

Mutations in the stereospecificity pocket and at the entrance of the active site of *Candida antarctica* lipase B enhancing enzyme enantioselectivity

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1 **Mutations in the stereospecificity pocket and at the entrance of**
2 **the active site of *Candida antarctica* lipase B enhancing enzyme**
3 **enantioselectivity**

4
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16
17
18 **Abstract**

19 Two different parts of *Candida antarctica* lipase B (stereospecificity pocket at the bottom of
20 the active site and hydrophobic tunnel leading to the active site) were redesigned by single-or
21 double-point mutations, in order to better control and improve enzyme enantioselectivity
22 toward secondary alcohols. Single point isosteric mutations of Ser47 and Thr42 situated in the
23 stereospecificity pocket gave rise to variants with doubled enantioselectivity toward pentan-2-
24 ol, in solid/gas reactor. Besides, the width and shape of the hydrophobic tunnel leading to the
25 active site was modified by producing the following single-point mutants: Ile189Ala,
26 Leu278Val and Ala282Leu. For each of these variants a significant modification of

1 enantioselectivity was observed compared to wild type enzyme, indicating that discrimination
2 of the enantiomers by the enzyme could also arise from their different accessibility from the
3 enzyme surface to the catalytic site.

4

5 **Keywords**

6 Lipase B from *Candida antarctica*; stereoselective catalysis; protein engineering;
7 stereospecificity pocket; substrate accessibility to the active site.

8

9

10 **1. Introduction**

11 It is now well-known that enantioselectivity of lipases can be modified by structural factors of
12 the substrates and of the enzyme. In this context, directed mutagenesis has been used a lot to
13 improve lipases enantioselectivity and to increase the understanding of how these enzymes
14 solve the task of distinguishing between enantiomers [1-4]. Especially, lipase B from *Candida*
15 *antarctica* (CALB) has been studied a lot by this way [5,6]. In this present work our goal has
16 been to better understand structural determinants of CALB enantioselectivity toward
17 secondary alcohols and to obtain more-selective mutants. Crystal structure analysis of CALB
18 reveals a rather narrow and deep active site with a small stereospecificity pocket flanked by
19 residues Thr42, Ser47 and Trp 104, which is at the origin of its enantioselectivity [7,8]. The
20 strong *R*-selectivity of CALB toward secondary alcohols is explained by the different binding
21 modes of the enantiomers in the stereospecificity pocket [9,10]. The fast-reacting enantiomer
22 positions its medium-sized substituent in the stereospecificity pocket and its large substituent
23 toward the active-site entrance. The slow-reacting enantiomer has to get into position in the
24 active site in the opposite way, in order to form a catalytically active binding mode, although
25 the largest substituent is not easily fitted in the stereospecificity pocket. This steric limitation

1 makes CALB very selective toward secondary alcohols with a medium substituent not larger
2 than an ethyl group and a large substituent bigger than an ethyl group.

3 Nevertheless, Rotticci *et al.* showed that steric interaction in the stereospecificity pocket is not
4 the only important factor for the discrimination between enantiomers. These authors produced
5 CALB variants with nearly isosteric mutations in the stereospecificity pocket (Thr42Val,
6 Ser47Ala and Thr42Val/ser47Ala) and obtained an increased enantioselectivity towards 1-
7 bromo-2-octanol and 1-chloro-2-octanol, when the single or double point mutation Ser47Ala
8 was used [6]. This effect was ascribed to a decrease of electrostatic repulsion between the
9 halogen atom of the substrate and the stereospecificity pocket induced by the mutation. In the
10 present paper, the effect of the same isosteric mutations in the stereospecificity pocket was
11 further studied in a solid-gas bioreactor, for the CALB-catalyzed resolution of linear
12 secondary alcohols (Scheme 1). Moreover, for these variants, variations in activation enthalpy
13 and entropy were measured, and the effect of water on selectivity was studied, in the same
14 type of reactor. This reactor includes the enzyme at a solid state forming a fixed bed, and
15 gaseous substrates and nitrogen as carrier gas percolating the enzymatic bed. The reactor
16 presents the advantage to independently control the thermodynamic activities of the
17 components present in the reaction medium by varying the partial pressure of each compound
18 in the carrier gas that percolates through the catalytic bed [11]. Moreover this type of reactor
19 permits the determination of the specific water addition in the medium, by adding water in the
20 gaseous stream at a controlled thermodynamic activity [12, 13].

21 Finally, we focused on the narrow and short tunnel leading to CALB active site, and bounded
22 by four hydrophobic residues, Ile189, Leu278, Ala282 and Ile285. As substrates have to cross
23 this tunnel before the formation of the second tetrahedral intermediate, we tested the ability of
24 these hydrophobic amino acids to play a role in the discrimination between enantiomers. By
25 using different single point mutations, the access of linear secondary alcohols to the active site

1 was made more difficult by increasing the volume of Ala282 at the active site entrance
2 (variant Ala282Leu), or easier by decreasing the volume of Leu278 or Ile189 (variants
3 Leu278Val and Ile189Ala). The only study mentioning the discrimination of the enantiomers
4 through their channeling from the enzyme surface to the catalytic triad, concerns
5 *Burkholderia cepacia* lipase in the case of the resolution of (*R,S*)-2-bromophenyl acetic acid
6 ethyl ester by transesterification in the presence of octanol [14]. Such an investigation has
7 never been performed in case of CALB lipase and is presented in this paper, by studying the
8 enantioselectivity of CALB variants with single-point mutations at the entrance of the active
9 site crevice.

