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► **To cite this version:**

H. J. Beine, F. Dominé, A. Ianniello, M. Nardino, I. Allegrini, et al.. Fluxes of nitrates between snow surfaces and the atmosphere in the European high arctic. *Atmospheric Chemistry and Physics Discussions*, 2006, 3 (1), pp.75-106. hal-00327794

HAL Id: hal-00327794

<https://hal.science/hal-00327794>

Submitted on 18 Jun 2008

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Fluxes of nitrates

H. J. Beine et al.

Fluxes of nitrates between snow surfaces and the atmosphere in the European high arctic

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Received: 24 October 2002 – Accepted: 4 December 2002 – Published: 6 January 2003

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Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

© EGU 2003

Abstract

Measurements of atmospheric and snow mixing ratios of nitrates and nitrites and their fluxes above the snow surface were made during two intensive campaigns during spring time 2001 at Ny-Alesund, Svalbard as part of the EU project “The Nitrogen Cycle and Effects on the oxidation of atmospheric trace species at high latitudes” (NICE).

At this coastal site close to the unseasonably unfrozen fjord of the measured nitrogen species only HNO_3 showed a significant flux above the snow surface; a mean deposition of $-8.7 \text{ nmol h}^{-1} \text{ m}^{-2}$ was observed in late April/early May 2001. These fluxes may be due to the reaction of HNO_3 with sea salt, and especially NaCl , or may be simply uptake of HNO_3 by ice, which is alkaline because of the sea salt in our marine environment. During snowfall periods dry deposition of HNO_3 may contribute up to 10% of the N budget in the snow; however, the main source for N is wet deposition in falling snow.

The surface snow at Ny-Alesund showed very complex stratigraphy; the NO_3^- mixing ratio in snow varied between 65 and 520 ng g^{-1} , the total NO_3^- content of the snowpack was on the order of 2700 ng cm^{-2} . In comparison the atmospheric boundary layer column showed a NO_3^- content of only 8 ng cm^{-2} . The limited exchange, however, between the snow and the atmosphere was attributed to low mobility of NO_3^- in the observed snow.

Contrary to other Arctic sites (i.e. Alert, Nunavut or Summit, Greenland) deposition of sea salt and crustal aerosols in this marine environment made the surface snow alkaline; snow NO_3^- was associated with heavier cations and was not readily available for physical exchange or photochemical reactions.

1. Introduction

Recent findings of NO_x and HONO production in snow interstitial air show that photochemical production of NO_x in or above snow surfaces is sufficient to alter the composi-

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

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tion of the overlying atmosphere (Honrath et al., 1999; Ridley et al., 2000). The current understanding of this process points towards the absorption of some NO_y species by snow surfaces, the presence and subsequent reduction of NO_3^- in a surface phase followed by photochemical release of NO , NO_2 , and HONO (Honrath et al., 2000; Dubowski et al., 2001). The release of NO_x and HONO was quantified during spring-time of 2000 at Alert (Canada, $82^\circ 27' \text{ N}$, $62^\circ 32' \text{ W}$) to be on the order of $40 \text{ nmol m}^{-2} \text{ h}^{-1}$ each (Beine et al., 2002a; Zhou et al., 2001). While the release of active N has been investigated and quantified, little is known about the sources of snow nitrate and other interactions of NO_3^- with the snow layer and the atmosphere. At Alert 2000 it was found that the predominant source for snow NO_3^- was wet deposition during snowfall (Ianniello et al., 2002). Snow experiments have been performed at Alert, showing that the mobility and thus availability for photochemical reactions of snow NO_3^- depend on the temperature, type and acidity of snow (Beine et al., 2002b).

We measured during the spring of 2001 at Ny-Alesund, Svalbard during two intensive campaigns surface fluxes of nitrogen and other atmospheric trace species, detailed aerosol physical and chemical characteristics, as well as snow physics and chemistry.

This paper is part of a series of papers describing the results from the European Community research program “The Nitrogen Cycle and Effects on the oxidation of atmospheric trace species at high latitudes” (NICE). The aim of this work is to help identify the atmospheric sources of snow- NO_3^- , and to help understand the mechanism of the exchange reactions between snow and the atmosphere and also within snow. Other aspects of these exchanges include the chemical composition of aerosols, which are discussed by Teinilä et al. (2002), and physical aerosol characteristics are shown by Nyeki et al. (2002); further papers are in preparation.

2. Experimental

2.1. The fieldsite at Ny-Alesund

The fieldsite was chosen ca. 1.5 km up-fjord of the Ny-Alesund International Arctic Research and Monitoring Facility (78°54' N, 11°53' E) at Svalbard (see Beine et al., 2001a for a map of the location). The sampling location was at sea level, ca. 20 m next to the fjord that unseasonably remained open throughout the arctic winter. The local wind field is documented for this specific location through SODAR measurements; during springtime the majority of the local flow arrived from easterly directions (Beine et al., 2001a).

Measurements were made during two intensive campaigns; a “dark intensive”, 21 February – 8 March (DOY 52–67), when the sun was below the horizon, and a “light intensive” during 24 h sunlight from 24 April – 19 May (DOY 114–139).

We used a wind-vane and anemometer connected to shut off valves to discontinue sampling when the air flow arrived from a 120° sector around the town of Ny-Alesund or under certain low-flow conditions. This shut-off system protected the denuder/filter measurements, as well as the impactor, and some of the physical aerosol measurements.

2.2. Denuder/filter measurements

Measurements of atmospheric nitrogen species were made with the C.N.R. – IIA denuder/filter system (and ion chromatographic (IC) analysis) at two sampling altitudes above the snow surface: 2 cm (L) and 180 cm (H). 12 h and 24 h long samples were collected at both heights, and analyses were subsequently carried out usually within 12 h at the Italian research station “Base Artica CNR Dirigibile Italia” in Ny-Alesund.

The annular denuder/filter pack sampling method and the analytical procedures were described previously in detail (Possanzini et al., 1983; Allegrini et al., 1987; Beine et al., 2001b). A brief account is given here for sampling of NO_3^- and NO_2^- species; other

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

species such as HCl, HBr, SO_4^{2-} etc. were measured simultaneously with this method. Nitrous acid (HONO) and nitric acid (HNO_3) in the gas phase were sampled with a combination of sodium fluoride (NaF), sodium carbonate (Na_2CO_3) and alkaline carbon coated denuders (two of each). This denuder line is followed by a cyclone ($2.5\ \mu\text{m}$ aerodynamic diameter cut point) and a filter pack for the collection of atmospheric aerosols, such as nitrate (NO_3^-) and nitrite (NO_2^-) in the coarse and fine fractions, respectively. The filter pack included a 47-mm Teflon filter (Gelman Teflon, $1\text{-}\mu\text{m}$ pore size) for the collection of particles, a Nylon filter (Nylosorb Gelman, 47 mm, $1\text{-}\mu\text{m}$ pore size) and two Na_2CO_3 -glycerol impregnated paper filters (Whatman 41) for the collection of evaporated species from the front Teflon filter. The samples were analyzed by ion chromatography (Dionex IC mod. DX120) using Dionex AS12 column for anions and Dionex CS12 column for cations.

