Molecular Design and Control Over the Morphology of Self-Assembled Films on Ionic Substrates

To cite this version:

HAL Id: hal-01702354
https://hal.archives-ouvertes.fr/hal-01702354
Submitted on 15 Feb 2018

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L’archive ouverte pluridisciplinaire HAL, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d’enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.
Molecular Design and Control Over the Morphology of Self-Assembled Films on Ionic Substrates


Achieving control over formation of molecular films on insulating substrates is important for designing novel 2D functional materials and devices. To study the main factors governing successful control, organic molecules with interchangeable polar functional groups, a variable length aromatic body, and flexible hydrocarbon chains are designed, synthesized and then deposited on the (001) surfaces of bulk sodium chloride, potassium chloride, and rubidium chloride. The deposited structures are imaged using noncontact atomic force microscopy and modeled using density functional theory. The results show that it is possible to form large-scale, highly ordered, 2D, porous molecular domains (>10⁴ pores), which are stable at room temperature, and to control the size of the 2D pores. Alternatively, it is possible to form line structures or droplets (through molecular dewetting) by altering the molecular structure or changing the substrate lattice constant. Theoretical calculations explain the balance of the molecule–molecule and molecule–surface interactions and the structure and thermodynamic stability of the grown films.

1. Introduction

Thin films composed of organic molecules are of wide interest in surface science, catalysis,[1] lubrication,[2,3] and molecular electronics.[4–6] Controlling the structure and properties of these films at molecular scale is of critical importance for designing novel functionalities of the overall material, which are determined by the substrate as well as the film. A high level of control over various molecular structures has been achieved on metal surfaces. For example, functional end groups have been incorporated into porphyrin molecules in order to fabricate 1D or 2D structures,[7] highly ordered mixed phases of different molecules have been created,[8,9] and polymerization has been induced to guarantee a good lateral conductivity in these films.[10–14] Highly controlled molecular self-assembly on semiconductors has also been demonstrated.[15] However, applications, such as optoelectronic and molecular devices, require the use of bulk insulators in order to avoid the coupling of molecular orbitals with the surface.[16] The formation of organic layers on insulators has been achieved in the past by either strengthening the molecule–surface (M–S) interaction by means of appropriate anchoring groups,[17,18] or by reinforcing the molecule–molecule (M–M) interaction.[19,20] Several attempts have also been made to influence the balance between M–S and M–M interactions.[21–23] However, achieving control over structural characteristics of large-scale supramolecular networks, such as e.g., pore size, proved challenging. One of the main difficulties for achieving self-assembly on insulating surfaces is concerned with relatively weak M–S interactions that lead to molecular diffusion and the dewetting of molecular layers.[24,25] For instance, π-stacking M–M interactions can dominate weak M–S interactions.[24,25] Two main strategies have been proposed in order to circumvent these problems. The first one concerns substrate choice and modification. Surface patterning of ionic surfaces by e-beam irradiation has been shown to limit diffusion[25] while surfaces with high surface energy, such as calcite, prohibit dewetting.[26] A second strategy involves educated molecular design. Some molecules are able to form wire structures through π-stacking,[27,28] other examples include stabilization of the molecular layers via hydrogen bonding,[19] and a few experiments report on the use of polar groups to enhance adsorption onto ionic substrates.[17,29,30] However, there is no single set of design criteria for the morphology and stability of such films that can satisfy a wide range of possible applications. For example, molecular electronics applications may require that the molecules self-assemble in large, highly ordered, and defect-free domains.

A. Amrous, Dr. F. Bocquet, Dr. L. Nony, F. Para, Prof. C. Loppacher
Aix-Marseille Université
CNRS, IM2NP UMR 7334
13397, Marseille, France
E-mail: christian.loppacher@im2np.fr
S. Lamare, Prof. F. Palmino, Dr. F. Cherioux
Institut FEMTO-ST, Université de Franche-Comté
CNRS, ENSMM
Besançon, France
D. Z. Gao, Dr. F. Federici Canova, Dr. M. B. Watkins, Prof. A. L. Shluger
Department of Physics and Astronomy
University College London
Gower Street, London, WC1E 6BT, UK
E-mail: david.gao.10@ucl.ac.uk
F. Federici Canova, Prof. A. L. Shluger
WPI-Advanced Institute for Materials Research
Tohoku University
2–1–1 Katahira, Aoba-ku, Sendai 980–8577, Japan
DOI: 10.1002/admi.201400414
contrast, self-healing and the ability to smooth asperities are more important for lubricants, while catalysts should provide accessible and effective reaction sites.

