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Alkali-glass behavior in honeycomb-type layered $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ solid-solution

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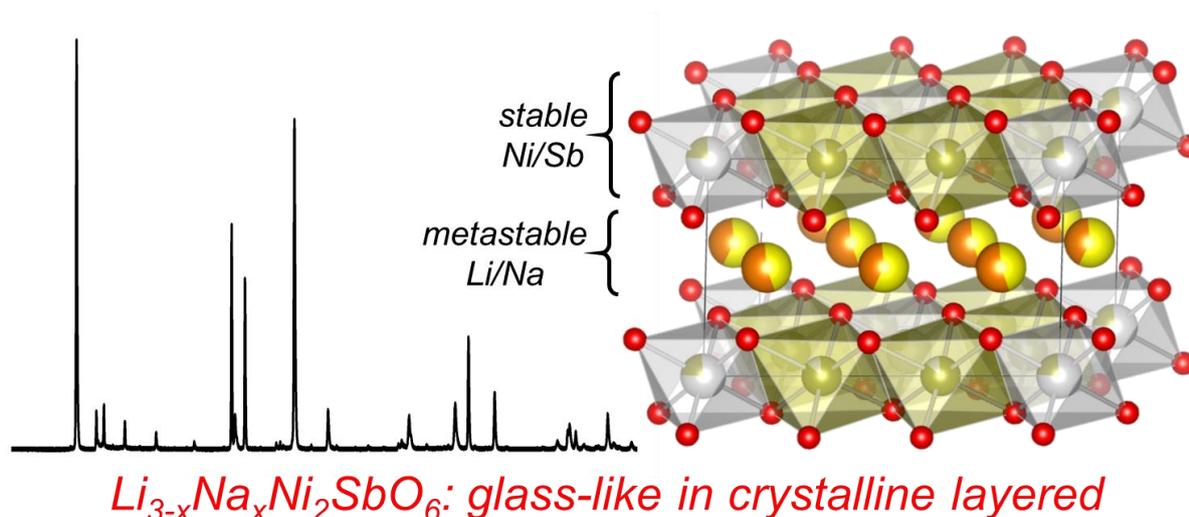
Keywords:

Layered oxides, honeycomb ordering, solid-solution, alkali mixed occupancy

Abstract

Layered oxides compositions $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ have been prepared by solid-state synthesis. A complete solid-solution is evidenced and characterized by X-ray and neutron diffraction, as well as ^7Li and ^{23}Na solid-state nuclear magnetic resonance spectroscopy. The transition metal layer is characterized by the classic honeycomb $\text{Ni}^{2+}/\text{Sb}^{5+}$ ordering, while a more uncommon randomly mixed occupancy of lithium and sodium is evidenced within the alkali interslab space. *In situ* X-ray diffraction and DFT calculations show that this alkali disordered feature is entropically driven. Fast cooling appears then as a synthesis route to confine bidimensional alkali glass within crystalline layered oxides.

Graphical abstract / TOC graphic



TOC synopsis

Structural refinement and solid-state NMR validates a complete solid-solution for honeycomb-type layered oxides $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$. An uncommon randomly mixed occupancy of lithium and sodium is evidenced within the alkali interslab space, this feature being entropically driven.

Fast cooling appears then as a synthesis route to create alkali glass layer within crystalline layered oxides.

Introduction

Ionic radii of lithium and sodium cations are significantly different in chalcogenides or halides.¹ As a consequence, there are in literature very few cases of mixed lithium/sodium occupancy in same crystallographic sites in transition metal oxides. Examples are vanadates $\text{Li}_{1-x}\text{Na}_x\text{VO}_3$ ² or bronzes $\text{Li}_x\text{Na}_y\text{V}_3\text{O}_8$, the latter being widely investigated as insertion-type negative electrode materials for rechargeable lithium-ion batteries,³ or niobates perovskites $\text{Li}_{1-x}\text{Na}_x\text{NbO}_3$ with piezoelectric properties.⁴⁻⁶ Regarding the important family of alkali transition metal layered oxides, the crystallographic structure offers a large playground for solid-state chemists and physicists due to the impressive richness of chemical compositions that can tailor different properties. The lamellar organization is usually described by a stacking of edge-shared MO_6 octahedra slabs, with intercalation of alkali ions in the interslab space. Chemical substitutions on the transition metal layers are possible thanks to similar ionic radii in octahedral symmetry for cations, enabling to tune some physical or chemical features. The best example might be the emerging layered compositions $\text{LiNi}_{1-x-y}\text{Co}_x\text{Al}_y\text{O}_2$ or $\text{LiNi}_{1-x-y}\text{Mn}_x\text{Co}_y\text{O}_2$ which exhibit enhanced electrochemical performance compared to the standard electrode material LiCoO_2 .⁷⁻
⁹ On the contrary, cases of alkali ions mixed occupancy within the interslab spaces are very scarce. To the best of our knowledge, only two examples have been reported, single crystals of rhomboedral titanate $(\text{Na}_{0.73}\text{Li}_{0.24})(\text{Li}_{0.32}\text{Ti}_{0.68})\text{O}_2$ studied by Nalbandyan,¹⁰ and nickelates $\text{Li}_{1-x}\text{Na}_x\text{NiO}_2$ investigated separately by Matsumura and Holzappel in early 2000s.¹¹⁻¹³

In the frame of discovering new layered dual-alkali transition metal layered oxides with possible interesting physical features, this work investigates intermediate compositions

between $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$ (or $\text{LiNi}_{2/3}\text{Sb}_{1/3}\text{O}_2$ and $\text{NaNi}_{2/3}\text{Sb}_{1/3}\text{O}_2$, respectively). These compounds are structurally related to O3-type¹⁴ LiCoO_2 , with an additional honeycomb-type ordering of the Ni^{2+} and Sb^{5+} cations.^{15–18} A complete solid-solution could be obtained by ceramic route one-pot synthesis from carbonates or oxides precursors. The lithium/sodium mixed occupancy is confirmed by structural refinement and solid-state NMR spectroscopy. The alkali ions random distribution is not thermodynamically stable as decomposition is observed by *in situ* XRD during either heating or slow cooling. These results pave the way to design crystalline layered compounds with 2D confined alkali glass.

