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Research Article

Tetraethyl Orthosilicate Coated Hydroxyapatite Powders for Lead Ions Removal from Aqueous Solutions

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The goal of this study was to synthesize and characterize a porous material based on tetraethyl orthosilicate (TEOS) coated hydroxyapatite (HApTh) after removal experiments of Pb^{2+} ions from aqueous solutions. In order to study the morphology and composition, the samples obtained after removal experiments of Pb^{2+} ions from aqueous solution with the initial Pb^{2+} ion concentrations of the aqueous solutions were $0.1\text{ g}\cdot\text{L}^{-1}$ (HApTh-50) and $0.9\text{ g}\cdot\text{L}^{-1}$ (HApTh-450) have been investigated by scanning electron microscopy (SEM) equipped with an energy dispersive X-ray spectrometer (EDS), Fourier transform infrared spectroscopy (FTIR), and transmission electron microscopy (TEM). Removal experiments of Pb^{2+} ions were carried out in aqueous solutions with controlled concentration of Pb^{2+} . After the removal experiment of Pb^{2+} ions from solutions, porous hydroxyapatite nanoparticles were transformed into HApTh-50 and HApTh-450 due to the adsorption of Pb^{2+} ions followed by a cation exchange reaction. The obtained results show that the porous HApTh nanopowders could be used for Pb^{2+} ions removal from aqueous solutions.

1. Introduction

One of the major problems encountered in the public health area worldwide is the poisoning with various heavy metals. Researchers around the world have turned their attention to finding new effective and cost-efficient methods for depollution, considering the fact that heavy metals are nonbiodegradable, having the tendency to accumulate in living bodies, leading to disorders of different functions and to serious diseases [1, 2]. Lead (Pb) is a very toxic heavy metal, found in the earth's crust with an average concentration of 16 mg/kg in soils [3, 4]. On the other hand, lead has been extensively used in various industries, being a constituent in building materials, pipes, lead-acid batteries, bullets, and paints. In a report released by the Agency for Toxic Substances and Disease Registry (ATSDR), in 2007, the harmful effect of lead on the human nervous and reproductive systems was emphasized. Contamination of wastewaters represents a major concern

because this is one way of lead bioaccumulation in the food chain [3]. Recent studies have revealed that children are more susceptible to lead poisoning, their bodies being able to absorb around 50% of inhaled or ingested lead and the effects being more pronounced and long-lasting, compared to adults [5]. Moreover, the effects of chronic lead exposure to various functions of the human body have already been reported [3, 6]. In this context, many researchers have focused their attention to developing different methods for dangerous heavy metals removal from wastewaters. Among the methods already used, chemical precipitation, membrane filtration, ion-exchange, and adsorption could be mentioned [7–9]. Of all these methods, the most preferred one is the adsorption of heavy metals ions from aqueous solution using cost-efficient materials [7–9]. Previous studies have revealed that apatites are able to successfully remove lead from aqueous solutions [4, 10]. The best known member of the apatite family is hydroxyapatite (HAp). Synthetic hydroxyapatite,

$\text{Ca}_{10}(\text{PO}_4)_6(\text{OH})_2$, has been used in the last decades in many biomedical applications as coating for orthopedic or dental implants or as filling material for various injuries or defects, being similar to the natural mineral component of the human bones and teeth [11]. Furthermore, two of the most appealing properties for environmental applications that HAp possesses are its ability to adsorb complex organic materials and the capacity of heavy metals ion-exchanging [11–14]. Previous studies have shown that Pb^{2+} ions exchange rapidly with Ca, this process inducing morphological changes in the surface region [15–17]. In order to increase the adsorption ability of hydroxyapatite, an increase in its porosity must be obtained. Therefore, doping hydroxyapatite with silicon ions may be the key to increase the porosity, thus creating a better material for lead removal from aqueous solutions. Recent studies have shown that the addition of tetraethoxysilane (TEOS) can induce the formation of large textural pores [18]. Consequently, in order to increase the lead adsorption capacity of hydroxyapatite, it must be doped with a silicon based compound, such as TEOS.

