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**Comprehensive
carbonaceous
aerosol
characterization**

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A new comprehensive approach to characterizing carbonaceous aerosol with an application to wintertime Fresno, California PM_{2.5}

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Abstract

Fine particulate matter (PM_{2.5}) samples were collected during a three week winter period in Fresno (CA). A composite sample was characterized by isolating several distinct fractions and characterizing them by infrared and nuclear magnetic resonance (NMR) spectroscopy. More than 80% of the organic matter in the aerosol samples was recovered and characterized. Only 35% of the organic matter was water soluble with another third soluble in dichloromethane and the remainder insoluble. Within the isolated water soluble material, hydrophobic acid and hydrophilic acids plus neutrals fractions contained the largest amounts of carbon. The hydrophobic acids fraction appears to contain significant amounts of lignin type structures, spectra of the hydrophilic acids plus neutrals fraction are indicative of carbohydrates and secondary organic material. The dichloromethane soluble fraction contains a variety of organic compound families typical of many previous studies of organic aerosol speciation, including alkanes, alkanols, alkanals and alcanoic acids. Finally the water and solvent insoluble fraction exhibits a strong aromaticity as one would expect from black or elemental carbon like material; however, these spectra also show a substantial amount of aliphaticity consistent with linear side chains on the aromatic structures.

1 Introduction

Carbonaceous material is an important component of ambient particulate matter. In many environments carbon contributes several $\mu\text{g}/\text{m}^3$ to ambient aerosol loadings (Malm et al., 2004). Despite its importance on global (climate) and local (visibility, health effects) scales (Kanakidou et al., 2005; Turpin et al., 2000), organic material has only in the past decade or so become a topic of systematic investigation.

For many years specific, usually toxic, species including polycyclic aromatic hydrocarbons or dioxins were studied in the particulate phase. More detailed speciation studies followed, characterizing some 200 species, including so called molecular markers

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used for aerosol source apportionment (e.g. Rogge et al., 1993; Schauer et al., 1996). Despite the large number of species quantified in source apportionment studies, they frequently account for less than 20% by mass of the total organic matter present (Brown et al., 2002), the remainder being largely uncharacterized.

In recent years some efforts have also been directed towards comprehensive approaches to characterizing particulate organic matter, focusing on bulk compound properties rather than on detailed molecular speciation. This general approach promises to provide more insight into modeling carbonaceous aerosol behavior in the atmosphere and, hopefully, to point out fruitful directions for future molecular speciation efforts. Some of the approaches taken include investigating solubility (Decesari et al., 2001), structure (Gelencser et al., 2000), functionality (Decesari et al., 2000; Graham et al., 2002; Maria et al., 2002; Reff et al., 2005; Suzuki et al., 2001), and molecular size (Kalberer et al., 2004). Several studies investigated the fraction of organic particulate matter that is water-soluble to better understand the hygroscopic growth potential of these particles and likely impacts on cloud and haze formation. In some cases the water soluble organic carbon (WSOC) was further analyzed by chromatographic and/or spectroscopic techniques (Chang et al., 2005; Decesari et al., 2000; Duarte et al., 2004, 2005; Graham et al., 2002; Kiss et al., 2002; Sullivan and Weber, 2006a; Sullivan and Weber, 2006b).

Early organic matter characterization efforts by Havers and co-workers concluded that carbonaceous aerosol, or a fraction thereof, shows properties similar to humic substances and the terminology HUmic Like Substances (HULIS) was introduced (Havers et al., 1998). Many efforts have focused since on demonstrating similarities or differences between aerosol carbon and soil carbon, as extensively discussed by Graber and Rudich (2006). Recent work has also suggested that polymeric or oligomeric material, potentially from secondary atmospheric formation, might comprise a substantial component of ambient particulate matter (Gao et al., 2004; Kalberer et al., 2004; Samburova et al., 2005a; Samburova et al., 2005b; Tolocka et al., 2004).

As our understanding of the nature of carbonaceous particles has improved, some

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investigators have attempted to formulate simplified models of organic aerosol composition that would be useful conceptually as well as for use in both laboratory and numerical modeling studies. Among these efforts is the work by Fuzzi et al. (2001) and McFiggans et al. (2005).

5 In order to continue improving our understanding of carbonaceous aerosol, and to suggest fruitful directions for future organic speciation efforts, an approach previously developed for organic matter characterization in surface waters (Leenheer et al., 2000) and applied recently for fog organic matter (Herckes et al., 2007) was applied to atmospheric aerosol. Aerosol samples were collected in Fresno, California. A composite
10 sample was characterized by fractionating and isolating organic matter into several classes based upon molecule size, solubility, polarity, and acid/base/neutral characteristics. Resulting organic isolates were then characterized by infrared and ^{13}C -nuclear magnetic resonance (NMR) spectrometry and elemental analyses. Observed characteristics of the aerosol are reported here and compared with previous organic aerosol
15 models.

2 Experimental methods

2.1 Sampling

Ambient aerosol samples were collected at five locations in the city of Fresno, California from 24 December 2003 until 15 January 2004. A detailed description of the
20 sampling protocols and concentrations of organic source marker species is given elsewhere (Gorin et al., 2006). In brief, the five sites included an urban site (downtown Fresno), three residential sites (California State University, Fresno campus; Fresno Pacific University campus and Clovis) and one site located in an industrial park in south Fresno near state highway 99. High-volume collectors were operated at all five sites to
25 collect 20-h samples at a nominal flow rate of $1.13\text{ m}^3/\text{min}$, periodically checked with a Thermo Andersen high-volume sampler calibration kit. The instruments were operated

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from 02:00 p.m. to 10:00 a.m., local time, to allow sufficient time for daily filter exchange by a single site operator at all sites. The period of collection might bias the 20-h average sample concentrations somewhat higher than a 24-h average as local diurnal trends in winter PM_{2.5} concentrations demonstrate that from 10:00 a.m. to 02:00 p.m. concentrations are typically at a minimum (Chow et al., 1999). All samplers were equipped with a Tisch Series 231 PM_{2.5} impactor plate, providing simultaneous coarse particle samples (on a slotted pre-filter) and PM_{2.5} samples. At two sites (Drummond and Clovis) the high-volume collector was equipped with a PM₁₀ inlet upstream of the PM_{2.5} impactor. In the current study we only focused on PM_{2.5} samples.