For chromatographic analyses, control samples, consisting of 100 ng/ml calibration solution of anions (Cl^- , Br^- , NO_3^- , NO_2^- , SO_4^{2-}) and cations (Na^+ , NH_4^+ , K^+ , Mg^{2+} , Ca^{2+}), were analyzed every seven samples in order to re-calibrate the ion chromatograph. The variation in the concentration of the control samples was generally of the order 0.5–1%. Calibration solutions are prepared every week from 1000 mg/l standard solutions (MERCK).

Field blanks were measured every 3 days, these blank values are shown in Table 1. To optimize blank values and detection limits of the analytical method and the sampling technique, special care was taken in the field and in the laboratory to the quality control/assurance procedures (Perrino et al., 1990; Ianniello et al., 1999). The detection limits (L.O.D.) of the analytical technique (Table 1) are calculated as the analyte signal which is three times the standard deviation (3σ) of a very low concentration measurement. The L.O.D. for denuder measurements of various particulate and gaseous species (Table 1) are evaluated on the basis of the L.O.D. of the analytical method and of the standard deviations of field blanks. The precision is given here as relative standard deviation (RSD). In the case where field blanks were below the analytical detection limit, the detection limits of species were calculated as the 3σ variation of

IC baseline noise. Data below the 3σ detection limits are shown with their numerical value.

2.3. Snow sampling

Snow was sampled regularly for ionic analyses and physical measurements during the spring campaign, with an emphasis on surface snow. The snow stratigraphy was very complex in the spring because of frequent snowfalls that deposited small amounts of snow, and of extensive erosion, remobilization, and (more rarely) accumulation by wind. Moreover, surface melting was also observed twice in the spring, which added further heterogeneity to the snowpack. Because of these features, it was often difficult to identify and monitor a given snow layer. To overcome this difficulty, the snowpack was very carefully observed twice a day or more near the atmospheric sampling site. Layer identification was made on the spot, by observing the crystals with a magnifying glass. When in doubt, crystals were sampled and observed with a photomacrographic equipment in a cold room at -20°C at the Japanese research station "Rabben". For ionic analyses, triplicate samples were taken using polyethylene gloves, with the usual precautions to avoid sample contamination, and were analyzed within 12 h using the IC system described above.

2.4. 3D meteorology and derivation of fluxes

To derive the characteristics of the surface-atmosphere interaction and to determine the turbulent fluxes we used a sonic anemometer (Metek USAT 1). Additionally, profile measurements of air temperature and wind speed were performed, to be able to compute fluxes of chemical species. The instrument was set up ca. 5m from the denuder measurements. The anemometric data were stored in a PC after electronic processing (two axis rotations and the computation of 15 min averages of the main physical quantities and standard deviations along with the momentum and sensible heat flux determination). The derivation of HNO_3 fluxes from this mixed sonic and gradient tech-

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

nique (Sozzi et al., 1998) is described in the Appendix.

3. Results and discussion

Numerous data have been obtained for NICE, and only some of the most important points will be selected for discussion. A field campaign with related objective was also performed at Alert, Nunavut, during the spring of 2000 in two intensives (dark and light) (Beine et al., 2002a, b; Dominé et al., 2002; Cabanes et al., 2002; Ianniello et al., 2002). Comparisons between the present data set and those reported from the ALERT 2000 campaign must thus be made. This discussion will focus on the atmospheric $\text{HNO}_3/\text{NO}_3^-$ partitioning, on sources of NO_3^- ions to the snow, and on exchanges of reactive nitrogen species between the snow and the atmosphere.

3.1. Nitrogen species

Figure 1 shows the measurement results for gaseous HONO and HNO_3 and total particulate NO_2^- and NO_3^- as box-and-whisker plots at both inlets for both intensives. The median values for HONO, NO_2^- , and HNO_3 were below 5 pmol mol^{-1} during the dark intensive at both inlets, NO_3^- showed a median mixing ratio of 21 and 22 pmol mol^{-1} at the lower and upper inlet, respectively (at typical dark conditions of 1 atm and -20°C this corresponds to 62 and $65 \text{ ng/m}^3 \text{ NO}_3^-$, respectively). HONO and NO_2^- increased only slightly until the light intensive, their mixing ratios were still below 5 pmol mol^{-1} at either inlet. NO_3^- increased by about 1/3 to median values of 28 and 30 pmol mol^{-1} . The most pronounced change between the two intensives can be observed in HNO_3 ; the median mixing ratio increased 5–6 fold to 11 and 18 pmol mol^{-1} at the lower and upper inlet, respectively. As seen in Fig. 1, the scatter of each measured species did not vary significantly with either sampling height or time.

The difference in mixing ratios Δ between the two sampling heights ($\Delta = \text{L-H}$) was statistically significant only for HNO_3 (Fig. 2). During the light intensive the median

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

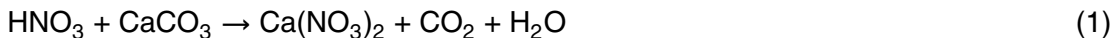
Print Version

Interactive Discussion

Δ_{HNO_3} was $-5.5 \text{ pmol mol}^{-1}$, which implies deposition to the surface.

3.2. Aerosol composition

Numerous data on aerosol composition have already been reported by Teinilä et al. (2002), and we will focus here on aspects relevant to nitrate chemistry. It is important to determine the composition of NO_3^- bearing particles to assess NO_3^- reactivity. The data obtained by Teinilä et al. using a small deposit area impactor (SDI) with 12 collecting stages indicate that NO_3^- is mostly found on coarse particles. For particles with an aerodynamic diameter $>1.7 \mu\text{m}$, we found a good correlation between NO_3^- and Ca^{2+} ($R^2 >0.7$) indicating that NO_3^- is mostly associated with Ca^{2+} , probably because of the facile reaction between HNO_3 and CaCO_3 :



that has often been suggested from aerosol measurements (e.g. Prospero et al., 1981) and confirmed by laboratory measurements (e.g. Hanisch and Crowley, 2001). The remaining part of the NO_3^- is found on particles between 0.8 and $1.7 \mu\text{m}$, which are probably the smallest sea-salt particles. The Cl^-/Na^+ ratio of those particles is much less than the marine ratio of 1.8 . This could be due to a significant Na^+ contribution from crustal dust, however, the Ca^{2+} data of our size-segregated aerosol measurements (Teinilä et al., 2002) indicate that there is probably not sufficient crustal contribution in this size range to explain the depletion, which then is mostly caused by reactions with SO_4^{2-} and NO_3^- :



A correlation between Na^+ and NO_3^- that would confirm the presence of NO_3^- on those small sea salt particles is difficult to evidence, however, because of the preponderance of SO_4^{2-} in the fractionation process. The involvement of NO_3^- in the fractionation process is evidenced by its presence in the coarse mode, with a peak around $2 \mu\text{m}$.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

The data obtained using the denuder system were used to calculate the ionic balance of the particles. Particles with aerodynamic diameters below and above $2.5\mu\text{m}$ were studied separately. The ionic balance of particles, $(\text{Cl}^- + \text{NO}_2^- + \text{Br}^- + \text{NO}_3^- + 2\text{SO}_4^{2-}) - (\text{Na}^+ + \text{NH}_4^+ + \text{K}^+ + 2\text{Mg}^{2+} + 2\text{Ca}^{2+})$, in units of neq m^{-3} , are shown in Fig. 3.

