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pH and Eh effects on phosphorus fate in constructed wetland's sludge surface deposit

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Abstract

The objective of the present study was to assess the influence of extreme pH and redox potential (Eh) conditions on phosphorus (P) retention within the surface sludge deposit layer of a vertical flow constructed wetland (VFCW) where phosphorus was captured by FeCl₃ injection. Series of 27 successive batch leaching tests were conducted under acidic, alkaline or reductive conditions using a representative sludge sample taken from an 8-year old VFCW plant. Experiments were followed by monitoring the pH and Eh variations and analysing the releases of P and other selected elements into the solutions. The sludge material was also analysed before and after leaching, using solution ³¹P NMR spectroscopy and sequential chemical extractions, in order to evaluate dissolutions of both organic and inorganic P-bearing species and their respective contributions to P release. The correlations between the monitored variables were analysed and visualized through principal components analyses (PCA). Results showed a very good stability of P retention in the sludge deposit and a relatively good acid-buffering capacity of the sludge, revealing that the risk of accidental P release into the environment would be extremely low during the real plant operation.

Keywords:

vertical flow constructed wetland; surface deposit; phosphorus leaching; environmental engineering; pH; redox

1. Introduction

Due to its role in eutrophication, its fertilizing value and the progressive depletion of its natural resources, phosphorus (P) has become a key element to control in environmental policies worldwide (Cordell et al., 2009; Reijnders et al., 2014). Domestic wastewaters constitute significant P sources, originating mostly from human metabolism and detergents (respectively estimated around 1.2 g and 1 g of P per person and per day). Increasing research efforts have therefore been focused in the recent years on phosphorus recovery from domestic wastewaters (Brod et al., 2015a; 2015b; García-Albacete et al., 2012; Weigand et al., 2013) and sewage sludge (Kahiluoto et al., 2015; Peccia and Westerhoff, 2015; Wilfert et al., 2015). The development of chemical and biological treatments to enhance the fertilizing value of organic waste and/or sequester P in soil while maintaining its availability to crops has even been claimed as a major research need (Sharpley et al., 2015).

Constructed wetlands processes have proven well adapted to the treatment of various types of wastewaters (Vymazal and Kröpfelová, 2009). The efficiency of classical vertical-flow constructed wetland (VFCW) processes is very good in the treatment of carbon and ammonia from small communities' domestic wastewaters. As for phosphorus, different alternatives are possible but unlike C and N, P can only be captured within the system and not removed from it or degraded into innocuous forms. In the system studied here (AZOE-NP[®], patent EP1857419A1; WO2012150296), dissolved phosphates are precipitated by injection of FeCl₃ before the wastewater is percolated through the VFCW. The precipitates thus formed, mixed with the suspended solids already present in the wastewaters, thereby progressively accumulate as a sludge deposit layer at the surface of the VFCW (Kim et al., 2013).

Over the years of operation of VFCW, environmental or operational factors may induce stresses on pH and Eh. For example, along with domestic wastewaters, some plants may receive agricultural and/or industrial wastewaters such as winery effluents whose pH is

relatively acidic (Masi et al., 2002; Shepherd et al., 2001) or brewery effluents whose pH is rather alkaline (Fillaudeau et al., 2006). Constructed wetlands may also be exposed to acidic rainfalls with pH reported to range from 4.4 to 5.7 in eastern France (Sanusi et al., 1996) and even as low as 2.9 in some parts of China (Du et al., 2014). Finally, temporary ill-functioning of the system resulting in the flooding of the filters may install reductive conditions within the sludge layer (Kim et al., 2015a).

These variations of pH and/or Eh are known to influence soil/plant/microorganism systems (Husson, 2012). For example, the dynamics of trace metals and organic matter (Blossfeld et al., 2010; Grybos et al., 2009) and P release (Castaldi et al., 2010; Christophoridis and Fytianos, 2006; Kjaergaard et al., 2012; Maine et al., 1992) can be affected. Very few studies are available however on the effect of such factors in constructed wetland (Braskerud et al., 2005), especially in vertical flow systems.

The present study was conducted to evaluate potential risks of P release during the operation of VFCW plants and gain better knowledge on the potential fate of phosphorus when the sludge should be dredged from the plant to avoid clogging of the filters after 10-15 years of operation. The scientific objectives were to (i) determine the influence of extreme conditions of pH and Eh on phosphorus release and (ii) characterize the effect of P leaching on its distribution and speciation within the sludge.