Experimental section

$\text{Li}_3\text{Ni}_2\text{SbO}_6$, $\text{Na}_3\text{Ni}_2\text{SbO}_6$ and intermediate compositions $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ ($0 \leq x \leq 3$) have been prepared by conventional solid-state synthesis. Dried alkali carbonates (Li_2CO_3 , Aldrich, 99.997 % or Na_2CO_3 , Sigma-Aldrich, 99.5 %) are thoroughly mixed with nickel and antimony oxides (NiO , Prolabo and Sb_2O_3 , Merck, 99 %), with a nominal alkali excess of 5-10 wt.% to compensate the high-temperature volatility. The resulting mixtures are then pelletized and heated for 5 hours at 900 °C in air on a platinum crucible. A rapid quench in oil bath ends the heat treatment, and the resulting pellets are then washed with ethanol and dried before storage.

X-ray diffraction (XRD) was performed on a Bruker D8 Discovery diffractometer with $\text{K}\alpha_1$ Cu radiation, from 5 to 120 ° (16 h of total scan). Sample powders were sieved below 100 μm for analysis. Moreover, the thermal stability of the samples was followed *in situ* by XRD on a PANalytical Empyrean with $\text{K}\alpha$ Co radiation and equipped with an Anton Paar heating chamber.

Neutron powder diffraction (ND) was carried out at room temperature out at Institut Laue Langevin on D2B high resolution powder diffractometer (1.594 Å wavelength). For that, 3 g of

as-prepared powder of the composition $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$ have been sealed in a cylindrical vanadium tube.

^{23}Na MAS NMR spectra were recorded on a Bruker 300 Avance spectrometer at 79.49 MHz (7.05 T magnet) using a standard Bruker 2.5 MAS probe at a spinning frequency of 30 kHz and a recycle time $D_0 = 0.5$ s, long enough to avoid T_1 saturation effects. A one rotor period synchronized Hahn echo sequence was used with a 90° pulse length of 1.5 μs . The external reference was a 1 mol.L⁻¹ NaCl aqueous solution. ^7Li MAS NMR spectra were recorded on a Bruker Avance III spectrometer at 38.93 MHz (2.35 T magnet), a standard Bruker 2.5 MAS probe at a 30 kHz spinning frequency. A Hahn echo sequence was used with a 90° pulse of 1.6 μs with a relaxation time of 1 s. The 0 ppm external reference was a 1 M LiCl aqueous solution.

Density functional theory calculation (DFT+U) used the monoclinic description of Zvereva¹⁸ with two alkali layers ($C 2/m$ space group) as starting point. About hundred different crystallographic structures have been generated with a random distribution of lithium and sodium ions over the 24 available positions. All structures have been fully relaxed using the plane wave code “Vienna Ab-Initio Simulation Package” with PAW pseudo-potentials and the PBE functional.^{19–21} A $U_{\text{eff}} = 4$ eV is considered to take into account electron-electron interactions on nickel.²² The forces on the atoms were converged to $5 \cdot 10^{-3}$ eV/Å with a plane-wave energy cut-off of 600 eV and a well converged set of K-points.

Results and Discussion

An alkali ions mixed occupancy

XRD patterns obtained with a final quenching at the end of the thermal treatment are gathered in Figure 1a for all compositions from $\text{Li}_3\text{Ni}_2\text{SbO}_6$ to $\text{Na}_3\text{Ni}_2\text{SbO}_6$. No extra peak could be assigned to impurities like binary oxides or alkali hydroxides, and the patterns could be indexed using the monoclinic symmetry (space group $C 2/m$). In comparison with a classic O3-type stacking (α - NaFeO_2 polytype), the $\text{Ni}^{2+}/\text{Sb}^{5+}$ honeycomb ordering induces superstructure peaks visible in the $19 - 32^\circ 2\theta$ region. They can be nicely observed in Figure 1b, with for Li-rich compositions a broadening that is reminiscent of stacking faults, as commonly observed in such honeycomb-ordered lamellar oxides.¹⁵⁻¹⁷

The refined cell parameters evolve regularly all along the chemical compositions, following a classical Vegard's law (Figure S1). The evolution of the corresponding monoclinic cell volume shown in inset of Figure 1c demonstrates a complete solid-solution behavior from $\text{Li}_3\text{Ni}_2\text{SbO}_6$ to $\text{Na}_3\text{Ni}_2\text{SbO}_6$. The question arising from this assumption is the localization of both lithium and sodium cations. Three structural models seem possible. They are pictured in Figure 2 for the particular compositions $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$. The first one is a mixed occupancy of lithium and sodium ions and a random in-plane distribution within. Following the structural description of Zvereva¹⁸ and Ma¹⁷, the slab thickness (basically including one transition metal and one alkali layers) is 4.856 Å and 5.344 Å for $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$. From lithium-rich to sodium-rich end members, the increase is about 10 % and comes from an important enhancement of the alkali layer (from 2.60 Å to 3.21 Å, hence + 23.5 %) whereas the transition metal layer is slightly contracted. In our case, the linear evolution of the slab thickness also shows an increase of about 10 % (Figure 1d) and consequently supports this structural assumption.