10

11 **2. Experimental**

12 **2.1 Chemicals**

13

14 Substrates and other chemicals were purchased from Sigma-Aldrich-Fluka Chemical Co.
15 They were of the highest purity available (98% minimum) and checked by gas
16 chromatography before use. Substrates were dried by distillation under argon prior to use and
17 stored under argon atmosphere and over molecular sieves. Solvents were purchased from
18 Carlo Erba. 1-methylpropylpropanoate, 1-methylbutylpropanoate, 1-methylpentylpropanoate,
19 1-ethylpentylpropanoate and 1-ethylhexylpropanoate were synthesized from corresponding
20 alcohol and propanoic acid in presence of *p*-toluene sulfonic acid and molecular sieves [15].
21 1-methylhexylpropanoate and 1-methylheptylpropanoate were obtained from the
22 corresponding alcohol and propanoic anhydride in pyridine at room temperature [16]. High
23 purity water was obtained with the Milli Q system (Millipore, France).

24

25 **2.2 Enzyme**

26

27 **2.2.1 Wild-type (WT) and stereospecific pocket mutants**

1 WT CALB and stereospecificity pocket variants (Ser47Ala, Thr42Val and Ser47Ala-
2 Thr42Val) were produced in the methylotropic yeast *Pichia pastoris*. These lipases were
3 expressed extracellularly and purified from the medium by hydrophobic interaction
4 chromatography, followed by gel filtration [6, 17].

5 Enzyme adsorption was performed onto 60–80 mesh Chromosorb P AW DMCS (Acid
6 Washed DiMethylChloroSilanized) (Varian, France). In a typical adsorption procedure for
7 solid/gas catalysis, WT enzyme (0.106 mg) or variant (0.85 mg) was dissolved in sodium
8 phosphate buffer (pH 7.5, 10 mM), and dry Chromosorb P AW DMCS (1g) was added to the
9 solution. The amount of immobilized enzyme was determined by measuring absorbance at
10 280 nm, by taking a molar extinction coefficient equal to 40690 M⁻¹ cm⁻¹. After vigorous
11 shaking, the preparation was left for one week under vacuum and over P₂O₅ at room
12 temperature.

13 **2.2.2 Mutants of the entrance crevice**

14 The three mutants Ile189Ala, Leu278Val and Ala282Leu were obtained through site-directed
15 mutagenesis [18]. The pGAPZα B plasmid was amplified by polymerase chain reaction
16 (PCR) with the mutagenic primers designed to be non-overlapping at the ends [19]. The
17 product of PCR was analysed on 1% agarose gel and was used to transform *E.coli* XL-1 blue
18 cells. Colonies containing the wished mutation were identified by sequencing and used to
19 produce the plasmid in large quantity (i.e.>20μg) using QIAPrep Spin miniprep (QIAGEN
20 GmbH, Hilden, Germany). Plasmids were linearized and used to transform *Pichia pastoris*
21 SMD 1168 H by electroporation. Transformants of *Pichia pastoris* were cultured in order to
22 produce CALB mutants. The enzyme contained in the supernatant was purified on Butyl
23 Sepharose FF column. Purified protein fractions were subjected to buffer exchange to 50 mM
24 potassium phosphate buffer, pH 7.5 and freeze dried.

1 0.53 mg enzyme was dissolved in 1mL of milliQ water, and dry Chromosorb P AW DMCS
2 (1g) was added to the solution. After vigorous shaking, the preparation was left for one week
3 under vacuum and over P₂O₅ at room temperature.

4 **2.3 Experimental setup for solid–gas catalysis**

5
6 The bioreactor used in this study has already been described in a previous publication [20].
7 Thermodynamic experiments were run from 308 to 363 K with immobilized enzyme
8 preparation (10 mg to 200 mg). All other experiments were made at 318 K. The total flow rate
9 passing through the reactor was set to 800 $\mu\text{mol}\cdot\text{min}^{-1}$ for thermodynamic study, 1000
10 $\mu\text{mol}\cdot\text{min}^{-1}$ for entrance crevice mutations effects study and 1150 $\mu\text{mol}\cdot\text{min}^{-1}$ for water effects
11 study. Methyl propanoate and secondary alcohols thermodynamic activities in the solid/gas
12 reactor were fixed at 0.1 and 0.05, respectively, except for experiments with heptan-3-ol and
13 octan-3-ol, for which thermodynamic activities for alcohol was fixed at 0.1 and 0.2 for methyl
14 propanoate.

15 **2.4 Chromatographic assays**

16
17 For the solid/gas system analyses, the vapor phase leaving the bioreactor was sampled by
18 using a loop (0.25 mL) on a six-way valve (Valco) maintained at 175 °C or 220 °C. Samples
19 were automatically injected into the split injector of a gas chromatograph (Agilent 6890N)
20 (split 2:1) equipped with a flame ionization detector (FID) for the detection of all products.
21 The column was a Chirasil-DEX CB (25 m, 0.25 mm i.d, 0.25 μm β -cyclodextrin;
22 Chrompack, France). Separation methods depended on the substrate used. Column
23 temperature was 55 °C for 20 min (1.5 mL/min, 180 °C, 180 °C) with butan-2-ol, 55 °C for
24 15 min, 3 °C/min, 85 °C (1.5 mL/min, 180 °C, 180 °C) with pentan-2-ol, or 50 °C for 30 min
25 (1.8 ml/min, 250 °C, 250 °C) when water was added, 60 °C for 15 min, 3 °C/min, 90 °C for 5
26 min (2 mL/min, 180 °C, 180 °C) with hexan-2-ol and hexan-3-ol, 60 °C for 15 min, 3 °C/min,

