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Recyclable Synthesis of Cs-ABW Zeolite Nanocrystals from Non-reacted Mother Liquors with Excellent Catalytic Henry Reaction Performance

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ABSTRACT

Nanoscale zeolites are very attractive for advanced applications such as in catalysis, adsorption, supports, sensors and carriers. Green synthesis of these materials is highly appreciated aiming to increase product yield while chemical waste and production cost can be reduced. Here we report high-yield, low-cost and green synthesis of nanosized Cs-ABW zeolite by recycling and using the non-reacted mother liquors waste solutions produced from the hydrothermal Cs-ABW zeolite synthesis. The nanocrystalline Cs-ABW solid product of each successive batch was collected and characterized while the non-reacted mother liquors were recovered and re-used for subsequent synthesis of nanozeolites. The synthesized Cs-ABW zeolite nanocrystals from three subsequent

cycles possessed fairly similar properties and they exhibited excellent catalytic performance in Henry (nitroaldol) reaction of benzaldehyde and nitroethane thanks to their basic sites, $(\text{Si-O-Al})^- \text{Cs}^+$, located at the external surface of the zeolite. This approach hence suggests a low-cost and eco-friendly preparation of Cs-ABW zeolite nanocrystals since less chemical waste is disposed while the product yield can be significantly improved which are ideal for industrial scale-up process.

Keywords: Zeolite; Cs-ABW; Recyclable synthesis; Nanocatalyst; Henry reaction

1. Introduction

Zeolites are hydrated aluminosilicate microporous materials that are widely used in catalysis, adsorption, and ion exchange [1–3]. Recently, zeolites in nanometer scale receive enormous scientific and commercial attentions due to their unique size- and shape-dependent properties (e.g. external surface area, diffusivity, colloid stability, electronic, etc.) [4–9]. While more than 250 types of zeolite known today, only 20 types of zeolite (ABW [10], AEI [11], AFI [12], AFO [13], ANA [14], *BEA [15], CHA [16], EDI [17], EMT [18], FAU [19, 20], GIS [21], LTA [22], LTJ [23], LTL [24], MEL [25], MER [26], MFI [27, 28], MOR [29], SOD [30]) have been successfully synthesized in nanometer scale.

Several efforts have been initiated to synthesize zeolites cost-effectively by optimizing the product yield while minimizing the generation of chemical wastes [31–33]. For instance, Lee et al. reported a method where the organic template can be dissociated, recombined and reused after the zeolite synthesis [31]. Since then, the synthesis of aluminophosphate zeolite materials by

reusing mother liquor wastes with minimal compensation is investigated [32, 33]. Nevertheless, the development of eco-friendly and cost-effective synthesis route of industrially important zeolite catalysts still remains a challenge.

ABW-type zeolite is a small pore (8-membered ring) zeolite with an opening of $3.4 \times 3.8 \text{ \AA}^2$ [34]. Particularly, nanosized ABW zeolite containing Cs^+ extraframework cation is of great interest due to its high basicity property that is demanded in heterogeneous catalysis. In addition, its small porosity is also ideal for the separation of small molecules in membrane technology besides being used as a host for immobilization of radioactive cesium [35]. Recently, nanosized Cs-ABW zeolite has been successfully synthesized free of organic template under relatively low pressure and temperature (180 °C, 22 bar) [10] as compared to the previous works (700-1200 °C, >1000 bar) [36]. As like other nanozeolites, the use of excessive structure-directing agent (CsOH in this case) for obtaining nanosized Cs-ABW with moderate crystalline solid yield is essential as it promotes supersaturation condition for allowing nucleation over crystal growth to take place. The desired solid products after the hydrothermal synthesis are then separated while the mother liquors that often contain non-reacted reagents are usually discarded as waste. This makes Cs-ABW nanozeolite synthesis not economical and environment-unfriendly due to the disposal of large amount of expensive, corrosive and harmful non-reacted chemicals, particularly CsOH.

To make use of the surplus chemicals effectively and further enhance the product yield of Cs-ABW nanozeolite, in this paper, a simple and eco-friendly method by recycling the non-reacted mother liquors to prepare Cs-ABW nanozeolite is reported. More importantly, the resulting nanosized Cs-ABW zeolites from three subsequent cycles of synthesis exhibit almost similar properties and they show excellent catalytic performance in Henry (nitroaldol) reaction.

2. Experimental

2.1. Recyclable synthesis of Cs-ABW zeolite nanocrystals

All the chemicals were purchased from commercial sources and used without further purification. First, clear aluminate solution was prepared by dissolving CsOH·H₂O (28.000 g, ≥99.5%, Sigma-Aldrich) and Al(OH)₃ (3.312 g, extra pure powder, Acros) in distilled water (21.504 g) under stirring (350 rpm) at 105 °C for 16 h. Next, the clear silicate solution was obtained by dissolving CsOH·H₂O (86.104 g) in distilled water (21.304 g) and HS-40 (12.750 g, Sigma-Aldrich). The aluminate solution was slowly mixed with the silicate solution to obtain a clear precursor hydrosol with a molar composition of 4SiO₂:1Al₂O₃:16Cs₂O:164H₂O. The mixture was subjected to hydrothermal heating at 180 °C for 4 h. The mother liquor after crystallization was separated *via* centrifugation (10000 rpm, 10 min) and the water from filtrate was partially evaporated at 55 °C for 96 h to enable chemical compensation analysis. From the wet centrifuge cake (zeolite product), very small amount of solid was taken and dried in order to estimate the mass of water in the wet centrifuge cake. Then, the centrifuged solid was further purified with distilled water using centrifugation (10000 rpm, 10 min) until the pH of the solid reached 8. The purified solid was dried at 65 °C overnight, and the mass of Cs-ABW zeolite solid (denoted A-1) was recorded.