5 A similar ion balance $((\text{MSA} + \text{Cl}^- + \text{NO}_3^- + 2\text{SO}_4^{2+} + 2\text{Ox}) - (\text{Na}^+ + \text{NH}_4^+ + \text{K}^+ + 2\text{Mg}^{2+} + 2\text{Ca}^{2+}))$ from the SDI impactor measurements (Teinilä et al., 2002) showed that acidic particles were found almost exclusively in the size range from ca. 0.15 to $1.06\mu\text{m}$ in a timeseries as shown in Fig. 3. Teinilä et al. (2002) showed that there is about the same mass of small (acidic) and large (alkaline) particles. Larger particles
10 are more efficiently filtered by snow; thus snow will retain the large alkaline particles, and become alkaline.

To sum up these aerosol chemistry data, it is reasonable to say that most of the particulate NO_3^- was probably in the form of $\text{Ca}(\text{NO}_3)_2$ on crustal aerosols, whose ionic balance suggest that they were not acidic. The remaining part of the NO_3^- may have
15 been located on sea salt aerosols, in the form of NaNO_3 , but this remains speculative as reaction (3) conceals a good correlation between Na^+ and NO_3^- .

3.3. Atmospheric $\text{HNO}_3/\text{NO}_3^-$ partitioning

During NICE, NO_3^- increased by about one third between winter and spring, while HNO_3 increased 5–6 fold to 11 and 18 pmol mol^{-1} at the lower and upper inlet, respectively. As seen in Fig. 1, the scatter of each measured species did not vary significantly with either sampling location or time. At Alert, the winter to spring increase was much greater for NO_3^- : 24.6 to 48.1 pmol mol^{-1} , and much smaller for HNO_3 : 2.0 to 8.5 pmol mol^{-1} (Ianniello et al., 2002). While the light conditions at Alert were similar to those during our intensives at Ny-Alesund, Alert was much colder, drier, and there was
25 much less air movement. The composition of aerosols and the ratios of $\text{HNO}_3/\text{NO}_3^-$ are a function of temperature, aerosol composition and acidity, as well as the kinetics

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

of production/destruction/deposition of both HNO_3 and NO_3^- . Since this system is very complex, we must limit ourselves to qualitative considerations.

At Ny-Alesund higher temperatures favor volatilization of HNO_3 from the aerosol; additionally, higher NO_x mixing ratios (Beine et al., 1997, 2002a) will lead to increased HNO_3 formation. Due to the stronger winds, it can be suggested that dry deposition of NO_3^- may be favored by the filtering of particles by snow. These elements will favor a higher $\text{HNO}_3/\text{NO}_3^-$ ratio at Ny-Alesund than at Alert. This may be counterbalanced by the alkaline or neutral character of the aerosol at Ny-Alesund, which will favor HNO_3 uptake and its conversion to NO_3^- . An ion balance at Alert suggested that the aerosol during spring time was acidic, and would thus favor the free acid. However, since the ratio $\text{HNO}_3/\text{NO}_3^-$ is greater at Ny-Alesund, we must conclude that the greater aerosol alkalinity is insufficient to compensate for the warmer temperatures and higher HNO_3 production from NO_x , and that these latter processes are probably predominant in the determination of $\text{HNO}_3/\text{NO}_3^-$.

3.4. Snow stratigraphy and NO_3^- concentration in snow

The snow stratigraphy was very complex. Most of the stratigraphic observations were made near the Amundsen mast (ca. 100 m to the South inland from the atmospheric sampling location). The accumulation at these coastal sites is less than on locations further inland because of strong wind erosion. Nine snowfall events were observed between 18 April and 9 May. Many of them were blown away by wind and did not contribute to accumulation. The 25 April (DOY 115) and 27 April (DOY 117) snowfalls were stabilized by the formation of a thin melt-freeze crusts and were monitored in detail. The 29 April (DOY 119) snowfall was over 5 cm thick and was not eroded until 8 May, and then only partially so, so it could also be studied over an extended period. The surface snow stratigraphy on 4 May (DOY 124) is shown in Fig. 4, to illustrate its complexity.

In general, the snow composition was found to be very variable within a given layer,

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

both vertically and horizontally. This is not due to contamination problems, as sampling in unperturbed areas yielded little variation (<5% between triplicate samples). (Moreover, the sampling and analysis procedures were similar to those used at Alert, where there was little wind, and where variations amongst triplicates were also <5%). The vertical variation can simply be attributed to varying atmospheric composition during precipitation. The horizontal variation is more difficult to explain, and several reasons can be suggested. First of all, many different layers were outcropping at any time, so that a new snowfall was in contact with layers that varied with location. Interactions between layers may lead to exchanges of solutes, and this may have resulted in horizontal variations. Second, wind was often blowing during precipitation, resulting in mixing between fresh snow and already deposited snow. Since snow chemistry varied during precipitation, and since wind results in irregular snow mixing, horizontal variations can be expected from these effects. These variations added noise to the data, even though we attempted to sample one given snowbank, and progressed within this snowbank with every sampling in an effort to minimize the effects of horizontal variations.

Figure 5 shows the variation of the NO_3^- concentration in the snow that precipitated on 27 April (DOY 117; Fig 5a.) and 29 April (DOY 119; Fig. 5b). The 27 April layer was covered by the 29 April snowfall and was not exposed again until 8 May, when only parts of it outcropped again. The analyses after 8 May show snow that was still covered. Of the 5 cm-thick 29 April snowfall both the middle and the top layer were sampled. The bottom part was not sampled to avoid contamination of the underlying 27 April layer that was richer in NO_3^- . The windpacked sublayer that formed on top of the 29 April layer on 4 May was sampled after its formation. This wind-pack formed by wind-induced sintering of surface snow, rather than by wind deposition. We observe no significant trend in NO_3^- concentrations in these layers, which is consistent with the role of sea-salt in the surface layer to prevent loss of NO_3^- . The sum of the two 27 April sublayers in Fig. 5a seems to remain constant, which suggests that exchange amongst sublayers may have taken place, without a net emission. The one significant observation is that the formation of the windpacked sub-layer was accompanied by an

increase in the NO_3^- concentration.

3.5. Atmospheric stability and HNO_3 fluxes

Figures 6a–c show the overall meteorological conditions during our experiment at Ny-Alesund. Figure 6d shows the friction velocity u^* ($u^* = \tau/\rho$, where ρ is the air density, and τ the in situ momentum flux), derived from the mixed micrometeorological measurements. No diurnal pattern was significant for u^* over the duration of the light intensive, except that the variability was slightly increased in the morning hours. While the majority of the u^* values fall below 0.15 ms^{-1} , values up to 0.3 ms^{-1} were observed corresponding to days with higher windspeeds.

Using both the measured Δ_{HNO_3} and the measured eddy diffusivity for the same time period, the atmospheric surface fluxes for HNO_3 were derived (Fig. 7). The average value for the deposition flux was $-8.7 \text{ nmol h}^{-1} \text{ m}^{-2}$ (and was statistically significant, $p = 0.0229$, $t = -2.3908$, $95\% \text{ CI} = \pm 7.46$). Two extreme values of -60 and $-80 \text{ nmol m}^{-2} \text{ h}^{-1}$ HNO_3 deposition were observed during the largest observed snowfall during our campaign (DOY 119). The winds were calm ($< 3 \text{ ms}^{-1}$), temperatures and relative humidities very high. The observed HNO_3 concentrations on this day, however, were the highest observed. The measurement at the upper line was 116 pmol/mol . This suggests that the air mass brought HNO_3 with the snow, and HNO_3 was scavenged by that snowfall and deposited to the surface.

