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Structure of synthetic K-rich birnessite obtained by high-temperature decomposition of KMnO₄.

I. Two-layer polytype from 800°C experiment.

Anne-Claire Gaillot¹, David Flot², Victor A. Drits¹,³, Alain Manceau¹, Manfred Burghammer², Bruno Lanson¹,*

1 – Environmental Geochemistry Group, LGIT – Maison des Géosciences, University of Grenoble – CNRS, 38041 Grenoble Cedex 9, France.
2 – European Synchrotron Radiation Facility, BP 220, 38043 Grenoble Cedex, France
3 – Geological Institute, Russian Academy of Sciences, 7 Pyzhevsky street, 109017 Moscow, Russia

* Author to whom correspondence should be addressed.

*Author: Bruno.Lanson@obs.ujf-grenoble.fr

Abstract

The structure of a synthetic potassium birnessite (KBi) obtained as a finely dispersed powder by thermal decomposition of KMnO₄ at 800°C was for the first time studied by single crystal X-ray diffraction (XRD). It is shown that KBi has a two-layer cell with $a = 2.840(1)$ Å, and $c = 14.03(1)$ Å, and space group $P6_3/mmc$. In contrast to the structure model proposed by Kim et al.,¹ the refined model demonstrates the sole presence of Mn⁴⁺ in the octahedral layers, the presence of 0.12 vacant layer sites per octahedron being responsible for the layer charge deficit. This layer charge deficit is compensated for 1) by the presence of...
interlayer Mn$^{3+}$ above or below vacant layer octahedra sharing three O$_{layer}$ with neighboring Mn$_{layer}$ octahedra to form a triple-corner surface complex ($^{VI}$TC sites), and 2) by the presence of interlayer K in prismatic cavities located above or below empty tridentate cavities, sharing three edges with neighboring Mn$_{layer}$ octahedra ($^{VI}$TE sites). As compared to the structure model proposed by Kim et al.,¹ this $^{VI}$TE site is shifted from the center of the prismatic cavity towards its edges. A complementary powder XRD study confirmed the structure model of the main defect-free KBi phase and allowed to determine the nature of stacking disorder in a defective accessory KBi phase admixed to the defect-free KBi.

**Introduction**

Birnessite is a phyllomanganate, that is a manganese oxide containing predominantly Mn$^{4+}$ cations assembled in layers of edge-sharing octahedra. A layer charge deficit arises from the presence within layers of Mn$^{3+}$ cations and/or vacant layer octahedra and is compensated for by the presence of interlayer cations which are typically hydrolyzable cations in natural phyllomanganates.²⁻⁹

Despite the relatively low natural abundance of manganese, birnessite is ubiquitous in nature and plays an essential role in the geochemistry of soils and oceanic nodules.⁸⁻¹⁰ This role originates from its remarkable cation exchange capacity,²⁰⁻²⁴ sorption,²⁵⁻³¹ and redox properties.³²⁻⁴² In particular, because of its high affinity for pollutants, such as heavy metals and organic contaminants, this mineral plays a pivotal role in their fate in contaminated waters and soils.¹⁶,⁴⁰⁻⁵³

Even though in nature birnessite is typically disordered and occurs in finely dispersed state, monomineralic birnessites with an enhanced three dimensional (3D) ordering are easily synthesized under a variety of laboratory conditions.¹⁻³,¹⁵,⁵⁴⁻⁶² Because some synthetic
varieties can be considered as analogs of the natural ones, they have been used to determine the structural mechanism of heavy metal sorption\textsuperscript{31,40,51,63,64} and to investigate the structural modification of birnessite as a function of pH.\textsuperscript{65}

Recently, synthetic birnessites have attracted additional attention, because of their electrochemical and magnetic properties.\textsuperscript{54-57,66-70} Birnessite species rank among the promising cathode materials for secondary lithium batteries if their lamellar framework is stable during insertion/de-insertion cycling. Such K-rich birnessite (KBi) varieties have been synthesized from the thermal decomposition of KMnO$_4$,\textsuperscript{1} as well as from the reduction of KMnO$_4$ under hydrothermal conditions.\textsuperscript{54,55}

A comprehensive structural and chemical knowledge of KBi varieties obtained under contrasting synthesis conditions is relevant to understanding their properties and to complementing the few studies that were devoted to the structural analysis of KBi.\textsuperscript{1,54,61} From these studies, idealized structure models have been determined for KBi. From the Rietveld refinement of the structure of KBi obtained by thermal decomposition of KMnO$_4$ at 800°C, Kim et al. concluded that their sample has a $P\bar{6}_3/mmc$ space group with a two layer periodicity, and chemical formula K$_{0.30}$MnO$_{2.11}$$\cdot$0.60 H$_2$O.\textsuperscript{1} The two-layer periodicity along the $c$ axis results from the regular alternation of octahedral layers rotated with respect to each other by 180° and related by a mirror plane. However, Kim et al.\textsuperscript{1} confined their attention "to obtain an approximate and simple structural model (space group, unit cell and layer structural model)" and many details of KBi actual crystal structure are still missing.

The present and the companion articles\textsuperscript{71,72} intend to provide new insight on the structure of high-temperature KBi varieties synthesized according to the protocol of Kim et al.\textsuperscript{1} These contributions describe the main crystal chemical features of KBi samples obtained at 700°C, 800°C and 1000°C and demonstrate the key influence of temperature on the KBi structure. For example, KBi samples synthesized at different temperatures differ by their i) layer
sequences, ii) sub- and super-cell parameters, iii) distributions of heterovalent Mn, and iv) distributions of interlayer K and vacant layer octahedra.

In the present article, the crystal structure of a KBi sample obtained at 800°C is studied by X-ray diffraction (XRD), X-ray absorption near edge structure (XANES) and extended X-ray absorption fine structure (EXAFS) spectroscopies, and thermal and chemical analyses. In contrast to previous structural characterizations of natural and synthetic birnessites, the structure refinement of our fine-grained KBi sample was performed for the first time on a single crystal using X-ray micro-diffraction, recently available on synchrotron radiation sources. KBi single crystal structural study was complemented by the simulation of the powder XRD pattern of the bulk sample. Powder diffraction confirmed the structure model of defect-free crystals and revealed the admixture of a defective KBi variety. The nature and content of stacking faults, which are common in birnessites, in this ancillary component were determined using a trial-and-error approach.

**Experimental Section**

**Experimental methods.** KBi sample was prepared by thermal decomposition of a fine-grained KMnO₄ powder (particle size < 50µm) following the procedure of Kim et al. Structural homogeneity of synthesis products was maximized by using flat crucibles containing a thin layer of KMnO₄ powder. Heating and cooling rates were set to 1°C per minute, pyrolysis occurring in the solid state for 5 hours in air, at 800°C. After decomposition and cooling, the final product was washed several times with bidistilled water (Milli-Q / 18.2 MΩ.cm⁻¹) to reduce pH from 12-13 to about 9-10 and to remove soluble by-products such as K₂MnO₄ and K₃MnO₄. The black crystals of this KBi sample synthesized at 800°C, and
hereafter referred to as KBi₈, were then freeze-dried and stored under room conditions for chemical and structural characterizations.

Morphology of KBi₈ particles was examined by scanning electron microscopy (SEM) on a JEOL JSM 6320F high resolution SEM equipped with a field emission electron gun. KBi₈ consists predominantly of large micrometer-sized crystals exhibiting well defined crystallographic faces (Figure 1). The average dimensions of the monocrysztals are 2 µm along the c axis and 6-10 µm within the ab plane, whereas the size of natural and synthetic low-temperature birnessites does not usually exceed 0.02 µm along the c axis and 1 µm within the ab plane.

Thermal analysis was carried out with a Netzsch Simultan STA409EP micro-analyzer with a heating rate of 10°C/min up to 1100°C. The weight loss due to structural water was measured by thermo-gravimetric analysis (TGA) and differential scanning calorimetry (DSC) on ~20 mg samples.

