility of the coordinated water molecule. The relaxivity of the gadolinium complexes is affected by change in oxidation state of the TEMPO arm, whereby r_1 decreases by about 15% at 10 MHz upon reduction. It must be stressed that the influence of the nitroxide's redox state is about 10 times greater at low field than at high field.

Finally, the complexes show only weak or no toxicity at all at concentrations up to 40 μM against M21 cells, which is inferred to a disfavoured cellular penetration. These results demonstrate that the original concept of redox-active lanthanide complexes has large potentialities for biological imaging (CEST, MRI). Further works are currently in progress in our laboratory to increase the CEST effect and incorporate sensitizing units for bimodal magnetic/optical imaging.

4. Experimental Section

4.1 Materials and methods

All chemicals were of reagent grade and were used without purification. High resolution mass spectra were recorded on a Waters Xevo G2-S QTof apparatus. The UV/Vis spectra were recorded on a Cary Varian 50 spectrophotometer. NMR spectra were recorded on a Bruker AM 400 (^{1}H at 400 MHz) spectrometer. Chemical shifts are quoted relative to tetramethylsilane (TMS). X-band EPR spectra were recorded on a Bruker EMX Plus spectrometer equipped with a Bruker Helium flow cryostat and a dual mode cavity. The relaxivity values r_1 were derived from the experimental longitudinal relaxation times T_1 of the water protons.[14] The T_1 values were measured from 20 kHz to 30 MHz with a commercial Spinmaster FFC 2000 Stelar relaxometer [49, 50] (Stelar s.r.l., Mede PV, Italy). The prepolarized (PP) and non-polarized (NP) sequences (see Fig. 1 of Ref. 49) were used below and above about 12 MHz, respectively. A high polarization field B_{pol} corresponding to a proton resonance frequency of 28 MHz was employed in the PP

experiments. For each magnetic field r_1 was calculated according to $r_1 = (R_1-R_{10})/[complex]$, where R_1 is the relaxation rate measured in the presence of the complex, R_{10} the relaxation rate measured in the absence of complex (typically 0.4 s^{-1}). The complex concentration is expressed in mM. Cyclic voltammetry curves as well as differential pulse voltammetry curves were recorded using a CHI 620 potentiostat. The measurements were performed in 0.5 mM CH_3CN solution containing 0.1 M tetra-*n*-butyl ammonium perchlorate (TBAP) as supporting electrolyte. Experiments were performed in a standard three-electrode cell under argon atmosphere. A glassy carbon disc electrode (3 mm diameter), which was polished with 1 mm diamond paste, was used as the working electrode. The auxiliary electrode is a compartmentalized platinum wire, while an $Ag/AgNO_3$ 0.01 M electrode was used as reference.

Luminescence spectra of the lanthanide complexes were recorded using a modular Fluorolog FL3-22 spectrometer Horiba-Jobin Yvon-Spex equipped with a double grating excitation monochromator and an iHR320 imaging spectrometer. Hamamatsu R928P and Hamamatsu R5509 photomultipliers were used for visible and NIR measurements, respectively. All spectra were corrected for detection and optical spectral response (instrumental functions) of the spectrofluorimeters. Quartz capillaries 4 mm in diameter were used. For the acquisition of the excitation and emission spectra in the NIR, a long pass coloured filter was always used at 870 nm to block the signal of the 2nd harmonics. The q value was determined by using the formula $q = 1.2\ [\Delta k_{obs} - 0.25 - 0.075q^{CONHR}]$,[51] which gives for the **$[Ln(L_1)]^{3+}$** series $q = 1.2\ [\Delta k_{obs} - 0.55]$ and for the **$[Ln(L_2)]^{3+}$** series $q = 1.2\ [\Delta k_{obs} - 0.48]$.

4.2 Synthesis

4-Amino-2,2,6,6-tetramethylpiperidine-1-oxyl, free radical was purchased from Sigma-Aldrich; 1,4,7,10-Tetraazacyclododecane-1,7-bisacetamide (**2**) was synthesized according to literature procedures[52] .