Total as well as organic and elemental carbon were determined on each individual filter using a thermo optical transmission method (Birch and Cary, 1996) on a Sunset Laboratories carbon analyzer. In order to provide enough particulate organic carbon for fractionation and analysis, a single composite sample was prepared by combining portions of each of the individual PM_{2.5} filters. Observations of aerosol composition, therefore, should be interpreted as being representative of a cross-section of Fresno urban and suburban environments during winter.

2.2 Preparative organic matter fractionation and isolation procedure

The overall sample processing is described in Fig. 1 and is patterned after protocols used previously to characterize surface waters (Leenheer et al., 2000, 2004). 96 quartz fiber filter sections, 1/8 of each collected PM_{2.5} sample, were divided into smaller portions that were inserted into six 250 mL Teflon centrifuge bottles. A solution of 200 mL of 0.01 M HCl was added to each centrifuge bottle, and the bottles were mixed by rotating on a GlasCol laboratory rotator overnight. The acid solutions were decanted and vacuum-filtered through a 47 mm, 1 μm porosity glass fiber filter. 200 mL of deionized water were added to each bottle, the bottles rotated for one hour, and the water solutions decanted, filtered, and combined with the acid solution. This water extract was vacuum rotary evaporated to approximately 20 mL. The distillate containing volatile organic acids was titrated to pH 8.0 with 0.1 M NaOH. This distillate was vacuum evap-

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orated and freeze-dried to isolate the sodium salts of volatile organic acids.

The 20 mL water extract was placed in a Spectra/Por 3 regenerated cellulose dialysis bag (29 mm diameter, 3500 dalton pore size cutoff). It was dialyzed for three days against 1.1 L of deionized water. The particulate fraction in the dialysis bag was freeze-dried to give a colloid fraction free from salts.

The glass fiber filters used to filter the water extract were added to the quartz fiber filters in the centrifuge bottles, and the filters were freeze-dried while in the centrifuge bottles. Methylene chloride (100 mL) was added to each centrifuge bottle and the sample was placed on the sample rotator overnight. The methylene chloride was vacuum-filtered through a 47 mm, 1 μ m porosity glass fiber filter, and was vacuum rotary evaporated to 5 mL. This 5 mL concentrate was transferred to a 5 mL glass vial and was blown down to dryness under a stream of nitrogen gas. This fraction is designated CH₂Cl₂-extractable organic matter.

The quartz fiber filters in the six centrifuge bottles and the glass fiber filters used to filter the water and methylene chloride extracts were transferred to six Spectra/Por 3 regenerated cellulose dialysis bags (29 mm diameter, 3500 dalton pore size cutoff), and each bag was dialyzed against 4 L of 0.2 M HF. The HF solutions were changed daily for four days until the filters were completely dissolved. The bags were then dialyzed against deionized water until the permeate specific conductance was less than 10 μ S/cm, and the particulate suspension inside the bags was freeze-dried to isolate the particulate organic matter fraction.

The 1.1 L of dialysis permeate from the water extract was passed through a 20 mL column (1 cm ID \times 28.5 cm) of MSC-1H cation exchange resin. The resin was rinsed with 50 mL of deionized water and then desorbed with 50 mL of 75% acetonitrile/25% water. This solution was evaporated and freeze-dried to isolate the hydrophobic neutral fraction. The MSC-1H column was then desorbed with 50 mL of 3 M NH₄OH which was evaporated and freeze dried to isolate the base fraction. The sample and rinse effluents from the MSC-1H column were then passed through a 20 mL column (1 cm ID \times 28.5 cm) of Amberlite XAD-8 resin followed by 50 mL of 0.01 M HCl rinse. The column

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was desorbed with 50 mL of 75% acetonitrile/25% water. This solution was evaporated and freeze-dried to isolate a hydrophobic acid fraction.

The pH of the sample and HCl rinse from the column effluent was adjusted to pH 5 with NaOH. The volume was then reduced to 100 mL by vacuum rotary evaporation in a large Buchi R-220 vacuum rotary evaporator. The sample was transferred to a small Buchi R-111 vacuum rotary evaporator and most of the NaCl, KCl, Na₂SO₄, and CaSO₄ was removed by a zeotropic distillation procedure (Leenheer et al., 2000) in which water from the sample is evaporated from glacial acetic acid, and inorganic salts are removed by vacuum filtration through a 1 μm glass fiber filter while hydrophilic acids and neutrals remain in solution. The filtrate volume was diluted with an equal volume of deionized water and 30 mL of 1 M barium formate was added to precipitate the remaining sulfate as barium sulfate which was removed by centrifugation. The precipitate was washed with 20 mL of 1.0M HCl, separated by centrifugation, and the supernatant was evaporated to dryness with acetonitrile to remove HCl. The dry wash residue was added to the supernatant from the first centrifugation step. Excess barium in the supernatant was removed by passing the supernatant through an 80 mL (1.5 cm inside diameter (ID) ×30 cm in length) MSC-1H cation-exchange resin column at a flow rate of 20 mL/min. The column effluent was vacuum evaporated to dryness by repeatedly adding anhydrous acetonitrile to the moist residue which removed water, HCl, acetic acid and formic acid as volatile azeotrope with acetonitrile. Fifty milliliters of 98% formic acid was added to the dry residue and evaporated to dryness. This process reduces nitric acid in the sample to volatile nitrogen dioxide which is removed during evaporation. Formic acid in the moist residue was removed by evaporation with acetonitrile, and 50 ml of methanol was added and evaporated to dryness to remove boric acid as volatile trimethyl borate. The residue was dissolved in 5.0 mL water and saturated LiOH was added to adjust to pH 12 to precipitate Li₃PO₄. The solution was vacuum filtered through a 1 μm glass fiber filter and excess lithium was removed with a 20 mL MSC-1H cation exchange resin column (1 cm inside diameter ×28.5 cm in length). The column effluent was freeze-dried to isolate a Hydrophilic acids plus