The goal of this study was to synthesize and characterize new porous material based on tetraethyl orthosilicate (TEOS) and hydroxyapatite (HApTh) after removal experiment of Pb^{2+} ions from aqueous solution. The obtained powders after removal experiment of Pb^{2+} ions from aqueous solution when the initial Pb^{2+} ion concentrations of the aqueous solutions were $0.1 \text{ g}\cdot\text{L}^{-1}$ (HApTh-50) and $0.9 \text{ g}\cdot\text{L}^{-1}$ (HApTh-450) have been investigated by scanning electron microscopy (SEM), transmission electron microscopy (TEM), and Fourier transform infrared spectroscopy (FTIR). Another objective was to investigate the removal of Pb^{2+} ions from aqueous solutions with different pH values using HApTh nanopowders.

2. Experimental Section

2.1. Materials. All the reagents including ammonium dihydrogen phosphate $[(\text{NH}_4)_2\text{HPO}_4]$ (Alpha Aesar), calcium nitrate tetrahydrate $[\text{Ca}(\text{NO}_3)_2\cdot 4\text{H}_2\text{O}]$ (Alpha Aesar), ethanol, and tetraethyl orthosilicate (TEOS) with 99.999% purity were purchased from Sigma-Aldrich and lead nitrate $[\text{Pb}(\text{NO}_3)_2]$ with 99.5% purity was purchased from Merck and used without further purification.

2.2. Synthesis of Hydroxyapatite/TEOS Nanocomposites. Hydroxyapatite/TEOS nanocomposites (HApTh) were prepared using tetraethyl orthosilicate and hydroxyapatite. The hydroxyapatite $(\text{Ca}_{10}(\text{PO}_4)_6(\text{OH})_2)$ nanoparticles were prepared by setting the atomic ratio of Ca/P at 1.67 in accord with [19, 20]. The hydroxyapatite (HAp) was immobilized into a tetraethyl orthosilicate foam using the technique reported in the literature [20]. HApTh was obtained when tetraethyl orthosilicate solution (1g HAp/10 mL) was dropped on the HAp powder. The mixtures were then stirred vigorously for 30 min until homogeneity was achieved. After forming stable structures, HApTh composites were allowed to dry at 80°C for 24 h in a vacuum oven for the excess solvent to evaporate. Finally, HApTh composite samples were then ground in order to obtain powders.

2.3. Samples Characterization. In order to investigate the composition and morphology of the samples an equipment FESEM HITACHI 4700 coupled with an energy dispersive X-ray attachment (EDAX/2001 device) was used. TEM studies were carried out using a FEI Tecnai 12 (FEI Company, Hillsboro, OR, USA) equipped with a low-dose digital camera from Gatan Inc. (Pleasanton, CA, USA). Small quantities of HApTh powder were dispersed in deionised water and deposited on a copper grid coated with carbon film. This technique allows a good definition of crystal morphology.

The FTIR spectra were acquired using a Spectrum BX spectrometer. Pellets of 10 mm diameter for FTIR measurements were prepared (1% of the powder was mixed and ground with 99% KBr) by pressing the powder mixture at a load of 5 tons for 2 min. The spectra were registered in the range from 400 to 4000 cm^{-1} with a resolution of 4 and 128 times scanning.

The Raman spectra were registered with a Renishaw InVia dispersive Raman spectrometer (2012), equipped with a Leica DM microscope and one laser source at 514 nm (gas-type), Spectra Physics Ar ion laser (20 mW). The samples were analyzed using the 514 nm laser with a power at 0.2 mW and 1800 L/mm gratings. The spectral range covered for all spectra is $100\text{--}2000 \text{ cm}^{-1}$ with a resolution below 2 cm^{-1} . Prior to each reference measurement, the instrument was calibrated on the internal Si-reference standard ($520.6 \pm 0.1 \text{ cm}^{-1}$).