2. Materials and Methods

2.1. Origin and composition of sludge samples

Sludge samples were taken from the surface deposit layer of a full-scale AZOENP[®] plant located in Vercia (Jura, France) in operation for 8 years at the time sampling was done. It was designed to treat on a regular basis domestic wastewaters from the Vercia township, but also to receive winery effluents in autumn. Planted with *Phragmites Australis*, its maximum

capacity corresponds to 1,100 people equivalent (PE) and previous studies reported its good performance (Kim et al., 2015b; Kim et al., 2014). The average depth of sludge layer accumulated on the first VFCW stage of this plant since its implantation was around 16 cm, corresponding to a rate of accumulation of about 2 cm per year.

Primary sampling operations were conducted using an excavator (power shovel) to extract the entire depth of the deposit layer from various spots of the surface of the first filter. The samples were then treated manually to sort out rhizomes/roots and coarse fractions (gravels), and subsequently mixed by quartering before being transported to the laboratory. There, an aliquot fraction was spread in a few-cm thick layer, dried at 35°C and then manually crushed to a particle size below 1 mm. The crushed dry materials were then stored at 4°C until the beginning of experiments and analyses. According to the elemental analysis results of this sludge, total contents of P, Fe, Ca, Al and Mn were found to be 2.57, 3.78, 4.54, 3.06 and 0.03% w/w of dry matter, respectively. Additional information relative to the sludge from the same plant was reported previously by Kim et al. (Kim et al. 2013, Kim et al. 2015c).

2.2. Experimental leaching protocol and leachates analyses

The influence of pH and redox potential (Eh) on water-extraction of phosphorus from the sludge was investigated following a sequential leaching protocol. Ten grams of dry sludge were mixed in a 250 mL centrifuge tube (polypropylene copolymer, Nalgene®) with 100 mL of leachant at each sequence. The following aqueous leachants were used.

- [Blank]: deionized water considered as reference conditions,
- [Acidic]: 0.01 M hydrochloric acid aqueous solution,
- [Alkaline]: 0.01 M sodium hydroxide aqueous solution,
- [Reduced]: 0.05 M sodium ascorbate aqueous solution (Chatain et al., 2005).

Blanks and assays were all made in triplicates. The tubes were closed with a screw cap and placed on a rotary shaker at 9 rpm at a temperature of $16 \pm 5^\circ\text{C}$. After 24 h, the tubes were rested for 5-10 minutes to allow settling of the suspensions before measurement of redox potential (Radiometer Analytical MC3051Pt combination metal electrode; Ag/AgCl reference, HACH), pH (electrode Sentix[®]; WTW) and temperature. The tubes were then centrifuged at 9000 rpm (ca 13000 g) for 20 min. The supernatants were collected using disposable syringes, filtered through 0.45 μm Sartorius Minisart filters and stored at 4°C before analyses. Each tube was filled again with fresh leachant to compensate the volume of supernatant extracted, and a new sequence was started. A total of 27 successive leaching sequences were performed. At the end of the experiment, sludge material was collected from each tube, freeze-dried, finely grinded and stored at 4°C until being analyzed.

Each collected supernatant was analyzed for determination of dissolved molybdate reactive phosphorus concentration (MRP, mostly orthophosphates) according to the blue-molybdate method (Murphy and Riley, 1962) using a ShimadzuUV-2450 spectrophotometer. For the other analyses, supernatants of 3 replicates were mixed except if phosphorus concentrations revealed particular samples. Total concentrations of dissolved aluminum (Al), calcium (Ca), iron (Fe) and manganese (Mn) were determined by ICP OES (Ultima 2; Horiba Jobin Yvon SAS).

2.3. Sequential phosphorus fractionation

The fractionation protocol proposed by Golterman (1996) was followed.

Aliquots of the sludge sample were extracted at a solid to solution ratio of 1g/100mL according to the following steps:

- 1 extraction of 2h with deionized water,
- 4 extractions of 2h with 0.05 M Ca EDTA + 1 % of Na dithionite; (pH around 7.5),

- 2 extractions of 16h with 0.1 M Na EDTA (pH 4.5),
- 1 extraction of 30min with 0.25 M H₂SO₄,
- 1 extraction of 30min with 1 M NaOH at 90 °C,

P extracted at each step was respectively considered to be associated to the following pools: water soluble phosphate (H₂O~P), iron bound phosphate (Fe~P), calcium bound phosphate (Ca~P), acid soluble organic phosphate (ASOP, P_{orgacid}) and alkali soluble organic phosphate (P_{orgalk}).