In the second model, lithium and sodium ions always occupy the same interslab space but following an in-plane ordering (either chains, stripes, cluster, *etc.*). The last model also lies on alkali ions ordering, lithium and sodium ions occupy distinct interslab spaces. For the intermediate composition $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$, such arrangement would result in alternate Li^+/Na^+ layers, as it has been already evidenced for the unique case of $\text{OP4-Li}_x\text{Na}_y\text{CoO}_2$.²³ It is important to point out that, in the two latter models, the mean slab thickness evolution (determined from the out-of-plane c parameter) could be also assumed.

In alkali transition metal layered oxides, alkali cations out-of-plane ordering like in $\text{OP4-Li}_x\text{Na}_y\text{CoO}_2$ or alkali cations / vacancies in-plane ordering like in $\text{Na}_{1/2}\text{CoO}_2$ or $\text{K}_{4/7}\text{CoO}_2$ lead to significant superstructure peaks in X-ray diffraction pattern.^{23–25} The two last abovementioned models involve cationic ordering and the corresponding simulated patterns clearly evidence superstructure peaks (Figure S2). The absence of such peak in our high-resolution XRD pattern in Figure 1 is a first clue to discard these two hypotheses and consequently to support the model of random Li/Na distribution despite their different ionic radii (0.76 and 1.02 Å in octahedral symmetry, respectively).¹

In order to get local information on lithium and sodium environments, solid-state NMR has been further carried out for all the compositions, and corresponding ^7Li and ^{23}Na spectra are shown in Figure 3. The peak shifts for all compounds results from the Fermi contact interaction between the paramagnetic low spin Ni^{2+} ($t_{2g}^6e_g^2$) ions and the probed nucleus. This interaction occurs through the chemical bonds and therefore denotes the local electronic structure. Whatever the structural hypotheses presented above, the two alkaline A^+ ions in the system are surrounded by four Ni^{2+} as first cation neighbors (90° A-O-Ni interaction) and four Ni^{2+} as second cation neighbors (180° A-O-Ni interaction) with slightly different distances and angles.

For $\text{Li}_3\text{Ni}_2\text{SbO}_6$, two distinct narrow peaks at 492 and 420 ppm and a larger set of signals around 300 ppm are observed by ^7Li MAS NMR. According to the literature, based on their relative intensities, the two first are assigned to lithium ions respectively in $2d$ and $4h$ Wyckoff positions of the space group $C 2/m$, *i.e.* within the interslab space in octahedral insertion sites. The third broad contribution around 300 ppm could be assigned to stacking defects, similarly to stacking faults observation in the Li_2MnO_3 reported by Grey and coworkers.²⁶ Indeed, the relative intensity of the contribution at 300 ppm compared to the two other sharp peaks depends on the synthesis conditions: when $\text{Li}_3\text{Ni}_2\text{SbO}_6$ is prepared with a final slow cooling, the two first peaks are clearly sharper and the broad contribution almost disappear (Figure S3). On the opposite side of the Figure 3, ^{23}Na NMR spectrum for $\text{Na}_3\text{Ni}_2\text{SbO}_6$ exhibits two contributions located at 1440 ppm and 1300 ppm with no second order quadrupolar lineshape, contrary to the spectrum reported by Ma *et al.* using a lower magnetic field (200 MHz spectrometer) and a higher spinning frequency (50 KHz).¹⁷ Based on the relative peak intensities, we can also assign those two peaks to Na ions in the $2d$ and $4h$ sites respectively. Interestingly, the difference between quenched and non-quenched $\text{Na}_3\text{Ni}_2\text{SbO}_6$ is very weak as both XRD pattern and NMR data look very similar (Figure S4). Thanks to higher weight and larger ionic radii of sodium compare to lithium and in agreement with the observation of superstructure peaks shapes in inset of Figure 1, $\text{Na}_3\text{Ni}_2\text{SbO}_6$ seems to ordered more easily than $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and is therefore less sensitive to quenching.

Intermediate compositions $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ can be analyzed by both ^7Li and ^{23}Na NMR. At a first glance, broader resonance peaks are clearly evidenced. In both cases, the broadness increases when compositions moves away from $\text{Li}_3\text{Ni}_2\text{SbO}_6$ or $\text{Na}_3\text{Ni}_2\text{SbO}_6$. Moreover, a weak but significant peak shift also accounts for variation of the (Li – O and Na – O) interatomic distances. Shifting toward lower ppm in ^7Li NMR spectra corresponds to a decrease of the electronic spin transfer from low spin Ni^{2+} to Li nuclei that can be attributed to an increase of

the Li – O distances, whereas displacement to higher ppm in ^{23}Na NMR spectra is attributed to a decreasing of the Na – O distances. Therefore, the NMR study tends to demonstrate that the alkali layers in $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ compositions do not correspond to an intergrowth of pure Li and Na layers but do contain both alkaline ions. Indeed, sodium ions present in a Li layer will increase the local Li-O distance compared to $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and lithium ions present in the Na layer will decrease the local Li-O distance compared to $\text{Na}_3\text{Ni}_2\text{SbO}_6$. Moreover, the overall broadening of the NMR lines results from a larger distribution of environments as locally distances and angles will vary depending on the amount of surrounding Li/Na ions. In addition, it worth noting the decrease of third contribution at 300 ppm in the ^7Li spectra with higher sodium amount. This accounts for less stacking faults for Na-rich compositions, which is logical when stated that $\text{Na}_3\text{Ni}_2\text{SbO}_6$ is less sensitive to quenching than $\text{Li}_3\text{Ni}_2\text{SbO}_6$.