1 90 °C for 5 min (2 mL/min, 200 °C, 200 °C) with heptan-2-ol and heptan-3-ol, 70 °C for 15
2 min, 3 °C/min, 100 °C for 5 min (2 mL/min, 200 °C, 200 °C) with octan-2-ol, and 90 °C for 17
3 min (2 mL/min, 200 °C, 200 °C) for octan-3-ol. Nitrogen was used as the carrier gas with a
4 constant flow rate (between 1.5 and 2 mL.min⁻¹ according to experiments). Hydrogen and air
5 were supplied to the FID at 35 and 350 mL.min⁻¹ respectively. Quantitative data were
6 obtained by integration on an Agilent 3396 Series III integrator or on GC Chemstation Rev
7 B.03.02. The external calibrations of the substrates were carried out by programming a range
8 of their partial pressures in the bioreactor, and by gas chromatography analysis. For the
9 products (methanol and chiral esters), an internal calibration was carried out by using methyl
10 propanoate and corresponding alcohol as internal standards. For accurate determination of *E*-
11 values the vapour phase leaving the bioreactor was condensed. Elution peaks of (*R*)-pentan-2-
12 ol and (*S*)-pentan-2-ol were identified by using commercial pure (*R*)-pentan-2-ol. Retention
13 time of the reaction products was obtained by adding chemically synthesized esters (*R*, *S*), as
14 described in paragraph 2.1. The absolute configuration of the reaction products was
15 established by esterification of commercial pure alcohol enantiomers.

16 **2.5 Determination of the enantiomeric ratio**

17
18 The enantiomeric ratio, *E*, was calculated from *ee_s* and *ee_p* [21,22]. The *E* values were based
19 on the average of three measurements for each condition; in all cases the standard deviation
20 was less than 10%. The equations $RT \ln E = -\Delta_{R-S}\Delta G^*$ and $\Delta_{R-S}\Delta G^* = \Delta_{R-S}\Delta H^* - T\Delta_{R-S}\Delta S^*$
21 were used to calculate enthalpic and entropic components of *E* [23,24].

22 **2.6. Shape of active site and entrance tunnel**

23 The crystal structure of native CALB was obtained from the Protein Data Bank [25] (PDB
24 entry: 1TCA; resolution 1.55 Å [26]). All molecular modeling visualizations, mutations and
25 calculations were performed with CHARMM force field [27] using the program-package

1 Discovery Studio (DS) version 1.6 or 2.1 (Accelrys Inc.). Systematic rotamer search was
2 performed on crystal structure using Richardson rotamer library [28]. Conolly surface [29] of
3 the sum of all possible rotamers for a residue was calculated at a probe radius of 1.4 Å.
4 Structure of second tetrahedral intermediate formed (*R*)-heptan-3-ol: complex enzyme-(*R*)-1-
5 ethylpentylpropanoate shown in Figure 4, was obtained from a structure previously obtained
6 with (*S*)-hexan-3-ol: complex enzyme-(*S*)-1-ethylbutylpropanoate, by manually modifications
7 followed by Steepest Descent minimizations ($\text{RMSD} \leq 0.01$). This last structure was obtained
8 from crystal structure of CALB by manual construction, after removal of crystallographic
9 water molecules, followed by MD simulation carried out by adopting a 12 Å non-bound
10 spherical cut-off, using the isothermal-isochoric ensemble (NVT) [30] and without solvents.
11 The Standard Dynamics Cascade Protocol of DS 1.6 was used with a time step of 1 fs. The
12 complex was energy minimized (2500 Steepest descent and 2500 adopted Basis NR) and
13 temperature of the system was raised from 50 to 400 K over 2 ps. After a short equilibration
14 (1 ps) a production run was carried out for 10 ps at 400 K. Backbone atoms were kept fixed
15 during simulation, in order to keep the original structural integrity.

16

1 3. Results and Discussion

2 3.1 The effect of two point mutations in the stereospecificity pocket of CALB

3
4 Acyl-transfer reaction rates, in case of catalysis by native or modified CALB, were
5 experimentally measured with methyl propanoate as acyl donor and pentan-2-ol as racemic
6 nucleophile, in a continuous solid gas reactor at 318 K, at water activity close to zero. The
7 acyl-transfer activity decreased by a factor 3 and 4, when using respectively Ser47Ala and
8 Thr42Val mutants, compared to WT CALB, for which a value of 30 μmol of product min^{-1}
9 and mg^{-1} of pure and free enzyme was found. It decreased by a factor 13 when using double
10 mutant Ser47Ala-Thr42Val. These results can be compared with the ones obtained by Rotticci
11 *et al.* with tributyrin as substrate in aqueous media. In this case activity decreased by a factor
12 1.8, 6 and 3 compared to WT CALB when using respectively Ser47Ala, Thr42Val and
13 Ser47Ala-Thr42Val mutants [6].