The long-term aim of our work is to design molecules that can form stable functional supramolecular structures on bulk insulators at room temperature. To achieve that, the chosen substrate and molecule should have the right balance of M–M and M–S interactions in order to inhibit dewetting and impose a regular molecular structure. In order to characterize the produced monolayers (MLs) with molecular resolution we employ noncontact atomic force microscopy (NCAFM) as this provides a nondestructive method to study the delicate structure of molecular layers, which has been successfully applied in the past to image individual molecules, aggregates, and films on insulating surfaces.[17,31–42] However, while a number of supramolecular structures have been observed,[35–37,39,40] the mechanisms of structure formation are not fully understood. Moreover, studies demonstrating high levels of control over the structure formation on insulators are still rare due to the subtle balance of M–M and M–S interactions.

In this study, we designed, synthesized and deposited organic molecules with interchangeable polar functional groups, a variable length aromatic body, and flexible hydrocarbon chains. The ability to modify each portion of these modular molecules allows us to study the competing interactions within these systems and improve our understanding of structure–property relationships. Sodium chloride (NaCl), potassium chloride (KCl), and rubidium chloride (RbCl) were chosen as ionic substrates to provide different lattice constants as an additional control parameter. The deposited structures were imaged using NCAFM at room temperature and modeled using density functional theory (DFT). Our theoretical studies allowed us to describe the M–M and M–S interactions, and elucidate the structure and thermodynamic stability of the grown films. The results show that using the designed molecules we are able to form large-scale, highly ordered, 2D, porous molecular domains, which are stable at room temperature, and to control the size of the 2D pores. We can also form line structures or droplets (through molecular dewetting) by altering the molecule or changing the substrate lattice constant and, consequently, the M–M and M–S interactions.

2. Results

2.1. CDB: Designing a Modular Molecule (or Building Block)

The main objective of this work is to grow large-scale supramolecular frameworks on insulators via molecular deposition. To provide specific M–M and M–S interactions, 1,4-bis(cyanophenyl)-2,5-bis(decyloxy)benzene (CDB) molecules, based on a triphenyl core ended by two cyano substituents and substituted by two lateral decyloxy chains (see Figure 1A), have been designed.

Lateral decyloxy chains have been observed to promote supramolecular self-assembly by van der Waals (vdW) interactions between parallel aligned chains (interdigitation).[43] A triphenyl core contributes to M–M interactions through π-stacking.[44] If only the interdigitation and π-stacking M–M interactions were present, then 2D line structures would probably be formed. The added cyano groups, which are believed to be involved in hydrogen bonding to enhance M–M interactions,[45] also provide site-specific M–S interactions,[17] allowing other types of film structure to be stabilized.

In order to investigate the role of cyano groups in the balance of M–M and M–S interactions, the two cyano groups have been substituted by vinyl (1,4-bis(vinylphenyl)-2,5-bis(decyloxy)benzene) groups or methyl (1,4-bis(tolyl)-2,5-bis(decyloxy)benzene) groups. The role of π-stacking has been highlighted by elongating the aromatic core from three to five rings to form the 1,4-bis(4″-cyanobiphenyl)-2,5-bis(decyloxy)benzene (CBDB) molecule.

2.2. CDB on KCl(001): Self-Assembled Highly Ordered Porous Domains

Targeting potential nanoelectronics applications, we would like to achieve the controlled formation of large, defect-free 2D
layers, with molecules preferentially adsorbed on identical substrate sites in order to preserve the molecules’ electronic and optical properties. The control parameters for supramolecular structures are the size of the central aromatic part of the molecule, the choice of its end groups, as well as the lattice constant of the substrate constant. The most striking structure observed in this work is an extended 2D porous network obtained by depositing one ML of CDB molecules on KCl(001), shown in Figure 2A.