In order to confirm the hypothesis of a random distribution of both lithium and sodium ions in the same interslab space suggested by the NMR analysis, structural refinements have been performed on high-resolution XRD and neutron powder diffraction patterns data for the particular composition $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$. Indeed with four different chemical species (Li, Na, Ni, Sb) possibly distributed among four crystallographic sites ($2d$ and $4h$ in the interslab space, and $2a$ and $4g$ in the metal layers), a unique solution cannot be found without *i*) combining X-ray with neutrons, that offers a nice contrast ($b_{\text{Ni}} = +10.3$ fm, $b_{\text{Sb}} = +5.57$ fm, $b_{\text{Li}} = -1.90$ fm, $b_{\text{Na}} = +3.63$ fm), and *ii*) imposing some constraints that are detailed as follows. We used as a starting model the $C 2/m$ monoclinic structural description of $\text{Li}_3\text{Ni}_2\text{SbO}_6$ by Zvereva *et al.*²⁷ The total number of Li, Na Ni and Sb was imposed to the nominal composition, with Li and Na being distributed on the $2d$ and $4h$ sites, and Ni and Sb on the $2a$ and $4g$ sites. All atomic positions and the relative occupancy of Li/Na and Ni/Sb were freely refined. An isotropic temperature factor was refined and imposed equal for a same chemical species/Wyckoff site.

The result of the combined XRD/neutron refinement is shown in Figure 4 and the corresponding table of atomic positions and resulting structure are presented in Table 1 and Figure 5, respectively. Although the overall Ni/Sb honeycomb pattern is confirmed within the transition metal slab with $4g$ and $2a$ sites mainly occupied by Ni and Sb, respectively; a certain Ni/Sb mixing is obtained in the structural refinement. Because the relative occupancies of Ni and Sb were not refined in previous structural refinement, this relative disorder was not reported in $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$. As stacking faults characterized quenched layered samples, it is reasonable to assume that the high value of Ni/Sb mixing is over-estimated, as reported for a similar lamellar compound $\text{Li}_4\text{FeSbO}_6$.²⁸ For comparison, structural refinement for quenched $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$ from ND also ends up with some Ni/Sb mixing. Interestingly the Ni/Sb mixing regularly increases from $\text{Li}_3\text{Ni}_2\text{SbO}_6$ to $\text{Na}_3\text{Ni}_2\text{SbO}_6$ (Figure S5 and Table S1).

Regarding the interlayer space, the alkali layer is occupied by both lithium and sodium cations and the refinement indicates a Li/Na mixing very close to the statistical value of 50 %. Lithium and sodium cations are therefore randomly distributed and seem to not form any kind of ordering patterns that could be induced by the strong ionic radii difference and/or by the neighbor Ni/Sb honeycomb-like ordering. To the best of our knowledge, this is the first example of a complete solid-solution from Li to Na layered oxide with no structural change and no biphasic domain like for $\text{Li}_{1-x}\text{Na}_x\text{NiO}_2$.^{11–13}

A glass-like behavior

After demonstrating the mixing of lithium and sodium despite their very different ionic radii, it was important to investigate the formation process of this very uncommon crystallographic feature. First, it is worth pointing out the influence of the quenching process at the end the

thermal treatment. Indeed, a slow cooling rate down to room temperature leads to very different XRD patterns, with coexistence of $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$ as the main phases and with a very weak intermediate contribution from the mixed $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ (see Figure S3 for a nominal composition “ $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$ ”). In their pioneering investigation on layered $\text{Li}_{1-x}\text{Na}_x\text{NiO}_2$, Matsumura and coworkers also briefly noticed that quenching was mandatory for obtaining a monophasic product.¹¹ In this current study, we decided to have much deeper grasp on the details of the stability of the obtained monophasic compositions by combining XRD and DFT approaches.

The thermal evolution of quenched $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$ has been monitored by *in situ* XRD (Figure 6). At 300 °C, the phase decomposes as its main diffraction peak (001) splits into two peaks accounting for $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$ domains. At higher temperatures, the synthesis conditions of a monophasic domain are gathered again and the two peaks merge into one contribution accounting for (001) of $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$. During the slow cooling down, a second splitting occurs, showing the reformation of $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$ domains. This result clearly highlights the role of the final quench for stabilizing at room temperature a monophasic composition with a random distribution of alkali ions in the interslab layer. Such thermal feature strongly recalls the behavior of the layered oxide OP4- $\text{Li}_x\text{Na}_y\text{CoO}_2$, for which lithium and sodium are not randomly mixed but alternately ordered one layer over two in octahedral and prismatic sites, respectively.²³

Following the peak position evolution of the $\text{Li}_3\text{Ni}_2\text{SbO}_6$ and $\text{Na}_3\text{Ni}_2\text{SbO}_6$ domains can help in understanding the mechanism involved in the formation of a monophasic domain $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$ at high temperature. First, it is important to keep in mind the thermal expansion that significantly increases the slab thickness and thus displaces the (001) peaks towards low angles (*i.e.* from 20.1 ° at room temperature to 19.7 ° at 900 °C for the (001) peak of $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$). This feature seems enhanced for $\text{Li}_3\text{Ni}_2\text{SbO}_6$ domains (from 21 ° at

400 °C to 20.6 °C at 700 °C), whereas the trend is different for the peak related to $\text{Na}_3\text{Ni}_2\text{SbO}_6$ domains, which slightly shifts to higher angles. The slab thickness contraction (expansion) accounts for lithium (sodium) diffusion into the $\text{Na}_3\text{Ni}_2\text{SbO}_6$ ($\text{Li}_3\text{Ni}_2\text{SbO}_6$) alkali layers, in agreement with the above NMR analysis (Figure 3). The underlying mechanism consequently involves a diffusive alkali exchange between domains, respect to a nucleation process for which only a shift due to thermal expansion would have been observed. The slow cooling monitoring shows a similar peak splitting and thus confirms that phase decomposition is thermodynamically favored. To briefly summarize, the Li/Na random distribution mimics an alkali-glass behavior, however confined within a crystalline matrix.