Total K and Mn contents were determined using a Perkin-Elmer Optima 3000 ICP-AES after digestion of about 8 mg of birnessite powder in 200 mL of 1% HNO₃ / 0.1% NH₃OH.Cl matrix. The mean oxidation state of manganese in birnessite was determined by potentiometric titration using (NH₄)₂Fe(SO₄) Mohr salt and sodium pyrophosphate.

KBi chemical and structural formulae. The chemical formula of KBi obtained by combining the K/Mn ratio (y) and the mean oxidation state, 2x, may be written as KₓMnO₇, where w = (2x + y) / 2. If m, n, p and q represent the amounts of Mn⁴⁺, Mn³⁺, Mn²⁺ and vacant layer sites, respectively, per octahedron these parameters are related by the following equations: q = 1 - 2 / w, m + n + p = 1, and 2p + 3n + 4m = 2x. Because the initial oxidation state of Mn in KMnO₄ was +7, it is reasonable to assume that KBi₈ sample does not contain Mn²⁺ (p = 0). Experimental support to this hypothesis is given in the XANES Results section. Accordingly, KBi₈ chemical formula may be written as:
\[ K^{+}_y Mn^{4+}_{2x-3} Mn^{3+}_{4-2x} O_{(2x+y)/2} \]  \hspace{1cm} (1)

which can be transformed into the structural formula:

\[ K^{+}_{2y/w}(Mn^{4+}_{(4x-6)/w} Mn^{3+}_{8-4x/w})_{1-2/w} O_2 \]  \hspace{1cm} (2)

Finally, the contents of structural water and hydroxyl groups can be added to this structural formula from the amounts of water released during thermal treatment.

**X-ray diffraction.** Data collection - KBi₈ monocrystal. Microdiffraction patterns on a KBi₈ monocrystal were recorded at the Microfocus beamline ID13 (European Synchrotron Radiation Facility - ESRF, Grenoble) using the ID13 microdiffractometer.⁷⁵,⁷⁶ The X-ray beam produced using an undulator is monochromatized by a Si(111) double monochromator, focused by an ellipsoidal mirror, and finally collimated to ~4 μm in diameter.

KBi₈ powder was first dispersed onto a kapton foil and diffraction images were collected from several individual grains to select the best candidates for a full data collection procedure. Selected crystals were then mounted on a very fine borosilicate glass fiber. The combined use of an in-line microscope and of fine-tune motors allows one to center the selected crystal on the rotation axis within ~1 μm. As a second step, visualization of the focal spot on a scintillator using the same microscope allows one to put the selected crystal exactly into the beam.

Complete data collection was then performed on a ~3 x ~3 x ~1 μm³ crystal with the oscillation technique (φ scans), using a two-dimensional mar CCD detector (~130mm diameter, 2048 x 2048 pixels and a 0.06445 x 0.06445 mm² pixel size) and an Oxford Cryostream device to keep the sample temperature at ~100 K. Data sets were recorded for two detector positions (0° and 40° tilts), first with the glass fiber axis approximately parallel to the sample rotation axis and then with the glass fiber mounted in a special sample holder allowing a 30° tilt with respect to the sample rotation axis. All four data sets were recorded (20-30s counting times per frame) with a 6° oscillation range and a sample-detector distance of ~42
mm (0° detector tilt) or ~ 73 mm (40° detector tilt). A minimum of 72 frames (446° sample rotation) were collected for each data set. The detector was tilted to 40° in the vertical plane to reach a ~0.55 Å resolution at λ = 0.729 Å. High-resolution was necessary because of the limited number of unique reflections in the space group previously reported for KBi.

Structure solution and refinement. Recorded frames were indexed and the reflections integrated using the XDS software suite. Lattice parameters (a = 2.840(1) Å, c = 14.030(6) Å) were refined from 1523 reflections with I > 6σ(I). Even though minimum because of the small crystal size, absorption correction was applied to obtain similar integrated intensities from symmetry-equivalent reflections. A total of 6727 reflections were recorded for a maximum resolution of 2θ = 82.8°. Finally, the structure model was refined from 153 unique reflections (out of 155 possible; completeness of 98.7%) with SHELXL-97.

Data collection - KBi powder. Powder XRD pattern was recorded using a Bruker D5000 powder diffractometer equipped with a Kevex Si(Li) solid state detector and CuKα1+2 radiation. Intensities were recorded at a 0.04° 2θ interval, from 5 to 90°, using a 40 sec counting time per step. A rotating sample holder was used to minimize the effect of preferential orientation.

Simulation of powder XRD patterns. One of the effective ways to determine the actual structure of layered minerals, and more especially of defective ones, is the calculation of powder XRD patterns using the mathematical formalism described by Drits & Tchoubar combined to a trial-and-error fitting procedure. This method has been successfully used to determine the structure of natural and synthetic birnessites. Details on the program used to simulate XRD patterns and on the fitting procedure are given by Drits et al.

Calculations were restricted to 10θ reflections (35°-65° 2θ CuKα range) because these lines are most sensitive to some important structural parameters of layered minerals, including order/disorder in their stacking sequences, and site occupancies in both layer and interlayer.
The background was assumed to be linearly decreasing over the angular range considered, and preferred orientation of particles was considered as a variable parameter. Quality of fit was estimated over the simulation range using the usual $R_{wp}$ parameter.

**XANES and EXAFS spectroscopy.** Mn K-edge XANES and EXAFS spectra were collected at the LURE synchrotron radiation laboratory (Orsay – France) on the D42 spectrometer. The positron energy of DCI storage ring was 1.85 GeV and the current between 200 and 300 mA. The incident X-ray beam was monochromatized with a channel-cut Si(331) crystal. The absolute energy was pinned to the first inflection point of pure Mn (6539 eV). XANES spectra were normalized to unit step in the absorption coefficient from well below to well above the edge. EXAFS spectra were recorded over the 6400-7300 eV range. Samples were prepared for transmission measurement as homogeneous thin mounts having an absorption edge jump lower than 1.0. They were oriented to the magic angle in the X-ray beam to eliminate any texture effects originating from possible preferential orientation of KBi crystals.

Experimental data were analyzed using the WinXAS software package. Radial structure functions (RSFs) were calculated using a Bessel function to minimize the intensity of side lobes resulting from truncation effects in Fourier transforms of EXAFS spectra. Manceau showed that the intensity of side lobes with this function represents about 5% of the intensity of main structural peaks. This analytical treatment enhances the sensitivity to weak EXAFS contributions and, in the case of phyllosilicates, allows one to detect low amounts of Mn-Mn pairs from corner-sharing octahedra (i.e. interlayer Mn cations located above/below vacant layer octahedra). Mn-O and Mn-Mn distances and the number of atoms in nearest O (CN$_O$) and Mn (CN$_{Mn}$) coordination shells were determined using experimental phase shift and amplitude function calculated from stoichiometric $\lambda$-MnO$_2$, in which Mn$^{4+}$ is surrounded by six O atoms at 1.91 Å and six nearest Mn atoms at 2.84 Å. Typical uncertainty on interatomic distances is ±0.02 Å and that on coordination numbers is ±1.0. RSFs are not
corrected for EXAFS phase-shifts, causing peaks to appear at shorter distances \((R + \Delta R, \text{with} \Delta R \sim -0.4 \text{ Å})\) relative to the true near-neighbor distances \((R)\).

**Results**

**Structural formula.** DTA curve of KB\(_{18}\) shows two sharp endotherms at 148°C and 925°C (Figure 2). The low-temperature one corresponds to the loss of weight related to interlayer H\(_2\)O, the amount of which is 6.7%, whereas the high-temperature feature possibly results from the reduction of Mn\(^{4+}\) cations, which are stabilized by the presence of interlayer K cations. The DTA curve also contains a small endotherm at \(~370^\circ\text{C}\) probably related to the dehydroxylation of layer OH groups and corresponding to the departure of 1.9% H\(_2\)O (Table 1). The loss of weight within the 60-120°C range is likely due to the departure of adsorbed water.