In the course of deposition, the molecules initially decorate substrate step edges and then form large (>300 \times 300 \text{ nm}^2) homogeneous domains. Areas with different molecular domain orientations are mostly observed at step edges, as shown in Figure 2A. We note that NCAFM imaging of weakly bound molecules on insulating substrates at room temperature is very challenging as it must be performed with extremely weak tip–sample interactions (note that the vertical contrast in Figure 2B is only 10 pm) in order to avoid manipulation of the molecules. A careful evaluation of the substrate orientation (see inset of Figure 2B) and the molecular layer in Figure 2B reveals a highly ordered square network in a point-on-point epitaxy with the substrate. Large defect-free domains (>10^4 molecules) are observed on most areas of the surface. Following the notation conventions given in Figure 1B, the measured lattice constant of the molecular film is \(a_{1m} = a_{2m} = (2.18 \pm 0.06) \text{ nm}\) and \(\beta = 90^\circ\), which is within less than the 3% of the experimental error to the square point-on-point epitaxial matrix with the substrate of [46]:

\[
\begin{pmatrix}
  a_{1m} \\
  a_{2m}
\end{pmatrix} = M \begin{pmatrix}
  a_{1x} \\
  a_{2x}
\end{pmatrix}, \quad \text{with} \quad M = \begin{pmatrix}
  3.5 & -0.5 \\
  0.5 & 3.5
\end{pmatrix}
\]

yielding \(a_{1m} = a_{2m} = 2.23 \text{ nm}\). Figure 2C shows a proposed model of this structure, where the molecules are kept in their gas-phase geometry (as calculated by DFT) and the hydrocarbon decyloxy chains are drawn straight. To further investigate the structure and stability of the observed porous structure we performed DFT calculations.

### 2.3. Calculation of the Molecule–Surface Interactions

In order to better understand the balance between the M–M and M–S interactions, we used (dispersion corrected) DFT according to the following plan. The adsorption of an individual CDB molecule on clean ionic surfaces was first investigated by carrying out geometry optimizations starting from different initial configurations. The molecule was then split into smaller molecular fragments, which were used to study the individual M–M and M–S interactions in the system. Using the results from these DFT simulations, we then benchmarked and employed a mixed quantum mechanics and molecular mechanics (QM/MM) [47,48] embedding scheme. This allowed us to reduce the number of quantum atoms in our simulations and thus increase the surface area and compute the adsorption energies of various ML structures. To compare the obtained structures with the experimental NCAFM images we simulated theoretical images using a virtual AFM (vAFM).

To characterize the M–S interaction on KCl(100) we performed Mulliken population analysis, produced charge density difference plots of the optimized adsorption geometry of a single CDB molecule, and studied the projected electronic density of states of the system. The highest occupied molecular orbital (HOMO) of the molecule lies nearly 2 eV above that of the surface while the lowest unoccupied molecular orbital (LUMO) of the molecule lies nearly 2 eV below that of the surface, effectively decoupling their electronic properties. These results demonstrated no significant charge transfer between the...
molecule and the KCl surface, indicating that the molecule is mainly physisorbed and that the electronic structure of the molecule is largely preserved. The adsorption energy can be partitioned into several contributions. The electronic interaction of the polar (oxygen, cyano) groups are about 0.7 eV, as calculated by a standard generalized gradient approximation (GGA) DFT functional. Meanwhile, vdw interactions contribute 2.4 eV to the total adsorption energy of 3.1 eV. The adsorption configuration of the molecule is determined mainly by the interaction between the cyano groups (separated by 1.65 nm from each other, as shown in Figure 1) and the K cations of the surface. This anchoring of the molecule translates into a preference in orientation along the (110) direction in the lowest energy configuration shown in Figure 3.

In addition to this lowest energy monomer adsorption geometry, the CDB molecule can be rotated around one end group in either direction in order to fit the other cyano end group above other surface cation sites without significant loss of adsorption energy. This is due to the fact that, although the distance between the cations slightly increases, the molecule is flexible enough to still efficiently anchor to the surface.

In order to further elucidate the different M–S and M–M contributions involved in the CDB adsorption, we carried out DFT calculations of molecular fragments (cyanophenyl, decyloxy chains) on alkali halide surfaces. The strongest M–S interaction comes from the CN functional group of cyanophenyl adsorbing to cation sites on the KCl surface. The fragment CN group adsorbs above a cation site while the attached benzene ring adsorbs parallel to the surface plane with combined adsorption energy of 0.7 eV. These results are consistent with the previous calculations that give the interaction of a physisorbed benzene ring with the KCl(001) surface of 0.5 eV and the interaction of a cyano group with KBr(001) of 0.2 eV. The adsorption energy of the CN group increases above anion vacancies, divacancies, and at step edges and kinks where it can interact with more cations at once.