In parallel, DFT calculations have been performed on about 100 structures with $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$ stoichiometry, within random ordering of Li/Na in the alkali layer. Figure 7 shows the formation energy of each ordered phase respect to the pure lithiated or sodiated compounds and the variation of the formation energy as a function of the number of Li-Na Nearest Neighbor (NN) bonds. As the relative formation energy is positive, all the random ordered phases are shown to be thermodynamically instable and should decompose following the equation: $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6 \rightarrow \frac{1}{2} \text{Li}_3\text{Ni}_2\text{SbO}_6 + \frac{1}{2} \text{Na}_3\text{Ni}_2\text{SbO}_6$.

The most stable order consist in the out-of-plane ordered/segregated Li/Na phases (at ~ 3 meV/atoms) without NN Li-Na, the energy increasing barely linearly with the number of Li-Na NN bonds. In this situation, the stabilization of any ordered phase does not appear as feasible except if configuration entropy, thus Li/Na inter-mixing is considered. In fact, the latter can be evaluated within the Stirling approximation to be ~ 20 meV/atoms at 900 °C far enough compared to the relative energy between ordered phases. Consequently, the most probable hypothesis consists in an entropically stabilized Li/Na intermixing in agreement with the synthesis conditions and thermal behavior, structural refinement and NMR analysis, and characteristic to glassy-like systems.

Conclusion

A complete solid-solution from layered $\text{Li}_3\text{Ni}_2\text{SbO}_6$ to $\text{Na}_3\text{Ni}_2\text{SbO}_6$ has been obtained by conventional solid-state synthesis ended by a final quench. On the transition metal layer, the honeycomb Ni/Sb cation ordering is maintained all along the solid-solution; however with a partial Ni/Sb mixing. Focusing on the alkali layer, configuration entropy leads to Li/Na randomly occupied crystallographic positions.

It is believed that other crystalline transition metal oxides exhibiting 2D confined alkali glass behavior could now be prepared thanks to a final quenching step, and thus opening the door towards unexpected physical properties.

Supporting Information

The Supporting Information section is available free of charge at DOI: xxx.

Notes

The authors declare no competing financial interest.

Acknowledgment

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Figure and Table captions

Figure 1. (a) XRD patterns of $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ compositions obtained after a final quench, (b) restricted 2θ window to highlight the honeycomb superstructure peaks, and evolution of (c) the monoclinic cell volume and (d) the slab thickness (calculated by $c \times \sin(\beta)$).

Figure 2. Structural hypotheses for the intermediate composition $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$. Nickel and antimony cations are green and grey, while lithium and sodium cations are in yellow and orange, respectively.

Figure 3. Solid-state ^7Li and ^{23}Na NMR spectra of as-quenched $\text{Li}_{3-x}\text{Na}_x\text{Ni}_2\text{SbO}_6$ compositions, with blue dotted lines showing the peak shifting from end members ($\text{Li}_3\text{Ni}_2\text{SbO}_6$ or $\text{Na}_3\text{Ni}_2\text{SbO}_6$) to intermediate compositions. The symbol (*) accounts for spinning sidebands.

Figure 4: Combined Rietveld refinement of X-ray (top) and neutron (bottom) diffraction patterns of $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$. The red crosses, black continuous line and bottom green line represent the observed, calculated, and difference patterns, respectively. Vertical blue tick bars mark the reflection positions.

Figure 5: (a) Structural model for $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$ resulting from the combined X-ray/neutron Rietveld refinement and (b) in-plane views showing the Ni/Sb and Li/Na intermixing.

Figure 6. Thermal evolution of the quenched $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$ monitored by in situ XRD.

Figure 7. (a) Relative formation energy in meV/atom of randomly Na-Li ordered phase respect to the decomposition in pure $\text{Li}_3\text{Ni}_2\text{SbO}_6 + \text{Na}_3\text{Ni}_2\text{SbO}_6$ phases. All considered phases are not stable versus decomposition since they have positive energy. (b) Relative formation energy in meV/atom as a function of the average number of Na-Li Nearest Neighbor bonds per alkali ions, showing that as more Na-Li are mixed, as high is the energy of the phase.

Table 1: Structural parameters for $\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$, deduced from the combined Rietveld refinement of the XRD and neutron powder diffraction patterns.