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neutrals fraction.

2.3 Chemical characterization

2.3.1 Fourier Transform-Infrared (FT-IR) Spectrometry

Infrared spectra were collected using 1–2 mg of organic matter fraction isolates in potassium bromide pellets. A Perkin Elmer System 2000 FT-IR, using an infrared source with a pulsed laser carrier and a deuterated triglycine sulfate detector, scanned from 4000 to 400 cm^{-1} , averaging 10 scans at 1.0 cm^{-1} intervals with a resolution of 4.0 cm^{-1} . All spectra were normalized after acquisition to a maximum absorbance of 1.0 for comparative purposes. For peak frequency assignments for organic compound classes see Table 1; assignments for inorganic solutes commonly found in water samples are given in Table 2.

2.3.2 Solid state cross polarization magic angle spinning ^{13}C -Nuclear Magnetic Resonance (NMR) Spectrometry

Solid State Cross Polarization Magic Angle Spinning (CPMAS) ^{13}C NMR spectra were obtained on 10 mg or greater masses of organic matter fractions. Freeze-dried samples were packed in ceramic rotors. CPMAS ^{13}C -NMR spectra were obtained on a 200-megahertz (MHz) Chemagnetics CMX spectrometer with a 7.5-mm-diameter probe. The spinning rate was 5000 Hz. The acquisition parameters included a contact time of 5 ms, pulse delay of 1 s, and a pulse width of 4.5 μs for the 90° pulse. Variable contact time studies and comparison of liquid-state with solid state ^{13}C -NMR spectra by Wershaw and coworkers (Wershaw et al., 2000) indicate these are the optimum parameters for quantitatively determining different organic carbon structural group contributions to the NMR spectra. The small fraction masses and hydroscopic nature of certain sample fractions created difficulties in obtaining ^{13}C -NMR spectra. Sample runs of two to three days were required to obtain ^{13}C -NMR spectra with adequate signal to noise ratios.

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Organic structural assignments for peaks in the CPMAS¹³C-NMR spectra are given in Table 3.

2.3.3 Elemental analyses

Elemental analyses (C, H, N) and moisture content of organic matter fractions were conducted by Huffman Laboratories, Golden, Colorado (Huffman and Stuber, 1985). The small fraction masses created some difficulties in obtaining complete elemental and moisture analyses. Moisture and nitrogen contents were omitted when there was insufficient material available to analyze.

3 Results and discussion

The fractionation procedure described above isolates carbon in the samples from inorganic salts, mineral components, and quartz fiber filters. The sample is fractionated into more homogeneous fractions based upon molecule size, polarity, volatility, and acid/base/neutral properties. The infrared and ¹³C-NMR spectra and elemental analyses yield complementary organic structural information at a compound-class level for each of the isolated fractions. Both the experimental approach and resulting information are unique, yielding original information on the composition of organic matter in Fresno winter aerosol.

Fraction weights and elemental analyses of organic matter fractions from the composite Fresno aerosol sample are presented in Table 4. A bar diagram of the total organic carbon fractionation is presented in Fig. 2. The largest fractions of carbon are represented by the particulate organic matter fraction and the CH₂Cl₂ extractable organic matter fraction, at 34 and 31%, respectively. The hydrophilic acids plus neutrals fraction contained 12% of the carbon and the hydrophobic acids fraction 10%. Other fractions individually contained 1–6% of the carbon recovered in the fractionation scheme.

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The total mass of organic carbon recovered by the organic matter fractionation procedure was 85.2 mg. The total organic carbon mass the quartz fiber filter composite was determined by the Sunset carbon analyzer measurements to be 102.3 mg (total carbon mass was 122.4 mg). Therefore, the organic carbon recovery of the fractionation procedure is 83%. This recovery percentage is approximate because of the various corrections and estimates for certain fractions presented in Table 4. Moisture contents were not determined for these fractions so that sufficient fraction mass would be left for nitrogen determinations. Recovery of 83% of the organic carbon (70% if one considers the total carbon) in the aerosol sample is quite good considering the following two factors. First, the organic carbon mass of the sample is exceptionally small for a total organic carbon fractionation of this type. Handling and transfer losses tend to increase percentage carbon losses as the sample mass decreases. Second, semivolatile neutral compounds, such as alcohols, aldehydes, and ketones, might be lost during the fractionation and isolation procedure. The recovery seen here is also comparable to other “comprehensive” approaches to particulate organic carbon characterization (Chang et al., 2005; Decesari et al., 2000; Graham et al., 2002).