14 The enantioselectivity of these different enzymes toward pentan-2-ol was also measured in
15 the solid gas reactor. The *E*-value significantly increased when using Ser47Ala and Thr42Val
16 mutants, reaching respectively 198 and 236, instead of 100 with WT enzyme [12]. On the
17 other hand, double mutant Ser47Ala-Thr42Val presented an enantioselectivity toward pentan-
18 2-ol very close to that of the WT enzyme. Furthermore, a thermodynamic analysis of
19 enantioselectivity was performed by considering both activation enthalpy and entropy (Table
20 1). It can be seen that the preferred enantiomer, (*R*)-pentan-2-ol, was favored by enthalpy ($\Delta_{R-S}\Delta H^* < 0$) but disfavored by entropy ($\Delta_{R-S}\Delta S^* < 0$) in case of WT enzyme and Ser47Ala variant.
21
22 In case of Thr42Val and Ser47Ala-Thr42Val variants, enantioselectivity is favored both by
23 enthalpy and entropy.

24 These different results show that single- or double-point mutations in the stereospecificity
25 pocket have very significant effects both on activity and selectivity of CALB. Moreover,
26 thermodynamic analysis reveals that these effects are complex, as mutations lead to

1 modifications of both activation enthalpy and entropy differences between *R*- and *S*-
2 enantiomers. The importance of the entropy contribution in defining enantioselectivity is once
3 again shown [31], but the exact molecular mechanism of how entropy and flexibility are
4 related to enzyme activity and selectivity is still elusive.

5 In the present study, the effects of mutations cannot be ascribed to modifications of the size of
6 the stereospecificity pocket, as threonine and valine on the one side are isosteric, and serine
7 and alanine on the other side, have similar molar volumes. Therefore the mutations did not
8 lead to major changes in the size of the stereospecificity pocket (Figure 1). Counterexample
9 can be found in literature in which mutation effects is directly related to a change of the
10 stereospecificity pocket size: the Trp104Ala mutation of CALB increases very significantly
11 the size of the stereospecificity pocket, and this redesigned pocket was found to accommodate
12 much larger groups than the WT enzyme. This change transformed the strongly *R*-selective
13 WT CALB into an *S*-selective mutant for the resolution of 1-phenylethanol [32].

14 Besides, the change in hydrophobicity of the stereospecificity pocket of CALB caused by the
15 Thr42Val and the Ser47Ala mutations cannot either be considered as a reason for variation of
16 enantioselectivity toward pentan-2-ol, as it was shown that substrate selectivity of these
17 mutants was very similar to that of the WT enzyme, for the acylation of symmetrical
18 secondary alcohols (propan-2-ol, pentan-3-ol and heptan-4-ol) [18]. On the contrary, in case
19 of resolution of halohydrins through CALB-catalyzed acylation, the change of hydrophobicity
20 promoted by Thr42Val and Ser47Ala mutations affected the CALB selectivity toward the
21 halodrin, by decreasing electrostatic repulsion between substrates and enzyme [6].

22 The increase of enantioselectivity toward pentan-2-ol afforded by single-point mutations
23 (Thr42Val and Ser47Ala) may originate from modification of the H-bond network between
24 amino acid residues in the vicinity of the stereospecificity pocket, and different spatial
25 degrees of freedom and mobility of these residues and also of substrate substituent positioned

1 in this pocket. The higher E -value for the single-point mutants compared to that of the WT
2 could arise from higher mobility of the non mutated residue, which now is not hydrogen
3 bonded and has more mobility. If it moves more than in the WT, the specificity pocket will be
4 smaller, then the S -enantiomer will not fit so well and E goes up. The H-bond network
5 progress and mobility of the residues located in the stereospecificity pocket, when using
6 variants Thr42Val and Ser47Ala, compared to WT CALB, will be the topic of future studies
7 by molecular dynamics simulation.

8
9

10 **3.2 The effect of water on enantioselectivity of the stereospecificity pocket variants.**

11
12 The effect of water thermodynamic activity (a_w) on WT CALB-catalyzed acyl-transfer
13 reaction was previously extensively studied by using a solid/gas reactor [12,13]. It was shown
14 that water had a pronounced effect on CALB activity and enantioselectivity. E -value
15 increased from 100 at water activity close to zero, in case of the resolution of pentan-2-ol at
16 318 K, to a maximum of 320, at water activity of 0.2. In this part, we were intending to
17 compare the effect of water on the enantioselectivity toward pentan-2-ol at 318 K, of
18 Thr42Val, Ser47Ala and Thr42Val-Ser47Ala variants with the results previously obtained for
19 the WT enzyme. Results are shown in Figure 2. First, enantioselectivity increases in all cases
20 up to $a_w = 0.1$ for Ser47Ala, $a_w = 0.08$ for Thr42Val, and $a_w = 0.3$ for Thr42Val-Ser47Ala to
21 reach the respective following maxima: $E = 449$, 567 and 289. Above a_w values
22 corresponding to maximal E , enantioselectivity decreases, after a plateau for Ser47Ala variant
23 in the range of a_w from 0.1 to 0.3.

24 So it appears that the enhancement of E promoted by water is significantly reduced in case of
25 double-point variant compared to WT enzyme, whereas it is strongly enhanced in case of
26 single-point variants.