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Figure 1

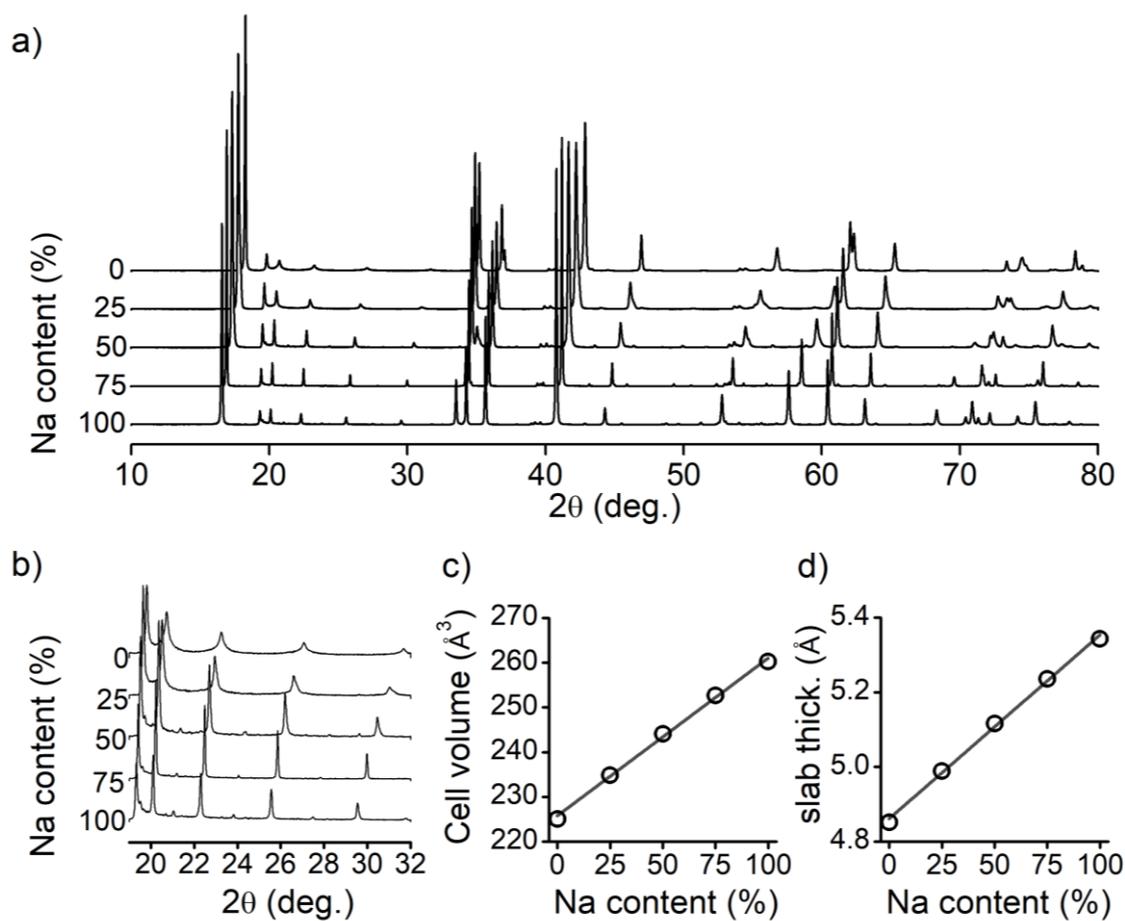


Figure 2

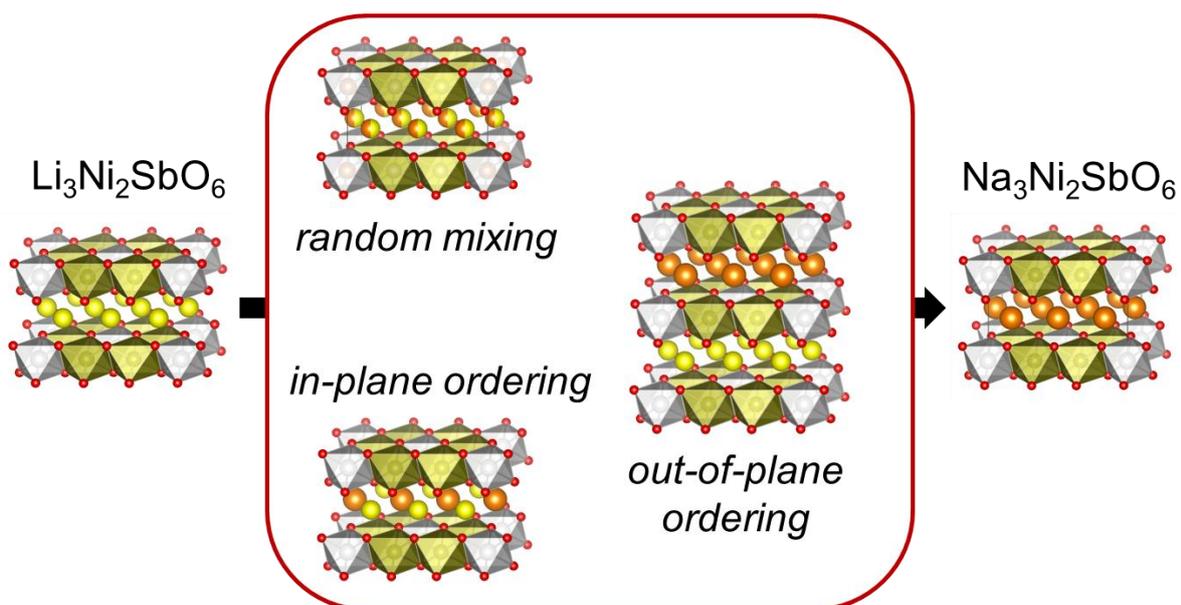


Figure 3

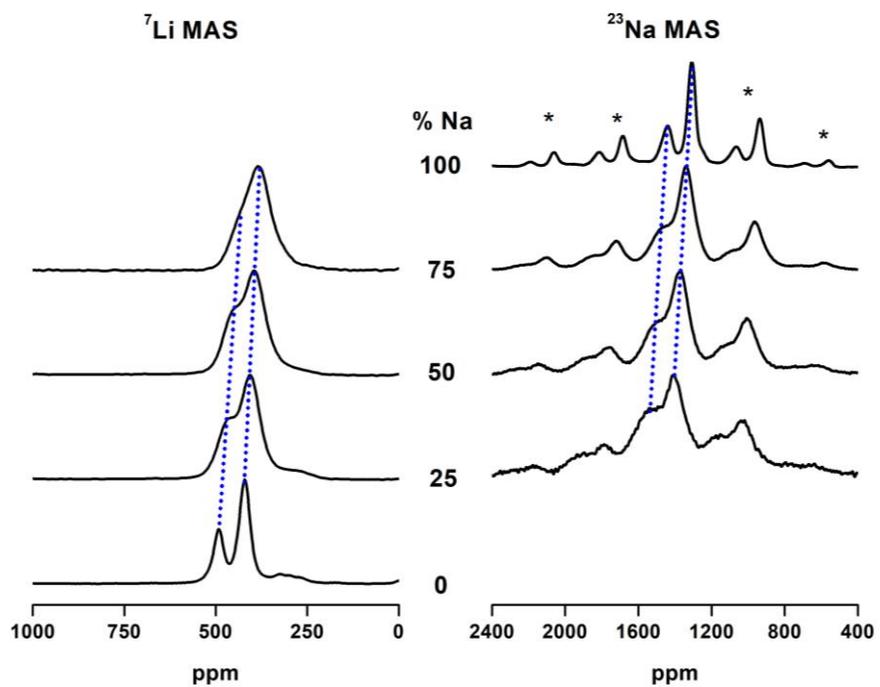


Figure 4

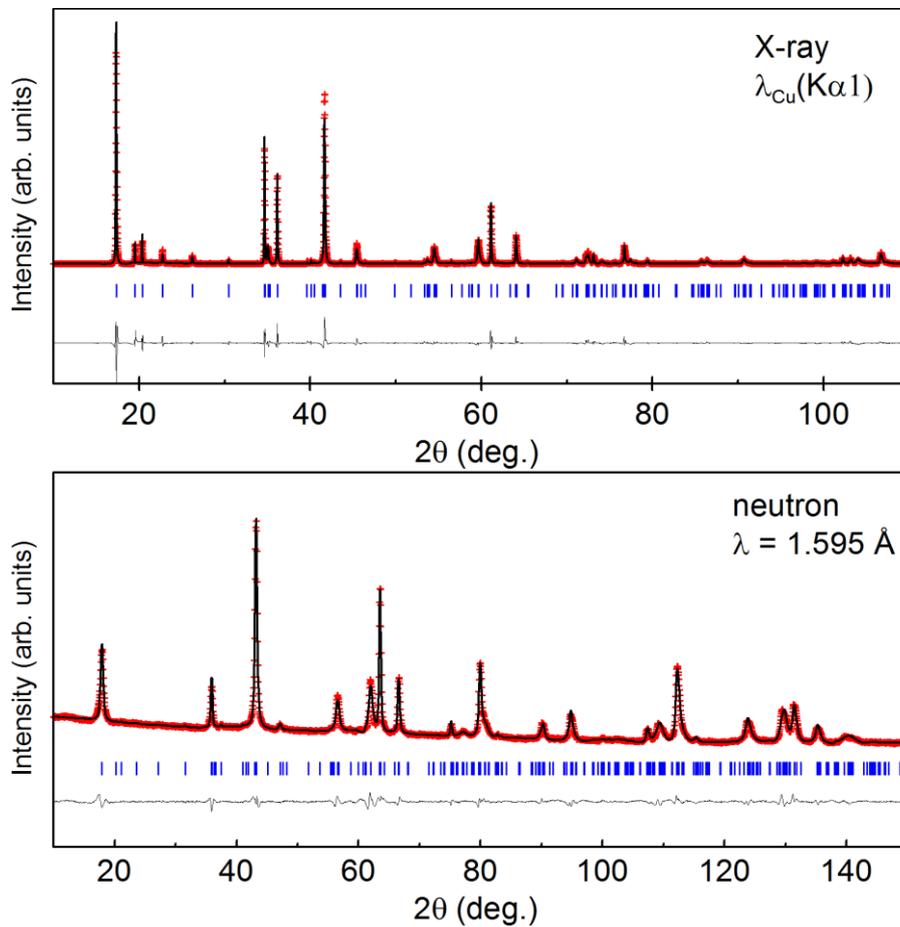


Figure 5

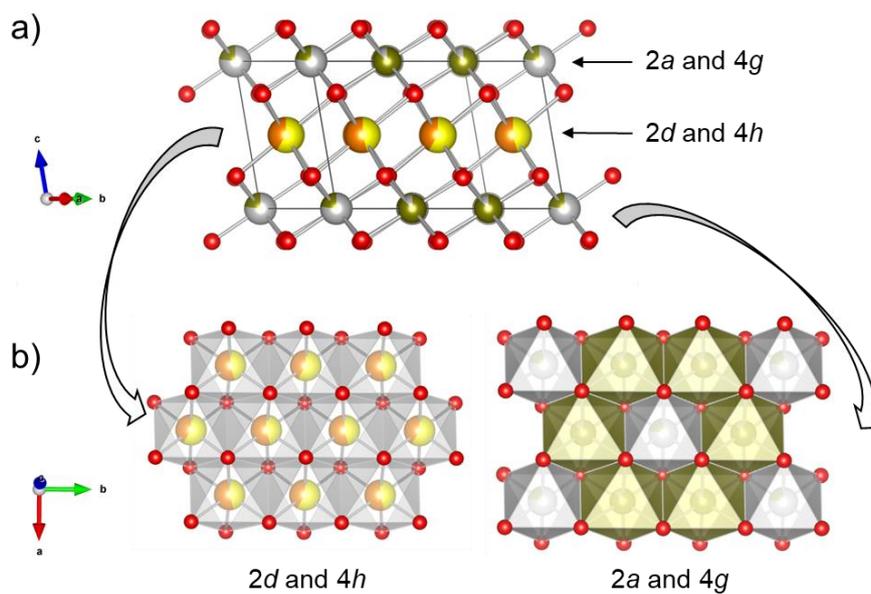


Figure 6

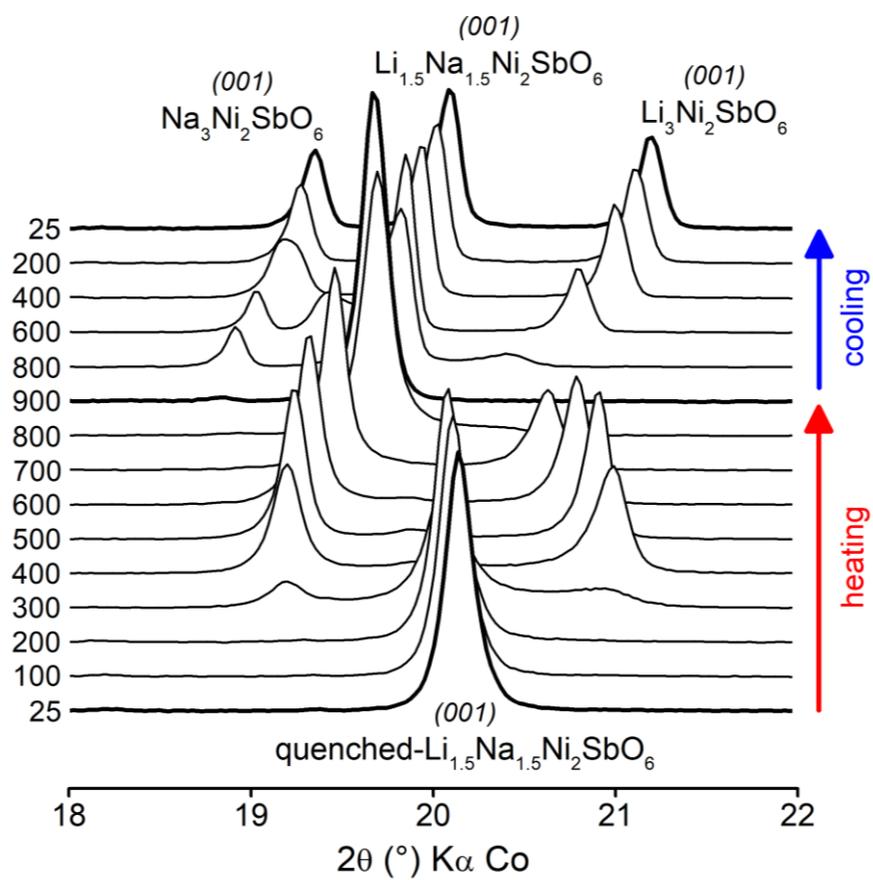


Figure 7