The water soluble fractions account for 29.4 mg of the 85.3 mg of isolated material suggesting a WSOC fraction of 35%. This percentage compares reasonably well with an independent determination of dissolved organic carbon (DOC) in a filter leachate which estimated 25% of the organic carbon is soluble in water. The fraction of WSOC in Fresno aerosol samples is at the low end of the typical range of 35–80% commonly reported in the literature (Alves et al., 2002; Decesari et al., 2001; Graham et al., 2002; Kiss et al., 2002). Highly polluted urban areas, however, tend to have a lower fraction of water soluble organic particulates since freshly emitted carbonaceous particles, especially from combustion processes, tend to be more hydrophobic than processed aerosol particles or secondary organic species. The absence of chemically aged particles is probably enhanced by reduced photochemistry in the winter time period studied here. Second, Fresno experiences frequent fog and haze conditions in winter and saw a substantial amount of precipitation and drizzle during the study period (Collett et al.,

2007; Gorin et al., 2006). Under these conditions, the water soluble species are preferentially scavenged and removed from the atmosphere. Finally, fresh wood combustion is a major source of the ambient particle particulate matter observed in Fresno during the study period (Gorin et al., 2006). The main combustion source is fireplace and woodstove emissions. Flaming combustion, common in fireplaces, exhibits a WSOC fraction as low as 4% (Novakov and Corrigan, 1996) while smoldering conditions, like those in large biomass burning events commonly yield a much higher WSOC fraction (Graham et al., 2002).

The particulate organic matter fraction is the largest fraction in the aerosol sample (Fig. 2). The isolated fraction appeared black, having substantial (35.6%) aromatic (elemental) carbon content based upon the ^{13}C -NMR spectrum of Fig. 3. The elemental carbon content determined by ^{13}C -NMR spectrometry is 10.2 mg, about half the content (20.1 mg) of elemental carbon directly determined on the quartz fiber filters by the NIOSH method (Birch and Cary, 1996). The infrared spectrum (Fig. 4) of this fraction is free from any bands indicating inorganic substances such as the quartz fiber filter in which the aerosols were entrained. The sharp band at 2920 cm^{-1} and the sharp peak at 30 ppm in the ^{13}C -NMR spectrum are indicative of methylene groups in straight chain hydrocarbons. The band at 1700 cm^{-1} and small peak at 174 ppm indicate the presence of carboxylic acid groups. The broad band near 1620 cm^{-1} in the infrared spectrum indicates aromatic and hydrogen-bonded quinone groups.

The large difference in “elemental carbon” measured by the thermo-optical method in individual filters and the aromatic (elemental) carbon content measured in the particulate organic carbon fraction is intriguing. It is unlikely that such a large difference (approximately a factor of 2) can be explained by uncertainties associated with constructing the filter composite or by losses in the isolation procedure. More likely, a large fraction of what is measured as EC by the thermo-optical method is not a pure graphite (aromatic) structure but rather contains a substantial amount of other refractory organic material. EC is only operationally defined and its amount can vary largely with the analytical method used (Chow et al., 2001; Currie et al., 2002). The fact that a

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large aliphatic component (shown in the ^{13}C -NMR spectrum of Fig. 3) in the particulate organic matter fraction was not extractable with methylene chloride suggests that the aliphatic hydrocarbons are covalently bonded with the black carbon component, perhaps as the result of radical coupling reactions. Unfortunately our understanding of the chemical structure of soot in general and aerosol EC in particular is poor. This is not a problem unique to the aerosol science community; black carbon has also recently been found to be overestimated in soils and sediments because of methodological artifacts (Simpson and Hatcher, 2004).

The CH_2Cl_2 -extractable organic matter fraction is the second highest organic carbon mass fraction (see Fig. 2). The extractable organic matter in the sample consisted mainly of straight chain aliphatic hydrocarbons. These aliphatic hydrocarbons may be attached to carboxylic acid (1710 cm^{-1}), ester (1735 cm^{-1}), and aromatic (1620 cm^{-1}) structures which contribute small peaks to the infrared spectrum of this fraction (Fig. 4). Many aerosol organic speciation studies use methylene chloride or similar solvents to extract the organic fraction prior to GC/MS analysis. Hence solvent extractable fractions are well characterized and it has been shown that straight and branched alkanes, alkanols, alkanolic acids, and alkanals are substantial components of particulate organic matter (Alves et al., 2001; Fraser et al., 2002; Limbeck and Puxbaum, 1999; Rogge et al., 1993; Schauer et al., 1996). A more detailed discussion of the methylene chloride extractable fraction of these Fresno aerosol samples is given elsewhere (Gorin et al., 2006) and several other studies also addressed the solvent extractable fraction of Fresno aerosol (Poore, 2002; Schauer and Cass, 2000). Silicone bands (1260 , 1100 , 1053 and 804 cm^{-1}) are also present in the methylene chloride extract fraction as a result of contamination by silicone lubricant used on the seals in the vacuum rotary evaporator. The silicone bands are a minor mass contaminant as the organic carbon percentage of these silicone lubricants is about 33% whereas the organic carbon percentage of this fraction is 72.2% (Table 4). The organic carbon contribution by the silicone lubricants was estimated in Table 4 by measuring the ratio of the peak intensity at 2960 cm^{-1} (methyl groups in silicones) to the peak intensity at 2920 cm^{-1} (methyl-

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lene groups in sample hydrocarbons) and calculating the relative contributions of silicone versus sample components based upon absorptivity ratios of standard compound spectra published by Pouchert (Pouchert, 1985). The organic carbon contribution with the silicones was calculated as 3.13 mg.