The separated mother liquid, which contained residual inorganic substances and water, was used for the subsequent cycles of synthesis. The chemical composition of the mother liquor and zeolite product was first analyzed (see Section 2.2) before subjecting to chemical compensation with the required amount of chemicals (Table 1). For example, 87.008 g of non-reacted mother liquor after analyzing with ICP-OES was added with 3.038 g of Al(OH)₃, 28.031 g of CsOH·H₂O,

36.803 g of water and 5.811 g of HS-40 so that a clear hydrosol with the same molar composition of $4\text{SiO}_2:1\text{Al}_2\text{O}_3:16\text{Cs}_2\text{O}:164\text{H}_2\text{O}$ was obtained. The crystallization, purification and compensation procedures were repeated and the solid products obtained after second and third subsequent syntheses were named as A-2 and A-3, respectively.

2.2. Characterization

The XRD patterns of the nanocrystalline Cs-ABW samples were collected on a Bruker-AXS D8 X-ray diffractometer with CuK_α radiation operating at 40 kV and 10 mA ($\lambda = 1.5418 \text{ \AA}$, scan speed of $0.2^\circ/\text{min}$, step size of 0.02°). The crystallite size was calculated by Scherrer equation (Eq. 1):

$$L = \frac{K\lambda}{B\cos\theta} \quad (\text{Eq. 1})$$

where L is the average size of the crystals (nm); K is the Scherrer constant, taken as 0.94; λ is the X-ray wavelength (1.5418 \AA); B is the full width at half maximum intensity (FWHM), after subtracting the line broadening (radians) and θ is the Bragg's angle.

The chemical composition of crystalline solids and non-reacted mother liquors was determined using an Varian Vista MPX ICP-OES inductively coupled plasma-optical emission spectrometer (ICP-OES). The samples were first mixed with HF solution to dissolve any non-soluble silica and alumina. H_3BO_3 was then added into the prepared solution to prevent interferences of fluoride in the ICP-OES measurement. Nitrogen adsorption-desorption was performed on a Micrometrics ASAP 2010 analyzer at $-196 \text{ }^\circ\text{C}$. The samples were first degassed under vacuum for 12 h at $300 \text{ }^\circ\text{C}$. The specific surface area of the samples was calculated using the BET equation. The pore size distributions and the average pore size of the zeolites were

calculated from the adsorption branch of the experimental isotherm using the Barrett–Joyner–Halenda (BJH) model. The total pore volume of the solids was computed from the nitrogen adsorbed volume at $P/P_0 = 0.993$. High-resolution transmission electron micrographs were recorded with a field emission TECNAI G² 20 S-TWIN TEM microscope (FEI, 200 kV). The particle size and colloidal stability of Cs-ABW nanocrystals in water were measured using dynamic light scattering (DLS) and zeta potential (ζ) analyses on a Malvern Zetasizer Nano Series. The infrared (IR) spectra of the solid was recorded with a Perkin Elmer spectrometer (System 2000) using KBr pellet technique.

2.3. *Catalytic experiment*

Nitroaldol condensation (Henry reaction) of benzaldehyde and nitroethane was carried out following this procedure: Activated ABW zeolite (0.500 g), benzaldehyde (2 mmol, 99%, Merck) and nitroethane (40 mmol, 97%, Merck) were added into a 10 mL-quartz vial. The vial was capped and subjected to microwave heating at 180 °C using a Monowave 300 microwave reactor (Anton Paar, 850 W output, stirring rate 400 rpm). The zeolite catalyst was separated while the reaction solution was withdrawn and injected into a GC-FID chromatograph (Agilent/HP 6890 GC, HP-5 capillary column) and a GC-MS spectrometer (Perkin-Elmer Clarus 500 system) for chromatography analyses.

3. **Results and discussion**

The recyclable synthesis of nanosized Cs-ABW zeolite is carried out with an initial molar ratio of $4\text{SiO}_2:1\text{Al}_2\text{O}_3:16\text{Cs}_2\text{O}:164\text{H}_2\text{O}$ under hydrothermal condition at 180 °C for 4 h. After crystallization, the solid product is recovered *via* high speed centrifugation; 8.758 g of solid

product (A-1) with 81.85% solid yield is obtained upon drying (Table 1). The non-reacted mother liquor is collected and subjected to hydrothermal treatment using the same condition but no solid product is formed.

The solution is concentrated for analysis prior to undergo chemical compensation and subsequent synthesis cycles. The mass of concentrated waste mother liquor is 87.008 g and the pH of the solution is 14. Hence, it indicates that there is still CsOH remain in the waste mother liquor. The chemical composition of the initial and waste mother liquors is characterized with ICP-OES spectroscopy and the chemicals needed for compensation are calculated and summarized in Table 1. As shown, nearly 24% of CsOH, 90% of Al and 45% of Si sources have been consumed during the crystallization process. As a result, the waste mother liquor does not crystallize any solid product due to insufficient Al and Si nutrients. When the calculated amount of chemicals are compensated, 6.779 g of solid product (A-2) equivalent to 63.36% solid yield is obtained when the second cycle synthesis is performed.

The non-reacted solution after second cycle of synthesis is separated and further subjected to partial water evaporation where 85.571 g of solution remains. The resulting solution has a pH of 14 and the elemental spectroscopy study shows that nearly 23% of CsOH, 74% of Al and 38% of Si sources are required for chemical compensation to allow the crystallization of Cs-ABW zeolite nanocrystals; 6.291 g of solid (A-3) or 58.79% solid yield is produced after the third cycle of hydrothermal synthesis process (180 °C, 4 h).

The crystalline phase of the solids (A-1, A-2 and A-3) is confirmed by XRD analysis (Fig. 1). As shown, the XRD pattern of all synthesized solids exhibits intense peaks at $2\theta = 18.80^\circ$, 26.94° , 27.52° , 32.90° and 38.78° , which are attributed to the (200)/(110), (211), (202)/(112), (310)/(020) and (312)/(022) planes of the ABW-type zeolite, respectively [34]. No extra crystalline

phases such as ANA or tridymite are detected in the XRD patterns, indicating that only nanocrystalline Cs-ABW zeolite is formed in those three samples. The XRD peaks also show almost similar peak intensity and peak broadening indicating that the nanosized Cs-ABW zeolite produced have nearly same crystallite size with same crystallinity [37]. According to Scherrer equation, the crystallite sizes of A-1, A-2 and A-3 are 33.1, 30.8 and 29.6 nm, respectively.