All extractions were carried out under continuous rotary shaking. Following extraction, each suspension was centrifuged at 4000 rpm (ca 3400 g) for 15 min then an aliquot fraction of the supernatant was analysed for dissolved MRP and total P determination. Total P concentration was measured by ICP-AES (Ultima 2; Horiba JobinYvon SAS). The difference between total P and MRP was referred to as molybdate unreactive phosphorus (MUP) which could include complex inorganic P forms, non-hydrolysed organic P and colloidal P (Cade-Menun and Liu, 2014; Haygarth and Sharpley, 2000).

2.4. Solution ³¹P nuclear magnetic resonance spectroscopy

Solution ³¹P NMR was used to seek information about the organic phosphorus species (Liu and Davis, 2014). Phosphorus was firstly extracted from the sludge material by mixing samples of freeze-dried sludge in a solution containing 0.25 M NaOH and 50 mM Na₂EDTA. Mixing was conducted for 16h at room temperature with a solid/liquid ratio of 1g/20mL. After centrifugation at 4000 rpm (ca 3400 g) for 15min, the supernatants were filtered (0.45 µm; Sartorius Minisart). Each extract was then divided into two aliquots. The first one was used for determination of total P concentration and the other aliquot was freeze-dried and subsequently homogenized by manual grinding to a fine powder.

For sample preparation, 100 mg of each freeze-dried sample were dissolved into 900 μL of NaOH/EDTA solution and 100 μL of D_2O . ^{31}P NMR spectra were recorded at 30°C using a BrukerAvance III HD 400 MHz spectrometer equipped with a 5-mm QNP probe with z gradients and operated at 161.972 MHz. Experiments were recorded using 4.33 μs pulse (60°), a delay time of 2s, an acquisition time of 0.255s and a number of scans of 4096. Data were processed using the 5 Hz line broadening. Chemical shifts were indirectly reference to H_3PO_4 solution (at $\delta = 0$ ppm).

As paramagnetic ions could interfere with subsequent solution ^{31}P NMR spectroscopy, a treatment of NaOH–EDTA extracts with a cation-exchange resin was performed on concerned sample (Turner, 2008). The Chelex-100 cation exchange resin (50–100 mesh; Sigma-Aldrich) which correspond to a styrene divinyl benzene copolymer with a high affinity for divalent cations (at $\text{pH} > 4$; especially transition metals such as iron and manganese), was chosen for this treatment. NaOH–EDTA extracts containing high concentration of paramagnetic ions were shaken with Chelex resin (1:20 solid to solution ratio) for 1h then proceeded as described above.

3. Results and Discussion

3.1 Evolution of pH and Eh according to the nature of the leachant

Figure 1 (a) shows the evolutions of pH and redox potential (Eh) along the successive leaching sequences. For deionized water series (DW, blank), almost no variations of pH (6.6 ± 0.1) and Eh (416 ± 36 mV/SHE) were observed. With hydrochloric acid (A), the highest levels of Eh were recorded (518 ± 78 mV/SHE). The pH decreased only slightly over the first 10 sequences (from 6.2 to 5.8), revealing the buffer capacity of the sludge, and then dropped to 3.6 in the rest of the experiment. With sodium hydroxide however (B), pH started to rise from the very beginning of the experiment (pH 7.3) up to 11.7 at the end of the experiment

revealing a smaller buffer capacity of the sludge against alkaline conditions. Under this condition, Eh values (272 ± 66 mV/SHE) were significantly lower than observed in blank and under acidic conditions. With sodium ascorbate (reducing conditions), pH variations were small (5.8 ± 0.2) along the experiment as observed in the blank (deionized water). The recorded levels of Eh (-176 ± 113 mV/SHE) were the lowest observed, revealing the installation of anoxic to anaerobic conditions.