5 The volatile acid fraction in aerosol samples has substantially less percentage mass (Fig. 2) than in fog samples collected in this region (Herckes et al., 2007). This is not surprising given the high solubility of weak gas phase organic acids in high pH (6–8) Fresno fogwaters. The aerosol volatile acid fractions is also substantially different in that acetic acid (measured as acetate), which is the dominant volatile acid in
10 the fogwater samples (Ervens et al., 2003), is a major component (peak at 26 ppm in the ^{13}C -NMR spectrum), but longer-chain (methylene peak at 30 ppm), and branched chain (methine peak at 38 ppm and methyl peak at 14 ppm) acids are present as well. The broad peak at 72 ppm might be indicative of volatile ether acids and the peak at 127 ppm indicates the presence of volatile aromatic acids. Many low molecular weight
15 acids, including formic and acetic acid along with branched and aromatic species have been reported in speciation studies of atmospheric particulate matter (Falkovich et al., 2005; Souza et al., 1999). The NMR peak at 171 ppm is indicative of sodium carbonate which is also indicated by the broad band at 1440 cm^{-1} in the infrared spectrum of Fig. 4. Sodium carbonate might result from drying this fraction at too high a pH value.

20 The mass of the colloid fraction, representing just 2% of the total recovered carbon, was sufficient only for an infrared spectrum (see Fig. 4). Inorganic sulfates are indicated by bands at 1125 and 607 cm^{-1} ; silicates are indicated by bands at 1095 and 460 cm^{-1} . Organic-structure bands at 1720 , 1665 , 1627 , 1513 , and 1413 cm^{-1} are suggestive of lignin structures (Faix, 1991). The hydrophobic neutral fraction contributes the smallest mass to the total organic carbon fractionation (1%), and only an
25 infrared spectrum (Fig. 4) was obtained. Organic-structure bands at 1720 , 1630 , 1515 , 1280 and 1113 cm^{-1} are again suggestive of lignin structures (Faix, 1991); the differing position of these bands relative to the similar bands reported in the colloid fraction may reflect a greater degree of oxidation and degradation of lignins in the hydrophobic neu-

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tral fraction. A small, sharp band at 1260 cm^{-1} suggests the presence of nitrate esters. Bands for nitrite esters and nitroso groups were not found in this fraction, in contrast to their presence in analysis of a Fresno composite fog sample collected over the study period (Herckes et al., 2007).

5 The hydrophobic acid fraction was isolated as a very “sticky” fraction which precluded obtaining a ^{13}C -NMR spectrum; however, an infrared spectrum (Fig. 4) and elemental analyses (Table 4) were obtained. Sharp bands in the infrared spectrum at 1630 and 1280 cm^{-1} indicate nitrate esters. The nitrogen content of this fraction (1.95%) is consistent with the presence of nitrate esters. Bands at 1515 , 1335 , and 1120 cm^{-1} may indicate again degraded lignin components (Faix, 1991).

10 The colloid, hydrophobic neutral, and hydrophobic acid fractions all show aromatic structures suggestive of lignin precursors. Humic-like substances derived from lignin-type precursors were reported to form in model cloud water (Hoffer et al., 2004). It is interesting that the lignin signal is stronger in the aerosol sample than in a fogwater sample collected at this location (Herckes et al., 2007) where fulvic acid was found to be derived from terpenoid precursors. Possibly, the phenol groups in lignin are more susceptible to oxidative degradation than in terpenoids where the phenol content is much lower. Organic nitrate esters were also found in the hydrophobic neutral and hydrophobic acid fractions, consistent with earlier findings of Garnes and Allen (2002).

15 The hydrophilic acid plus neutral fraction contains the majority of the water-soluble organic carbon (Fig. 2). The infrared spectrum (Fig. 4) has a strong C-O band at 1050 cm^{-1} and only a moderate carboxylic acid band at 1720 cm^{-1} . The ratio of these two bands indicates that neutral components rich in aliphatic alcohols are greater than carboxylic acid components. The ^{13}C -NMR spectrum (Fig. 3) indicates the fraction is very rich in carbohydrates (anomeric carbon peak at 102 ppm and alcohol peak at 20 72 ppm) this could be indicative of carbohydrates and anhydrosugars. The presence of a substantial amount of anhydrosugars, including levoglucosan, is consistent with separate measurements made on individual sample filters in a joint speciation study (Gorin et al., 2006) in which levoglucosan concentrations averaged $0.6\text{ }\mu\text{g}/\text{m}^3$ and wood com-

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bustion emissions were demonstrated to be a dominant contributor to the collected aerosol. Previous studies have found levoglucosan, a cellulose pyrolysis product, at concentrations up to 4.8% of the total organic carbon present (Schauer and Cass, 2000).

Both the infrared and the ^{13}C -NMR spectra for the hydrophilic acid plus neutral fraction indicate the presence of structural methyl groups (2960 cm^{-1} band and 21 ppm peak) that are not typically found in carbohydrate structures. Methylene hydrocarbon structures are indicated by the 2920 cm^{-1} band and 30 ppm peak. No aromatic carbon structures are found in this fraction. A possible explanation for the presence of branched methyl groups in this fraction is the formation of secondary organic aerosols through photooxidation of isoprene (Claeys et al., 2004). A number of methyl branched-chain polyols are formed through this atmospheric reaction, and their structures are consistent with spectral signatures in this fraction.

The base fraction is a relatively minor component of the Fresno aerosol (4% of recovered carbon), but it is a fraction that is very rich in nitrogen (Table 4) with an atomic C:N ratio of 3.5. The infrared spectrum (Fig. 4) of the base hydrochlorides does not show any indication of ammonium chloride; thus, the nitrogen content of this fraction is derived from the sample and not the ammonia eluent. The broad bands at 1720 , 1600 , and 1500 cm^{-1} are typical for a mixture of amino acids and peptides (Pouchert, 1985). The plateau between 3200 and 3000 cm^{-1} are N-H stretching modes of these amino acids and peptides. The atomic C:N ratio also supports the amino acid and peptide assignment. Again the presence of these species is not surprising in the nitrogen rich environment of Fresno. Zhang and Anastasio have shown high concentrations of free and total amino acids in different locations in the Central Valley of California (Zhang and Anastasio, 2003; Zhang et al., 2002). Separate measurements of dissolved organic nitrogen in Fresno fog samples collected during the study period confirm the importance of organic nitrogen in the area (Collett et al., 2007).