3.6. Sources of NO_3^- ions to the snow

Surface snow at Ny-Alesund showed NO_3^- mixing ratios of 65–520 ng/g. Three distinct sources of nitrate contribute to this signal; 1. The major fraction of NO_3^- is contained in precipitating snow; 2. HNO_3 is scavenged during precipitation; and 3. $\text{HNO}_3 / \text{NO}_3^-$ is dry deposited independently of precipitation. The first two mechanisms are illustrated by the 29 April snowfall event which was mentioned above: The nitrate content in the falling snow corresponds to the surface snow layers later sampled. Neither HNO_3 nor

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

Fluxes of nitrates

H. J. Beine et al.

dry particle NO_3^- deposition were large at Ny-Alesund, so that the main NO_3^- source for snow seems to be wet deposition in falling snow. A similar conclusion was reached for Alert 2000 (Ianniello et al., 2002). The largest observed source of HNO_3 was deposition during the major snowfall on DOY 119. During that snowfall 5 cm snow of density 0.124 g cm^{-3} fell with a NO_3^- content of 200 ng/g. Assuming that the high HNO_3 deposition of -60 to $-80 \text{ nmol m}^{-2} \text{ h}^{-1}$ (see above) lasted for 24 h, the contribution of gaseous HNO_3 to the NO_3^- content in the fresh snow is only ca. 10%.

Dry deposition of HNO_3 , though small, was statistically significant at Ny-Alesund on most days (Fig. 7). A significant difference between Alert and Ny-Alesund is the sea salt content of the snow. At Ny-Alesund, the median Cl^- was about 1500 ng/g. At Alert, it was about 250 ng/g before a storm event on 25–29 April 2000. That storm brought in marine air, increasing Cl^- to about 1500 ng/g. It can then be hypothesized that the reaction (2) led to HNO_3 deposition to the snow (Fenter et al., 1994). The rate obviously depends on the concentration of NaCl, and thus would be much more efficient at Ny-Alesund, relative to pre-storm Alert. Indeed, for the 3 days prior to the storm, the snow at Alert was a small source of HNO_3 ; during the storm it was a sink of about equal strength (up to $30 \text{ nmol m}^{-2} \text{ h}^{-1}$, comparable to the observed NO_x and HONO emissions, Beine et al., 2002a). After the storm no HNO_3 fluxes were detected anymore.

Fresh snow usually contains little sea salt, and is slightly acidic (the snow ion-balance shows an acidity of about $4 \mu\text{eq/L}$). Wind deposition of sea salt, which is alkaline (Stumm and Morgan, 1970) turns the snow alkaline (the snow ion-balance shows alkalinity of 10 to $40 \mu\text{eq/L}$). Therefore surface snow that has been exposed to sea salt is a good sink for HNO_3 . Deeper snow layers may remain acidic and release HNO_3 , which immediately gets trapped in the alkaline surface layer.

Under the crude assumption that our observed mean HNO_3 fluxes hold true for an entire winter season (September to May), the accumulation of snow NO_3^- from atmospheric HNO_3 would be on the order of $50 \mu\text{mol m}^{-2}$. If this affected the surface layer

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

(top 5 cm) of the snow, of mean density of 0.2 g cm^{-3} , the increase was 310 ppbw (ng/g) over the entire season. This translates to about 1 ppb per day, which is a change virtually impossible to detect in the snow.

Dry deposition of NO_3^- particles were not observed in statistically significant quantity at Ny-Alesund. However, we see evidence that the observed surface snow layers gain Na^+ and Ca^{2+} over time; though more Na^+ than Ca^{2+} . This difference in the efficiency of dry particle capture may depend on the particle nature. A NaCl particle is likely to dissolve upon contact with the snow surface, and will thus easily stick. Crustal CaCO_3 , on the other hand, does dissolve much slower, and is likely to bounce off. Thus dry deposition during increased wind deposits preferably soluble ions. Since NO_3^- is connected with crustal particles, the observed gradients (Fig. 2) were too small to be observed with our sampling methods.

Additionally, as yet undocumented chemical transformations of NO_y reservoir species such as peroxyacetyl nitrate (PAN) or alkyl nitrates to NO_3^- in the snow phase (Ford et al., 2002) may be possible. Beine and Krognnes (2000) showed an anticorrelation of PAN mixing ratios with the relative atmospheric humidity at Ny-Alesund, but no further evidence is available that PAN, or alkyl nitrates, lead to a nitrate signal in snow, either directly or indirectly.

In conclusion; most of the nitrate signal in snow is determined at deposition, and little modification takes place afterwards. Dry deposition of HNO_3 occurs, and is enhanced by the presence of NaCl.

3.7. Exchanges of reactive nitrogen species between the snow and the atmosphere

Significantly, no emissions of HONO were observed during the spring of 2001 next to the unseasonably unfrozen fjord at Ny-Alesund. HONO emissions were observed at both Alert and Summit (Zhou et al., 2001; Honrath et al., 2002) and were attributed to the photolysis of the nitrate ion contained in snow. By extension we assume that NO_x , also produced by nitrate photolysis in a ratio $\text{NO}_x:\text{HONO}$ of $\sim 1:1$ (Beine et al.,

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

2002a), was not emitted by the snowpack. Further, the surface snow content of NO_3^- was invariable with time. These observations are in contrast to our findings at Alert during the spring of 2000, where NO_3^- in a surface snow layer of 30 mm thickness decreased by ca. $1.8 \mu\text{mol m}^{-2}/\text{day}$ over a period of 2 weeks of no new snow fall (Beine et al., 2002a). Thus, even though snow at Alert (Beine et al., 2002b; Dominé et al., in preparation), Summit (Honrath et al., 2002), and Ny-Alesund showed similar nitrate contents, the resulting emissions appeared very different at Ny-Alesund. Above we have discussed how snow composition, and the effect of sea-salt may affect HNO_3 uptake. In the following we will investigate how snow composition may affect nitrate reactivity.

Snowpack NO_3^- had the potential to affect atmospheric composition. The NO_3^- concentration in the snow at Ny-Alesund was 65 to 520 ppbw, with a total NO_3^- content of the snowpack on the order of 2700 ng/cm^2 (or $435 \mu\text{mol m}^{-2}$) near the Amundsen mast. This value was up to 5 times higher in areas that were not wind-swept, and where snow accumulation reached up to 2 m. This is 10 times more than the value from atmospheric HNO_3 deposition alone. Atmospheric NO_3^- concentrations of ca. 200 ng/m^3 (ca. 70 pmol mol^{-1} ; as HNO_3 and particulate NO_3^-) add up to a column boundary layer (300 m) content of only 8 ng/cm^2 . Thus even a release of NO_3^- by the snowpack that would result in undetectable changes in snow composition, had the potential to considerably increase boundary layer NO_3^- , HONO or NO_x concentrations.