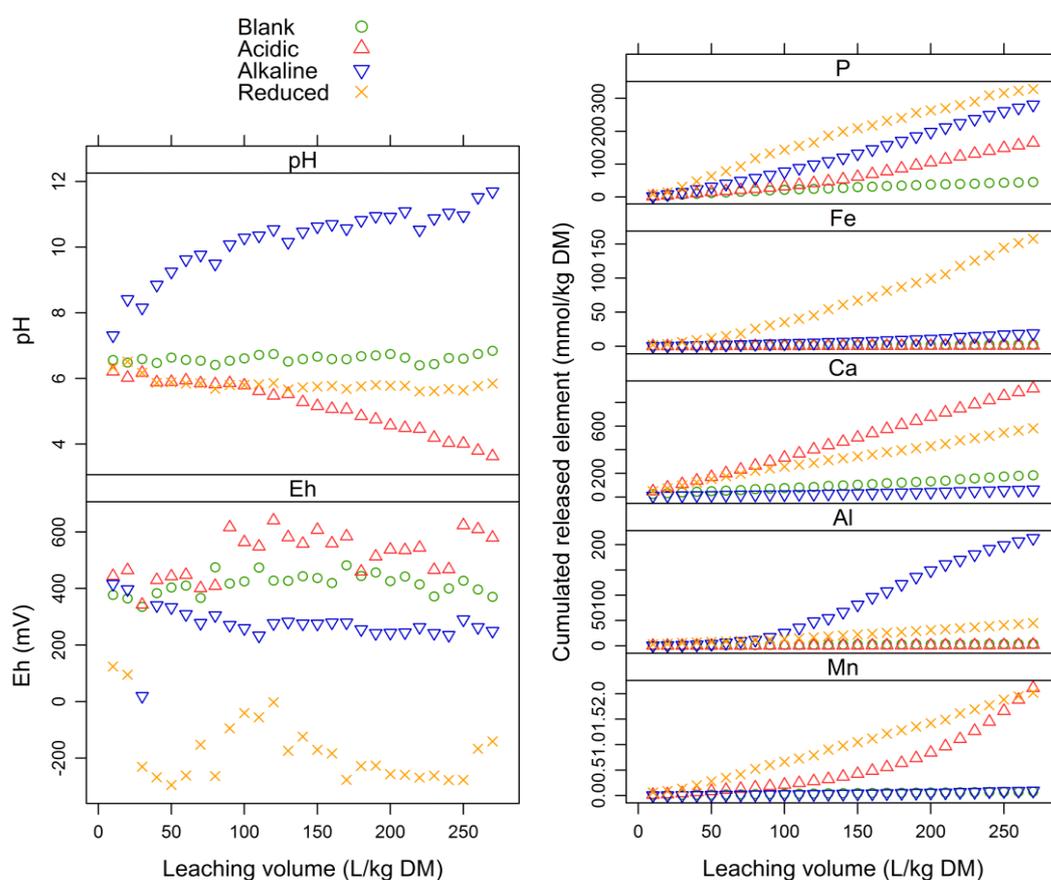


Figure 1. (a) Evolution of pH and redox potential [Eh] and (b) Cumulated release of selected elements along the successive leaching sequences in deionized water (blank) and under acidic (hydrochloric acid), alkaline (sodium hydroxide) and reduced (sodium ascorbate) conditions.

The buffer capacity of soils is known to be controlled by organic matter and clay contents (Csoma, 2010; Murphy 2014). In the present case, the influence of organic matter can be considered predominant due to the very high organic matter of the sludge (59%, Kim et al., 2013). Csoma (2010) investigated the acid and alkaline buffer capacities of 20 organic and 12 mineral materials. Regarding organic materials (peats, composts, organic manures) the author clearly observed that alkaline buffer capacity was smaller than acid buffer capacity which is in full agreement with our observations.

3.2 Cumulated releases of monitored elements

Figure 1 (b) and Table 1 show the cumulated amounts of released elements (P, Fe, Ca, Al and Mn) expressed as mmol of each element per kg of dry matter (DM) of sludge samples. Over the whole experiment, a cumulated volume of 2700 mL of leachate was collected for each condition tested (i.e. 270 L kg¹ of dry matter of sludge).

Table 1. Cumulated release of different elements over the 27 successive leaching sequences in deionized water (blank) and under acidic (hydrochloric acid), alkaline (sodium hydroxide) and reduced (sodium ascorbate) conditions.

	Blank	Acidic	Alkaline	Reduced
Element	mmol.kg ¹ (%)	mmol.kg ¹ (%)	mmol.kg ¹ (%)	mmol.kg ¹ (%)
P	45 (5.5)	165 (19.9)	281 (33.7)	329 (39.5)
Fe	2.5 (0.4)	0.8 (0.1)	19 (2.8)	158 (23.3)
Ca	184 (16.2)	919 (81.1)	61 (5.4)	582 (51.4)
Al	3.5 (0.3)	2.0 (0.2)	213 (18.8)	45 (4.0)
Mn	0.08 (1.5)	2.1 (41.8)	0.09 (1.8)	2.0 (40.0)

Phosphorus leaching from sludge was found to be low in the blank, not exceeding over the whole experiment 5.5% of the total P content, confirming that P species were poorly extractable in water considering the very high cumulated extraction volume. This result was in

agreement with the low content of water-extractible P (1.5% of total P) determined by sequential extraction by Kim et al. (2015). Over the whole experiment, the extent of P release was higher under alkaline (34% of total P content) than acidic conditions (20% of total P content) which may be related to the lower alkaline buffer capacity of the sludge. The strongest leaching was observed under reduced conditions (39.5% of total P content), indicating a specific influence of Eh, stronger than that of pH within the ranges studied here.