Chemical analyses showed that the atomic K/Mn ratio calculated from ICP data and the mean Mn oxidation state are equal to 0.24 and 3.92, respectively (Table 1). Using these values and the structural water weight losses the following structural formula is deduced:

\[
\text{K}^{+}\cdot\text{Mn}^{4+}_{0.885}\cdot\text{Mn}^{3+}_{0.077}\cdot\text{H}^{0.038}\cdot\text{O}_{2}^{2}\cdot 0.60\text{H}^{2}\text{O} \tag{3}
\]

**Structure refinement of a KB\(_{18}\) microcrystal.** Since our KB\(_{18}\) sample has similar powder XRD pattern and unit-cell parameters as those of the sample synthesized by Kim et al.,\(^1\) the same two-layer structure model can be hypothesized. An idealized scheme of this model is drawn in Figure 3. The close-packing notation for this model is:

\[
\text{AbC a’ CbA c’ AbC…}
\]

where A and C represent the positions of layer oxygen atoms (O\(_{\text{layer}}\)), b the Mn\(_{\text{layer}}\) positions, and a’ and c’ the positions of interlayer K cations which are located above and
below empty tridentate layer cavities in the interlayer prisms formed by the $O_{layer}$ from adjacent layers (irregular dashed line - Figure 3).

The origin of the unit cell is assumed to be in the b site. Because of the hexagonal layer symmetry, $O_{layer}$ in the ab plane form equilateral triangles with their side lengths equal to the $a$ parameter. Triangles formed by $O_{layer}$ in the A and C sites are rotated around the c axis by $(2n+1)60^\circ$. Therefore, adjacent layers in the structure should be related by a 63 axis passing through $Mn_{layer}$, by a mirror plane $m$ passing through the c axis and the long diagonal of the layer unit-cell, and by a glide plane $c$ passing through the short diagonal of the layer unit-cell. In addition, the structural model may have centers of symmetry located for example in the b sites and a mirror plane $m$ parallel to the ab plane and passing through interlayer K. After integration of recorded intensities, possible space groups were $P3_{1}c$, $P-3_{1}c$, $P6_{3}mc$, $P-6_{2}c$, and $P6_{3}/mmc$ in agreement with the absence of reflections with $hh\bar{2}h\ell$ in the experimental XRD pattern. The structure model was first refined using the most symmetrical $P6_{3}/mmc$ space group in agreement with the structure model proposed. The other four space groups were subsequently assessed and lead to structure models similar to that obtained using $P6_{3}/mmc$ space group. In particular, for the 3 non-centrosymmetric space groups it was impossible to reject the presence of a center of symmetry and the high symmetry $P6_{3}/mmc$ space group was thus preferred.

The structure model proposed by Kim et al. was used to start our refinement. In this model all atoms, except $O_{layer}$, occupy special positions with $Mn_{layer}$ and K located in the 2a (0,0,0) and 2c (1/3,2/3,1/4) positions, respectively. Only the z-coordinate of $O_{layer}$, which is located in the 4f site (2/3,1/3,z), needed to be refined in this initial model. It was first assumed that the layer thickness is 2.00 Å ($z_{O_{layer}} = 0.070$) as determined for different synthetic birnessites. Occupancies of $Mn_{layer}$ (0.96), $O_{layer}$ (2.00) and K (0.24) were set
according to the structural formula determined for KBi₈ (Eq. 3). After this initial refinement, 
R₁ was equal to 8.9% (GoF = 2.74) for strong reflections [Fo > 4σ(Fo)].

After this initial attempt, the strongest peak in the difference-Fourier map lay in the
interlayer region with coordinates (0,0,0.1523). The distance from this position to the three
nearest O₈ layer is 2.00 Å, this distance being similar to the Mn₈ layer-O₈ layer bond length. From the
chemical composition of KBi₈, the additional peak was assigned unambiguously to interlayer
Mn cations (Mninterlayer). To provide an octahedral coordination to Mninterlayer water molecules
were introduced in the interlayer space in 2c sites (1/3,2/3,1/4), like interlayer K, the resulting
< H₂O-Mninterlayer> being 2.14 Å. The resulting average distance between Mninterlayer and
coordinated oxygen atoms (O₈ layer and H₂O - 2.07 Å) is compatible with <Mn³⁺-O> reported in
Mn³⁺-bearing layer Mn oxides and oxi-hydroxides (2.040-2.045 Å). It is likely that
Mninterlayer results from the migration of Mn³⁺layer into the interlayer. This migration should be
accompanied by the formation of the same amount of vacant layer octahedra capped by
Mninterlayer as evidenced by the distance from the Mninterlayer site to the (0,0,0) position
(2.14 Å), which is unrealistically short if the two sites are occupied simultaneously. From
KBi₈ structural formula (Eq. 3), it is possible to deduce the number of vacant layer sites (0.12
per octahedron), and that of Mninterlayer and of associated H₂O (0.08 and 0.24, respectively).
The new structural formula writes:

\[ K^+_{0.231} Mn^{3+}_{0.077} (Mn^{4+}_{0.885} O_{0.115}) O_2 \cdot 0.60 H_2O \]  

(4)

Including these two interlayer positions into the structure refinement significantly
improved the quality of fit (R₁ = 8.2% - GoF = 1.22). From DTA results, the amount of H₂O
is 0.60 per octahedron. 0.24 of these H₂O molecules are coordinated to Mninterlayer, and it is
likely that the remaining 0.36 H₂O molecules are coordinated to K. These latter H₂O
molecules will be hereafter referred to as O3, whereas O₈ layer and H₂O molecules bond to
Mninterlayer will be referred to as O1 and O2, respectively. Similarly, Mnlayer and Mninterlayer will
be referred to as Mn1 and Mn2. Initial location of O3 sites was determined from the strongest peak observed at this stage in the difference-Fourier map (0.2054,0.1027,1/4 - Site 6h) which lay in the interlayer region. In addition, interlayer K was moved from its original 2c position to a 6h one because of its unrealistically high thermal displacement factors (Ueq = 0.18). K1 and O3 positions were further refined and R$_1$ parameter was decreased to 5.2% (GoF = 0.78).

Further improvement of the fit was achieved by introducing anisotropic thermal displacement parameters for Mn1 and K. As a result, R$_1$ decreased to 2.75% (GoF = 1.09) for the 137 strong reflections [Fo > 4σ(Fo)], the value of this parameter being 3.3% when taking into account all 153 reflections. In the final step, 14 parameters were refined. Refined structural parameters of KBi$_8$ are listed in Table 2, atomic positions of layer and interlayer species are schematized in Figure 4, and selected interatomic distances are given in Table 3.

**Indexing and simulation of the powder XRD pattern.** Because it was difficult to find a monocrystal suitable to carry out the above structure refinement, that is a crystal showing defined diffraction spots at very high resolution, one may wonder about the representativity of the selected crystal. In addition to assessing the representativity of the studied monocrystal, and to characterizing the sample heterogeneity, a complementary structural study by powder XRD was undertaken on the same sample to determine the nature and content of stacking faults which are common in birnessites. Reflections of KBi$_8$ powder XRD pattern (Figure 5) were indexed using a two layer hexagonal (2H) unit-cell with $a = 2.845$ Å, $c = 14.235$ Å, $\alpha = \beta = 90^\circ$, and $\gamma = 120^\circ$. These parameters are similar to those obtained by Kim et al. ($a = 2.842$ Å, $c = 14.16$ Å, $\alpha = \beta = 90^\circ$, and $\gamma = 120^\circ$), and are consistent with our structure refinement.\textsuperscript{1}

Even though the experimental XRD pattern recorded for KBi$_8$ exhibits a low background intensity and sharp reflections, the bulk structure of this sample was not refined using the Rietveld technique but rather determined by a trial-and-error fitting procedure on the powder
XRD pattern. Indeed, the presence of significant and asymmetrically broadened tails near the 10ℓ reflections of the main KBi₈ phase impeded the use of the Rietveld method (Figures 5, 6). This broadening is most likely induced by the presence in the sample of a second birnessite-like phase having a disordered structure. As a consequence, the complete structure determination should include this disordered KBi₈d phase in addition to the main ordered phase (KBi₈o).