The removal performance of Pb^{2+} ions by the HApTh powders was investigated by batch experiments, monitoring the change of Pb^{2+} ion concentration in the aqueous solutions. For these experiments, 5 g of HApTh nanocomposites were added to 500 mL aqueous solution with various initial Pb^{2+} ion concentrations and pH values in accord with Jang et al. [21]. The initial Pb^{2+} ion concentrations of the aqueous solutions were controlled and the values were set in the range $0.1\text{--}1.5 \text{ g}\cdot\text{L}^{-1}$ by dissolving lead nitrate $[\text{Pb}(\text{NO}_3)_2]$ in deionized water. The pH values of aqueous solutions with controlled initial Pb^{2+} ion concentrations were adjusted from 3 to 6 by adding small amounts of 0.1 M HCl standard solution. For all experiments the solution was stirred constantly for 24 h by a mechanical stirrer at room temperature.

3. Results and Discussions

Scanning electron microscopy (SEM) was used to characterize the morphology and elemental composition of synthesized powders after lead removal experiments. SEM images presented in Figure 1 showed that lead incorporation into HApTh samples affects the particle shape. The crystal clusters with needle or rod-like shapes were observed in SEM micrographs for the HApTh-50 and HApTh-450 samples. The clusters with needle or rod-like shapes were distinguished when the Pb^{2+} ion concentration from the aqueous solutions was $0.9 \text{ g}\cdot\text{L}^{-1}$ (HApTh-450 samples).

The EDX analysis indicates that calcium, phosphorus, silicon, and oxygen are the major constituents of the samples. The presence of Pb is also observed.

Figure 2 shows the TEM image of HApTh powders after the sorption experiment. These nanoparticles exhibit