4-(2-chloroacetamido)-2',2',6',6,'-tetramethylpiperidine-1'-oxyl, free radical (0)

To a solution of 4-amino-tempo, free radical (0.200 g, 1.17 mmol) in CH_2Cl_2 (7 mL) was added Et_3N (310 µL, 2.2 mmol, 1.9 eq). Chloroacetyl chloride (140 µL, 1.8 mmol, 1.5 eq) in CH_2Cl_2 (5 mL) was added dropwise at 0°C. After the addition was complete, the reaction mixture was allowed to warm to room temperature and stir for 6 hours. The reaction mixture was quenched with 3 mL of water. The organic phase was washed with HCl (1M) (2x10 mL) and $NaHCO_3$ sat (2x20 mL). The organic phase was dried using $MgSO_4$ and the solvent was removed under reduced pressure. The product was purified by column chromatography (CH_2Cl_2/MeOH; 95/5; v/v). The solvent was removed under reduced pressure to obtain the title compound as a red solid (85%). MS-ESI *m/z* $[M+H]^+$: 247.08.

1-(2',2',6',6,'-tetramethyl-1'-oxyl-4'-piperidyl)-amide-1,4,7,10-Tetraazacyclododecane (1)

Under argon atmosphere, to a solution of cyclen glyoxal (0.536 g, 2.8 mmol) in THF (10 mL) was added 0 (0,750 g, 3.0 mmol, 1.1 eq) in THF (5 mL). The reaction mixture was stirred at room temperature for 24 hours. A precipitate was formed, isolated by filtration and washed with diethyl ether. The resulting solid was dissolved in THF (20 mL) and o-phenylenediamine (0.108 g, 1.0 mmol, 1.1 eq) was added. The mixture was stirred at room temperature for 7 days. The residue was then filtrated and rinced with diethyl ether and the desired product was obtained in 33% yield. ^{1}H NMR (400 MHz, MeOD) δ = 3.10-2.80 (m br, 16H), 1.80 (d br, *J* = 12.5 Hz, 2H), 1.46 (t, *J* = 13.3 Hz, 2H), 1.17 (d, *J* = 5.6 Hz, 12H); ^{13}C NMR (100 MHz, MeOD) δ = 78.5, 62.7, 58.9, 50.3, 44.7, 31.0, 19.2; IR: 3275, 3038, 2975, 2930, 2822, 1653, 1534, 1458, 1362, 1243, 1179, 1112, 929, 731; HR-MS *m/z* calcd for $C_{19}H_{40}N_6O_2$, $[M+H]^+$: 384.321; found, 384.321.

1-(2',2',6',6,'-tetramethyl-1'-oxyl-4'-piperidyl)-amide-1,4,7,10-Tetraazacyclododecane-4,7,10-bisacetamide (L_1)

The monosubstituted cyclen 1 (0,334 g, 0.87 mmol, 1eq) was dissolved in 20 mL of MeCN. 2-bromoacetamide (0.419 g, 3.0 mmol, 3.5 eq), potassium carbonate (0.420 g, 3,0 mmol, 3.5 eq) and potassium iodide (0.540 g, 3.3 mmol, 3.8 eq) was added. The reaction was stirred at 50°C for 5 days. The solution was filtrated and the filtrate was evaporated under reduced pressure. The product was purified by an alumina column chromatography (EtOH/NH_4OH) to afford the title compound as an orange solid (0.587 g, 99%). ^{1}H NMR (400 MHz, D_2O) δ = 3.26-2.10 (m br, 24H), 1.82 (d, J = 11.3 Hz, 2H), 1.42 (m, 2H), 1.16 (m, 12H); ^{13}C NMR (125 MHz, D_2O) δ = 177.0, 174.8, 174.7, 145.2, 130.3, 69.8, 69.1, 61.0, 57.6, 56.9, 55.9, 55.7, 17.0; IR: 3375, 3175, 2966, 2836, 1661, 1549, 1457, 1384, 1328, 1284, 1110; HR-MS *m/z* calcd for $C_{25}H_{49}N_9O_5$ $[M]^+$: 555.385; found, 555.385.