Although the details of the photolysis mechanism of NO_3^- photolysis in natural snow are not clear, it has been suggested that photolysis will be facilitated if NO_3^- is located on the surface of snow crystals (Dubowski et al., 2001). If NO_3^- is dissolved in the crystalline lattice of ice, or trapped inside aerosol particles, recombination of products due to the cage effect will considerably reduce emissions. HNO_3 readily adsorbs onto snow crystals (Abbatt, 1997, Zondlo et al., 1997) while its solubility into the ice crystalline lattice is low (Thibert and Dominé, 1998). Thus, if NO_3^- in snow is associated to H^+ , it can be suggested that it will be located on the surface of snow crystals and readily

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

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available for photolytic activation and physical exchange. However, while at Alert snow nitrate decreased with time along with the specific surface area (SSA) of snow (Beine et al., 2002b; Cabanes et al., 2002), we observed no decrease in snow nitrate for the 29 April snowfall at Ny-Alesund despite a factor of 2 decrease in snow SSA in just 3 days. This leads to the suggestion that NO_3^- in snow was not located on the surface of snow crystals, and that it was not associated with H^+ . This is confirmed by the ionic measurements that lead to the conclusion that the surface snow was most of the time alkaline, because of sea salt and crustal aerosol deposition.

The chemical form of snow NO_3^- may be constrained further; in particular it is essential to determine whether it is in the form of NH_4NO_3 , which is fairly unstable and can thus lead to facile physical exchange (Mozurkewich, 1993), or associated with Ca^{2+} , Mg^{2+} , or Na^+ , leading to stable forms of NO_3^- .

The median snow NH_4^+ concentrations from different layers of various snowfalls at Ny-Alesund was only 31 ng/g; this amounted to 2.5% of all available equivalent cations. Additionally, Teinilä et al. (2002) showed that no correlation existed between NH_4^+ and NO_3^- in the atmospheric aerosol, but that NH_4^+ was associated with SO_4^{2+} , and NO_3^- with Ca^{2+} . Hence the majority of NO_3^- in aerosols was present as $\text{Ca}(\text{NO}_3)_2$, which is fairly stable. This preferential association of NO_3^- with Ca^{2+} is frequently observed in marine air when crustal aerosols are present (e.g. Zhuang et al., 1999). In snow, determining the chemical form of NO_3^- is more difficult, because we lose the size-dependent composition. However, the chemistry of the 29 April snow layer illustrates interesting changes. The initial composition of that snow was calcium-poor (10 to 20 ppbw), and it was acidic. The composition of the top sublayer changed, and Ca^{2+} reached about 50 ppbw after a couple of days, presumably because of the dry deposition of crustal aerosol, while the snow became alkaline (by about $5 \mu\text{eq/L}$). With the formation of the wind crust, Ca^{2+} increased to 300 ppbw, while the alkalinity increased to 10–40 $\mu\text{eq/L}$. Na^+ was the most abundant cation, around 2000 ppbw, while there were 300 ppbw of Mg^{2+} and Ca^{2+} . Since we do not know whether the particles deposited onto the snow

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[I◀](#)[▶I](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

retained their individuality or whether they became dissolved in snow, it is difficult to determine the form of NO_3^- in snow. However, considering that the snow is alkaline, and that the major cations are Na^+ , Ca^{2+} and Mg^{2+} , the most reasonable hypothesis is that in snow, NO_3^- is associated with one of these ions and is therefore fairly inert with regards to physical exchanges. It is possible that aerosol $\text{Ca}(\text{NO}_3)_2$ remains in the particle form; crustal aerosols have been observed inside ice crystals in ice cores (Fuji et al., 1999). With snow metamorphism, particles can then become imbedded inside ice crystals, thus becoming isolated from the atmosphere and therefore unreactive.

Alternatively, ionic species may interact with ice to form ionic mixtures that could even be liquid, and that would spread on the ice crystal surfaces. (Koop et al., 2000). In that case, NO_3^- may be potentially reactive. However, release of surface NO_3^- as HNO_3 would require an acidic medium, which appears incompatible with the increase in snow alkalinity after snowfall. Moreover, such HNO_3 emissions were not observed, and the arguments just developed, together with the observations, lead to the conclusion that NO_3^- in snow at Ny-Alesund is (i) not reactive towards physical exchange with the atmosphere, because it is associated with counter cations such as Na^+ or Ca^{2+} . (ii) not reactive towards photolysis, probably because it is not located on the surface of snow crystals; it may be either trapped inside particles or, less likely, trapped inside the ice crystalline lattice. The potential for exchanges may thus not be as high as suggested by the partitioning of nitrate between the atmosphere and the snow.

4. Conclusions

At Ny-Alesund during the spring of 2001 at a coastal site close to the unseasonably unfrozen fjord only deposition fluxes of HNO_3 were detected. Other measured nitrogen species showed no significant fluxes which may be in part due to the measurement uncertainties and long integration times of the used denuder technique, and atmospheric turbulence and mixing. These HNO_3 fluxes may be due to the reaction of HNO_3 with

sea salt, and especially NaCl, or may be simply uptake of HNO_3 by ice, which is alkaline because of the sea salt in our marine environment.

At Alert and Summit, the snow was acidic. NO_3^- was thus mainly associated with H^+ . It was probably located on the surface of snow crystals and was available for physical exchanges and for photochemical reactions, that led to HONO and NO_x emissions by the snowpack. At Ny-Alesund, although fresh snow was often acidic, deposition of sea salt and crustal aerosols in this marine environment made the snow alkaline, and NO_3^- was then associated with heavier cations. Although it is not clear whether NO_3^- was actually incorporated in the snow crystals or remained within aerosol particles, we suggest that it was not readily available for physical exchange or photochemical reactions. If such reactions did take place, formation and/or release of products was hindered by cage effects or trapping in solids.

These results lead to the conclusion that understanding NO_3^- exchanges between the snow and the atmosphere will require more efforts in determining what contributes to the NO_3^- signal in snow (Dibb et al., 1998). We also need to understand nitrate localization and reactivity in snow, and whether it interacts with NH_4^+ or heavier cations inside particles or dissolved in ice crystals. It is then clear that understanding the impact of snow NO_3^- on atmospheric composition will require microphysical investigation of the snow, and in particular how aerosol particles interact with snow. It is crucial to determine whether soluble species remain within aerosol particles as a separate phase, or whether they become dissolved in ice.

Finally, we have shown here that reactivity of species in the snow phase will depend on their chemical state, and on other species, that influence reactivity (e.g. pH). This confirms recent suggestions by Dominé and Shepson (2002) that the mechanism and magnitude of emissions of other species, such as HCHO, by snow would depend on its overall composition. A larger number of locations with different atmospheric and snow compositions will have to be studied before the overall impact of snow on atmospheric chemistry can be estimated reliably.

Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

Appendix – Calculation of HNO₃ surface fluxes

Using a mixed sonic and gradient technique (Sozzi and Favaron, 1996), atmospheric stability and surface fluxes were determined.