**Ordered KBi₈o phase.** The atomic positions refined on the microcrystal (Table 2) were used as initial values for the simulation of the XRD pattern, and site occupancies were derived from the structural formula of KBi₈ (Eq. 4). Coherent scattering domains were supposed to have a disk-like shape with the mean disk radius equal to 550 Å, whereas mean and maximum numbers of layers in the disks were equal to 20 and 60, respectively. These parameters yielded a satisfactory agreement between experimental and calculated profiles (data not shown – R_wp = 12.7%). Optimal fit to the powder diffraction data (Figure 6a – R_wp = 11.8% - R_EXP = 3.0%) was obtained by slightly adjusting atomic positions. Note that the atomic positions obtained from this trial-and-error fit to experimental data (Table 4) are close to those obtained from the refinement on KBi₈ monocystal.

Even though the KBi₈d contribution to the diffracted intensity is not negligible (Figure 6b), the trial-and-error procedure provides reasonable constraints on the KBi₈o structure model. For example, if K is located in the center of the interlayer prisms (1/3,2/3,1/4), as proposed by Kim et al., intensities calculated for 100 and 104 reflections of KBi₈o are significantly lower than experimental ones (Figure 7a – R_wp = 14.2%). In contrast, intensities of these two reflections are significantly higher than experimental ones if K is located in the center of the prismatic faces (Figure 7b – R_wp = 16.7%), or on the edges of these prisms (not shown).
In addition, the intensity of 103 and 106 reflections strongly depends on the number of interlayer Mn cations, and these reflections may be used to constrain the distribution of Mn$^{3+}$ between layer (Mn1) and interlayer (Mn2) sites. If the calculation is performed assuming only 0.04 vacant layer site per octahedron (Eq. 3), the intensities calculated for 103 and 106 reflections are much lower than experimental ones (Figure 7c – \(R_{wp} = 18.1\%\)), the best distribution of intensities being obtained for the cation distribution from Eq. (4) as for KBi$_8$ monocrystal.

**Disordered KBi$_{8d}$ phase.** To complete the structural determination of the bulk KBi$_8$ sample, special attention was paid to the characterization of the disordered KBi$_{8d}$ phase. Because well-defined stacking faults are common in natural and synthetic birnessites,$^{63,65,80}$ diffraction features resulting from the random interstratification of similar layers with different interlayer displacements were considered. Introduction of such well-defined stacking faults in the simulation significantly broadens \(h\ell k\ell\) reflections, and shifts them from their nominal positions.$^{89}$ The first stacking mode considered was identical to that in the ordered KBi$_{8o}$ structure (2\(H\)), adjacent layers being rotated with respect to each other by 180°. In the second stacking mode, adjacent layers had the same orientation and were shifted with respect to each other by one third of the long diagonal of the layer unit cell. Following this translation, the origin of the upper layer had coordinates (1/3, 2/3) in the \(ab\) plane relative to the origin of the lower layer. Such a layer sequence corresponds to a three-layer rhombohedral (3\(R\)) periodicity, the close-packing notation of which is:

\[
^c\text{Ab}^a\text{C}_a^b\text{BaC}_c^a\text{BeA}_b^c
\]

where A,B,C and a,b,c, correspond to the O$_{layer}$ and Mn$_{layer}$ positions, respectively, whereas \(a'\), \(b'\), \(c'\) are the positions for interlayer K. Such a KBi phase with a 3\(R\) stacking sequence was obtained hydrothermally.$^{54,72}$ In randomly interstratified 2\(H\)/3\(R\) structures, 10\(\ell\)/01\(\ell\) reflections should be located between neighboring 10\(\ell\)/01\(\ell\) reflections of individual
2H and 3R phases (Figure 6c), the actual position depending on the relative proportions of each stacking sequence.\textsuperscript{89} Diffraction patterns calculated for such mixed layered structures allowed us to reproduce the broad asymmetric tails near the 10\(\ell\) reflections of the KBi\textsubscript{8o} pattern. The optimum fit to the experimental KBi\textsubscript{8} XRD profile shown on Figure 6a includes the contribution of a randomly interstratified phase containing 60\% of 2H fragments and 40\% of 3R layer pairs (Figure 6c) in addition to the contribution of the ordered KBi\textsubscript{8o} phase (Figure 6b). Composition and structure of layers and associated interlayers were assumed to be identical in both 2H and 3R layer fragments.

**XANES spectroscopy.** The position in energy and the shape of the Mn K-edge XANES spectra are characteristic of Mn valency as illustrated in Figure 8a by the comparison of Mn\textsuperscript{2+}, Mn\textsuperscript{3+}, and Mn\textsuperscript{4+}-reference compounds. The position of the main absorption edge occurs at \(\sim6562\) eV for tetravalent Mn, 6559-6560 eV for trivalent Mn, and 6551-6552 eV for divalent Mn. Since XANES spectra of mixed Mn-valency compounds are weighted average of these elementary contributions this spectroscopy should provide additional insight about the oxidation state of Mn in KBi\textsubscript{8}. Note, however, that Mn\textsuperscript{3+} spectra have a hump at 6550-6551 eV, which should not attributed to the presence of Mn\textsuperscript{2+} impurity.

The structures of synthetic triclinic Na-birnessite and hexagonal H-birnessite, hereafter referred to as NaBi and HBi, have been determined by X-ray and electron diffraction,\textsuperscript{5,65,90} and EXAFS spectroscopy.\textsuperscript{91} NaBi has a vacancy-free layer structure containing 69\% Mn\textsuperscript{4+} and 31\% Mn\textsuperscript{3+}, whereas the ideal cation composition of HBi at pH 5 is

\[
\text{Mn}\textsuperscript{2+}0.055\text{Mn}\textsuperscript{3+}0.233\text{Mn}\textsuperscript{4+}0.722 \text{ per octahedron.} \textsuperscript{65}
\]

Accordingly, the Mn absorption edge spectrum of NaBi (Figure 8b – solid triangles) is clearly left-shifted relative to the ideal pure Mn\textsuperscript{4+} (Figure 8b – dotted-dashed line). In spite of the large proportion of Mn\textsuperscript{3+} cations in the structure (0.31 per octahedron), no shoulder is visible at 6550-6551 eV. The edge crest for HBi-pH5 (Figure 8b – crosses) is slightly right-shifted relative to NaBi closer to 6562 eV, in
agreement with the higher proportion of Mn$^{4+}$ cations in the hexagonal form. Two shoulders are now visible, one at 6559 eV in agreement with the presence of Mn$^{3+}$ cations, and another at 6551-6552 eV which may be attributed to Mn$^{2+}$ owing to the absence of this feature in NaBi. The reduced normalized intensity of the edge resonance in HBi is also consistent with a higher mixed-valency of Mn in this sample.

It is then possible to assess Mn valency in KBi$_8$ by comparison with the Mn absorption edges exhibited by these two birnessite samples. First, in KBi$_8$, the position of the absorption edge (6562 eV) and the high normalized absorbance at the edge crest plead for a high proportion of Mn$^{4+}$ in agreement with the average oxidation state of Mn measured by titration in this sample (3.92). Also, no hump is observed at ~6552 eV as expected if Mn$^{2+}$ cations were present in significant amount. On the basis of titration measurement, the maximum proportion of Mn$^{2+}$ is 0.04. Since this value is close to that in HBi-pH5 (0.05), which exhibits a shoulder at ~6552 eV, we can conclude that Mn$^{3+}$ is the only low-valent form of Mn atom in KBi$_8$.