The extraction of iron was also very poor in water (0.4% of total Fe content) and almost nil under acidic condition in the range tested. A slight release was observed under alkaline conditions (2.8% of total Fe content) but much less than observed under reduced condition (23% of total Fe content), indicating that the Eh was the governing parameter for iron release.

The reduction of oxyhydroxides species was considered to be the most probable process.

Calcium behavior exhibited quite a different pattern than iron's. The leaching of calcium was found to be very pronounced under acidic conditions (81% of total Ca content) whereas only 5.4% of Ca was extracted under alkaline conditions and 16% in the blank. Under reduced conditions, a relatively high release of Ca was observed (51% of total Ca content). Since Ca species within wetlands are not known to be sensitive to redox conditions (Golterman, 1998), this observation may rather be attributed to the acidity of the reducing reagent (sodium ascorbate), and/or the ion exchange between sodium and calcium, than the Eh itself.

Al was not significantly released in the tested experimental conditions except in alkaline extractions (19% of total Al content), which may be attributed to its amphoteric properties.

On the other hand, Mn release was fairly high under reduced and acidic conditions (40 and 42% of total Mn content, respectively), as commonly observed in the literature (Chen et al., 2015).

3.3. Evolution of organic P forms

Organic forms represented a small but non negligible fraction (5%) of phosphorus in the sludge (Kim et al., 2015c). Due to the high content in organic matter in the sludge, the contribution of this pool to the leaching of phosphorus was investigated. Solution ^{31}P NMR spectra of NaOH-EDTA extracts are shown in figure 2 and relative percentages of organic P species in the solids given in Table 2. Literature data were used for peak assignments as follows: peaks between 5.3 and 6.6 ppm were assigned to orthophosphate (Cade-Menun et al., 2005; Newman and Tate, 1980), between 3 and 5 ppm to orthophosphate monoesters (Newman and Tate, 1980), between -1 and 2.5 ppm to orthophosphate diesters (Makarov et al., 2002) and between -4.5 and -3.5 ppm to pyrophosphate (Turner et al., 2003). Polyphosphates, usually found around -19 ppm (Turner et al., 2003), were not observed in any spectrum.

Table 2. Relative percentages of organic P species, determined by solution ^{31}P NMR spectroscopy for initial sludge (initial), sludge leached by deionized water (blank), hydrochloric acid (acidic), sodium hydroxide (alkaline) and sodium ascorbate (reduced). Results are expressed as mass % of extracted organic P.

Sample	P_{Orthomono} (%)	P_{Orthodi} (%)	P_{Phosphonate} (%)	Ratio (P_{diester}/ P_{monoester})
Initial	75.7	10.8	13.5	0.14
Blank	75.0	10.0	15.0	0.13
Acidic	66.2	15.6	18.2	0.24
Alkaline	97.8	- [†]	2.2	-
Reduced	70.7	18.9	10.4	0.27