1 Previously, the increase of enantioselectivity of WT CALB for pentan-2-ol at low water
2 activity, has been interpreted by specifically binding of a water molecule in the
3 stereospecificity pocket, leading to a reduction of its size and thereby rendering binding of *S*-
4 enantiomer still more difficult compared to the *R*-form [12]. The result obtained with double-
5 point variant Thr42Val-Ser47Ala showing a significant decrease of water effect, is in
6 accordance with this hypothesis, as the redesigned stereospecificity pocket is not water-
7 attractive anymore, because both mutations lead to the replacement of a polar hydroxyl group
8 by an apolar substituent (methyl or hydrogen atom). This fact can be confirmed by calculating
9 the water dissociation constant with specificity pocket (K_d), thanks to a Hill equation analysis
10 for $a_w < 0.1$. K_d for Thr42Val-Ser47Ala variant was found to be equal to 0.1, which is much
11 higher than the value of 0.03 previously obtained with WT enzyme. The Hill coefficient for
12 double mutant ($h = 0.55$) was consistent with one water molecule bound in the
13 stereospecificity pocket as for the WT enzyme [12]. Finally, the fact that maximal *E*-value
14 was obtained at higher water activity for Thr42Val-Ser47Ala variant compared to WT is also
15 consistent with a lower affinity for water.

16 On the contrary, in case of single-point variants, the Hill equation analysis gave much lower
17 K_d values than that obtained for the WT enzyme: $K_d = 0.004$ and 0.0002 for Ser47Ala and
18 Thr42Val respectively, with Hill coefficients being respectively equal to 1.45 and 2.24. These
19 results are in agreement with a high affinity for water for the stereospecificity pocket of
20 single-point variants, leading to a strong enhancement of *E*-value when going from water
21 activity close to zero to 0.1 for Ser47Ala or 0.08 for Thr42Val.

22 These very high *E*-values are promoted by water at very low activity, where water acts mainly
23 through specifically located effects. At higher water activity, entropic effects coming from
24 water molecules have been shown to become predominant and can lead to a decrease of *E*
25 [13].

1

2 **3.3 The effect of mutations at the entrance of the active site of CALB**

3

4 The crystallographic structure of CALB shows that the substrate binding pocket is an
5 elliptical, steep funnel. The pocket constitution is similar to other lipases with a hydrophilic
6 bottom of the funnel, which is formed by the catalytic Ser105 and residues of the oxyanion
7 hole, Thr40 and Gln106 [33]. Near the surface the pocket is bordered by hydrophobic
8 residues, which interact with organic solvents, but unlike to lipases, CALB have no lid and is
9 not submitted to interfacial activation. The active site crevice can be partitioned into two
10 sides, an acyl side and an alcohol side, where the corresponding parts of the substrate will be
11 located during catalysis [7]. The alcohol side is bounded by four hydrophobic residues,
12 Ile189, Leu278, Ala282 and Ile285. They form a narrow tunnel, which must be crossed by
13 substrate before the formation of the second tetrahedral intermediate. The volumes occupied
14 by these residues were calculated by systematic conformational search. This study showed
15 that these residues have side chains free to rotate and have no steric interaction with
16 neighboring residues (Fig 3a). Consequently a concerted motion of the side chains cannot be
17 observed, whereas such was the case with *Burkholderia cepacia* lipase, when the substrate
18 passes though the bottleneck formed by Val266 and Leu17 [14]. However figure 3a suggests
19 that these four hydrophobic residues, Ile189, Leu278, Ala282 and Ile285 can affect substrate
20 trajectory. Mutations were considered in order to understand the role of these residues in the
21 selectivity of CALB. To avoid electrostatic effects, only mutations in another aliphatic
22 hydrophobic residue were considered. The volume of residue 282 was expended by changing
23 native Ala into Leu, and volumes of positions 278 and 189 were decreased by respectively
24 changing Leu into Val and Ile into Ala (Figure 3b-d). Ile189Ala, Leu278Val and Ala282Leu
25 mutants were produced and their activity and enantioselectivity toward different chiral
26 alcohols tested in solid/gas bioreactor.

1 Acyl-transfer reaction rates were experimentally measured with methyl propanoate as acyl
2 donor and eight different linear secondary alcohols ranging from four to eight carbons as
3 racemic nucleophile, at 318 K and water activity close to zero. Results showed that activity
4 increased by a factor 1.1 to 2, when using Ala282Leu variant compared to WT CALB, except
5 for heptan-3-ol, for which it decreased; activity was similar or slightly increased when using
6 Leu278Val variant, except for hexan-3-ol and heptan-3-ol, for which it decreased, and activity
7 decreased with Ile189Ala mutant. As an example, in case of hexan-2-ol as nucleophile, a
8 value of 13.5 μmol of product. min^{-1} mg^{-1} of pure and free enzyme was found with WT
9 CALB, which was doubled to 26.1 with Ala282Leu, increased to 18.2 with Leu278Val and
10 decreased to 10.8 with Ile189Ala.

11 Table 2 shows the results for the enantiomeric ratio E for acylation of the eight chiral alcohols
12 (butan-2-ol, pentan-2-ol, hexan-2-ol, heptan-2-ol, octan-2-ol, hexan-3-ol, heptan-3-ol and
13 octan-3-ol), catalyzed by WT CALB or the different variants of CALB with mutations at the
14 entrance of the active site crevice. It appears that all mutations have a significant effect on E ,
15 promoting an increase of E up to a factor 1.5 or a decrease up to a factor 2.2. Nevertheless the
16 upward or downward trend for E is variable for each mutation, depending on the length of the
17 large substituent of the chiral alcohol. So another time, the analysis of the results for E
18 variations is complex and needs to consider many different structural causes.