Fig. 2 shows the particle size distribution of the nanosized Cs-ABW crystals produced from the three subsequent cycles of synthesis. The A-1 nanocrystals are sized from 25.2 nm to 86.4 nm and the average crystal size is 49.4 nm. The average crystal size decreases slightly to 46.1 nm and 43.4 nm for A-2 and A-3, respectively, but the particle size distribution remains unaffected which is in line with the XRD observation. The nanocrystals (10 wt%) are dispersed in distilled water and the resulting colloidal solution is investigated by using zeta potential technique for colloidal stability study. The zeta potential values for A-1, A-2 and A-3 suspensions are -52.6 mV, -53.7 mV and -52.7 mV, respectively, which are less than -30 mV. Hence, it indicates good physical colloidal stability of nanosuspensions due to strong electrostatic repulsion of individual particles over attractive forces which prevents particle aggregation [38]. As a result, no sedimentation of solid is observed in the colloidal suspensions even after one month of storage under room conditions.

The morphological properties of nanocrystalline Cs-ABW zeolites prepared from the three subsequent cycles of synthesis are studied by HRTEM microscopy. As shown in Fig. 3, the nanocrystals are very homogeneous in size distribution and the average crystal sizes (A-1: 48.6 nm, A-2: 44.7 nm, A-3: 42.1 nm) are also nearly identical to those determined from the DLS technique (Fig. 2A). In addition, the nanocrystals exhibit orthorhombic-like shape, which is different from that of micron-sized Li-ABW crystals adopting needle-like shape (Fig. 2a–2d) [39,

40]. The difference in morphological properties (size and shape) of nanosized Cs-ABW could be explained by the different synthesis recipe and crystallization condition by which supersaturation is achieved under very high alkalinity condition that favors nucleation over crystal growth [41]. Furthermore, low heating temperature and short crystallization time also facilitate spontaneous controlled nucleation and crystallization while lowering the crystal growth rate as compared to the previous heating temperature ($>700\text{ }^{\circ}\text{C}$) and crystallization time (46 h).

The framework composition of nanocrystalline A-1, A-2 and A-3 solids is determined by using ICP-OES spectroscopy. The Si/Al ratios of the samples are found to be identical and the ratios are close to unity similar to that of the theoretical value indicating homogeneous distribution of Al element in ABW framework. Furthermore, the Cs/Al ratios of the samples are also close to unity because equivalent amount of Cs^+ cation is required to counter balance the negative charge of the tetrahedral Al species in the ABW framework [10].

The nitrogen adsorption-desorption isotherms of A-1, A-2 and A-3 are measured to be type V, according to the IUPAC classification (Fig. 4) [42]. In addition, the hysteresis loops of type H1 are also obviously shown at $P/P_0 = 0.8\text{--}1.0$, indicating that the obtained nanosized Cs-ABW samples in three subsequent syntheses are meso/macroporous materials [43]. The Barrett-Joyner-Halenda (BJH) analysis also proves that these three samples have monomodal meso/macropores distribution ranging from 11 to 112 nm while the mean and average pore sizes of the samples decrease after subsequent cycles of synthesis. The generation of meso/macroporosities in the samples can be resulted from the textural porosities due to different stacking styles of the nanocrystals [44]. Meanwhile, very small amount of micropore surface areas ($<7.2\text{ m}^2/\text{g}$) are measured because nitrogen molecules ($\sigma = 3.64\text{ \AA}$) are hardly accessible to the micropores of Cs-

ABW zeolite ($\sigma = 3.4 \times 3.8 \text{ \AA}^2$) [45]. As a result, only meso/macropores measured as external surface area ($S_{\text{external}} = 26\text{--}31 \text{ m}^2/\text{g}$) are obtained.

Ar adsorption-desorption analysis is further used to study the porous structure of A-1, A-2 and A-3 where Ar atom ($\sigma = 3.40 \text{ \AA}$, $1.64 \times 10^{-24} \text{ cm}^3$) has smaller size and lower average electric dipole polarizability than N_2 molecule ($\sigma = 3.64 \text{ \AA}$, $1.76 \times 10^{-24} \text{ cm}^3$) [46]. It can be seen that all samples exhibit type V with low Ar uptake and very narrow hysteresis loop at high P/P_0 due to the presence of intergranular mesopores (Fig. 5). The S_{BET} of the samples are ca. $34 \text{ m}^2/\text{g}$, which are almost similar to those determined using N_2 probe molecule. Nevertheless, the total pore volumes (ca. $0.05 \text{ cm}^3/\text{g}$) are much lower than those determined using N_2 molecule (ca. $0.25 \text{ cm}^3/\text{g}$). Also, the average pore size determined using the BJH model is not possible because the volume adsorbed in the mesopore region is very low.

A scheme showing the total consumption of chemical reagents by both the conventional and recyclable synthesis approaches for synthesizing nanosized Cs-ABW zeolite in laboratory and industry scales is shown in Fig. 6. For laboratory scale, 3 batches of synthesis for producing 20 g of nanocrystalline Cs-ABW (ca. 6.67 g of zeolite are obtained each run) are considered using an 75-mL autoclave whereas a 100-L reactor is projected to produce 1 tonne of Cs-ABW zeolite (ca. 10 kg of zeolite obtained each run) *via* 100 industrial-scale synthesis batches. Furthermore, an assumption is also made that the amount of the raw materials used is proportional to the amount of zeolite produced as reported in Table 1. The results show that the HS-40 and $\text{CsOH}\cdot\text{H}_2\text{O}$ consumption can be reduced up to 26.6% and 40.7%, respectively, when recyclable synthesis method is applied (three cycles are used) as compared to the conventional synthesis pathway (three batches of synthesis are performed), whereas the consumption of $\text{Al}(\text{OH})_3$ is nearly similar for laboratory-scale synthesis (Fig. 6a). Hence, the production cost, which is the major concern in

industry, can be reduced considerably. The total consumption of raw chemicals at industrial scale on the basis of 1 tonne of nanosized Cs-ABW zeolite is hence estimated for both synthesis methods (Fig. 6b). As seen, the low consumption of costly and corrosive CsOH·H₂O (67.5%) and HS-40 colloidal silica (43.8%) is obvious, hence offering significant cost saving to industry and also environmental protection.