[†] Undetected

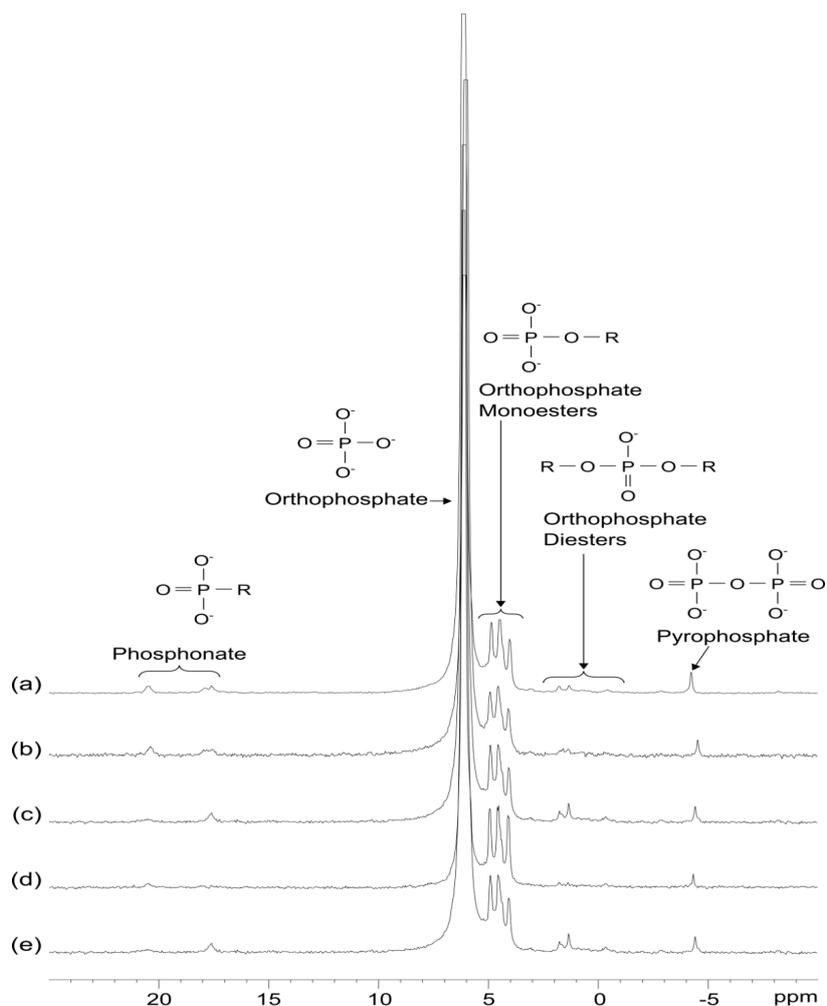


Figure 2. ^{31}P NMR spectra of NaOH-EDTA extracts of sludge samples (a) initial sludge before leaching, after leaching under (b) blank (deionized water), (c) acidic (hydrochloric acid), (d) alkaline (sodium hydroxide) and (e) reduced (sodium ascorbate) conditions.

Most of organic P extracted from the sludge solids before leaching consisted in P-monsters (75.7% of total organic P extracted), followed by phosphonate (13.5% of total organic P extracted) and diesters (10.8% of total organic P extracted).

Leaching under acidic or reduced conditions induced a strong increase of the diesters to monoesters ratio in the sludge (from 0.14 to 0.24 or 0.27, respectively), indicating that monoesters were more easily leached than diesters under such conditions. On the contrary, diesters were not observed anymore in the residual solids after alkaline leaching, which were

attributed to their alkaline hydrolysis (Leinweber et al., 1997; McDowell and Stewart, 2006). Probably for the same reasons, the proportion of phosphonates was slightly decreased by alkaline leaching.

3.4. Effect of leaching conditions on P distribution within the solids

In the leaching experiments, phosphorus input was entirely present within the sludge solids. It was therefore initially distributed exclusively within the different pools identified by P fractionation of the sludge shown in Figure 3. Fe~P species were initially predominant (60.9% of total P content) and Ca,Al~P fraction was analyzed in second highest proportion (28% of total P). Alkali soluble organic P fraction ($P_{\text{org-alk}}$; 4.0% of total P), water-extractible fraction ($H_2O\sim P$; 1.5% of total P), and acid soluble organic fraction ($P_{\text{org-acid}}$; 0.8% of total P) were analyzed in smaller proportions. Detailed information on P speciation in the considered sludge material was published in a recent article by the authors (Kim et al., 2015c).