The first step of an eddy-correlation process is the calculation of the perturbation value of each data point; by way of example, the vertical wind speed component w was considered, which can be written as (Stull, 1988)

$$w(t) = \bar{w} + w'(t), \quad (\text{A1})$$

where w is the measured value, \bar{w} is the mean value and w' is the fluctuation value. From a time series of data, the mean and the fluctuation values can be determined for each variable. To calculate heat fluxes, we need to know time series for the temperature, (for the sensible heat flux), and the specific humidity, (for the latent heat flux). Then the respective fluctuations are multiplied with each other, which yields the following time series:

$$\begin{aligned} &w'T'(t), w'T'(t + \Delta t), w'T'(t + 2\Delta t), \dots \text{ for the sensible heat flux,} \\ &w'q'(t), w'q'(t + \Delta t), w'q'(t + 2\Delta t), \dots \text{ for the latent heat flux.} \end{aligned} \quad (\text{A2})$$

Finally, the averages of the series $\overline{w'T'}$ and $\overline{w'q'}$ constitute the useful quantities for the computation of the latent and heat fluxes

$$\begin{aligned} H_0 &= \rho C_p \overline{w'T'} \\ LE &= \lambda \overline{w'q'} \end{aligned} \quad (\text{A3})$$

where ρ is the air density, C_p is the specific heat of the air at constant pressure and λ is the latent heat of sublimation. In the same way also the friction velocity can be determined by the covariance between w and u .

$$u_* = \sqrt{-\overline{u'w'}} \quad (\text{in streamline coordinates}) \quad (\text{A4})$$

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

Because our aim is to compute HNO_3 fluxes for the time integrated (12 h) HNO_3 samples a combined eddy covariance and gradient technique were applied. Following the Monin-Obukhov similarity theory, the vertical profile of a generic scalar in the surface layer can be represented by the same universal similarity function of air temperature.

5 First of all it is necessary to define the scale parameter for a generic gas concentration c_* :

$$c_* = -\frac{\overline{w'c'}}{u_*} \quad (\text{A5})$$

The similarity function for the vertical profile of a generic gas c will be given by the relation:

$$10 \quad c(z_2) - c(z_1) = \frac{c_*}{k} \cdot \left\{ \ln \frac{z_2}{z_1} - \Psi_H \left(\frac{z_2}{L} \right) + \Psi_H \left(\frac{z_1}{L} \right) \right\} \quad (\text{A5a})$$

where the universal similarity function is:

$$\Psi_H = \begin{cases} 2 \ln \left[\frac{1 + \sqrt{1 - 16z/L}}{2} \right] & \text{if } z/L < 0 \\ -17 \cdot [1 - \exp(-0.29z/L)] & \text{if } z/L > 0 \end{cases} \quad (\text{A5b})$$

and z/L is the stability parameter derived from sonic anemometer measurements. If we equate $\Delta Q = c(z_1) - c(z_2)$, the relation between ΔQ and the HNO_3 vertical flux will be:

$$15 \quad \overline{w'c'} = -u_* c_* = +\Delta Q \frac{ku_*}{\ln \frac{z_2}{z_1} - \Psi_H \left(\frac{z_2}{L} \right) + \Psi_H \left(\frac{z_1}{L} \right)} \quad (\text{A6})$$

Now we can obtain the eddy diffusivity for HNO_3 flux

$$D = \frac{ku_*}{\ln \frac{z_2}{z_1} - \Psi_H \left(\frac{z_2}{L} \right) + \Psi_H \left(\frac{z_1}{L} \right)} \quad (\text{A6a})$$

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

$$\overline{w'c'} = D\Delta Q \quad (\text{A6b})$$

Because the HNO_3 measurements were conducted for long time periods (12–24 h) we computed a mean for the micrometeorological parameters and the HNO_3 flux. The mean (efficacy) eddy diffusivity for the HNO_3 flux is:

$$\overline{D} = k \cdot \frac{u_*}{\ln \frac{z_2}{z_1} - \Psi_H \left(\frac{z_2}{L} \right) + \Psi_H \left(\frac{z_1}{L} \right)} \quad (\text{A7})$$

and the mean period vertical HNO_3 flux is equal to:

$$\overline{w'c'} = \overline{D} \cdot [c(z_1) - c(z_2)] \quad (\text{A8})$$

The surface fluxes were computed from the profile measurements using the flux gradient methodology. In order to validate this technique the results were compared with the eddy covariance technique applied to sonic anemometer data.

Acknowledgement. The authors would like to thank all the participants of the NICE experimental campaign, in particular R. Sparapani for the organization of NICE at Ny-Alesund, A. Felici for his technical help, and L. Ammiraglia, E. Colaiacomo, and F. DiBari for their tireless help in making the IC analyses of the denuder, filter and numerous snow samples. Thanks to M. Legrand (CNRS – LGGE) for discussions on snow chemistry. Funding for this project was received from the European Commission (EVK2-1999-00029 “NICE”). The participation of FD to this campaign was funded by the French Polar Institute (IPEV).

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Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

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Fluxes of nitratesH. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

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Fluxes of nitrates

H. J. Beine et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

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Fluxes of nitrates

H. J. Beine et al.

Table 1. Measurement of nitrite and nitrate with denuders/filters and IC analysis^a

Species	Field blank ^b measured by IC during the dark intensive ^c (pmol mol ⁻¹)	Field blank ^b measured by IC during the light intensive ^d (pmol mol ⁻¹)	Precision (%) ^e	3 σ detection limit (pmol mol ⁻¹)
HONO (gas)	0.18	0.19	1.5	0.55
NO ₂ ⁻ (particle)			1.2	0.57
HNO ₃ (gas)			2.57	0.27
NO ₃ ⁻ (particle)			3.0	1.05

^a = These values were calculated using the procedures shown in detail by Perrino et al. (2001) and Beine et al. 2001b).

^b = The field blank has the highest contribution to the overall blank. The laboratory and glass blanks are up to ca. 75% and 40% of the field blanks for NO₂⁻ and NO₃⁻, respectively.

^c = 24 h samples

^d = 12 h samples

^e = at ca. 10 pmol mol⁻¹

[Title Page](#)
[Abstract](#)
[Introduction](#)
[Conclusions](#)
[References](#)
[Tables](#)
[Figures](#)
[I◀](#)
[▶I](#)
[◀](#)
[▶](#)
[Back](#)
[Close](#)
[Full Screen / Esc](#)
[Print Version](#)
[Interactive Discussion](#)