General procedure for the synthesis of the complexes $[Ln(L_1)]^{3+}$:

The ligand L_1 was dissolved in water (10 mL) and the lanthanide salt was added to the solution. The pH was adjusted to 7 by the addition of sodium hydroxide (1 M). The reaction mixture was stirred and heated to 50°C for 1 day. Water was evaporated in reduced pressure. The yellow solid was dissolved in minimum of methanol and precipitated in diethyl ether. It was then dissolved in acetonitrile, filtrated and solvent was evaporated to give the desired compound **$[Ln(L_1)]^{3+}$** as an orange powder. **$[Yb(L_1)]^{3+}$**: from L_1 (100 mg, 0.15 mmol) and ytterbium triflate (93 mg, 0.15 mmol). Yield: 100%. ^{1}H NMR (400 MHz, D_2O) δ = 97.71, 94.42, 92.95, 19.51, 18.21, 16.86, 16.65, 15.97, 15.80, 15.24, 14.86, 12.31, 8.56, 4.12, 3.73, 1.57, 1.48, 0.13, -1.71, -24.16, -25.80, -26.42, -29.00, -31.27, -34.80, -53.92, -57.43, -58.00; IR (cm^{-1}): 3361, 2981, 2927, 2867, 1667, 1626, 1458, 1323, 1246, 1161, 1083, 1030, 924, 668, 639; MS-ESI *m/z* an appropriate isotope pattern was observed for $C_{26}H_{49}F_3O_8N_9SYb$ $[M+CF_3SO_3]^{2+}$: 439.14; found, 440.77. $[Eu(L_1)]^{3+}$: from L_1 (187 mg, 0.27 mmol) and europium triflate (164 mg, 0.27 mmol). Yield: 49%. IR (cm^{-1}): 3354, 3197, 2987, 1658, 1628, 1590, 1459, 1391, 1242, 1225, 1173, 1079, 1027, 639. **$[Gd(L_1)]^{3+}$**: from L_1 (30 mg, 0.06 mmol) and gadolinium triflate (34 mg, 0.06 mmol). Yield: 37%. IR (cm^{-1}): 3367, 3291, 3208, 2984, 2939, 2876, 1661, 1627, 1458, 1245, 1169, 1027, 634; HR-MS *m/z* calcd for $C_{26}H_{49}O_8N_9F_3GdS$ $[M+CF_3SO_3]^{2+}$,

431.130; found 431.131. **[Dy(L_1)]$^{3+}$**: from L_1 (100 mg, 0.15 mmol) and dysprosium triflate (92 mg, 0.15 mmol). Yield: 27%. ^{1}H NMR (400 MHz, D_2O) δ = 18.1, -2.6, -18.3; IR (cm^{-1}): 3331, 3172, 2982, 1655, 1621, 1588, 1454, 1403, 1322, 1272, 1244, 1080, 632.

4-propargyl-1,4,7,10-Tetraazacyclododecane-1,7-bisacetamide (3)

To a solution of 2 (1.31 g, 4.6 mmol) and NaOAc (0.804 g, 9.9 mmol, 1.2 eq) in EtOH (200 mL) was cooling at 0°C. Propargyl bromide (580 µL, 5.2 mmol, 1.2 eq) in EtOH (100 mL) was added dropwise at 0°C. After the addition was complete, the reaction mixture was allowed to stir for 4 days at 50°C. The solvent was removed under reduced pressure. The residue was taken up in MeCN and filtered. The solvent was removed under reduced pressure to obtain the title compound as a brown oil (93%). ^{1}H NMR (400 MHz, MeOD) δ = 3.45 (d, J = 2.2 Hz, 1H), 3.38-3.36 (m, 2H), 3.30 (s, 2H), 3.19 (s, 2H), 2.76-2.56 (m, 16H); ^{13}C NMR (100 MHz, MeOD) δ = 176.4, 168.2, 77.9, 75.6, 59.2, 53.4, 53.0, 51.3, 46.8, 39.7, 30.7; IR: 3283, 2962, 2842, 2689, 2464, 2062, 1654, 1586, 1457, 1372, 1341, 1118, 973, 758; HRMS (m/z): [M+H]$^+$ calcd for $C_{15}H_{29}N_6O_2$, 325.235; found, 325.235.