**EXAFS spectroscopy.** EXAFS data for KBi$_8$ are first compared to those of NaBi and HBi. Then, interatomic distances and number of atoms in nearest shells around Mn atoms are calculated from the simulation of EXAFS data.

In the NaBi structure Mn$^{3+}$ cations are segregated in rows along the b axis, each Mn$^{3+}$ row alternating with two Mn$^{4+}$ rows in the perpendicular direction. Mn$^{3+}$ octahedra are elongated because of the Jahn-Teller distortion and have a unique azimutal orientation resulting in an orthogonal layer symmetry and, at the local scale, in a distribution of Mn-Mn distances across octahedral edges (Table 5).$^{91}$ The ideal structural formula of HBi at pH 4 is $\text{H}^+\cdot_{0.333}\text{Mn}^{3+}\cdot_{0.123}\text{Mn}^{2+}\cdot_{0.043}(\text{Mn}^{4+}\cdot_{0.722}\text{Mn}^{3+}\cdot_{0.111}[\cdot_{0.167}]\text{O}_2(\text{OH})\cdot_{0.013}\cdot_{0.49}\text{H}_2\text{O},^{65}$ where cations within the parentheses are in the layer and those outside are in the interlayer. Interlayer Mn$^{3+}$ results from the migration of layer Mn$^{3+}$ during the NaBi-to-HBi transformation at low pH.
As a result, Mn-Mn distances in the layer are less scattered in HBi than in NaBi, and the presence of octahedrally coordinated Mn$^{2+}$ and Mn$^{3+}$ above/below vacant layer sites leads to the presence of a corner-sharing Mn-Mn shell at about 3.52 Å (Table 5).

These structural differences between NaBi and HBi are well apparent on EXAFS spectra and RSFs (Figure 9a). Comparison of the modulus and imaginary parts of the Fourier transforms of NaBi and HBi indicates that the nearest Mn-Mn shell at $R + \Delta R = 2.5$ Å, corresponding to edge-sharing contributions, is significantly expanded in NaBi because of the higher amount of layer Mn$^{3+}$, whereas HBi possesses a characteristic peak at $R + \Delta R = 3.10$ Å from the corner-sharing Mn-Mn shell. Another difference between NaBi and HBi is the decrease of the amplitude of the first RSF peak in HBi due to the wider distribution of Mn-O distances caused by the mix of Mn atoms in layer and interlayer sites and their coordination to only O ligands in NaBi and to O, OH and H$_2$O ligands in HBi. EXAFS spectra and RSFs obtained for KBi$_8$ and HBi are essentially the same, indicating that Mn has a similar local structure in the two phyllomanganates. The two spectra have the same frequency but a significantly distinct shape in the 4-7 Å$^{-1}$ range (arrows - Figure 9b). This observation suggests that the mean Mn-O and Mn-Mn distances are the same in the two structures, but that Mn has a slightly different coordination chemistry. The comparison of RSFs shows that the first peak has a higher amplitude and the third peak a lower amplitude in KBi$_8$ compared to HBi. More coherent Mn-O distances account for the enhancement of the first RSF peak whereas the decreased number of Mn$_{\text{layer}}$-Mn$_{\text{interlayer}}$ pairs, which is consistent with a lower proportion of Mn$^{3+}$ in the interlayer of KBi$_8$ as indicated by XRD, explains the reduction of the third RSF peak.

Results from the least-squares fits of the contribution to EXAFS spectra of Mn-Mn shells for KBi$_8$ are compared to NaBi and HBi in Table 5. Good agreement between theory and data over the fit range was obtained assuming a single shell at 2.87 Å for KBi$_8$, and two shells at
2.88 Å and 3.52 Å for HBi. The short distance shell corresponds to \( \text{Mn}_{\text{layer}} - \text{Mn}_{\text{layer}} \) pairs and the higher distance to \( \text{Mn}_{\text{layer}} - \text{Mn}_{\text{interlayer}} \) pairs. The addition of a \( \text{Mn}_{\text{layer}} - \text{Mn}_{\text{interlayer}} \) shell in the simulation for KBi\(_8\) did not improve statistically the spectral fit, although XRD results and the examination of RSFs gave supportive evidence for its existence. The average number of \( \text{Mn}_{\text{layer}} - \text{Mn}_{\text{interlayer}} \) pairs can be calculated from the structural formula as the weighted sum of the different coordination environments:

\[
\text{CN}_{\text{corner}} = \sum_i W_i \text{CN}_i, \tag{5}
\]

where \( i \) refers to Mn site, \( W_i \) to Mn site occupancy, and \( \text{CN}_i \) to the number of Mn neighbors for site \( i \). For KBi\(_8\) (Eq. 4), \( \text{CN}_{\text{corner}} = 0.885 \times (0.077/0.115) + 0.077 \times 6.0 = 1.1 \), and for HBi, \( \text{CN}_{\text{corner}} = 0.833 \times 1.0 + 0.167 \times 6.0 = 1.8 \). The former value is equal to the detection limit of corner-sharing Mn octahedra in phyllomanganates by EXAFS spectroscopy,\(^{40}\) and the latter almost twice as high.

**Discussion**

The structure model proposed for KBi\(_8\) differs substantially from that refined by Kim et al.\(^1\) First, in the new model interlayer K is not located in the center of a prismatic cavity (2/3,1/3,1/4) but in a split 6h position (0.227,0.113,1/4). Second, interlayers of KBi\(_8\) contain \( \text{Mn}^{3+} \) and \( \text{K}^{+} \) cations, instead of only \( \text{K}^{+} \) in the former model. Third, the manganese layer contains only \( \text{Mn}^{4+} \) cations and a significant amount of vacant sites. Finally, the positions of interlayer water molecules and their occupancies were identified. These structural features are discussed below.

**KBi\(_8\) layer.** As in other phyllomanganates, the length of shared edges (2.567 Å - all distances are from the single crystal refinement) is significantly shorter than that of unshared edges (2.840 Å). The Mn-O bond length (1.914 Å) is consistent with the sole presence of
Mn$^{4+}$ cations in KBi$_8$ layers as this value is typical for MnO$_2$ compounds. The bond-valence calculation presented in Table 6 is also consistent with this conclusion. As a result, vacant octahedra represent the sole source of layer charge deficit in KBi$_8$. The presence of vacant layer sites is supported by micro-crystal XRD structure refinement, trial-and-error simulation of powder XRD pattern and EXAFS spectroscopy. Given that in layered structures vacant and occupied octahedral sites have different size and shape, atomic positions listed in Table 2 are average values. This approximation may account for the small deviations observed between calculated and expected bond valence values.

**KBi$_8$ interlayer.** According to XRD and EXAFS results, Mn$_{\text{interlayer}}$ (Mn2) is located above/below vacant layer sites. The three O$_{\text{layer}}$ (O1) atoms from these vacant octahedra provide half of the Mn2 octahedral coordination which is completed by H$_2$O molecules (O2). Mn2 site is shifted from the center of the octahedron towards O1 atoms, leading to $\langle d(\text{Mn2-O1}) \rangle$ (1.989 Å) shorter than $\langle d(\text{Mn2-O2}) \rangle$ (2.152 Å). These distances are consistent with Mn$^{3+}$-O and Mn$^{3+}$-H$_2$O bond lengths, respectively, whereas they would be unrealistically short for Mn$^{2+}$-O and Mn$^{2+}$-H$_2$O distances, since Mn$^{2+}$ is ~0.1 Å bigger than Mn$^{3+}$. The mean Mn2-O distance (2.07 Å) is slightly larger than $\langle \text{Mn}^{3+}\text{-O} \rangle$ (2.04 Å) reported for Mn$^{3+}$-bearing oxides and oxi-hydroxides. In addition, interlayer O2 atoms form weak H-bonds with nearest O$_{\text{layer}}$ atoms from the adjacent layer ($\langle d(\text{O1-O2}) \rangle$ across the interlayer equals 3.006 Å - Table 3).