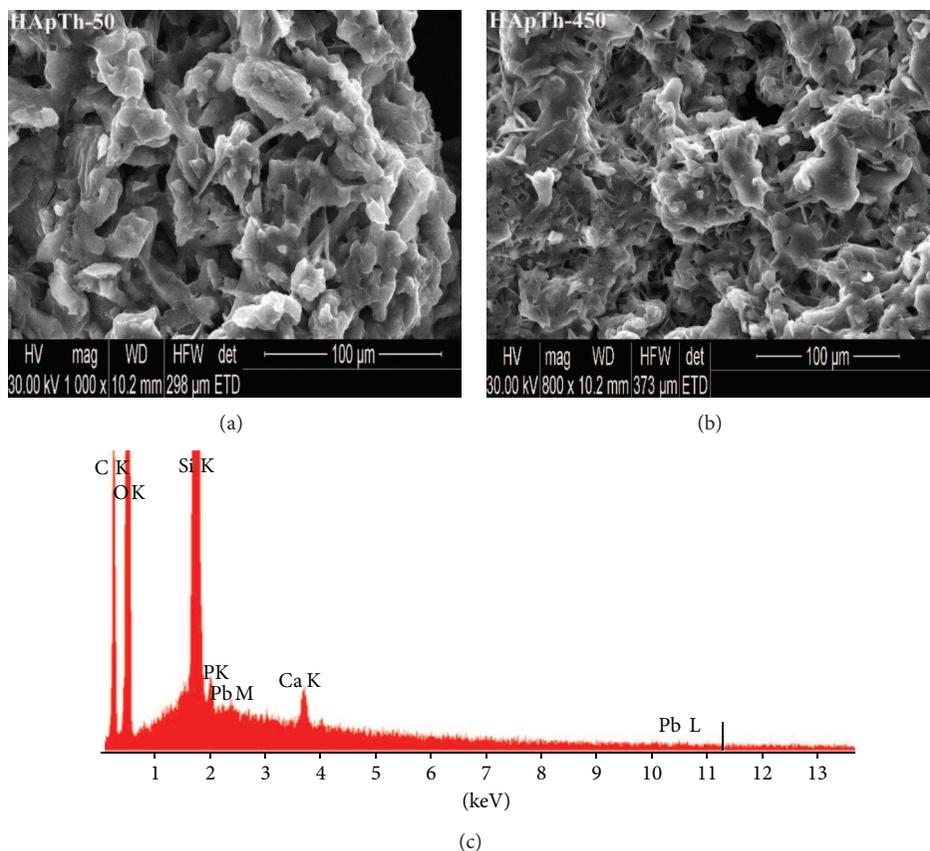


FIGURE 1: SEM micrographs with EDX spectra for the (a) HApTh-50 and (b) HApTh-450 samples.

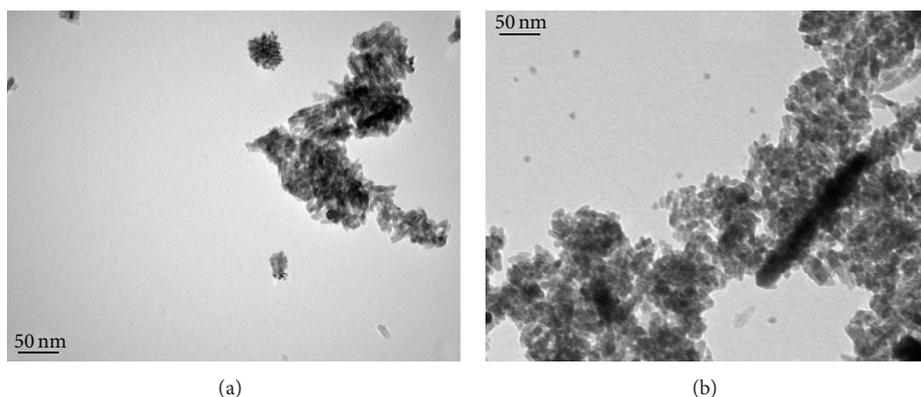


FIGURE 2: TEM micrograph of HAp after Pb²⁺ removal experiments for: (a) HApTh-50 and (b) HApTh-450 samples.

a needle-like morphology usually observed on HAp powders obtained by coprecipitation method [22]. Moreover, in TEM images it is easy to observe the formation of a new phase (pyromorphite) with a plate-like morphology (especially in the case of HApTh-450 samples). This behavior suggests that the incorporation of lead in the HApTh matrix stimulates local calcium enrichment.

These images also confirm the hypothesis that the main mechanism of lead immobilization is first of all based on the dissolution of HApTh followed by the precipitation of phases with higher lead content [23].

The Fourier transform infrared spectroscopy provides valuable information about the short-range ordering of materials. In Figure 3 are presented the FTIR spectra of the obtained powders after Pb²⁺ removal experiments. The interaction between lead (II) and HApTh structure led to some modifications of infrared vibrations of the functional groups. This behavior can be explained by the reaction mechanism which involves the HApTh dissolution in acid environment and HApTh Pb crystallisation [24].

In all the spectra the presence at around 640 cm⁻¹ of OH vibration peak could be noticed. The broad peaks in the

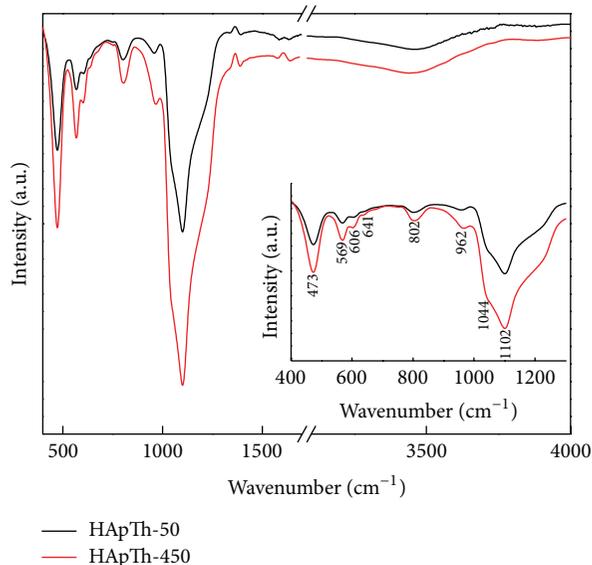


FIGURE 3: FT-IR spectra of HApTh-50 and HApTh-450 samples.

regions $1600\text{--}1700\text{ cm}^{-1}$ and $3200\text{--}3400\text{ cm}^{-1}$ are attributed to the hydroxyl groups [25–27].