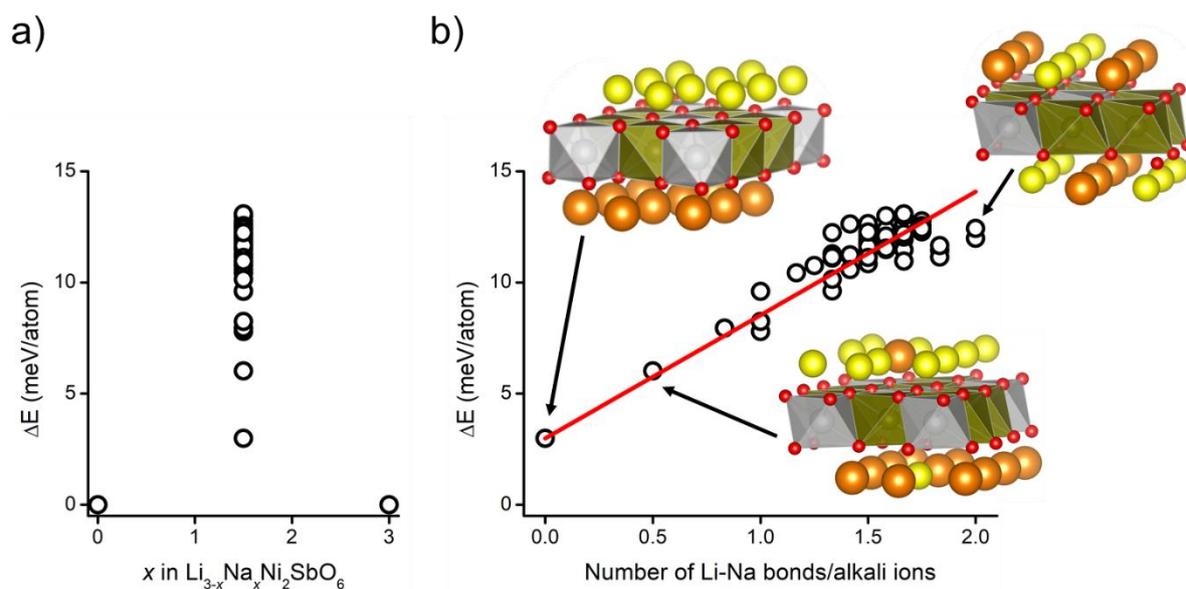


Table 1

$\text{Li}_{1.5}\text{Na}_{1.5}\text{Ni}_2\text{SbO}_6$

Space Group: $C 2/m$

$a = 5.2434(2) \text{ \AA}$, $b = 9.0997(3) \text{ \AA}$, $c = 5.4034(6) \text{ \AA}$, $\beta = 108.794(8) \text{ deg}$.

$V = 243.641(5) \text{ \AA}^3$

Atom	Wyckoff site	x	y	z	$B_{\text{iso}} (\text{\AA}^2)$	Occ
O1	$4i$	0.736(2)	0	0.209(2)	0.50(6)	1
O2	$8j$	0.2351(16)	0.1584(9)	0.2230(12)	0.50(6)	1
Ni1/Sb1	$4g$	0	0.3330(4)	0	0.02(3)	0.882(2) (Ni) 0.118(2) (Sb)
Sb2/Ni2	$2a$	0	0	0	0.02(3)	0.764(2) (Sb) 0.236(2) (Ni)
Li1/Na1	$4h$	0	0.1707(15)	$\frac{1}{2}$	0.48(13)	0.461(8) (Li) 0.539(8) (Na)
Li2/Na2	$2d$	0	$\frac{1}{2}$	$\frac{1}{2}$	0.48(13)	0.578(8) (Li) 0.422(8) (Na)

XRD: Bragg R-factor = 8.61 %, Rf-factor = 6.23 %

Neutrons: Bragg R-factor = 2.81 %, Rf-factor = 2.30 %