19 To analyze these data, the following postulate can be formulated: modifications of the
20 entrance crevice to the active site by mutations of amino acids, will predominantly affect the
21 trajectory of the S -enantiomer, as this substrate has to enter the active site with its large
22 substituent in the front, in order to react. Indeed, according to the model proposed by Haeffner
23 *et al.*, the S -enantiomer has to put its large substituent into the stereospecificity pocket,
24 located at the bottom of the active site, in order to form catalytically active second tetrahedral
25 intermediate [9]. This last covalent intermediate is considered to be a good model of the

1 transition state intermediate in case of CALB-catalyzed resolution of chiral nucleophile by
2 acyl-transfer [9].

3 As shown in the following discussion, this postulate provides a rational explanation for most
4 results. In some cases however, other structural explanations have to be given, like a
5 modification of the *R*-enantiomer trajectory also, an effect of the mutation on the structure and
6 energy state of second tetrahedral intermediate or an interaction between acyl part and residue
7 189, which are very close together. In any case, a complete structural explanation of the effect
8 of mutations on *E* very likely involves a combination of these different individual effects.
9 Molecular dynamics simulations are currently in progress to study substrate trajectory in WT
10 and variants of CALB.

11 According to the above postulate, Leu278Val mutant is supposed to make access of *S*-
12 enantiomer to the active site easier, because of the smaller size of Val compared to Leu. This
13 hypothesis is verified (lower *E*-value than WT enzyme) for hexan-2-ol, heptan-2-ol, octan-2-
14 ol, heptan-3-ol and octan-3-ol. So, if the large substituent of linear secondary alcohol includes
15 four or more C, then the postulate is verified. If the large substituent includes three C or less,
16 the observed effect is not in agreement with the postulate: higher *E*-value than WT enzyme
17 (butan-2-ol and hexan-3-ol) or similar *E*-value (pentan-2-ol). In these cases, an easier access
18 to the active site for the *R*-enantiomer could explain the effect.

19 Ile189Ala mutant is also supposed to ease access of *S*-enantiomer to the active site. This
20 hypothesis is verified (lower *E*-value than WT enzyme) for all alcohols except butan-2-ol. So
21 if the large substituent of linear secondary alcohol includes three or more C, then the postulate
22 is verified.

23 For the Ala282Leu variant, our postulate implies a more difficult access of the *S*-enantiomer
24 to the active site, because of the steric hindrance of the crevice caused by Leu282 replacing
25 Ala. The postulate is verified (higher *E*-value than WT enzyme) for butan-2-ol, pentan-2-ol

1 and hexan-3-ol. So, if the large substituent of linear secondary alcohol includes three C or
2 less, then the postulate is verified. If the large substituent includes more than three C (hexan-
3 2-ol, heptan-2-ol, octan-2-ol, heptan-3-ol, octan-3-ol), the observed effect is opposite to the
4 one predicted by the postulate: lower *E*-value than WT enzyme. In these cases, a more
5 difficult access to the active site for the *R*-enantiomer could explain the effect or also an
6 interaction between the large substituent of the *R*-enantiomer and the residue Leu282, during
7 formation of the second tetrahedral intermediate. Figure 4 shows an example of such an
8 interaction in case of the second tetrahedral intermediate formed with (*R*)-heptan-3-ol: (*R*)-1-
9 ethylpentylpropanoate covalently attached to Ser105, with WT or Ala282Leu CALB.

10

11

12 **4. Conclusion**

13 In this paper, thanks to single-point mutations, new amino acid positions important for CALB
14 enantioselectivity were identified at the entrance of the active site and offered altered
15 enantioselectivity. Moreover, the crucial role of residues situated in the stereospecificity
16 pocket was confirmed and resolution of pentan-2-ol could be improved, by using variants of
17 this pocket. The present investigation is a new demonstration of the possibilities offered by
18 protein redesign, for improving kinetic resolution using CALB. It also highlights that entropic
19 phenomena, including enzyme and substrate flexibility, are crucial for understanding
20 molecular mechanism of enzyme activity and selectivity, and remain to be accurately
21 elucidated.

22

23 ***Acknowledgements***

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2 *esters.*

3

4 **5. References**

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Table 1

Thermodynamic components of the enantiomeric ratio, E , and activity A (μmol of product $\text{min}^{-1} \text{mg}^{-1}$ of free enzyme) for acylation of pentan-2-ol with methyl propanoate as acyl donor, catalyzed by WT *Candida antarctica* lipase B and the three variants of the stereospecificity pocket. All experiments were performed at a_w equal to 0 in the gaseous stream of the solid/gas reactor. For measurements of both E - and A - values the thermodynamic activity of methyl propanoate and alcohol were fixed at 0.1 and 0.05, respectively.

Type of enzyme	A	E at 318 K ^[a]	$\Delta_{R-S}\Delta G^*$ at 318 K [kJ mol ⁻¹] ^[b]	$\Delta_{R-S}\Delta H^*$ [kJ mol ⁻¹] ^[c]	$-T\Delta_{R-S}\Delta S^*$ [kJ mol ⁻¹] ^[d]	$\Delta_{R-S}\Delta S^*$ [J K ⁻¹ mol ⁻¹] ^[c]	N ^[e]
WT ^[f]	30	101±5	-12.2±0.1	-18±1.2	6±1.3	-19±4	18
Thr42Val	7.5	236±6	-14.4±0.2	-11±0.7	-3±0.5	11±2	16
Ser47Ala	10	198±19	-14±0.6	-20±1.0	6±0.9	-19±3	11
Thr42Val- Ser47Ala	2.3	98±13	-12.1±0.4	-11±0.7	-1±0.6	3±2	13

^[a] Standard deviations were calculated with three values of E , and correspond to 95% confidence intervals.