The specific bands of (PO_4^{3-}) phosphate groups characteristic to hydroxyapatite structure were observed at 568 cm^{-1} , 637 cm^{-1} , 605 cm^{-1} , 964 cm^{-1} , and $1000\text{--}1100\text{ cm}^{-1}$ [28, 29]. The 964 cm^{-1} band can be associated to the ν_1 nondegenerated symmetric stretching mode of P–O. The vibrational bands at 605 cm^{-1} and 568 cm^{-1} are attributed to the triply degenerated ν_4 vibration of O–P–O bond, and the band at 472 cm^{-1} may be attributed to the ν_2 bending of O–P–O bond. Moreover, the weak band observed around 1500 cm^{-1} could be assigned to the A-type carbonate ion substitutions, whereas the bands at $1422\text{--}1456\text{ cm}^{-1}$ (ν_3) can be attributed to B-type CO_3^{2-} substitutions [30–32].

Moreover, it is well known that the vibrations bands (stretching and bending) assigned to the Si–O–Si bonds appear in the $1100\text{--}900\text{ cm}^{-1}$ and $550\text{--}400\text{ cm}^{-1}$ spectral regions. In the FTIR spectra (Figure 3) it is obvious that the stretching and bending bands of SiO_4 groups are overlapping with the band of the PO_4 group [30, 31, 33–37].

According to Ping et al. [38] the vibration band at 806 cm^{-1} (HApTh-450 sample) could be attributed to Si–O stretching of dimer silicate chains, indicating the start of silicate polymerization.

The Raman spectra of the HApTh-50 and HApTh-450 powders are reported in Figure 4. The Raman vibrations bands at 959 and 924 cm^{-1} are attributed to the ν_1 PO_4 vibration modes. In HApTh-50 samples, only the 959 cm^{-1} band appears, whereas the HApTh-450 samples exhibit two bands, one strong at 959 cm^{-1} and the other band at 924 cm^{-1} . This behavior is caused by the $A_g\text{-}E_{2g}$ splitting in the hexagonal C_{6h} symmetry [38, 39].

For the all samples, the vibrational spectra exhibit a strong molecular character associated with the internal modes of the tetrahedral PO_4^{3-} . The ν_2 bands (426 cm^{-1} and 450 cm^{-1})

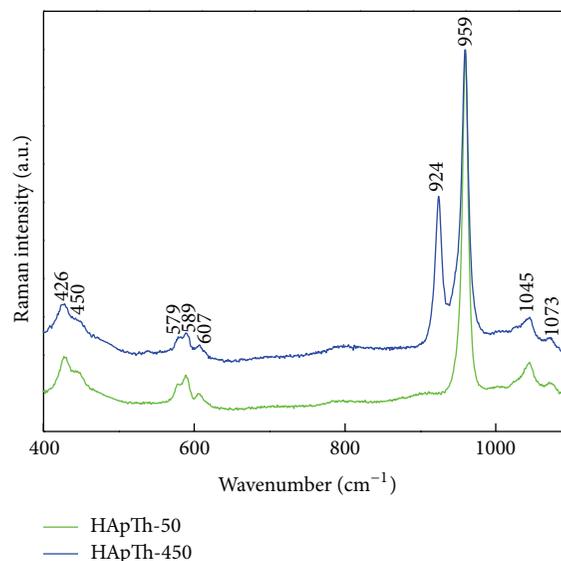


FIGURE 4: Raman spectra for the HApTh-50 and HApTh-450 samples.

are attributed to the O–P–O bending modes; ν_3 vibration bands (1073 cm^{-1} and 1045 cm^{-1}) correspond to asymmetric P–O stretching and the ν_4 bands (579 cm^{-1} , 589 cm^{-1} and 607 cm^{-1}) are attributed mainly to O–P–O bending.