© EGU 2003

Fluxes of nitrates

H. J. Beine et al.

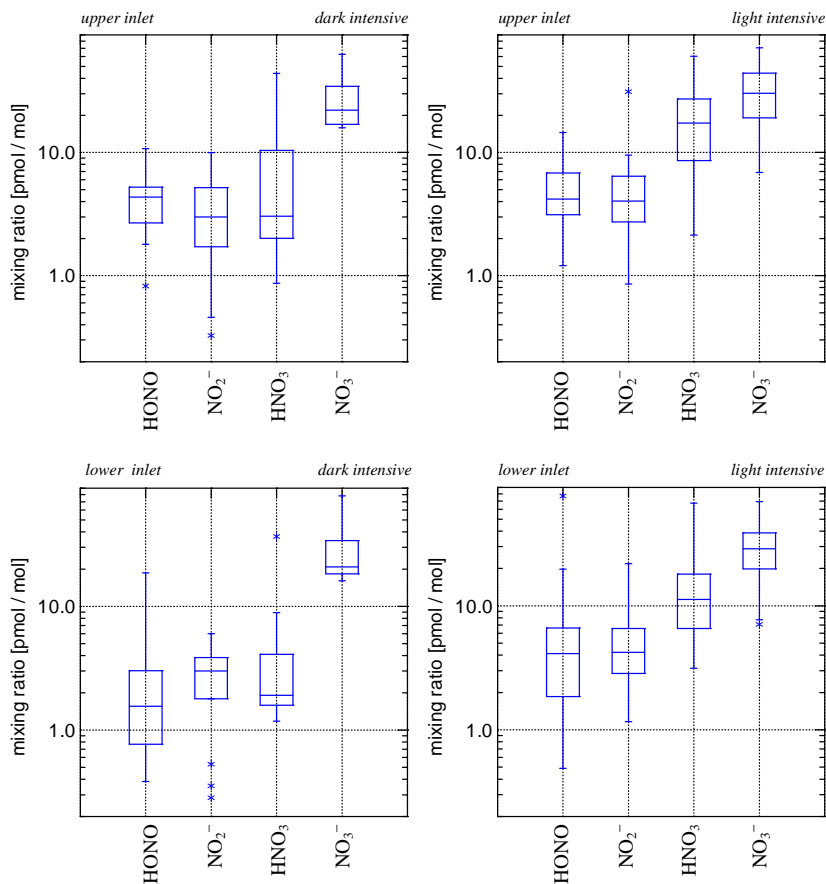


Fig. 1. Box-and-whisker plots of the nitrogen measurements at the lower and upper inlet (lower and upper row, respectively) during the dark and light intensive (left and right column, respectively). In the box the middle line shows the median, the box includes 50% of the distribution, and the whiskers extend to 95% of the distribution (Systat 9, 1999).

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

© EGU 2003

Fluxes of nitrates

H. J. Beine et al.

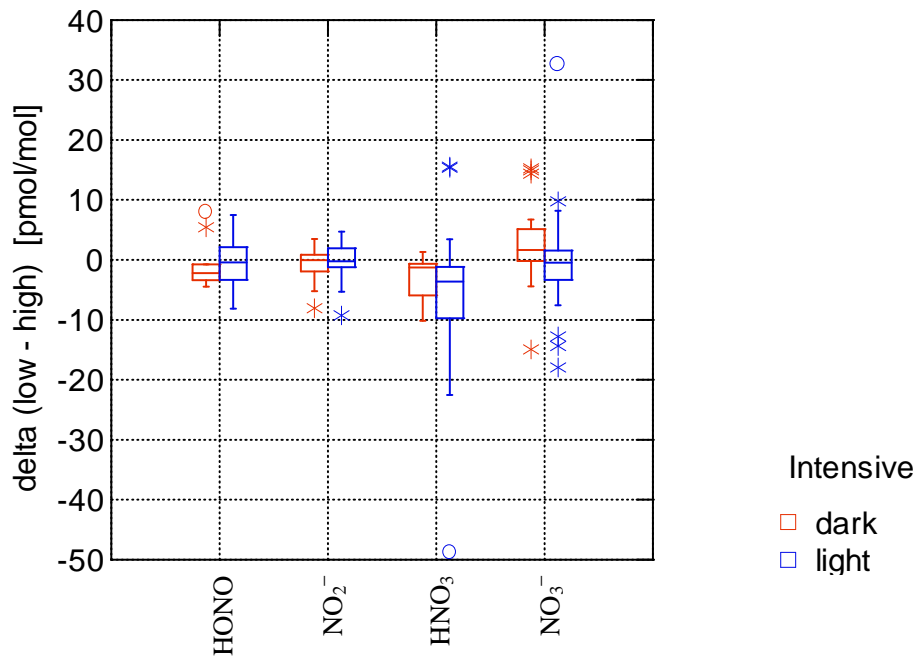


Fig. 2. Δ (low – high) of nitrogen species during the two intensives. Only Δ_{HNO_3} was statistically significantly different from zero. During the dark intensive the mean Δ_{HNO_3} was $-2.90 \text{ pmol mol}^{-1}$ (95% confidence interval from -5.56 to -0.24 , $p = 0.0356$), during the light intensive Δ_{HNO_3} was $-5.49 \text{ pmol mol}^{-1}$ (95% confidence interval from -9.39 to -1.59 , $p = 0.0072$).

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

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Fluxes of nitrates

H. J. Beine et al.

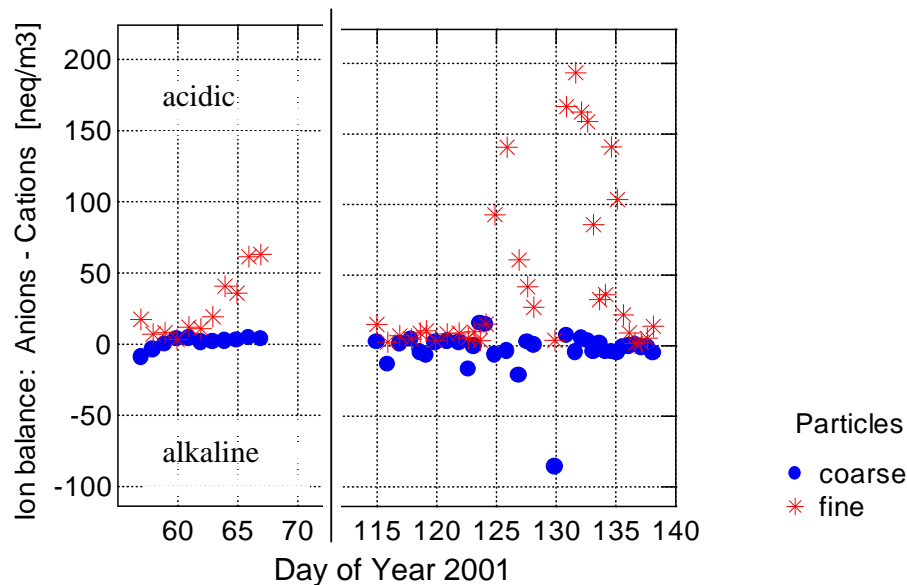


Fig. 3. Ion balance for the high line during both intensives at Ny-Alesund: $(\text{Cl}^+ + \text{NO}_2^- + \text{Br}^- + \text{NO}_3^- + 2 \text{SO}_4^{2-}) - (\text{Na}^+ + \text{NH}_4^+ + \text{K}^+ + 2 \text{Mg}^{2+} + 2 \text{Ca}^{2+})$ (neq m^{-3}) for coarse ($> 2.5 \mu\text{m}$ aerodynamic diameter) (blue circles) and fine ($< 2.5 \mu\text{m}$) particles (red stars).

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

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Fluxes of nitrates

H. J. Beine et al.

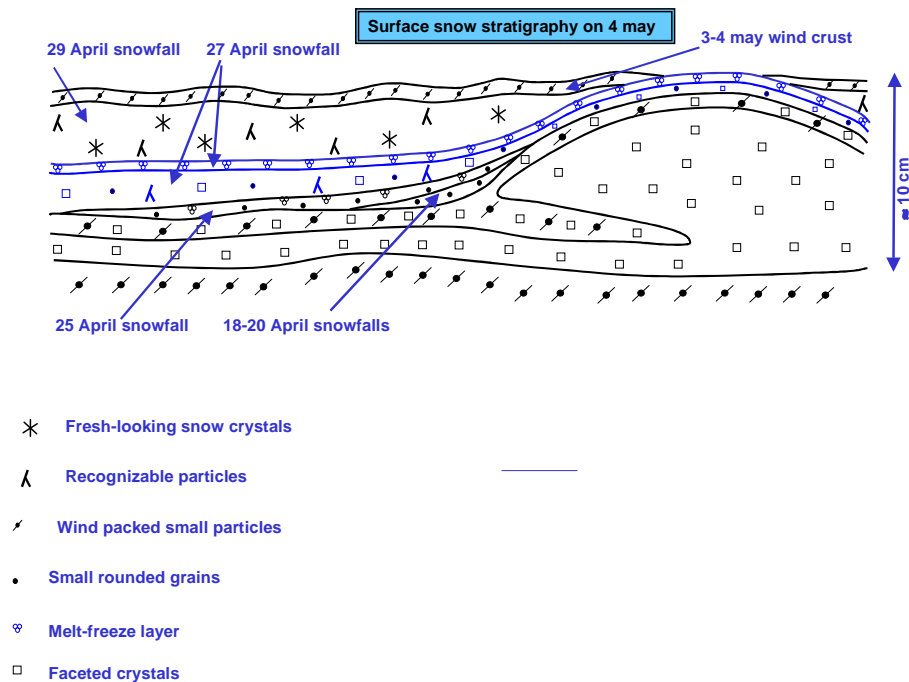


Fig. 4. Surface stratigraphy of the snowpack near the Amundsen mast. The total thickness of the snowpack is about 40 cm.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