1-(2',2',6',6,'-tetramethyl-1'-oxyl-4'-piperidyl)-amide-4-propargyl-1,4,7,10-Tetraazacyclododecane-1,7-bisacetamide (L_2)

To a solution of 3 (1.39 g, 4.3 mmol, 1 eq) in EtOH (100 mL), K_2CO_3 (1.18 g, 8.6 mmol, 2 eq), KI (1.46 g, 8.8 mmol, 2.1 eq) and 1 (1.18 g, 4.8 mmol, 1.1 eq) was added. The reaction mixture was stirred for 5 days at 60°C. The solution was filtered on celite, rinsed EtOH and the solvent was removed under reduced pressure. The product was purified by an alumina column chromatography (AcOEt then EtOH/NH_4OH). The solvent was removed under reduced pressure to obtain the title compound as an orange solid (88 %). Slow evaporation of water led to orange crystals which were determine by X-ray crystallography. Reduction with sodium ascorbate was used to characterized the ligand. ^{1}H NMR (400 MHz, D_2O) δ = 4.53 (s, 2H), 4.22 (t, J = 10.5 Hz, 1H), 3.56 (br s, 8H), 3.02 (s, 16H), 1.94 (d, J = 13.0 Hz, 2H), 1.60 (t, J = 13.0 Hz, 2H), 1.28 (d, J = 16.4 Hz, 12H); ^{13}C NMR (100 MHz, D_2O) δ = 78.4, 69.6, 62.5, 43.1, 29.9, 19.6; IR: 3360, 3283, 3194, 3123, 2977, 2941, 2825, 1653, 1568, 1456, 1363, 1320, 1108, 647; HRMS (m/z): [M+H]$^+$ calcd for $C_{26}H_{48}N_8O_4$, 536.379; found, 536.379.

General procedure for the synthesis of the complexes [Ln(L_2)]$^{3+}$:

The ligand **L_2** was dissolved in water (20 mL) and the lanthanide salt was added to the solution. The pH was adjusted to 6.5 by the addition of sodium hydroxide (1 M). The reaction mixture was stirred and heated to 50°C for 2 days. Water was evaporated in reduced pressure. Purification by precipitation in diethyl ether gives the desired compound **[Ln(L_2)]$^{3+}$** as an orange powder. **[Eu(L_2)]$^{3+}$:** from **L_2** (300 mg, 0.45 mmol) and europium triflate (272 mg, 0.45 mmol). Yield: 83%. ^{1}H NMR (400 MHz, MeOD) δ = 18.1, -2.1, -6.3, -7.3, -9.0, -10.9, -12.7, -17.0; IR (cm^{-1}): 3369, 3244, 2958, 1667, 1606, 1454, 1233, 1172, 1029, 647; ESI-MS *m/z* an appropriate isotope pattern was observed for [M]$^{3+}$: $C_{26}H_{47}N_8O_4Eu$, 229.43; found, 229.39. Elemental analysis: calculated for $C_{26}H_{47}N_8O_4EuI_2Cl.4H_2O$; C: 29.76, H: 5.30 N: 10.68, observed C: 29.71 H: 5.30 N: 10.62. **[Yb(L_2)]$^{3+}$:** from L_2 (300 mg, 0.45 mmol) and ytterbium triflate (281 mg, 0.45 mmol). Yield: 96%. ^{1}H NMR (500 MHz, MeOD) δ = 77.8, 65.8, 57.9, 46.5, 36.9, 22.4, -5.3, -15.7, -17.4, -32.5, -53.2, -83.1; IR (cm^{-1}): 3376, 3270, 2975, 2945, 2876, 1670, 1627, 1463, 1243, 1150, 1081, 1028; ESI-MS *m/z* an appropriate isotope pattern was observed for $C_{26}H_{47}N_8O_4Yb(CF_3SO_3)_4$, [M+4 CF_3SO_3]$^-$: 1305.12; found, 1304.89. **[Gd(L_2)]$^{3+}$:** from L_2 (300 mg, 0.45 mmol) and gadolinium triflate (274 mg, 0.45 mmol). Yield: 84%. IR (cm^{-1}): 3379, 3257, 2972, 2868, 1664, 1615, 1455, 1167, 1080, 1019, 632. **[Dy(L_2)]$^{3+}$:** from L_2 (300 mg, 0.45 mmol) and gadolinium triflate (276 mg, 0.45 mmol). Yield: 87%. IR (cm^{-1}): 3364, 3274, 2969, 2874, 1667, 1618, 1459, 1246, 1174, 1079, 1024, 635.