In order to evaluate the impact of the lead concentration in the aqueous solution, absorption experiments were performed. HApTh solutions with concentration varying from 0.1 to $1.5\text{ g}\cdot\text{L}^{-1}$ at pH 5 were used. The measurements were performed on 500 mL solution (pH 5) with an initial Pb^{2+} ion concentration of $63\text{ mg}\cdot\text{L}^{-1}$. In Figure 5 the adsorption efficiency of Pb^{2+} ions as a function of the Pb^{2+} concentration in the solution is presented. It was observed that the removal efficiency is dependent on the initial Pb^{2+} concentration. For a lead concentration of $0.2\text{ g}\cdot\text{L}^{-1}$, the removal efficiency reached 98.6%, showing that the adsorbent material (HApTh) has a strong affinity to Pb^{2+} ions. For Pb^{2+} concentration ranging from $0.4\text{ g}\cdot\text{L}^{-1}$ to $1.5\text{ g}\cdot\text{L}^{-1}$, the removal efficiency was around 100%. This behavior could be explained by the fact that the Pb^{2+} ions were completely removed from the solution.

Figure 5 shows the measured lead concentration in the solution after the reaction with HApTh has taken place. For the studies on the effect of the solution pH, a solution containing $0.9\text{ g}\cdot\text{L}^{-1}$ of lead was selected (Figure 6). A solution containing $563\text{ mg}\cdot\text{L}^{-1}$ of Pb^{2+} ions was obtained from 0.9 g of $\text{Pb}(\text{NO}_3)_2$ in 1 L of distilled water.

A removal efficiency of 100% was achieved for pH 6. On the other hand, the removal efficiency of Pb^{2+} ions was 98.5% at pH 3. The removal efficiency of Pb^{2+} ions was 99% for pH 4. When the pH value was set to 5 the removal efficiency of Pb^{2+} ions was 99.5%. The present studies have shown that the removal of Pb^{2+} ions from aqueous solution by hydroxyapatite/TEOS nanocomposites was greater than 98.6% at all pH values of the aqueous solution used for the removal of Pb^{2+} ions.

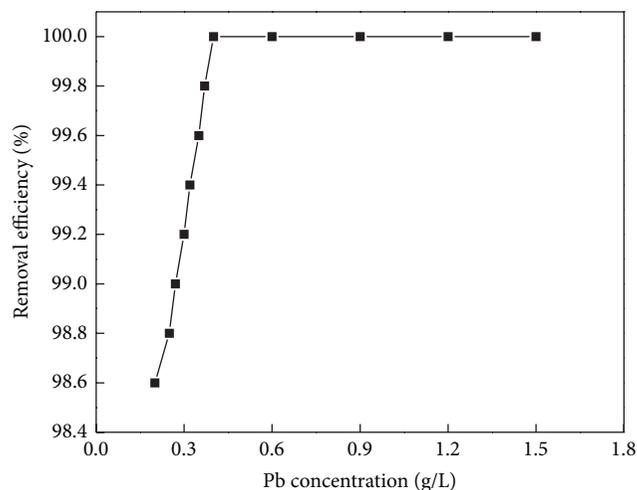


FIGURE 5: Effect of the Pb^{2+} concentration on the removal of the Pb^{2+} ions by HApTh.