^[b] Values calculated from $-RTL\ln E$. Standard deviations were obtained from the standard deviations of E .

^[c] $\Delta_{R-S}\Delta H^*$ and $\Delta_{R-S}\Delta S^*$ were calculated by using the least square method, with between 11 and 18 measurement of E as a function of T . Standard deviations correspond to 99% confidence intervals.

^[d] $-T\Delta_{R-S}\Delta S^*$ values were calculated at 318 K. Standard deviations were obtained from standard deviations of $\Delta_{R-S}\Delta S^*$.

^[e] Total number of E measurements.

^[f] Results of this line have already been published in a previous publication [12].

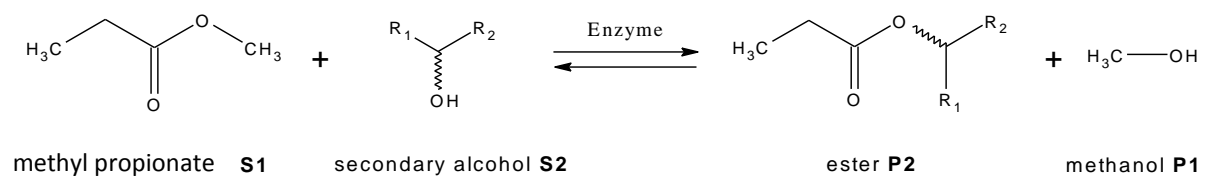
Table 2

Enantiomeric ratio $E^{[a]}$ and activity A (μmol of product min^{-1} and mg^{-1} of free enzyme) for acylation of chiral secondary alcohols with methyl propanoate as acyl donor at 318 K, in solid/gas reactor, catalyzed by WT *Candida antarctica* lipase B and variants of CALB obtained by mutations of amino acids at the entrance of the active site crevice. All experiments were performed at a_w equal to 0 in the gaseous stream of the solid/gas reactor. Thermodynamic activity of methyl propanoate and alcohol were fixed at 0.1 and 0.05, respectively, except for heptan-3-ol and octan-3-ol (0.2 and 0.1), for measurements of both E - and A -values.

Structure	WT		A282L		L278V		I189A	
	E	A	E	A	E	A	E	A
	4.6±0.1	n.d.	4.9±0.1	20.7	5.8 ±0.2	20.5	6.9±0.4	12.3
	100±6	17.7	130±1	22.6	98±8	29.1	87±1	12.6
	297±9	13.5	228±8	26.1	216±7	18.2	211±6	10.8
	476±30	11.7	337±2	20.9	299±18	11.1	437±25	4.8
	435±18	6.7	298±11	11.1	227±3	7.2	400±10	4.3
	78±5	8.7	103±8	9.8	83±3	6.4	45±3	2.8
	318±15	2.4	208±14	2.2	203±15	1.8	192±7	1.7
	523±31	3.1	245±15	4.1	210±15	3.3	199±14	2.0

^[a] Standard deviations were calculated with 5 to 12 values of E from two condensates, and correspond to 95% confidence intervals.

Scheme 1



Legends to Figures

1

2 **Figure 1:** Stereospecificity pocket of CALB; a) WT; b) double mutant Thr42ValSer47Ala.

3 **Figure 2:** Influence of thermodynamic water activity, a_w on enantioselectivity at 318 K for
4 the acylation of pentan-2-ol with methyl propanoate as acyl donor, catalyzed by WT CALB
5 and three variants Thr42Val, Ser47Ala and Thr42ValSer47Ala, in continuous solid/gas
6 reactor.

7 **Figure 3:** Shape of the entrance tunnel. a) WT with occupancy of hydrophobic residues:
8 Ile189, Leu278, Ala282 and Ile285 (surface of all rotamers); b) mutant Ala282Leu; c) mutant
9 Leu278Val; d) mutant Ile189Ala; occupancy of mutants in black grid.

10 **Figure 4:** Structure of WT and Ala282Leu CALB with (*R*)-1-ethylpentylpropanoate
11 covalently attached to Ser105 (second tetrahedral intermediate formed with (*R*)-heptan-3-ol),
12 obtained by MD simulation described in section 2.6; occupancy of mutants in black grid, (*R*)-
13 1-ethylpentylpropanoate in dark grey.