The second significant contribution is between the O atoms of the decyloxy chains and the surface. However, in the full, unfragmented molecule the central phenyl ring adsorbs in a tilted configuration to minimize steric interactions with adjacent rings. This means that only one O atom can interact with the surface at a time, resulting in a net interaction that is much weaker than the CN interactions with surface K ion sites.

Finally, the rumpling of ionic surfaces results in the cations being displaced inside, towards the bulk and the anions slightly elevated above the surface plane, in agreement with the prior calculations and experiments. However, this does not appreciably affect the adsorption of hydrocarbon chains as the differences in energy between various chain configurations at the surface are within 0.02 eV. Therefore, at room temperature the hydrocarbon chains should be quite mobile in contrast to the more stable anchoring groups.

2.4. Calculation of the Molecule–Molecule Interactions

In addition to the M–S interactions described above, another major contribution affecting the growth of CDB supramolecular structures on KCl(001) comes from the M–M interactions. Since the electronic properties of the molecule do not change significantly as a result of adsorption to the KCl(100) surface, the M–M interactions can be studied in the gas phase. As expected, CN functional groups repel each other. However, CN groups interact favorably with the hydrocarbon decyloxy chains of adjacent molecules; this latter interaction energy is 0.2 eV at an optimal distance of 0.36 nm when the CN and CH₃ groups are positioned end to end. This value is in agreement with prior studies of the CN-alkyl chain interaction (0.12 eV with a distance of 0.45 nm).

The benzene rings that make up the core of the molecule can also attract each other in a side-by-side configuration on the surface with 0.1 eV of interaction energy with an optimal distance of 0.28 nm between the centers of the rings. This is consistent with prior studies that also estimate π-stacking of parallel-displaced benzene rings can contribute up to 0.1 eV to M–M interaction. Finally, depending on the distance between rows on the alkali halide surface, the hydrocarbon chains may either attract due to vdw forces or repel each other at close range. This interaction is also on the order of 0.1 eV at the optimal distance of 0.5 nm on the KCl(100) surface. It falls off to zero as the chains move farther away, and quickly becomes repulsive at distances shorter than 0.4 nm.

2.5. Calculation of the CDB Monolayer Structure

We are now in a position to propose more detailed atomistic models of the CDB ML. Using the periodicity of the film formed on KCl(001) from experimental data as a constraint, we constructed periodic MLs of CDB on KCl(001) by placing
molecules in various configurations on the surface and allowing the structures to relax. The lowest energy configurations are shown in Figure 4.

Since the CDB molecule is not symmetric, films can be either enantiopure (Figure 4A), racemic (Figure 4B), or somewhere in between. The two possible enantiopure as well as the racemic structures, however, were found to be energetically equivalent. In the ML configuration each CDB molecule gains 0.5 eV via interactions with neighboring molecules. However, due to distortions from the ideal adsorption geometry, ML configurations shown in Figure 4 are only 0.3 eV more stable per molecule than isolated CDB molecules on KCl(001), meaning that the total adsorption energy per molecule is increased to 3.4 eV.

To try and confirm the pore structure of the ML, and make direct link between theory and experiment, attempts were made to image the ML structure at large tip–surface distances to minimize the layer distortion by the tip. Although, generally this has not been possible due to a lack of signal, rarely images with measurable contrast were obtained (such as the one shown in Figure 5C). To simulate an NCAFM image of the theoretical structure using a vAFM, we assumed that at long tip–substrate distance the tip can be well represented by a point dipole corresponding to a silanol group or KCl cluster at the end of the tip (see Methods for further discussion).

A theoretical image of the most stable atomistic structure of the CDB ML on KCl(001) (shown in Figure 5A) calculated using a vAFM, a point-dipole tip model of 10 Debye, and a tip–substrate distance of 0.8 nm is shown in Figure 5B.

The bright lines in the image correspond to the main body of the CDB molecule and the two CN functional groups. The hydrocarbon chains are not visible due to a lack of electrostatic interactions with the tip. This theoretical image can be directly compared with the experimental image of CDB on KCl(001) shown in Figure 5C obtained at large tip–surface separation. Note that for both structures presented in Figure 4 there are alternating pores with the CN-groups or the hydrocarbon chains looking inwards, which can be seen in both experimental and theoretical images as squares with different contrast—the four marked molecules in Figures 5A, and B surround a pore with CN-groups facing inwards. The qualitative agreement between theory and experiment suggests that the image interpretation is realistic and strengthens the suggestion that the pores are indeed empty.