Comparing our results to other studies using fractionation techniques must be done with caution as other studies use different procedures and hence isolated fractions

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are not directly comparable. In addition, wintertime Fresno is a particular environment influenced by substantial agriculture emissions combined with dust, biomass burning (domestic wood burning) and vehicle emissions; frequently processed by local radiation fogs. The resulting fine carbonaceous material is only poorly water soluble. However
5 comparing our results with the model composition proposed by Fuzzi et al. (2001) for water soluble organic aerosol, there are many similarities. In our case a fulvic acid type material could reasonably represent ~40% of the water soluble organic material as suggested by Fuzzi. Levoglucosan and polyol type material also have similar importance as in Fuzzi's model, although small organic acid molecules are less important in
10 our case. Finally our observations reveal a higher organic nitrogen content than proposed for the aerosol model, perhaps reflecting different source types characteristic of the Fresno region.

Similar processing and analysis of organic material in Fresno fogs collected on several days during the same study (Herckes et al., 2007) permits an interesting comparison of the water soluble organic carbon (WSOC) fraction studied here with characteristics of the organic material actually scavenged by Fresno fog drops. Some important similarities and significant differences are both seen. The main organic fraction in both fog and PM_{2.5} WSOC are the hydrophilic acids, not surprising given their solubility and the likelihood that significant carbonaceous material is contributed to the fog drops by
20 nucleation scavenging. On the other hand the volatile acid fraction, mostly formic and acetic acid, is more important in fog droplets. As discussed above, this is not surprising given the high vapor pressure of these compounds which keep them mostly in the gas phase, except when high pH liquid water is available to dissolve them. We also note that the fog organic matter contained a substantial amount of biological organic material (cell wall fragments, proteins, acetylaminosugars, . . .) while this was not seen in the PM_{2.5} aerosol fraction. This difference might result from fog scavenging of coarse particulate matter, rich in biological material, excluded from our PM_{2.5} analysis. Finally hydrophobic, lignin precursor-type material is less important in fog compared to PM_{2.5} WSOC and also appears to feature somewhat different structures in these two phases.
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This might reflect different formation pathways, although more investigation is certainly necessary to test such a hypothesis.

4 Conclusions

The organic composition of fine particulate matter was investigated in Fresno (CA) during a 3 week winter period. A composite sample was characterized by isolating several distinct fractions and characterizing them by infrared and nuclear magnetic resonance (NMR) spectroscopy.

More than 80% of the organic carbon in the aerosol samples was recovered in the extraction and separation process and characterized. A majority of the organic matter was insoluble in water. The small water soluble organic carbon fraction (35%) is consistent with local winter time conditions with fresh, local particle emissions from domestic wood burning and urban traffic as well as preferential removal of soluble material by frequent radiation fog episodes.

An insoluble carbon fraction, which did not dissolve in water or dichloromethane, showed a strong aliphatic character paired with the expected aromaticity of this fraction which might be expected to reflect the presence of black carbon. Our results suggest that the aerosol black carbon is likely made up by aromatic structures associated with linear hydrocarbon chains. The methylene chloride soluble carbon fraction featured many spectral characteristics, typical of species commonly identified in speciation studies using this solvent, including alkanes, alkanols and alkanolic acids.

The main water soluble organic carbon fraction, the hydrophilic acid fraction, exhibited characteristics suggestive of carbohydrates, compounds common in biomass emissions. Potential polyol type compounds, suggestive of the presence of secondary organic material, were also observed. Several other carbon fractions appeared to contain phenolic material like lignin precursors, not unlike some suggested humic like structures.

Comparisons of the composition of the water soluble fraction of Fresno PM_{2.5} organic

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material showed a number of similarities to an earlier organic aerosol model published by Fuzzi et al. (2001), although greater amounts of nitrogen were found in Fresno particles. Comparison of the same water soluble fraction of PM_{2.5} OC with OC in Fresno fog water revealed that both were dominated by hydrophilic acids. Uptake of water soluble, low molecular weight carboxylic acids and the presence of greater amounts of biological material distinguished the fog water organic composition from the water soluble organic carbon particles that likely acted as cloud condensation nuclei.

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Table 1. Infrared frequency bands for various structures in organic matter isolates.

Compound Class	Frequencies (cm^{-1}), Structure (ϕ =aromatic carbon ring), and Elements (C=carbon, H=hydrogen, N=nitrogen, O=oxygen)
Carbohydrates	3400–3300 (O-H), 1100–1000 (C-O)
Fulvic Acid	3400–3300 (O-H), 2700–2500 (COOH), 1760 (COOR), 1720(COOH), 1660–1630(ϕ -C=O), 1280–1150 (ϕ -O, COOH)
Aliphatic Hydrocarbons	2960 (CH_3), 2940 (CH_2), 1460 (CH_2), 1380(CH_3)
Aromatic Hydrocarbons	~1,600 (ϕ), ~1500 (ϕ), 700–900, ϕ -H
Proteins	1660 (Amide 1 band, N-C=O), 1,540 (Amide 2 band, N=C-O)
N-Acetyl Amino Sugars	1660 (Amide 1 band, N-C=O), 1,550 (Amide 2 band, N=C-O) 1380 (CH_3)
Lipids	1760 (COOR), 1720 (COOH), 2960 (CH_3), 2940 (CH_2), 1460 (CH_2), 1380(CH_3)
Nitrate Esters	1630 (N=O), 1280 (N-O)
Nitrite Esters	1650 (N=O), 1370 (N-O)
Nitroso Groups	1400–1300 (N=O)

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Table 2. Characteristic infrared spectral peaks of inorganic solutes (in KBr pellets.)