Table 6 shows that Mn$^{3+}$ cations in Mn2 sites are slightly undersaturated as they receive only 2.66 valence unit (v.u.). As a consequence, an additional refinement step was performed by setting the O2 site in a 6h position. The refined O2 position (0.212, 0.606, 1/4) results in a shorter Mn2-O2 distance (2.060 Å) and thus provides a better charge compensation to Mn$^{3+}$ cations, which then receive 2.94 v.u. However, this additional step did not improve the quality of fit (R$_1 = 2.74\%$ - GoF = 1.12). In addition, all attempts to describe the Jahn-Teller
distortion of Mn\(^{3+}\) octahedra by setting the Mn2 site, and possibly the O2 one, in a 12k position failed. This may be due to the random orientation of the elongated diagonal of the distorted Mn\(^{3+}\) octahedra and, consequently Mn2 and O2 positions are averaged in Table 2. Bond-valence calculations performed assuming the presence of Mn\(^{2+}\) in the Mn2 site lead to unrealistically high positive charge values (3 x 0.584 + 3 x 0.376 = 2.88 v.u.). The absence of Mn\(^{2+}\) inferred from bond valence calculations are in complete agreement with the Mn valency determined by XANES spectroscopy.

As can be seen in Figure 10, within the prisms formed by O\(_{\text{layer}}\) atoms from adjacent layers and topped on either side by empty tridentate cavities, K occupies with equal probability one of the three possible sites shown on Figure 4a. With respect to the position given by Kim et al. these three sites are shifted towards the edges of the prism, staying at equal distance from the two adjacent layers (z = 1/4).\(^1\) In this site, interlayer K is coordinated to the four O\(_{\text{layer}}\) atoms from the nearest face of the prism and to two H\(_2\)O molecules located in O3 sites, these six atoms defining a distorted octahedron (Figures 4, 10). <K-O1> and <K-O3> distances are equal to 2.907 Å and 3.250-3.428 Å, respectively, and their average (3.05 Å) is typical for K in six-fold coordination.\(^9\) The bond-valence calculation (Table 6) shows that this first octahedral shell compensates only part of the K\(^+\) charge (0.57 v.u.), the remainder being partially compensated by the next-nearest O-shell (0.19 v.u.). The short distance (2.743 Å) between H\(_2\)O molecules located in the O3 site and the nearest O\(_{\text{layer}}\) allows for the formation of strong H-bonds, with a favorable O1-O3-O1 angle of \(\sim 133^\circ\) (Figure 4).

**Migration of Mn\(^{3+}\) and origin of vacant layer sites.** Structure models of synthetic birnessite-like species obtained at high pH and containing contrasting proportions of Mn\(^{3+}\) indicate that the behavior of this cation depends on its concentration within the Mn layer. NaBi,\(^5,90\) CaBi,\(^80\) and KBi\(_{10}\) obtained from thermal decomposition of KMnO\(_4\) at 1000°C\(^71\) have high concentrations of Mn\(^{3+}\) (\(\sim 1/3\) of total Mn). In these three species, heterovalent
Mn\text{layer} cations are segregated in Mn\text{3+}- and Mn\text{4+}-rich rows, and all Mn\text{3+}-octahedra, which are elongated because of the Jahn-Teller distortion, present a unique azimuthal orientation. These two features reduce the steric strains that would result from the random distribution of Mn\text{3+}-octahedra, and lead to the departure from the hexagonal layer symmetry.

In contrast to NaBi, CaBi and KBi\text{10}, KBi\text{8} contains only 0.08 Mn\text{3+} cations per octahedron. It is thus likely that the local strains resulting from the distribution of Mn\text{3+} as isolated cations compel these cations to migrate from layer to interlayer. The pure Mn\text{4+}-containing layers are then devoid of strains and have a hexagonal symmetry. The presence of both K and Mn\text{3+} in interlayers contributes to the cohesion between successive KBi\text{8} layers.

**Local distribution of interlayer Mn and K.** The presence of vacant layer sites in KBi\text{8} results in an uneven distribution of O\text{layer} charges (Table 6) which is likely responsible for the shift of interlayer K off the center of interlayer prismatic cavities. Each of the three O\text{layer} forming the roof of a vacant layer octahedron are coordinated to only two Mn\text{4+}\text{layer} cations (Figure 10 – small open circles). These O\text{layer} atoms receive $0.647 \times 2 = 1.29$ v.u., whereas O\text{layer} coordinated to three Mn\text{4+} receives $3 \times 0.647 = 1.94$ v.u. (Figure 10 – small solid circles). On one side of the vacant layer sites, the charge deficit of O\text{layer} atoms is partly compensated by the presence of Mn\text{interlayer} as each O\text{layer} is coordinated to two Mn\text{4+}\text{layer} and one Mn\text{3+}\text{interlayer}, thus receiving $0.647 \times 2 + 0.539 = 1.83$ v.u. If no Mn\text{interlayer} is present, the neighboring prismatic cavity shares two edges with occupied Mn\text{layer} sites and one with the vacant layer site. As a result, K hosted in these prismatic sites will likely shift towards the undersaturated edge to balance partially the charge deficit of the two undersaturated O\text{layer}. According to our results this charge balance mechanism provides 0.123 v.u. to each O\text{layer} (Table 6). As each vacant octahedron shares undersaturated edges with three prisms on either side of the layer, it is probable that each vacant site is neighbored by one K\text{+} cation on each side of the layer, if not capped on one side by Mn\text{interlayer}. However, one may note that O\text{layer}
atoms coordinated by two Mn$_{\text{layer}}$ and one interlayer K remain strongly undersaturated (1.33-1.42 v.u.) and they likely hold a proton.

From the previous considerations, it is possible to estimate the number of K connected to vacant layer sites. First, 0.08 K should be bond to the 0.04 vacant octahedra per unit cell which are not capped by Mn$_{\text{interlayer}}$ (Eq. 4). Second, remaining vacant layer sites (0.08 per unit cell) share three corners on the one side with Mn$_{\text{interlayer}}$ octahedra and are neighbored by one K on the other side (Figures 4b, 10). Therefore, the 0.12 vacant layer sites should be neighbored by $0.04 \times 2 + 0.08 = 0.16$ K. The remaining 0.08 K are distributed within interlayer prisms which are not associated with vacant layer octahedra. In this case, the reason for the shift in position of K atoms is unclear. Note however that in 3R layer pairs the center of a tridentate cavity from one layer is superimposed with the Mn1 site of the adjacent layer. As a consequence, in 3R layer pairs the shift of K would decrease the electrostatic repulsion between interlayer K and Mn$_{\text{layer}}$ from the next layer.

**Origin of KBi$_{8d}$ phase.** The interstratified 3R/2H phase likely results from an heterogeneous heat distribution within the KMnO$_4$ layer during the synthesis. Gaillot et al. indeed showed that temperature is the main factor controlling the formation of KBi polytypes and interstratified phases. Specifically, thermal decomposition of KMnO$_4$ at 700°C leads to the formation of 3R/2H phases in which 3R layer pairs prevail.

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analyses, respectively). We are grateful to the LURE and to the ESRF (D42 and ID13 beamlines, respectively) for the provision of beam time, and to Agnès Traverse for her assistance during the EXAFS data collection.

**Supporting Information Available:** An X-ray crystallographic file (CIF). This material is available free of charge via the Internet at http://pubs.acs.org.

**Figure captions**

**Figure 1.** Scanning electron micrograph of KBi$_8$ particles.

**Figure 2.** Experimental DTA-TG and DSC traces obtained for KBi$_8$.

**Figure 3.** Idealized structure model for KBi$_8$ (modified from Kim et al. $^1$). a) Projection on the $ab$ plane. The upper surface of the lower layer is shown as light shaded triangles. O$_{\text{layer}}$ and Mn$_{\text{layer}}$ of this lower layer are shown as small solid and open circles, respectively. Large shaded circles = Interlayer potassium. b) Projection along the $b$ axis. Open and solid symbols indicate atoms at $y = 0$, and at $y = \pm 1/2$, respectively. Large circles = O$_{\text{layer}}$ atoms, small circles = Mn$_{\text{layer}}$ atoms. Dotted-dashed lines outline the interlayer prisms defined by the two empty tridentate layer cavities. The center of these prisms is shown by regular dashed lines. c) Central projection. Dotted-dashed lines as in Figure 3b.