2.6. Modifying the Molecule

To demonstrate control over the growth and structure of molecular films, we modified the CDB molecule in two ways. The first was to enlarge the central aromatic body of the molecule (see Figure 1A with \( R = \phi-CN \)). The idea was to confirm that this larger CBDB molecule is also able to form square porous networks and elucidate the effect that changing the size of the central body has on the network structure. Indeed, the geometrical conditions for forming a square porous network were again satisfied on KCl(001), as shown in Figure 6.

This modified molecule also forms large, defect-free domains in the shape of a 2D porous network. These domains are oriented differently with respect to the surface due to changes in commensurability and are composed of larger pores (pore size 4.8 nm² for CDB and 6.1 nm² for CBDB). The measured lattice constant of the CBDB molecular film is \( a_{1m} = a_{2m} = \)
For this group of molecules, the vertical M–S interaction between CN and CH₃ groups. M–M interactions with a large contribution from the interaction between anchoring end groups and cations on the ionic substrate. The formation of different molecular organizations is possible while anchoring the molecules to their preferred adsorption site, the porous network is formed. For the other investigated substrates (NaCl and RbCl), it is π-stacking and for some systems possibly interdigitation of the hydrocarbon decyl chains that determines the formation of 2D line structures. The experimental data are reported in Figure 8.

For all these structures, the experimentally determined parameters show that there is an overlap of the aromatic cores of the assembled molecules that corresponds well to the distances normally observed for π-stacking. Interdigitation of the hydrocarbon chains is not always very significant and the interline distances vary. This is most probably due to the fact

$$M = \begin{pmatrix} 2.5 & -3 \\ 3 & 2.5 \end{pmatrix}$$

(2.47 ± 0.06) nm and β = 90°. The closest epitaxial relation with the substrate is the one with an epitaxial matrix of:

$$a_{1m} = a_{2m} = 2.46 \text{ nm}$$

The experimental data are reported in Figure 8.

Figure 6. CBDB molecules on KCl(001). A) NCAFM large-scale topography image (frame edge = 120 nm; γ = −0.03 nN/nm, A₀ = 10 nm) showing two differently oriented molecular domains D₁ and D₂. The inset shows the FFT analysis which reveals the orientation of the two domains. B) Molecular scale image (frame edge = 40 nm, digital zoom of (A) showing the square porous molecular network as well as the orientation of the molecular lattice with respect to the substrate orientation: α = 90°, and θ = (−50 ± 4)°. C) Proposed structural model taking into account that all CN groups adsorb above the K ions: a₁m = a₂m = 2.46 nm, β = 90°, and α = atan(−3/2.5) = −50.2°. The closest epitaxial relation with the substrate is the one with an epitaxial matrix of:
that these chains are very mobile at room temperature and for the larger CBDB molecule in particular, there is a lot of room between individual π-stacked molecular rows. For the other functional end groups considered (i.e., vinyl and methyl termination), the main observation is that there are line structures for the vinyl groups on all investigated substrates and that there is a more or less rapid dewetting process for all methyl terminated molecules (data not shown).

3. Discussion

One of the major features governing assembly of 2D networks on insulating surfaces is the balance of relatively weak interactions that controls the supramolecular structure formation. Our experimental and theoretical results illustrate this fine balance at work. On the KCl(001) surface, CDB molecules form a defect free, porous 2D structure. In this system, the CN to surface–cation interaction is the strongest site specific M–S interaction component that plays a large role in the structure formation. As this interaction is weakened via a vinyl substitution of the CN group, the other contributions become more significant resulting in the preferential formation of a 2D line structure. Indeed, previous studies indicate that π-stacking should contribute significantly to M–M interactions.\[53-56] Furthermore, weakening of the anchoring group’s interaction with the surface results in reduction of adsorption strength and leads to dewetting. The molecules still adsorb to KCl(001) terraces and decorate step edges initially, as previously observed. However, as the M–M interaction becomes stronger than the M–S interaction, the formation of large domains of molecular crystals is favored over the formation of defect-free MLs.