14

1

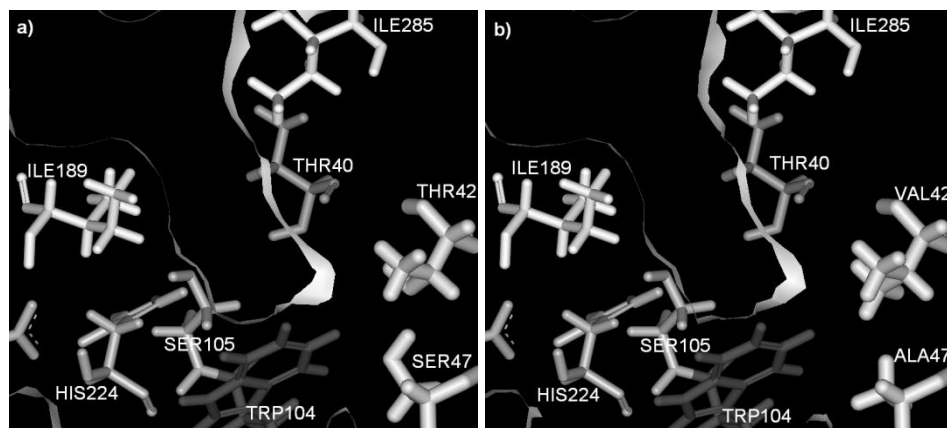
Figure 1 :

Figure 2 :

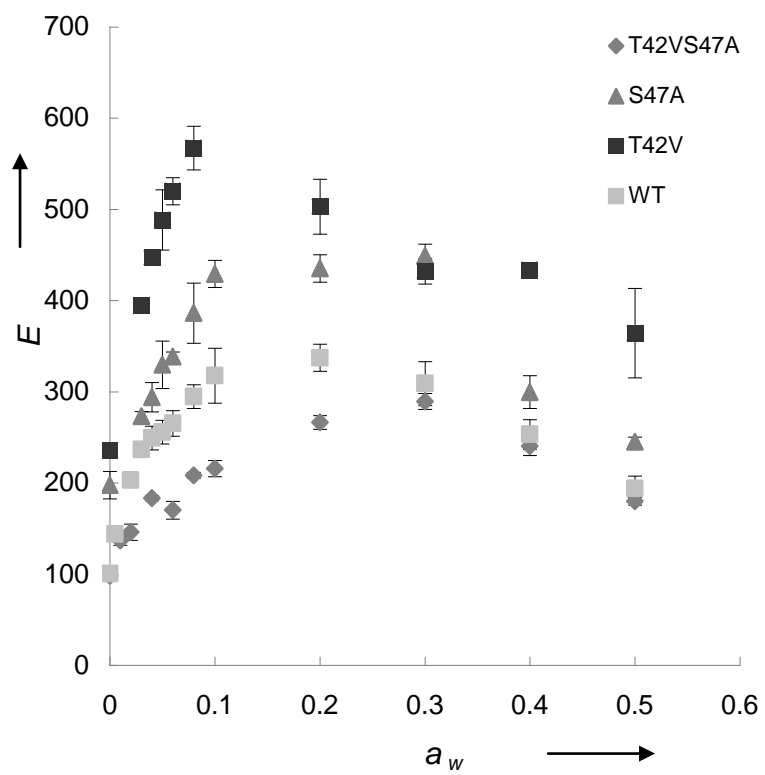
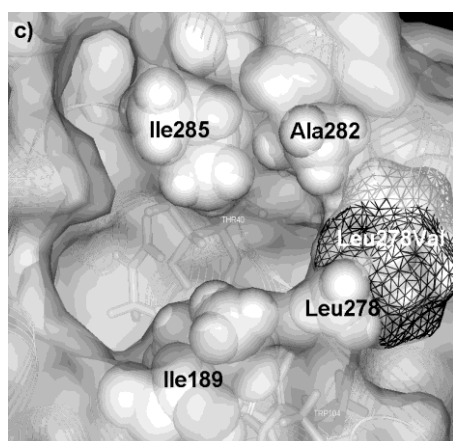
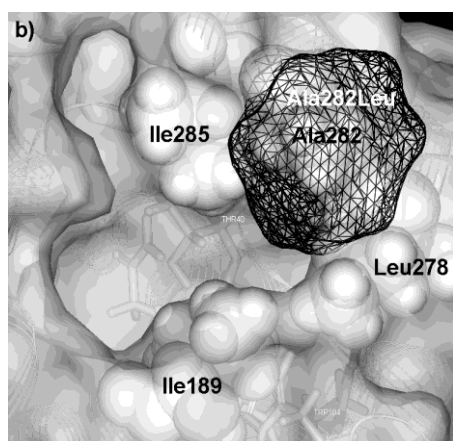
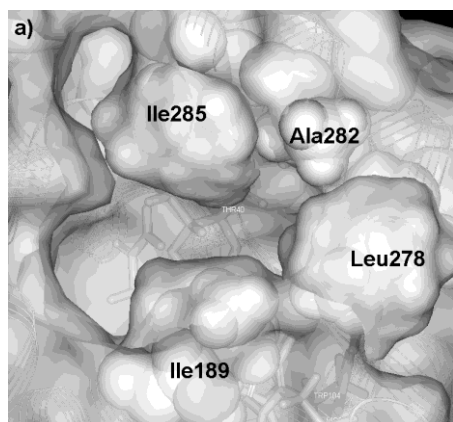


Figure 3



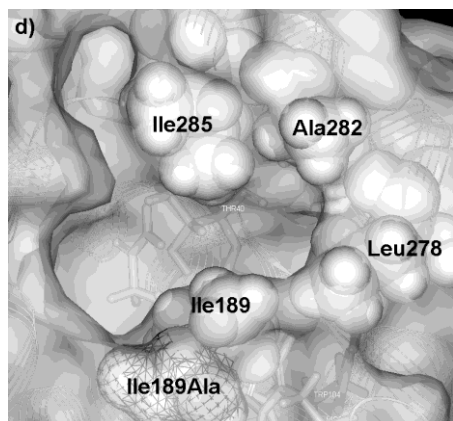


Figure 4 :