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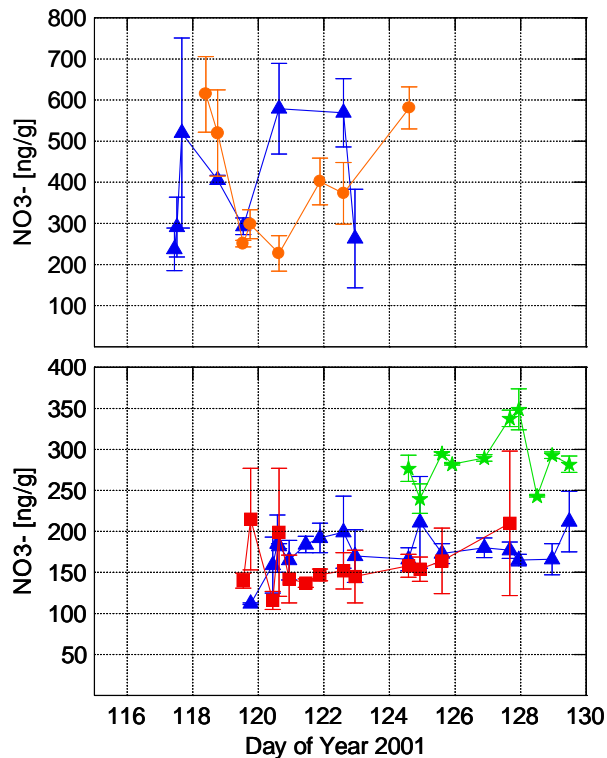


Fig. 5. (a) Time series of the NO_3^- content of the 27 April 2001 snowfall. The top (blue triangles) and bottom (orange circles) parts of the layer were sampled separately. **(b)**. Time series of the NO_3^- content of the 29 April 2001 snowfall. The top (blue triangles) and middle (red squares) layers were sampled separately. The very top of the layer became windpacked on 4 May 2001, and was subsequently sampled separately (green stars). Error bars are standard deviations in triplicate samples.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

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H. J. Beine et al.

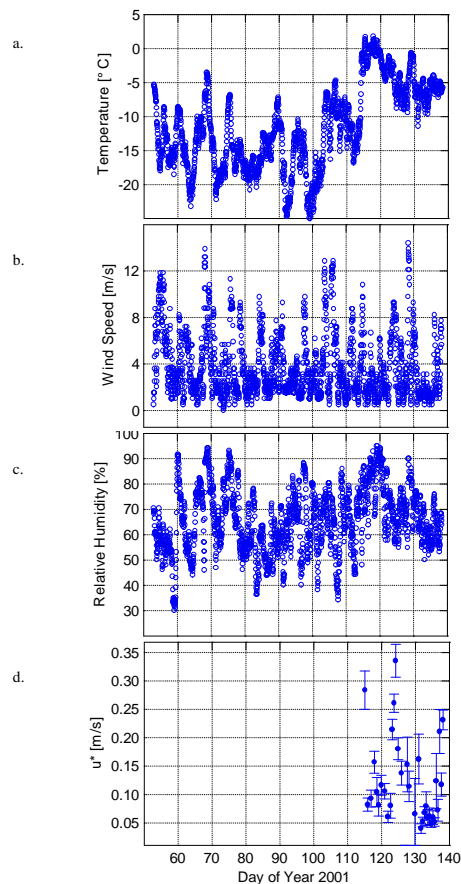


Fig. 6. (a)–(c) Meteorology at Ny-Alesund during spring 2001. (d) Variability of u^* during the light intensive. The dots show the mean value during each denuder sampling period, the error bars show the 95% confidence interval for that time period.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

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H. J. Beine et al.

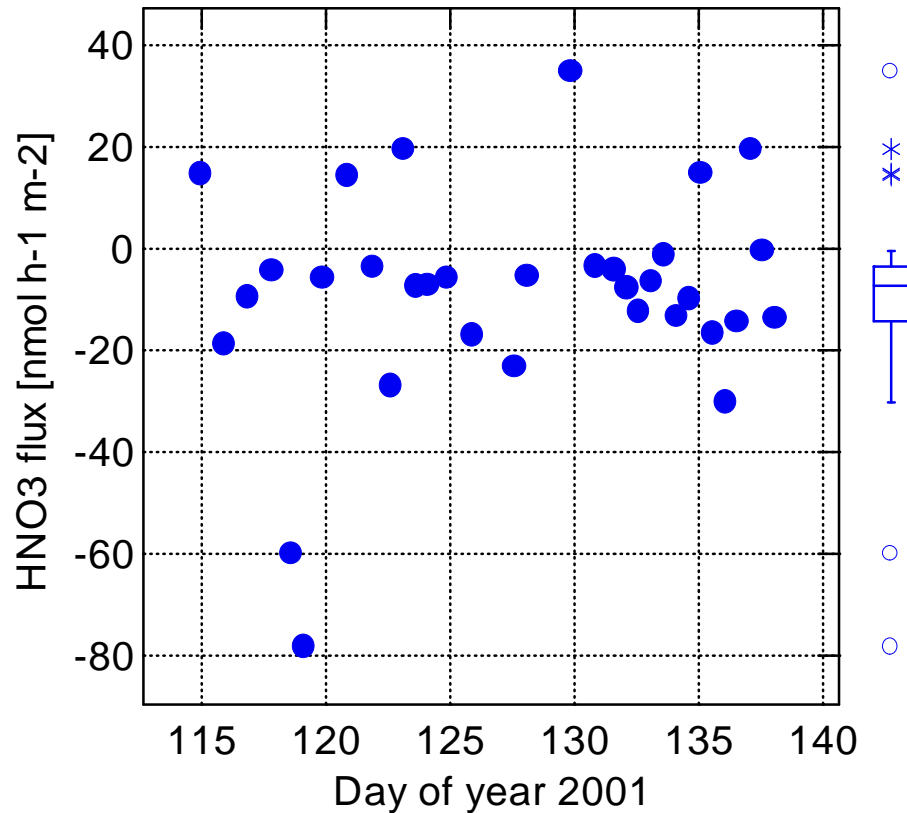


Fig. 7. HNO₃ flux above the snow surface derived from denuder measurements at 2 cm and 180 cm above the snow surface. The box-and-whisker plot to the right shows the entire distribution.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)