Since our experimental data depict regular networks that are free of defects, it is reasonable to infer that room temperature NCAFM provides information on the average position
rather than the instantaneous position of atoms on the surface. Mobile features, such as diffusing molecules, are typically not experimentally observable on the timescale of NCAFM image acquisition and must be stabilized before they can be observed. Averaging likely encompasses both full molecular motion and the motion of individual groups. For instance, the hydrocarbon chains of decyloxy groups can be stabilized by an interaction with the CN group of a neighboring molecule, however, in the proposed ML configuration the hydrocarbon chains are weakly digitated of hydrocarbon chains, and via an interaction between the hydrocarbon chains with the cyanophenyl rings. The strength of the latter interaction was confirmed by DFT calculations. To control the assembly, we modified the number of phenyl rings in the central part of the molecule, its polar end groups, as well as the ionic substrate. These results demonstrate that the delicate balance between M–M and M–S interactions can be used to control not only the orientation and spacing of π-stacked and interdigitated lines (2D line growth), but more interestingly the growth of a porous 2D network with adjustable pore size. A detailed understanding of the film-formation mechanism was obtained by means of theoretical calculations, which revealed the strength of the competing M–M and M–S interactions and elucidated the experimentally found porous structure. The most striking result of our work is that we could obtain a defect-free porous network of more than \( 10^5 \) molecules. Such porous networks are prospective candidates for optical and catalytic applications where either reactive clusters can be deposited inside the pores or functional groups added to the molecules. As all molecules within the network are adsorbed on identical substrate sites and the substrate is a bulk insulator, this guarantees that the molecule’s properties are entirely preserved.

5. Experimental Section

Molecule Synthesis: The starting point of the molecular synthesis was the CDB molecule, which consists of a central part with three phenyl rings, the outer ones being terminated by polar cyano end groups (Figure 1A). Two decyloxy chains were attached to the central ring. In order to control on the self-assembly of the molecule we chose to vary two parameters in the molecular design: first, the polar cyano groups were substituted by the less polar vinyl or the nonpolar methyl groups; second, the length of the central part was increased to five rings, the outer ones being terminated by polar cyano end groups.

4. Conclusions

In summary, we have successfully designed a group of molecules and achieved control over their assembly into highly ordered, defect-free, 2D molecular networks on ionic alkali halide surfaces. In order to avoid dewetting while still maintaining a high degree of ordering of the organic layer, polar anchoring groups were added to the molecule’s core to strengthen the M–S interaction. Flexible M–M interactions were achieved via π-stacking of phenyl rings of the molecule’s core, interdigitation of hydrocarbon chains, and via an interaction between the hydrocarbon chains with the cyanophenyl rings. The strength of the latter interaction was confirmed by DFT calculations. To control the assembly, we modified the number of phenyl rings in the central part of the molecule, its polar end groups, as well as the ionic substrate. These results demonstrate that the delicate balance between M–M and M–S interactions can be used to control not only the orientation and spacing of π-stacked and interdigitated lines (2D line growth), but more interestingly the growth of a porous 2D network with adjustable pore size. A detailed understanding of the film-formation mechanism was obtained by means of theoretical calculations, which revealed the strength of the competing M–M and M–S interactions and elucidated the experimentally found porous structure. The most striking result of our work is that we could obtain a defect-free porous network of more than \( 10^5 \) molecules. Such porous networks are prospective candidates for optical and catalytic applications where either reactive clusters can be deposited inside the pores or functional groups added to the molecules. As all molecules within the network are adsorbed on identical substrate sites and the substrate is a bulk insulator, this guarantees that the molecule’s properties are entirely preserved.
Molecules were deposited by evaporation from homebuilt crucibles for 1 h in order to obtain atomically clean substrates with large terraces. Quickly introduced into UHV and annealed to approximately 200 °C, the resulting white solid is purified by column chromatography (silica gel, CHCl₃). The organic solution was washed three times with 100 mL of water. The solvent was removed under reduced pressure. The crude solid was dissolved in 100 mL of Chloroform. The solvent was removed under reduced pressure. The resulting white solid is purified by column chromatography (silica gel, acetone/petroleum ether 1:1). The solid was recrystallized in acetonitrile to give a pure white solid (710 mg, 67%).