The catalytic behavior of A-1, A-2 and A-3 solids is also tested in Henry (nitroaldol condensation) reaction of nitroethane and benzaldehyde (as a model reaction) using microwave heating (170 °C, 850 W) where this mode of heating provides rapid and homogeneous heating which is different from the conventional oven and oil bath heatings [47]. The results reveal that all three series of nanocrystals synthesized from the recyclable synthesis approach have consistent catalytic activity and they are very active in this Henry reaction with ca. 81.5% conversion after 10 min of microwave heating thanks to their basic sites, (Si–O–Al)[–]Cs⁺, which are located at the external surface of the zeolite (Table 3: Entries 2–4). Their catalytic activity is also compared with sodium hydroxide, potassium hydroxide and sodium carbonate. Interestingly, those three nanocrystalline Cs-ABW zeolites exhibit better catalytic performance than the homogeneous catalysts (Entries 6–8). The catalytic performance of nanocrystalline Cs-ABW zeolites is also compared with those catalyzed by CsOH and Cs₂CO₃. The results show that Cs-ABW nanocatalysts have comparable activity as Cs₂CO₃ (82.0%) in Henry reaction (Entry 9). Meanwhile, CsOH, which is a very basic homogeneous catalyst, gives the highest conversion (92.4%) after 10 min of reaction but it is not reusable (Entry 10). In contrast, the Cs-ABW nanocatalysts can be reused and no significant change in catalyst reactivity is observed even after the 5th run of catalytic reaction (Entry 5), thus offering as an reusable, environment-friendly and promising solid base catalyst both in organic synthesis and chemical manufacturing.

4. Conclusion

In summary, nanosized Cs-ABW zeolite can be prepared in a more cost-effective and environment-friendly way *via* re-using the non-reacted reagents with slight compensation of needed chemicals. The nanocrystals prepared from three subsequent synthesis batches exhibit considerably similar crystallinity, crystallite size, surface area and chemical composition whereas their morphological structure remains preserved. The nanocrystalline Cs-ABW solids prepared from this approach are also colloidally stable in water even without further surface modification. Furthermore, the recyclable synthesized Cs-ABW zeolite nanocrystals also show strong catalytic activity in base-catalyzed nitroaldol condensation of nitroethane and benzaldehyde with ca. 81.5% conversion under microwave heating. More importantly, their catalytic performance is better than that of homogenous catalysts (NaOH, KOH, Na₂CO₃) and is comparable to Cs₂CO₃. Thus, the recyclable synthesis presented in this work is extremely beneficial for environment and industrial scale up process since the synthesis cost and disposal of hazardous and expensive chemical waste can be significantly reduced after the synthesis of nanocrystals, thus offering a promising option for green synthesis of aluminosilicate microporous materials.

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Author Contribution Statement

Tamara Mahmoud Ali Ghrear: Doing experiment

Cyril Vaultot, T. Jean Daou: Performing Ar adsorption, paper proof-reading, response to reviewer

Eng-Poh Ng, Svetlana Mintova, Tau Chuan Ling: Give idea for the entire project, writing manuscript, paper proof-reading, response to reviewer

Soon Huat Tan: Assist in catalytic reaction, writing manuscript, paper proof-reading

Boon Seng Ooi: DLS and zeta potential analysis and discussion, paper proof-reading

Figure captions

Fig. 1. XRD patterns of (a) A-1, (b) A-2 and (c) A-3 Cs-ABW zeolite nanocrystals.

Fig. 2. (A) DLS and (B) zeta potential curves of (a) A-1, (b) A-2 and (c) A-3 Cs-ABW zeolite nanocrystals.

Fig. 3. HRTEM images of (a–c) A-1, (d–f) A-2 and (g–i) A-3 samples at high and very high magnifications. Insets show the lattice fringes of single-crystal of Cs-ABW zeolite captured at selected square areas.

Fig. 4. Nitrogen adsorption-desorption isotherms and (Inset) pore size distributions of (a) A-1, (b) A-2 and (c) A-3 Cs-ABW zeolite nanocrystals.

Fig. 5. Argon adsorption-desorption isotherms of (a) A-1, (b) A-2 and (c) A-3 Cs-ABW zeolite nanocrystals.

Fig. 6. Total consumption of chemical reagents by conventional and recyclable synthesis methods for producing (a) 20 g of nanosized Cs-ABW zeolite in laboratory scale using 3 batches, and (b) 1 tonne of nanosized Cs-ABW zeolite in industry scale (estimated) using 100 batches.

Figures

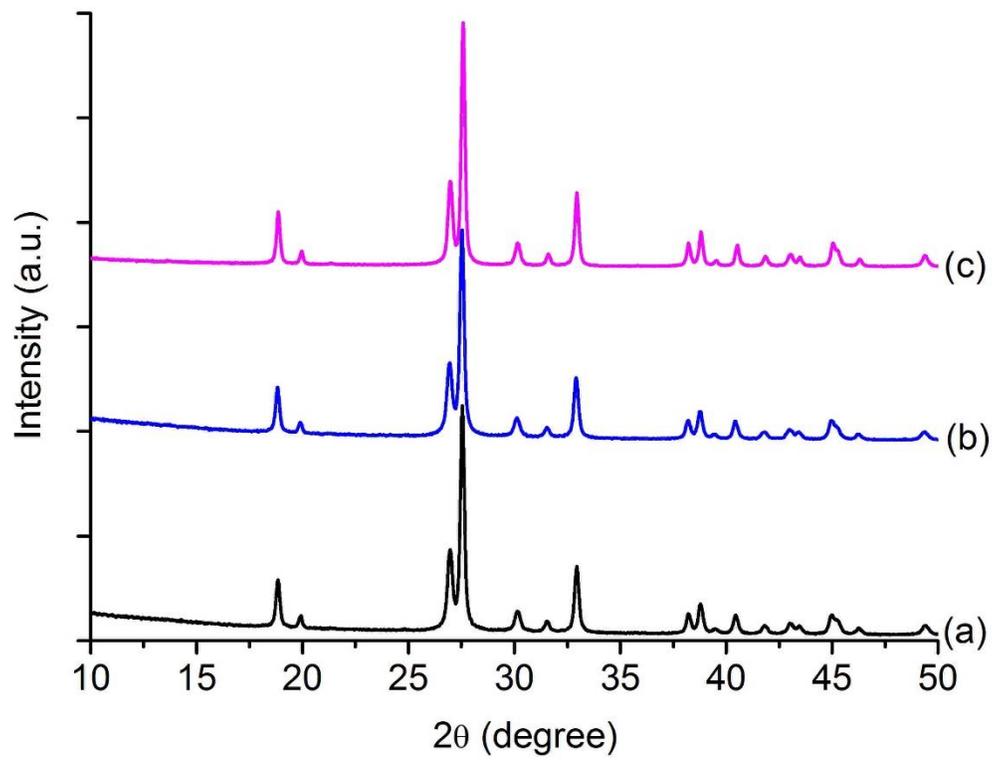


Fig. 1

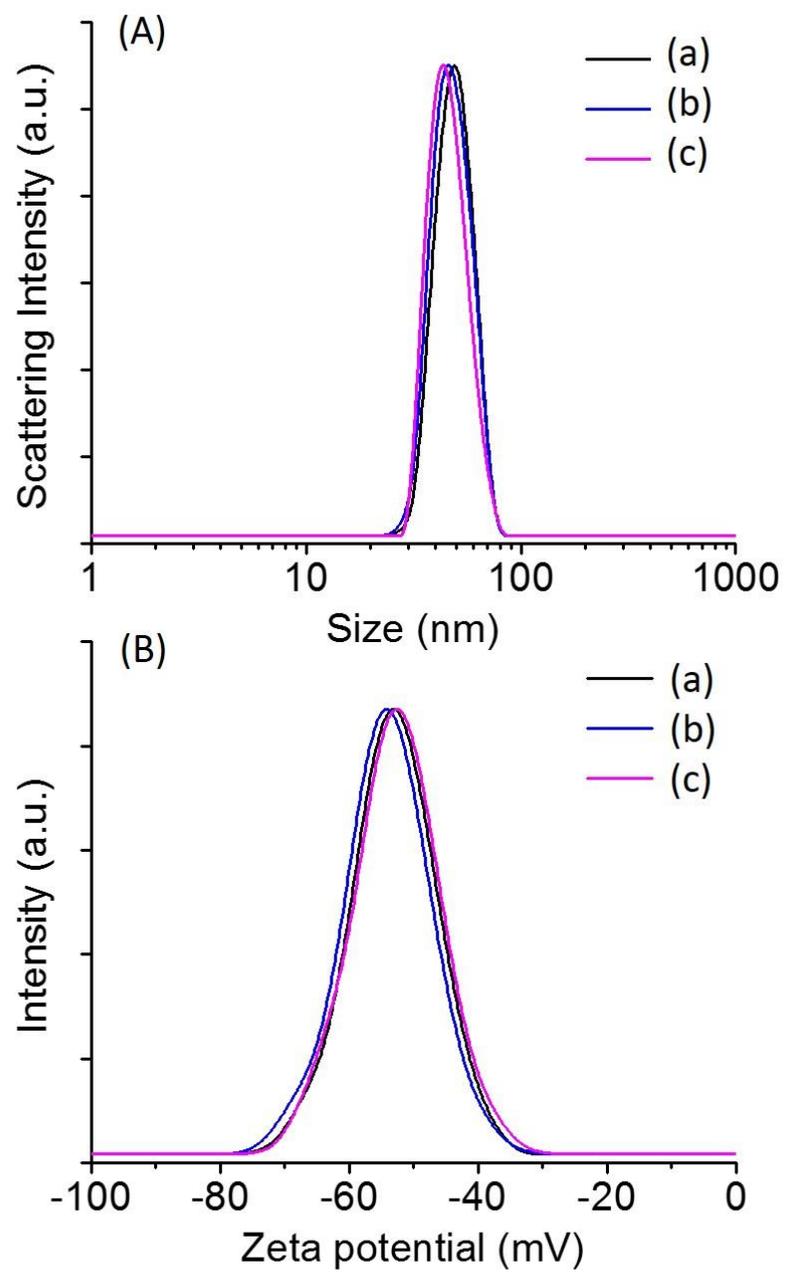


Fig. 2