**Figure 4.** Structure model for KBi$_8$. a) Projection on the $ab$ plane. Symbols as for Figure 3a. Mn1 and Mn2 cations are shown as large open circles. Interlayer H$_2$O molecules (O2 and O3) are shown as large open circles with a dashed outline. b) Projection along the $b$ axis. Symbols as for Figure 3b. Large circles = Mn$_{\text{layer}}$ atoms, small circles = O$_{\text{layer}}$ atoms, large squares = vacant layer sites. Dotted-dashed lines outline the interlayer prisms defined by the two empty tridentate layer cavities. The center of these prisms is shown by dashed
lines, and the arrow outlines the shift of K cations from this ideal position. O1-O3-O1 angle (~133°) is outlined by a solid line.

**Figure 5.** Experimental powder XRD pattern obtained for KBi₈.

**Figure 6.** Comparison between experimental (crosses) and calculated (solid line) XRD patterns for KBi₈. Structural parameters used for calculation are listed in Table 4. **a)** Optimum model. The calculated pattern is the sum of a periodic KBi₈₀ phase and of a mixed layered KBi₈%d phase (R_wp = 11.8%). **b)** Contribution of the KBi₈₀ phase. Arrows indicate the asymmetrically broadened tails near the 10ℓ reflections of the KBi₈₀ phase. **c)** KBi₈%d is a mixed-layered structure containing 2H structural fragments (dashed line - similar to KBi₈₀) and 3R layer pairs (solid line). The random interstratification of these two types of layer pairs (2H:3R = 60:40) induces the shift of the KBi₈%d contribution from ideal 2H positions as indicated by horizontal arrows and leads to the diffraction pattern shown as a solid bold line.

**Figure 7.** Comparison between experimental and calculated XRD patterns for KBi₈. Patterns as for Figure 6. Arrows indicate the main misfits between experimental and calculated patterns. **a)** K is located in the center of the interlayer prisms (1/3, 2/3, 1/4) as suggested by Kim et al.¹ (R_wp = 14.2%). **b)** K is located on the faces (1/6, 5/6, 1/4) of the interlayer prisms (R_wp = 16.7%). **c)** All Mn atoms are located in layer octahedral sites resulting in a K⁺₀.₂₃₁(Mn⁴⁺₀.₈₈₅Mn³⁺₀.₀₇₇₋₀.₀₃₈)O₂·₀.₆₀H₂O structural formula (R_wp = 18.1%).

**Figure 8.** XANES spectra for reference compounds with a single Mn valency (Mn²⁺, Mn⁴⁺, and Mn³⁺), and for birnessite samples with a mixed Mn valency. **a)** The Mn²⁺ references are MnCO₃, and Mn(OH)₂, the Mn³⁺ references are Mn₂O₃ and β–MnOOH, and the Mn⁴⁺ reference is λ-MnO₂. The position of some characteristic spectral features are outlined as dotted-dashed line (Mn⁴⁺), dashed lines (Mn³⁺) and solid line (Mn²⁺). **b)** XANES spectrum of KBi8 (solid line) together with low pH one–layer hexagonal birnessite (HBi,
dashed line) and triclinic birnessite (NaBi, open circles) references.\textsuperscript{5,91} Other patterns as in Figure 8a.

**Figure 9.** \(k^3\)-weighted EXAFS spectra and Fourier transforms (modulus and imaginary parts) of \(k\)-weighted EXAFS spectra for KBi\(_8\), and for low pH one–layer hexagonal birnessite (HBi) and triclinic birnessite (NaBi) references.\textsuperscript{5,91}

**Figure 10.** Structure model for KBi\(_8\). Projection on the ab plane. Symbols as for Figure 4a. Mn\(_1\) cations are shown as large solid circles. Dot-dashed lines outline the distorted octahedron defined by O1 and O3 coordinated to interlayer K. The lower surface of the layer is shown as dark shaded triangles. \(O_{layer}\) atoms coordinated to 3 Mn\(^{4+}\) cations are shown as small solid circles, whereas \(O_{layer}\) atoms coordinated to only 2 Mn\(^{4+}\) cations are shown as small open circles. Water molecules coordinating the Mn\(^{3+}\) cation in Mn\(_2\) position below the vacant layer site are shown.
Table 1. Main chemical parameters of KBi$_8$ sample

<table>
<thead>
<tr>
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<tbody>
<tr>
<td>Weight loss$^a$ due to interlayer H$_2$O</td>
<td>6.72%</td>
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<tr>
<td>Weight loss$^a$ due hydroxyl groups</td>
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<tr>
<td>K/Mn$^b$ ratio</td>
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<td>Mn mean oxidation state</td>
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$^a$ Weight losses correspond to the 148°C and 370°C endotherms observed on the DTA-TG curves.

$^b$ K/Mn ratio was measured by ICP-AES.

Table 2. Atomic positions and occupancies for KBi$_8$ microcrystal

<table>
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<th>x</th>
<th>y</th>
<th>z</th>
<th>$\zeta$(Å)</th>
<th>Occ.$^a$</th>
<th>Ueq x 10$^2$ (Å$^2$)$^b$</th>
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<td>2a</td>
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<td>0.89</td>
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<td>O1</td>
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<td>2/3</td>
<td>1/3</td>
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$^a$ = 2.840(1) Å, c = 14.03(1) Å. $^a$ Occupancies, which were set from the chemical analyses, are given for space group $P6_3/mmc$. $^b$ Anisotropic thermal displacements are U11 = U22 = 11.5(2)10$^{-3}$, U33 = 14.1(3)10$^{-3}$, U12 = 5.8(1)10$^{-3}$, U23=U13 = 0 for Mn1, and U11 = U22 = 158.0(166)10$^{-3}$, U33 = 8.2(18)10$^{-3}$, U12 = 109.8(176)10$^{-3}$, U23=U13 = 0 for K1.
Table 3. Selected inter-atomic distances$^a$ in KBi$_8$

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$^a$ All distances are given in Å.
Table 4. Atomic positions and occupancies for KBi$_8$ powder

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<th>x</th>
<th>y</th>
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<th>$\zeta$(Å)</th>
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<tbody>
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<td>0</td>
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<td>4f</td>
<td>2/3</td>
<td>1/3</td>
<td>0.070</td>
<td>1.00</td>
<td>1.00</td>
</tr>
<tr>
<td>Mn$_{\text{interlayer}}$ Mn2</td>
<td>4e</td>
<td>0</td>
<td>0</td>
<td>0.150</td>
<td>2.13</td>
<td>0.04</td>
</tr>
<tr>
<td>K</td>
<td>K1</td>
<td>0.24</td>
<td>0.76</td>
<td>1/4</td>
<td>3.56</td>
<td>0.08</td>
</tr>
<tr>
<td>H$_2$O</td>
<td>O2</td>
<td>1/3</td>
<td>2/3</td>
<td>1/4</td>
<td>3.56</td>
<td>0.24</td>
</tr>
<tr>
<td>H$_2$O</td>
<td>O3</td>
<td>0.225</td>
<td>0.113</td>
<td>1/4</td>
<td>3.56</td>
<td>0.12</td>
</tr>
</tbody>
</table>

$a = 2.845$ Å, $c = 14.235$ Å. $^a$ Occupancies, which were set from the chemical analyses, are given for space group $P6_3/mmc$. $^b$ Thermal factors B were not refined. The proportions of random stacking faults in KBi$_{8o}$ and KBi$_{8d}$ are 6% and 20%, respectively. KBi$_{8d}$ is a randomly interstratified mixed-layered structure of 2H and 3R layer pairs (60:40). Mean sizes of the coherent scattering domains along the c axis are 20 and 10 layers in KBi$_{8o}$ and KBi$_{8d}$, respectively.
Table 5. EXAFS parameters for Mn-Mn pairs in birnessite samples and in the reference \( \lambda \text{MnO}_2 \).