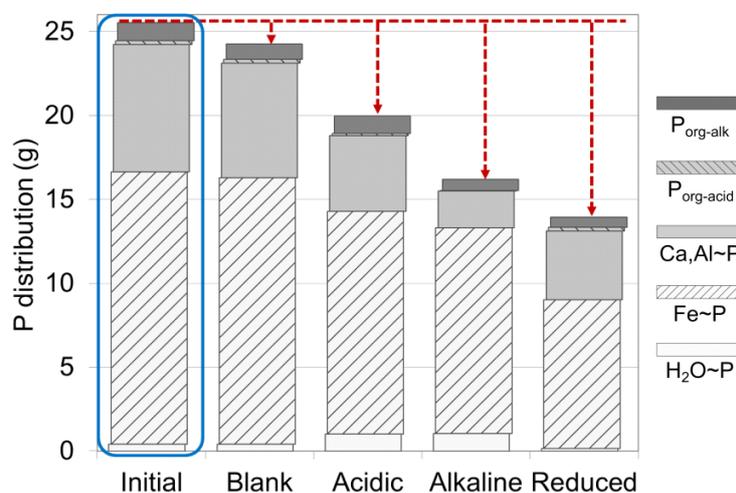


Figure 3. Distribution of P (in g) before leaching (initial sludge solids) and after leaching (residual solids + solution) in deionized water (blank), hydrochloric acid (acidic), sodium hydroxide (alkaline) and sodium ascorbate (reduced). The masses of P in the pools of the initial and residual solids were obtained by P fractionation protocol. Arrows indicate the cumulated masses of P leached into the solutions under the different conditions.

Figure 3 also shows the cumulated masses of phosphorus leached under the different conditions tested (arrows) and the distribution of phosphorus within the pools of the residual solids collected at the completion of the leaching sequences.

Phosphorus leaching in the blanks was very small as discussed above. Consequently, phosphorus distribution in the residual solids was very similar to that observed in the initial sludge solids.

The easily soluble P forms (fraction H₂O-P) were found to increase under acidic and alkaline conditions. This apparently surprising observation was attributed to the physical changes occurring within the sludge matrix due to physicochemical stresses from leaching sequences which probably eroded and weakened the sludge structure.

The masses of P in the Fe-P pool were found to decrease in all assays. The extent of the decreases was correlated with the cumulated masses of phosphorus leached. Hence, under reducing conditions where phosphorus leaching was the highest (close to 40%, Table 2), the Fe-P fraction was found to decrease by close to 45% as confirmed by the regular leaching of iron shown in Figure 1. This result showed that the reductive dissolution of ferric species and the concomitant release of adsorbed phosphates was a major process of phosphorus leaching. However, even though the Fe-P fraction was responsible in mass terms for most of phosphorus leached under reduced conditions, the proportion of P leached from this pool (close to 45%) was similar to the respective contribution of Ca,Al~P pool (46%) or even P_{orgalk} (46%). Leaching under reduced conditions therefore did not exclusively affect Fe-P fraction, partly due to the combined effects of the acidity of the reducing agent used.

As a matter of fact, under acidic conditions, the mass of P in Ca,Al~P fraction decreased by 41%, as confirmed by the regular leaching of calcium observed in Figure 3. A slight decrease of the mass of P in Fe~P fraction was also observed in Figure 3 (18%), suggesting that Fe-bound phosphorus within the initial sludge was at least partly in the form of phosphate

adsorbed onto iron hydroxides which were susceptible to dissolve under acidic conditions. However, the final pH of the acidic assays (below 4) was favorable for the re-precipitation of phosphates and ferric ions into iron phosphate (FePO_4), which would therefore be analyzed in the Fe~P fraction within the residual solids. This phenomenon may explain why the observed decrease of the mass of phosphorus in Fe~P fraction was small and the leaching of iron observed under acidic conditions (Figure 1) was low itself.

Finally, under alkaline conditions, the alkaline soluble organic P fraction decreased as expected. The mass of P also decreased within Fe~P fraction (24%) but a much stronger decrease of P was observed in Ca,Al~P fraction (close to 72%). A direct effect of alkaline conditions on the dissolution of P-bearing Fe or Ca species would be unlikely. Only very slight leaching of calcium and iron were indeed observed under alkaline conditions (Figure 1). The re-precipitation of these elements could not occur since the final pH (11.5) and their ranges of concentrations in the solutions were below the solubility product of Ca and Fe hydroxides.

Although calcium phosphate would be expected to precipitate at alkaline pH, Alvarez et al. (2004) highlighted that the presence of soil humic substances at a concentration around 2 g/L inhibited calcium phosphate precipitation. The humic substances content of the sludge studied here was analyzed at around 4% (Kania et al., 2015), thereby providing conditions that may induce this inhibition. Moreover, the observed leaching of Al during our assays (Figure 1) suggested a possible contribution of Al-P species dissolution. The strong decrease of Ca,Al~P fractions in the residual solids after leaching under alkaline conditions may therefore be attributed to these phenomena.

3.5. Factors of influence of P release or extraction

In the objective to estimate the possible correlations between the variables followed, experimental data were treated through PCA. Results showed that only the first two principal components had statistically significant Eigen values (> 1). These 2 axes accounted for 77.7% of the total variance. Figure 4 shows the position of the selected variables on the correlations circle (a) and a plot of individual data points (b) on the plane represented by the first two principal component axes.