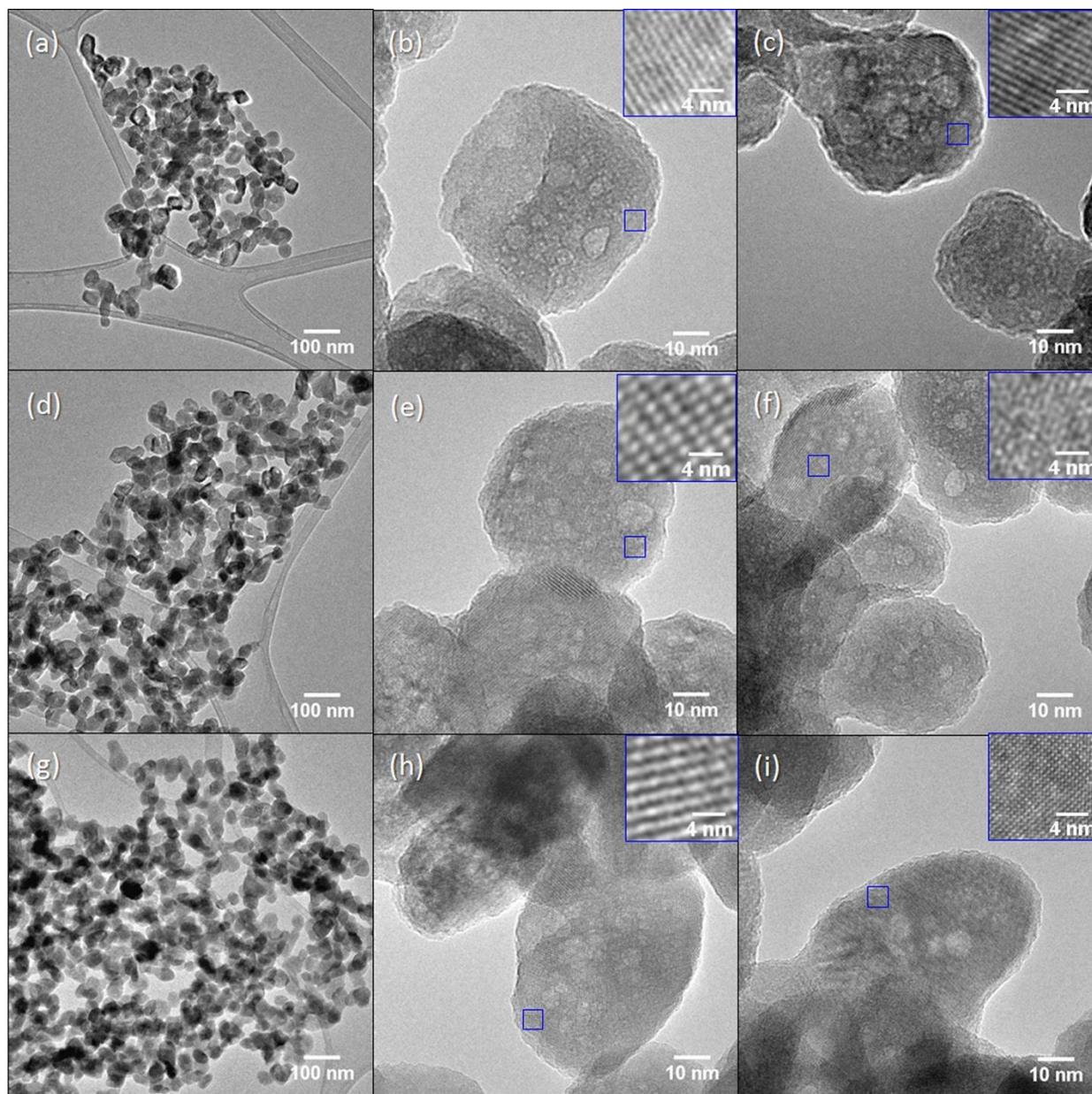


Fig. 3

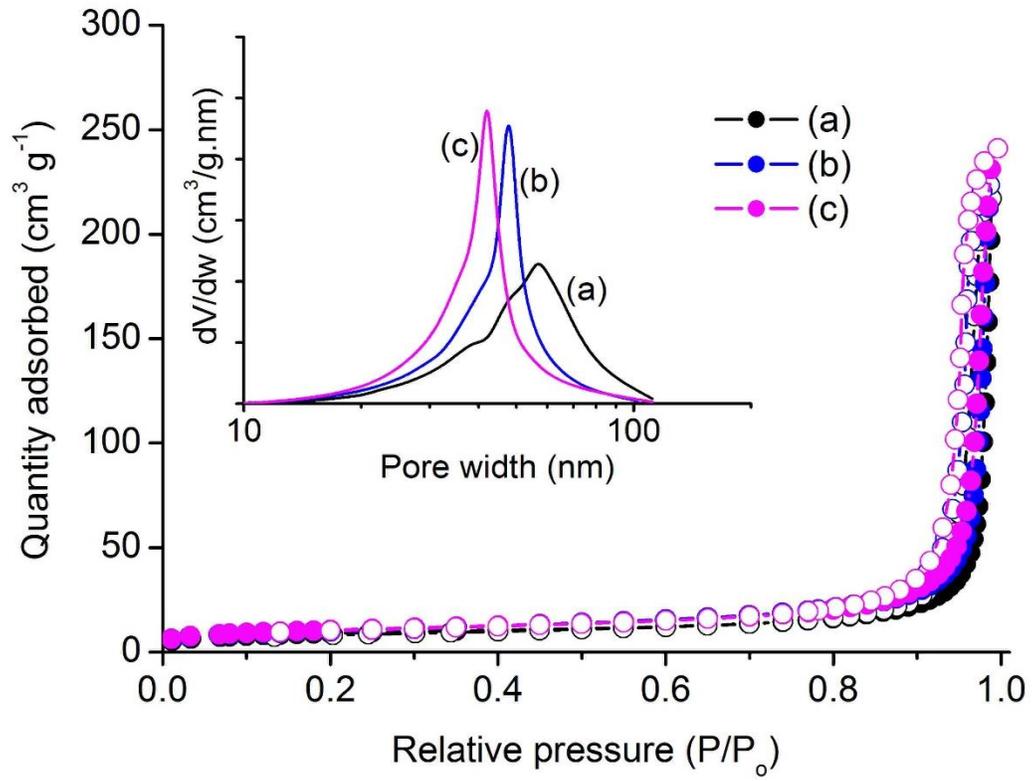


Fig. 4

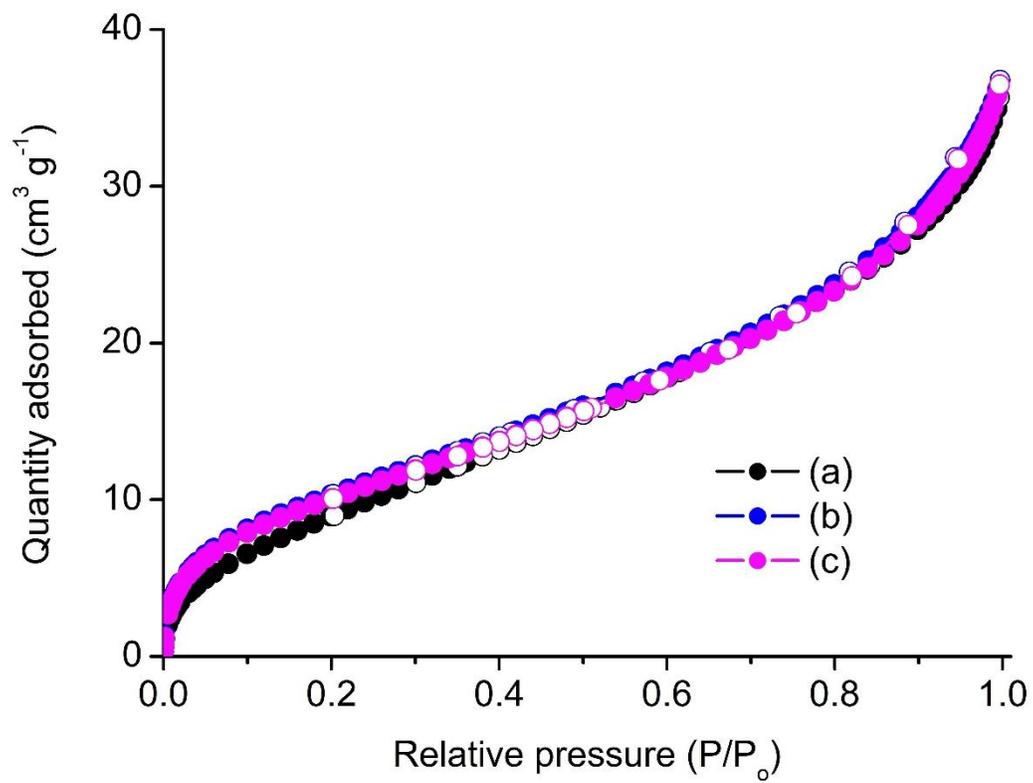


Fig. 5

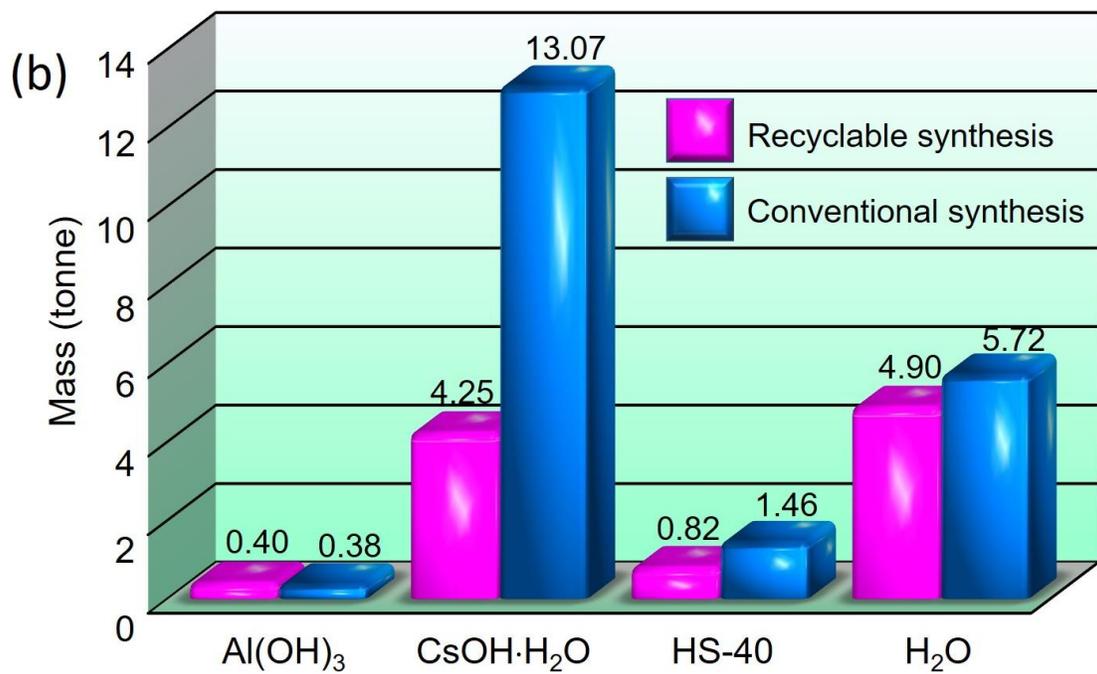
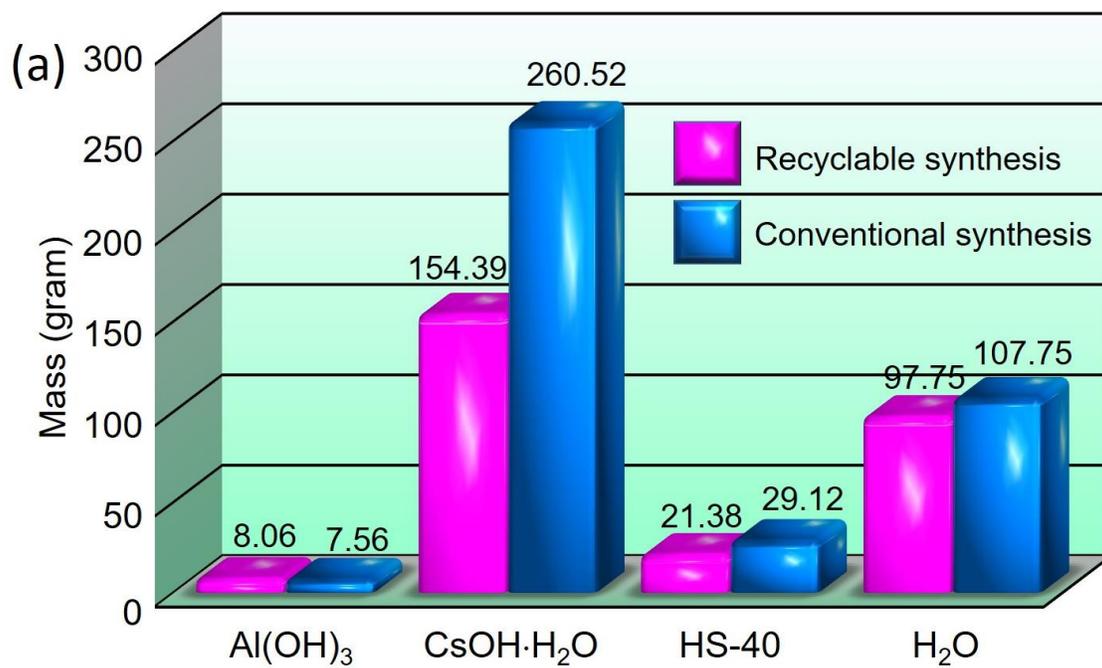


Fig. 6

Table captions

Table 1. Compensated chemicals added into non-reacted mother liquor for Cs-ABW synthesis cycles and Cs-ABW solid yield.

Table 2. Chemical compositions and textural properties of Cs-ABW zeolite nanocrystals.

Table 3. Catalytic nitroaldol condensation of benzaldehyde with nitroethane under microwave irradiation.^a

Tables

Table 1. Compensated chemicals added into non-reacted mother liquor for Cs-ABW synthesis cycles and Cs-ABW solid yield.

Cs-ABW synthesis	1 st cycle (A1)	2 nd cycle (A2)	3 rd cycle (A3)
Non-reacted mother liquor (g)	–	87.008 ^a	85.571 ^a
Added chemicals			
Al(OH) ₃ (g)	3.312	3.038	2.445
CsOH·H ₂ O (g)	114.104	28.031	26.386
HS-40 (g)	12.750	5.811	4.776
H ₂ O (g)	42.808	36.803	37.994
pH before compensation	–	14	14
Solid yield (g, wt.%) ^b	8.758, 81.85%	6.779, 63.36%	6.291, 58.79%

^aAfter evaporating water at 55 °C for 96 h.

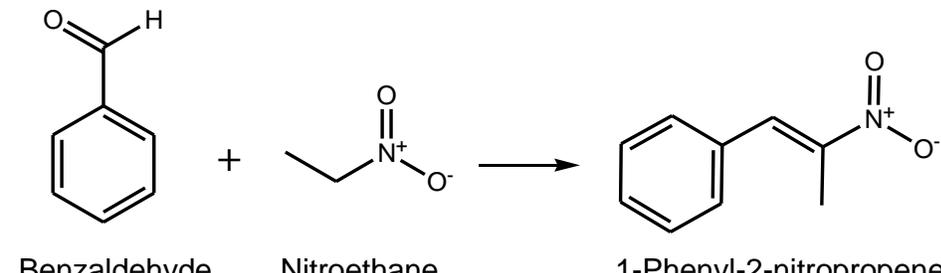
^bSolid yield (%) = $\frac{\text{Weight of dried solid (g)}}{\text{Initial weight of Al+Si+O+Cs (g)}} \times 100\%$ where Al is the limiting reagent and the total mass in the denominator theoretically produces Cs-ABW zeolite (Cs₄Al₄Si₄O₁₆) [34] with 100% solid yield.

Table 2. Chemical compositions and textural properties of Cs-ABW zeolite nanocrystals.

Sample	Si/Al ratio	S _{BET}		S _{micropore}		S _{external}		V _{Total}		Average pore size (nm) ^a		Crystal size (nm)	Zeta potential (mV)
		(m ² /g)		(m ² /g)		(m ² /g)		(cm ³ /g)		size (nm) ^a			
		N ₂	Ar	N ₂	Ar	N ₂	Ar	N ₂	Ar	N ₂	Ar		
A1	1.03	31.9	33.4	5.3	0.2	26.6	33.2	0.21	0.04	46.4	n.m.	49.4	-52.6
A2	1.02	36.2	34.7	4.9	0.1	31.3	34.6	0.25	0.05	39.8	n.m.	46.1	-53.7
A3	1.02	37.3	34.0	7.2	0.1	30.1	33.9	0.33	0.05	37.1	n.m.	43.4	-52.7

^aDetermined from BJH model; n.m.: Not measurable

Table 3. Catalytic nitroaldol condensation of benzaldehyde with nitroethane under microwave irradiation.^a



Benzaldehyde
Nitroethane
1-Phenyl-2-nitropropene

Entry	Catalyst	Conversion (%)		
		3 min	5 min	10 min
1	No catalyst	13.2	25.5	34.0
2	Cs-ABW (A-1)	56.9	70.5	80.8
3	Cs-ABW (A-2)	58.6	73.4	81.5
4	Cs-ABW (A-3)	60.7	75.6	82.6
5	Cs-ABW (A-3) ^b	59.5	74.8	81.8
6	NaOH	35.5	40.8	55.7
7	Na ₂ CO ₃	25.7	33.1	40.2
8	KOH	45.7	56.4	68.5
9	Cs ₂ CO ₃	75.3	78.5	82.0
10	CsOH	85.4	88.7	92.4

^aReaction conditions: Catalyst (1.98 mmol equivalent to 0.50 g Cs-ABW zeolite nanocrystals), benzaldehyde (2 mmol), nitroethane (40 mmol), microwave power 850 W, 170 °C. ^b5th run of Entry 4.

Graphical abstract