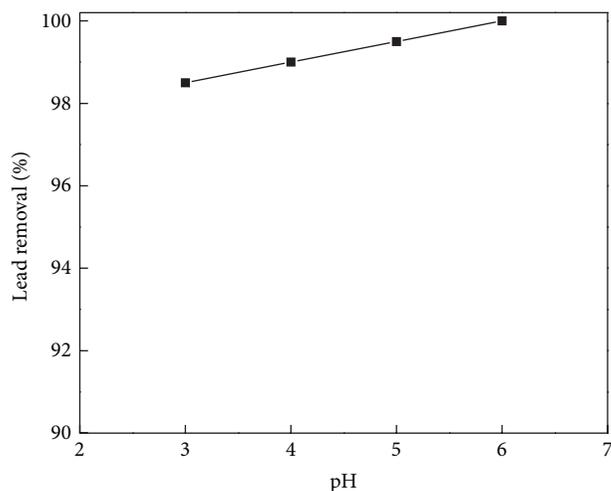


FIGURE 6: Effect of the solution pH on the removal of Pb^{2+} ions by HApTh.

According to these results, it can be highlighted the fact that the HApTh can really remove Pb^{2+} ions from aqueous solutions.

One of the most important parameters which must be considered in the absorption experiments is the pH of the solutions due to the fact that the pH of groundwater and surface waters varies between 5 and 7 [40]. The results of our studies have shown that the solid reaction products of aqueous Pb^{2+} with the HApTh are mainly pH-dependent, this being in good agreement with other results reported in literature [40].

For the assimilation of different metals from aqueous solutions by hydroxyapatite [17, 41–43] different processes have been proposed, such as metal complexation on the HAp surface, apatite dissolution followed by a new metal phase precipitation, and cation exchange. In previous studies [44], it has been shown that the formation of a stable lead apatite

such as $\text{Ca}_{(10-x)}\text{Pb}_x(\text{PO}_4)_6(\text{OH})_2$ by ion exchange mechanism where Pb^{2+} present in the solution replaces Ca^{2+} ions from the hydroxyapatite lattice is possible. Furthermore, the $\text{Pb}_{(10-x)}\text{Ca}_x(\text{PO}_4)_6(\text{OH})_2$ crystals with higher content of calcium are unstable [45]. Therefore, the HAp should be subject of a permanent process of dissolution and precipitation in order to obtain more stable materials with a higher lead percentage.

4. Conclusions

The main goal of this study was to synthesize a new porous nanocomposite material based on tetraethyl orthosilicate (TEOS) coated hydroxyapatite. In this study we investigated the ability of these new materials to remove Pb^{2+} ions from aqueous solutions with different concentration and pH values. After HApTh reacted with the solution containing Pb^{2+} ions, the lead ions were completely removed from the solution and the dissolution of $\text{Ca}_{10}(\text{PO}_4)_6(\text{OH})_2$ occurred. The HApTh powders exhibited the higher removal efficiency of Pb^{2+} ions at pH 6. The TEM studies confirm the hypothesis that the main mechanism of lead immobilization is first of all based on the dissolution of HAp followed by the precipitation of phases with higher lead content. This study showed that the HApTh nanopowders are promising materials for lead ions removal from aqueous solutions with different pH values, being able to be used in the future for depollution of wastewaters.

Conflict of Interests

The authors declare that there is no conflict of interests regarding the publication of this paper.

Acknowledgment

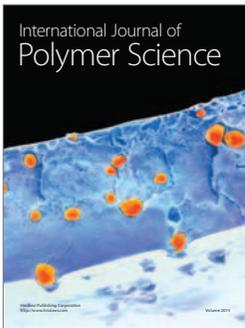
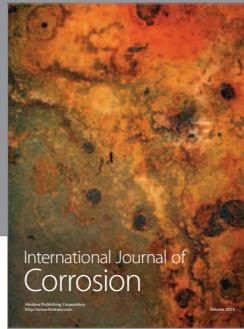
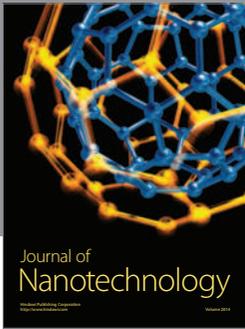
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