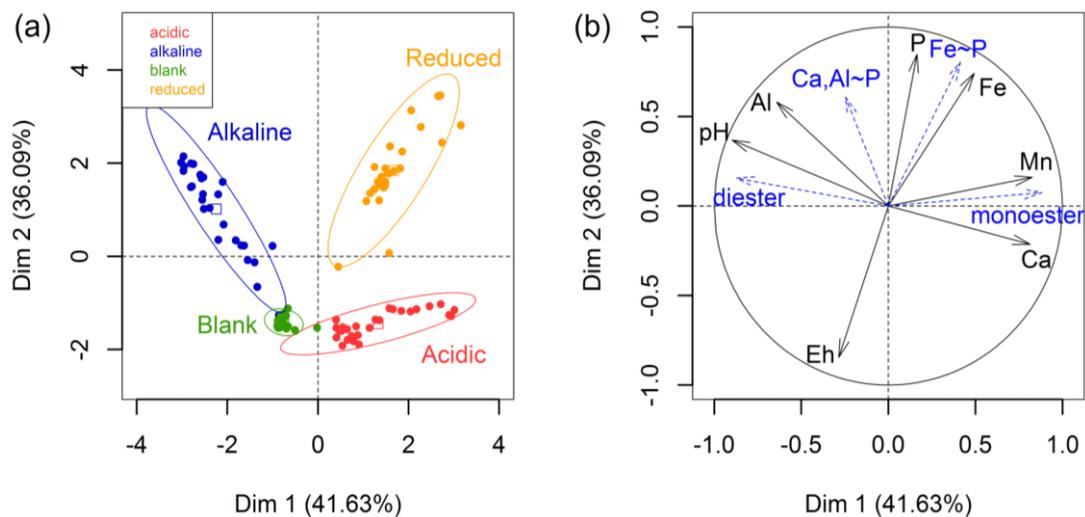


Figure 4. PCA plot of individual data (a) and variables plot on the correlation circle (b).

The plot of individual data points (Figure 4, a) showed that the different leaching conditions (blank, acidic, alkaline and reduced) resulted clearly in the identification of four statistically distinct groups. In a logical manner, the “acidic” and “alkaline” data groups were mainly identified by their correlation with pH whereas the “reduced” data group was mainly governed by redox conditions.

The first component, accounting for 41.6% of the total variance, was correlated primarily with pH, Ca, Al and Mn, confirming pH dependent releases of those elements (Al release under

alkaline pH; Ca and Mn releases under acidic pH). The second component, accounting for 36.1% of the variance, was correlated with Eh, Fe and P, confirming the strong Eh dependence of Fe release. The negative correlation between Eh and P release suggested that the reduction of P-bearing ferric species was a major mechanism of P release from the studied sludge.

The addition of four supplementary variables (Fe~P; Ca,Al~P; monoester; diester; Figure 4, b) into the correlation circle allowed to investigate the relationship between the dissolution of P bearing solid phases and releases of their constituents into the aqueous solution. Positions of those variables on the correlation circle indicated that the dissolution of solid phases may not be correctly appreciated by monitoring solely the transfer of their constitutive elements into the solution. Both solution and solid materials had to be analyzed to allow a relevant description of leaching phenomena.

4. Conclusions

The reversibility of P capture by FeCl₃ injection within the sludge deposit layer of an operating VFCW was assessed by studying adverse effects of pH and Eh stresses on a representative sludge sample taken from a VFCW plant treating mostly domestic wastewater according to the AZOE-NP[®] process. Series of 27 sequential batch leaching tests were conducted at a L/S ratio of 10 to evaluate the maximal extraction of P and other elements by submitting the sludge to a volume of water equivalent to 10 years of rain in the field (Vercia, France), and to extreme condition of pH and Eh depending on the leaching solutions used.

Chemical P fractionation on the sludge before and after leaching tests showed that inorganic phases were the major pool of P release under the tested conditions. Solution ³¹P NMR

confirmed that the organic P forms were not significantly affected, except under alkaline conditions. PCA results statistically validated experimental observations and highlighted the complementarity of solutions' and solids' analyses to effectively assess the leaching processes. Results from this study allowed to conclude that the risk of accidental release of phosphorus into the environment would be extremely low during the operation of the plant. They also suggested that phosphorus would be leached at a low but regular rate in case of land application of the dredged sludge, thereby potentially satisfying the needs of the crops without the risk of surface- or ground-waters contamination which might result from excessive phosphorus release.

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