<table>
<thead>
<tr>
<th>Sample</th>
<th>Fit interval (Å(^{-1}))</th>
<th>Shells</th>
<th>R (Å)</th>
<th>( \text{CN}^a )</th>
<th>( \sigma^2 ) (Å(^2))(^{b} )</th>
<th>( \Delta E ) (eV)(^{c} )</th>
<th>Resid.(^{d} )</th>
<th>Free var.(^{e} )</th>
</tr>
</thead>
<tbody>
<tr>
<td>( \lambda \text{MnO}_2 )</td>
<td>3.7 &lt; ( k &lt; 13.5 )</td>
<td>1</td>
<td>2.85</td>
<td>6</td>
<td>29 ( 10^{-4} )</td>
<td>0.8</td>
<td>11</td>
<td>4</td>
</tr>
<tr>
<td></td>
<td></td>
<td>2</td>
<td></td>
<td></td>
<td>2.88</td>
<td>3.6</td>
<td>33 ( 10^{-4} )</td>
<td>-0.2(^g)</td>
</tr>
<tr>
<td>NaBi(^f)</td>
<td>3.7 &lt; ( k &lt; 13.5 )</td>
<td>1</td>
<td>2.90</td>
<td>5.7</td>
<td>73 ( 10^{-4} )</td>
<td>-2.1</td>
<td>17</td>
<td>4</td>
</tr>
<tr>
<td></td>
<td></td>
<td>2</td>
<td></td>
<td></td>
<td>2.99</td>
<td>2.4</td>
<td>33 ( 10^{-4} )</td>
<td>-0.2(^g)</td>
</tr>
<tr>
<td>HBi(^f)</td>
<td>3.7 &lt; ( k &lt; 12 )</td>
<td>1</td>
<td>2.88</td>
<td>3.5</td>
<td>35 ( 10^{-4} )</td>
<td>0.5</td>
<td>21</td>
<td>4</td>
</tr>
<tr>
<td></td>
<td></td>
<td>2</td>
<td></td>
<td></td>
<td>2.88</td>
<td>3.7</td>
<td>37 ( 10^{-4} )</td>
<td>0.8(^g)</td>
</tr>
<tr>
<td>KBi(^f)</td>
<td>3.7 &lt; ( k &lt; 12 )</td>
<td>1</td>
<td>2.87</td>
<td>4.3</td>
<td>47 ( 10^{-4} )</td>
<td>-0.3</td>
<td>12</td>
<td>4</td>
</tr>
</tbody>
</table>

Typical uncertainty on interatomic distances is ±0.02 Å. \(^a\) Coordination number. The scaling factor \( S_0^2 \) was calculated to obtain \( \text{CN} = 6 \) in the reference \( \lambda \text{MnO}_2 \) (\( S_0^2 = 0.8 \)). Typical uncertainty on coordination numbers is ±1.0. \(^b\) Debye-Waller factor. \(^c\) Variation of the energy threshold treated as a single adjustable parameter for all sub-shells. \(^d\) Residual calculated from \( R = [\sum (k^3 \chi_{\text{exp}} - k^3 \chi_{\text{cal}})/\sum k^3 \chi_{\text{exp}}] \times 100 \). \(^e\) Number of variable parameters. In all fits, the number of independent variables, calculated by the Nyquist formula \( 2\Delta k \Delta R/\pi \),\(^{93}\) was equal to 7. \(^f\) EXAFS results are, within uncertainty, identical to those published by Silvester et al.\(^{91}\) \(^g\) Parameter varied but constrained equal for the two sub-shells.
Table 6. Empirical bond-valences (vu)\(^a\) for KBi\(_8\).

<table>
<thead>
<tr>
<th></th>
<th>O1(^b)</th>
<th>O1(^c)</th>
<th>O1(^d)</th>
<th>O2</th>
<th>O3</th>
<th>(\Sigma)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Mn1</td>
<td>0.647(^{x6}) →</td>
<td>0.647(^{x2}) ↓</td>
<td>0.647(^{x2}) ↓</td>
<td></td>
<td></td>
<td>3.88</td>
</tr>
<tr>
<td></td>
<td>(x^3) ↓</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Mn2</td>
<td>0.539(^{x3}) →</td>
<td></td>
<td></td>
<td>0.347(^{x3}) →</td>
<td></td>
<td>2.66</td>
</tr>
<tr>
<td>K1</td>
<td>0.041(^{x2}) →</td>
<td>0.123(^{x4}) →</td>
<td>0.042(^{x2}) →</td>
<td>0.049(^{x1}) →</td>
<td>0.76</td>
<td></td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>H(^+)</td>
<td>0.09(^e)</td>
<td></td>
<td>0.788(^{x2}) ↓</td>
<td>0.788(^{x2}) ↓</td>
<td></td>
<td></td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>(\Sigma)</td>
<td>1.94 - 2.03(^g)</td>
<td>1.83</td>
<td>1.29 - 1.42(^h)</td>
<td>1.92</td>
<td>1.61 - 1.62</td>
<td></td>
</tr>
</tbody>
</table>

\(^a\) Bond valences in v.u. (valence unit) calculated using the Valence for Dos program (v. 2.0 - http://www.ccp14.ac.uk/solution/bond_valence/index.html) and the parameters from Brese and O'Keeffe.\(^94\)  
\(^b\) O1 coordinated to 3 Mn\(^{4+}\) in Mn1. \(^c\) O1 coordinated to 2 Mn\(^{4+}\) in Mn1 and 1 Mn\(^{3+}\) in Mn2. \(^d\) O1 coordinated to 2 Mn\(^{4+}\) in Mn1. \(^e\) O3-H-O1 H-bond. \(^f\) O2-H-O1 H-bond. \(^g\) Depending on whether this O1 receives additional valence from interlayer K\(^+\) or H\(^+\) through H-bond. \(^h\) Depending on whether this O1 receives additional valence from interlayer K\(^+\).
References cited


(71)  Gaillot, A.-C.; Drits, V. A.; Lanson, B.; Manceau, A. Amer. Mineral., In preparation.

(72)  Gaillot, A.-C.; Drits, V. A.; Plançon, A.; Lanson, B. Amer. Mineral., In preparation.


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Figure 05: X-ray diffraction pattern showing intensity peaks for different crystallographic planes. The intensity scale is increased by a factor of 10 for angles between 30° and 90°.

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Gaillot et al. Fig. 05
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Gaillot et al. - Fig. 06
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Please print in two column format
Structure of synthetic K-rich birnessite obtained by high-temperature decomposition of KMnO₄. Part I. Two-layer polytype from 800°C experiment.

Anne-Claire Gaillot, David Flot, Victor A. Drits, Alain Manceau, Manfred Burghammer, Bruno Lanson

Graphic joined as EPS file

The crystal structure of microcrystalline phyllomanganate birnessite, which was determined from the first time from a single-crystal, consists of edge-sharing Mn⁴⁺O₆ octahedra building up layers. Vacant layer octahedra are capped on either side by interlayer Mn³⁺ cations (Mn2) having octahedral coordination. Layer charge deficit is compensated for by hydrated (O3) interlayer K⁺ cations in prismatic cavities. These cations are shifted from the center of these cavities to provide local charge compensation to O1 atoms.
A.-C. Gaillot, D. Flot, V.A. Drits, A. Manceau, M. Burghammer, B. Lanson
Structure of synthetic K-rich birnessite obtained by high-temperature decomposition of KMnO4.
PartI. Two-layer polytype from 800°C experiment.