Inorganic Solute	Characteristic IR Peaks (cm^{-1})
Boric Acid	3212, 2260, 1450, 1194, 548
Sodium Bicarbonate	2541, 1920, 1695, 1618, 1307, 1000, 837, 696
Sodium Carbonate	1440, 880
Sodium Nitrate	1385, 838
Phosphoric Acid	1007, 490
Disodium Hydrogen Phosphate	1159, 1074, 950, 860, 544, 521
Silicic Acid	1093, 964, 798, 468
Sulfuric Acid	1288, 1176, 1071, 1012, 889, 852, 617, 577, 455
Sodium Hydrogen Sulfate	1251, 1182, 1046, 865, 607, 577, 481
Sodium Sulfate	1122, 640, 608

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Table 3. Structural assignments for ^{13}C -NMR spectra.

[C, carbon; H, hydrogen; N, sulfur; O, oxygen; S, sulfur; ϕ , aromatic carbon ring]		
Chemical Linkage Between Elements	Compound Type	Chemical Shift Range (parts per million)
C-H	Aliphatic Hydrocarbon	0–55
C-N	Amines, Amides, Proteins	40–55
O-CH ₃	Methoxy Groups in Tannins and Lignins	55–60
C-O	Aliphatic Alcohols, Ethers, and Esters	60–90
O-C-O	Anomeric Carbon in Carbohydrates, Lactols	90–110
ϕ	Aromatic Carbon	95–165
ϕ -O	Aromatic Esters, Ethers, and Phenols	135–165
ϕ -SO ₃ H	Aromatic Sulfonic Acids	140–145
O=C-O, O=C-N	Carboxylic Acids, Esters, Amides	160–190
O=C-C=C	Flavones, Quinones	170–200
O=C-C	Aliphatic and Aromatic Ketones	190–220

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Table 4. Weight, and Elemental Analyses of Organic Matter Fractions from Fresno Aerosols (ND=Not Determined).

Fraction	mg	Percent Carbon	Percent Hydrogen	Percent Nitrogen	Milligrams Carbon	Corrected Carbon (mg)	% C
Particulate Organic Matter	49.4	60.83	4.65	2.37	30.05	28.8 ¹	34
Colloids	6.1	ND	ND	ND	ND	1.8 ^{1,2}	2
Volatile Acids, Na Salts	72.2	8.53	0.58	0.45	6.16	4.8 ^{1,3}	6
CH ₂ Cl ₂ tractable Organic Matter	Ex- 38.3	78.69	9.07	0.66	30.13	27.0 ⁴	31
Hydrophobic Neutrals	1.5	ND	ND	ND	ND	0.7 ^{1,2}	1
Hydrophobic Acids	19.3	45.56	5.24	1.95	8.79	8.5 ¹	10
Hydrophilic Acids plus Neutrals	48.7	23.50	3.62	1.81	11.44	10.5 ¹	12
Bases	9.9	33.37	5.66	11.14	3.30	3.1 ¹	4

1. Corrected for estimated moisture content. Moisture estimate based upon moisture content of same fraction in isolated fractions from Fresno fog samples (Herckes et al., 2007)
2. Corrected for estimated carbon percentages. Carbon percentage estimate based upon same fraction in isolated fractions from Fresno fog samples (Herckes et al., 2007)
3. Corrected by subtraction of inorganic carbonate percentage determined from ¹³C-NMR spectra of Fig. 3.
4. Corrected by subtraction of silicone organic carbon percentages estimated from the infrared spectrum of Fig. 4.

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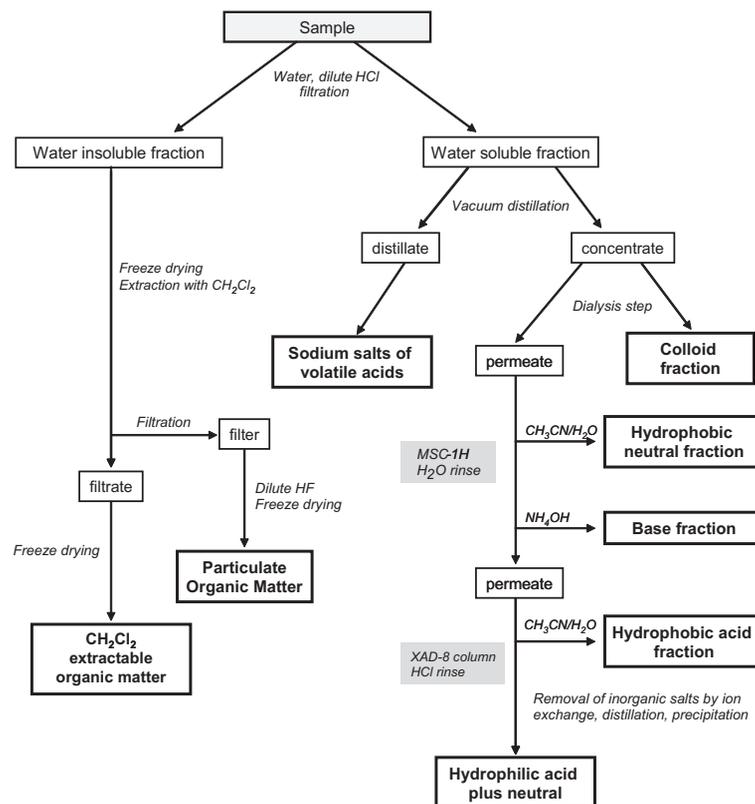
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**Fig. 1.** Overview of isolation scheme.

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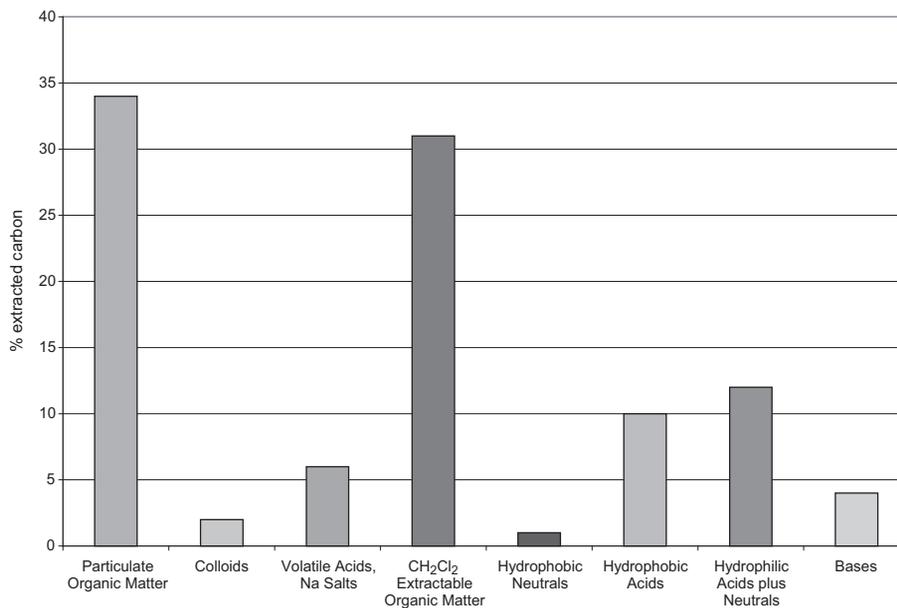
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**Fig. 2.** Total organic carbon fractionation of Fresno aerosols.[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

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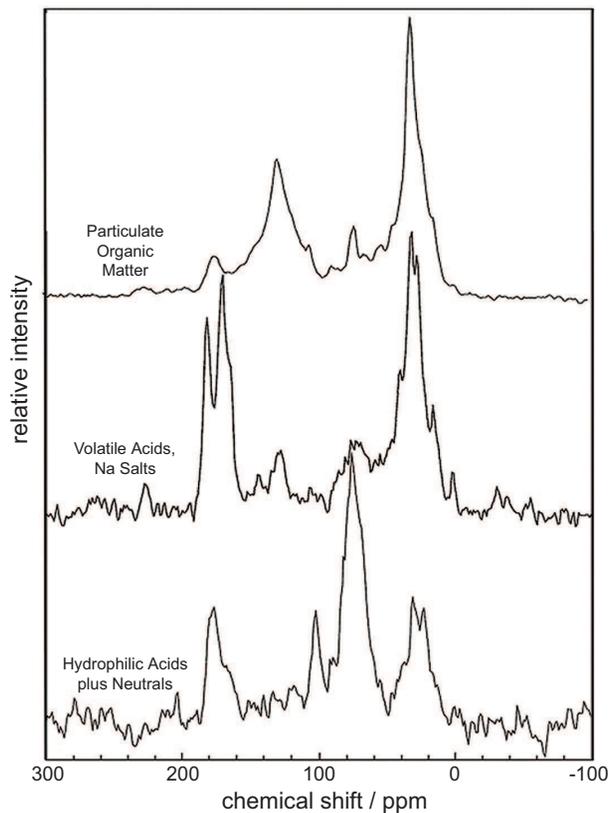
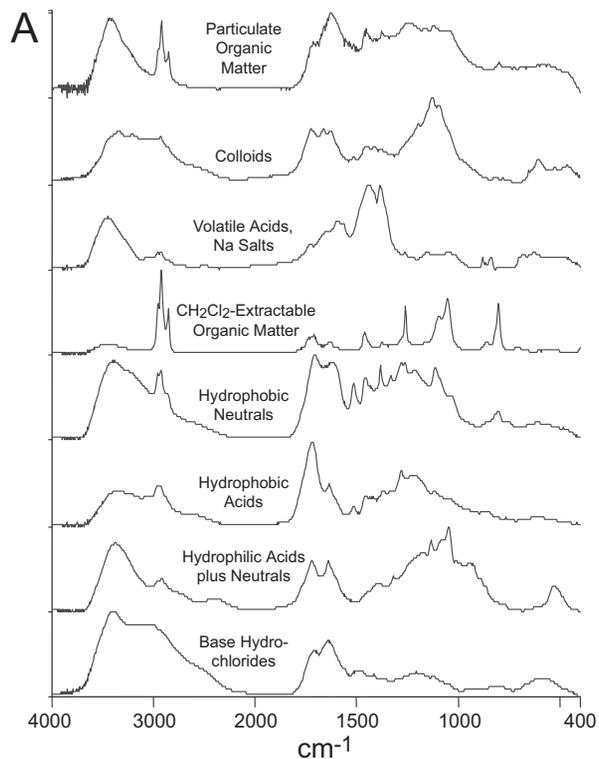


Fig. 3. ^{13}C -NMR spectra of organic matter fractions from Fresno aerosols.

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**Fig. 4.** Infrared spectra of organic matter fractions from Fresno aerosols.[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)