4.3 X-Ray diffraction. Single crystals were coated with perfluoropolyether, picked up with nylon loops and mounted in the nitrogen cold stream of the diffractometer. Mo-Kα radiation (λ = 0.71073Å) from a Mo-target rotating-anode X-ray source equipped with INCOATEC Helios mirror optics was used. Final cell constants were obtained from least squares fits of several thousand strong reflections. Intensity data were corrected for absorption using intensities of redundant reflections with the program SADABS. The structures were solved by Patterson methods and subsequent difference Fourier techniques. The OLEX software was used for the refinement.[53] All non-hydrogen atoms were anisotropically refined and hydrogen atoms were placed at calculated positions and refined as riding atoms with isotropic displacement parameters. CCDC-1997062 contain the crystallographic data for **[Na(L_2)]$^+$**(Cl^-); these data can be obtained free of charge via http://www.ccdc.cam.ac.uk/conts/retrieving.html.

4.4 DFT Calculations

Full geometric optimizations were performed with the Gaussian 9.0 program.[54] The B3LYP functional[55, 56] was used together with the 6-31g* basis set for the C,H,N atoms[57] and a pseudo potential (LanL2DZ)[58] for the central metal ion. Frequency calculations were systematically performed on the optimized structures in order to ensure that they correspond to a real energy minimum and not a saddle point. The relaxed potential energy surface scans were performed by varying dihedral angles.

4.5 Cell culture

Human melanoma cell line M21 was purchased from ATCC (Molshein, France). Cells were cultured in DMEM medium supplemented with 10% (v/v) fetal calf serum (FCS) and 2 mM glutamine (Thermo Fisher Scientific, Courtaboeuf, France). Cells were maintained at 37 °C in a 5% CO2- humidified atmosphere and tested to ensure freedom from mycoplasma contamination. All cell lines were used within 5-50 passages of thawing the original stocks.

4.6 MTT assay

M21 cells were seeded into 96-well plates (1.5×10^3 cells per well) in 100 µl of culture medium. After 24 h, cells were treated with the lanthanide(III) complexes at various concentrations. Following incubation for 24 h, 48 h or 72 h, 10 µl of a MTT (3-(4,5-dimethylthiazol-2-yl)-2,5-diphenyltetrazolium bromide) stock solution (Euromedex, Mundolsheim, France) in PBS at 5 mg ml^{-1} was added in each well and the plates were incubated at 37 °C for 2 h. To solubilize water-insoluble purple formazan crystals, SDS 10% / HCl 0.1% solution was used. After 24 h, absorbance was measured on an ELISA reader (Tecan, Männedorf, Switzerland) at a test wavelength of 570 nm and a reference wavelength of 650 nm.

4.6 EPR spectroscopy

M21 cells were seeded in 4-well plates (1×10^5 cells per well) in 1 mL of culture medium. After 24 h, cells were treated with the complexes at 200 µM concentration during 5 min, 10 min, 30 min or 1 h. Cell culture medium was recovered and M21 cells where harvested. A fraction of M21 pellets was then washed twice in cold PBS and incubated in lysis buffer (10mmol/L Tris–HCl pH 7.5, 120 mmol/L NaCl, 1 mmol/L EDTA, 1 mmol/L dithiothreitol, 0.5% Nonidet P-40, 0.05% sodium dodecyl sulfate, supplemented with protease inhibitors (Na_3VO_4 and NaF). After 10 min on ice lysates were centrifugated at 20,000 x g for 15 minutes, and soluble fraction was recovered. Cell culture medium, pellets of intact M21 cells or M21 cell lysates were then analyzed by EPR spectroscopy.

Conflicts of interest

There are no conflicts to declare.

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