**Sample Preparation:** The single crystal ionic substrates NaCl, KCI, and RbCl (MaTeK GmbH, 52428 Jülich, Germany) were cleaved ex situ, quickly introduced into UHV and annealed to approximately 200 °C for 1 h in order to obtain atomically clean substrates with large terraces. Molecules were deposited by evaporation from homebuilt crucibles at evaporation rates of 10 nm/s, which result in a 1 nm molecule landing per oscillation cycle. The WsXM software. All distances and directions were calibrated with respect to the ionic crystal surface from atomically resolved images of the substrate. The experimental parameters for the oscillating conditions of the cantilever are given in the form of the so-called normalized frequency shift, η, introduced by Giessibl[60] with η = Δf × k / f₀ and the energy dissipation E₀ = π × k × Δf₀ / Q in [eV/cycle].[61]

**Density Functional Theory:** DFT simulations to examine the interactions responsible for CDB adsorption and compare competing interactions between molecules on the surface were performed with the CP2K code using the PBE GGA density functional and a mixed Gaussian and plane waves basis set.[62,63] These calculations included geometry minimizations of fragments of the CDB molecule interacting in the gas phase and fragments of the molecule interacting with the KCl(001), NaCl(001), and RbCl(001) surfaces. The surface was represented using four atomic layers with one fixed layer at the bottom and three layers of relaxed atoms in order to reproduce the band gap, rumpling, and lattice constants. We employed semimicroscopic long-range dispersion corrections[64] in order to represent vdW interactions in the system. The MOLOPT[65] basis set was used to minimize basis set superposition error.

The lowest energy configuration of each molecule was computed by placing the molecule onto each surface in various positions and optimizing the geometry. These starting positions were obtained by rotating the molecule by 2° increments on the surface. The molecule was then divided into several main components by breaking bonds in order to examine the various contributions to adsorption energy. Each of these broken bonds was terminated with an H atom. The fragments studied include A) a hexane molecule representing the hydrocarbon chains, B) a benzonitrile molecule representing the functional groups of the larger CDB molecule, and C) a benzene molecule representing the core of the CDB molecule.

**Virtual Atomic Force Microscopy:** Theoretical NCAFM images were produced using a point dipole model of the NCAFM tip. For experimental images taken at large tip–sample distances, it is reasonable to attribute the observed image contrast to electrostatic interactions between a dipole moment present in the tip itself and the potential of the sample. This method allows us to produce a force field directly from DFT calculations, as has been discussed previously[68] and evaluated in the context of imaging small molecules on NiO(100). While the tips employed in this study are not metal coated tips as in the previous study,[68] they may still possess permanent dipole moments near the tip apex depending on tip termination. Si tips used in this work are likely to be terminated by inert silanol or silisilan groups,[68] which may result in a significant dipole moment at the tip apex. While the actual dipole moment of this tip termination is unknown, previous calculations suggest that the dipole moment of an isolated silanol group is of the order of 1.5 D. Fragments from the KCl surface are also very plausible tip contaminants.

In this data set, we do not have long-range force curves or large sets of scans over the bare surface so were not able to explicitly fit the macroscopic vdW contributions or the magnitude of the dipole moment.[68] The interactions between the macrotip and the sample were represented using a generic tip with a Hamaker constant of 5.94 × 10⁻¹⁰ J, a tip radius of 3.95 nm, and a tip angle of 16.42°. Since this contribution is macroscopic, it will have a large effect on the frequency shift observed and can provide useful information regarding the true tip–surface distance. In this work, however, we were only

**Figure 9.** Chemical sketches of tailored molecules.
interested in the observable image contrast at 0.8–1.0 nm tip–surface distances where a generic representation of the macrotip should be sufficient. Simulations were performed using the same parameters as in experiment: 150 kHz resonance frequency, a 30 N m⁻¹ spring constant, and 10 nm oscillation amplitude using the vAFM code. We simulated best agreement with the experimental images shown.

Acknowledgements

Financial support of this research by the Agence Nationale de la Recherche with the NANOKAN project (ANR-11-BST0004) is gratefully acknowledged. D.Z.G. was supported by Chevron Oronite LLC. M.W. is grateful to the Leverhulme Trust (grant F/07134/CK) for financial support. F.C.C. is supported by the MEXT WPI programme. The authors acknowledge the use of the HECTOR and the ARCHER high-performance computing facilities via our membership to the UK’s HPC Materials Chemistry Consortium, which is funded by EPSRC (EP/F067496).

Received: September 11, 2014
Revised: October 20, 2014
Published online:


