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INFLUENCE OF PREPARATION PROCESS ON PHYSICAL PROPERTIES AND
DEVITRIFICATION OF $\text{Li}_2\text{B}_2\text{O}_4$ (0.9) LiFe_5O_8 (0.1) GLASSES

C. Martin, C. Chaumont, J.P. Sanchez* and J.C. Bernier

*Département Science des Matériaux, E.N.S.C.S.⁺, 1 rue Blaise Pascal,
67008 Strasbourg Cedex, France***Centre de Recherches Nucléaires, B.P. 20 CR, 67037 Strasbourg Cedex, France*

Résumé: L'hypertrempe du mélange $\text{Li}_2\text{B}_2\text{O}_4(0,9)\text{-LiFe}_5\text{O}_8(0,1)$ a été réalisée en faisant varier la température du bain fondu et la vitesse des rouleaux. Les changements de propriétés physiques et du processus de dévitrification de ces verres résultent de la variation de la teneur en LiFe_5O_8 ou de l'apparition de Fe^{2+} . La nature de l'amorphe ne varie pas.

Abstract: Double roller quenching of $\text{Li}_2\text{B}_2\text{O}_4(0.9)\text{-LiFe}_5\text{O}_8(0.1)$ has been performed with various melt temperatures and roller speeds. The changes in physical properties or in the devitrification process of the amorphous samples are shown to be related to the LiFe_5O_8 content variations or to the Fe^{2+} appearance but not to structural changes of the amorphous state due to preparation processes

1 - INTRODUCTION

Structures and thermodynamic states of glasses are fixed by their composition but also by the freezing-in process at the glass transition T_g . Experimental and theoretical work has shown that glasses with same chemical composition can have different structures and physical properties (1), (2), (3), (4). We report here results on i) the influence of both the melt temperature and the roller speed (i.e. the cooling rate) on the physical properties of ternary oxides of nominal composition $\text{Li}_2\text{B}_2\text{O}_4$ (0.9) - LiFe_5O_8 (0.1) obtained by double roller quenching method ii) the influence of the preparation process on the crystallization behaviour of these glasses. Special attention has been paid to the behaviour of glasses heated at temperatures close to T_g . Physical properties of glasses of this composition have already been studied (5), (6). Curie Weiss behaviours with large negative paramagnetic Curie temperature are observed at high temperature ($>100\text{K}$), while depressed μ_{eff} are demonstrated at low temperature; a spin glass transition is evidenced nearly 3K.

2 - INFLUENCE OF MELT TEMPERATURE AND ROLLER SPEED ON THE PHYSICAL PROPERTIES OF THE GLASSES.

2.1 - Sample preparation.

A nominal composition melt $\text{Li}_2\text{B}_2\text{O}_4$ (0.9) - LiFe_5O_8 (0.1) has been quenched by double roller technique: first for six temperatures ranging from 850 to 1300°C with a constant peripheral speed velocity of 20 m/s.

Second, for constant temperature 1200°C but with various peripheral velocities from 1 m/s to 34 m/s.

The brittle flakes obtained are not entirely amorphous. Some of them are partially crystallized and contain small ferrimagnetic LiFe_5O_8 particles. The lightly ground quenched sample is sorted with a magnet and the non-magnetic part is checked by X-Ray diffraction and electron microscopy. The room temperature

magnetization curves of the non magnetic part show a linear field dependence in addition to a very weak ferrimagnetic component due to some LiFe_5O_8 crystallites. Therefore the crystallized LiFe_5O_8 content of every amorphous sample can be determined. The totality of LiFe_5O_8 is shown to crystallize after twelve hours annealing at 700°C . Magnetization measurements of the annealed samples allow the determination of the total LiFe_5O_8 content of each sample to be made.

2.2 - Melt temperature influence.

A previous paper (7) has established that the starting melt temperature influences the physical properties of the glasses through a change of chemical composition. Melt temperatures below 1100°C (which is approximatively the liquidus temperature) result in glasses with iron content (Fe^{3+} ions) below the nominal concentration, as expected from the semi-binary diagram of the $\text{Li}_2\text{B}_2\text{O}_4 - \text{LiFe}_5\text{O}_8$ system. On the other hand if the melt temperature is above 1100°C the iron concentration is close to the nominal one but the iron ions occur in both Fe^{3+} and Fe^{2+} (<20 %) oxidation states.

2.3 - Rollers speed influence.

Glasses obtained at different cooling rates by quenching the melt at 1200°C have the same chemical composition (close to the nominal one) but different thicknesses (Table 1). The amount of amorphous flakes increases with roller speed up to 20 m/s. Above this value the contact-time of the melt with the rollers is too short to ensure an efficient cooling.

Table 1:

Peripheral roller speed m/s	Flakes thickness (μm)	amorphous percent (wt %)	R.X.	Total LiFe_5O_8 content (mol %)	(wt%) crystallized LiFe_5O_8 fraction in the amorphous part
1	70	27	Amorphous	9.4	1.53
2	50	27	Amorphous	9.3	1.18
5	30	53	Amorphous	9.4	1.05
20	20	71	Amorphous	9.6	0.82
34	20	42	Amorphous	9.8	2.40

The local order of the amorphous samples has been investigated by Mössbauer spectroscopy. Assuming Lorentzian line shapes the Mössbauer spectra of each amorphous compound can be satisfactorily analyzed as a superposition of three quadrupole doublets with different isomer shifts and quadrupole splittings characterizing Fe^{3+} ions in both octahedral and tetrahedral coordinations (6). The asymmetry of the quadrupole doublet of Fe^{2+} has been introduced in the fitting program to take into account previous observations (8) (9). Mössbauer results demonstrate that the cooling rate has no effect on the nature of the local order (at least in the investigated range).

The evolution of the susceptibility versus temperature (from 4.2 to 300K) as well as the magnetization versus magnetic field (at 7 or 300K) are similar for all the amorphous samples (within the experimental errors) and agree with the previous conclusion.

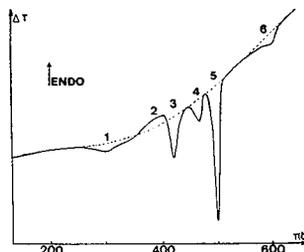
3 - DEVITRIFICATION PROCESS

Devitrification process has been investigated by DTA and X Ray diffraction. Devitrification proceeds by the crystallization of LiFe_5O_8 and lithium borate phases $\text{LiBO}_2\alpha$, $\text{LiBO}_2\gamma$ (respectively a monoclinic form and a tetragonal form of metaborates (10)) and one (or more) other borate phase, called LiBO_2X , that has not yet been referenced in the literature. This latter phase appears also during the devitrification process of 0.5 Li_2O -0.5 B_2O_3 glasses and its X ray diffraction diagram is very similar to the one observed during the dehydration process of $\text{LiBO}_2\cdot 8\text{H}_2\text{O}$ (to be published). This suggests that slight hydration could be responsible for the appearance of this phase.

A typical DTA curve performed in air ($2^\circ/\text{mn}$) is shown in figure 1 where the different DTA peaks are assigned as follows :

Fig.1.

- 1 : reoxidation of Fe^{2+} .
- 2 : endothermal effect due to the glass transition.
- 3 : crystallization of LiFe_5O_8 .
- 4 : crystallization of a Lithium borate phase ($\text{LiBO}_2\alpha$ or LiBO_2X).
- 5 : crystallization of LiBO_2X .
- 6 : transformation $\text{LiBO}_2\text{X} \rightarrow \text{LiBO}_2\alpha$.



The appearance of $\text{LiBO}_2\gamma$ in the X ray diffraction diagram of amorphous sample annealed at temperatures ranging from 450 to 580°C is due to the progressive transformation of $\text{LiBO}_2\alpha$ (X) to $\text{LiBO}_2\gamma$ as expected from the Li_2O - B_2O_3 diagram (10). The more developed exothermic peak corresponding to the LiBO_2X crystallization is shifted to lower temperatures when using finer crushed glass powder whereas the LiFe_5O_8 crystallization peak is almost not affected. This behaviour indicates that LiBO_2X and LiFe_5O_8 are respectively surface and bulk nucleated (11).

Glasses of the same composition, heated in non-oxidizing atmosphere exhibit the same DTA thermograph. In air, the LiFe_5O_8 crystallization peak is shifted to higher temperatures. This behaviour could result from oxidation of Fe^{2+} taking place just before LiFe_5O_8 crystallization.

4 - TRANSITION ZONE

4.1 - First results

The endothermal effect (Fig. 1) corresponding to T_g occurs at the same temperature (within experimental error (4°C)) for all samples with the same composition quenched with various cooling rates. Glasses heated just above T_g , show a softening point.

4.2 - Experimental.

Several amorphous samples of the same origin (quenched from 1100°C) have been heated ($2^\circ/\text{mn}$) in the DTA apparatus, at different temperatures in the transition zone (352, 379, 394, 400 and 405°C) and then, quenched in air.

4.3 - Results and discussion.

The three annealed samples at 379, 400 and 405°C show up a decrease in susceptibility below 180K with respect to the non annealed one's. The room temperature magnetization curve of the glass heated at 352°C is identical to the non-annealed sample. The magnetization curve of the amorphous sample annealed at 394°C exhibits a paramagnetic susceptibility (i.e. the slope of the linear part of the curve) slightly lower than for the non-annealed sample. Magnetization curves

of the annealed samples at 400 and 405°C reveal a curvature indicating a superparamagnetic behaviour with a very smooth saturation (Fig.2). Mössbauer spectrum, at 4.2K, (Fig.3) of the annealed sample at 379°C is identical to virgin-sample whereas the glass annealed at 394°C presents a better resolved hyperfine structure (Fig.3). It is concluded that the spin glass freezing temperature (θ) has increased with thermal treatment. The spectra of the annealed samples at 400 and 405°C (Fig.3) indicate an evolution toward a superparamagnetic behaviour in agreement with magnetization results.

Fig.2.
Room temperature magnetization curves of:
-virgin sample: *-*
-annealed samples at
394°C: □
400°C: ●

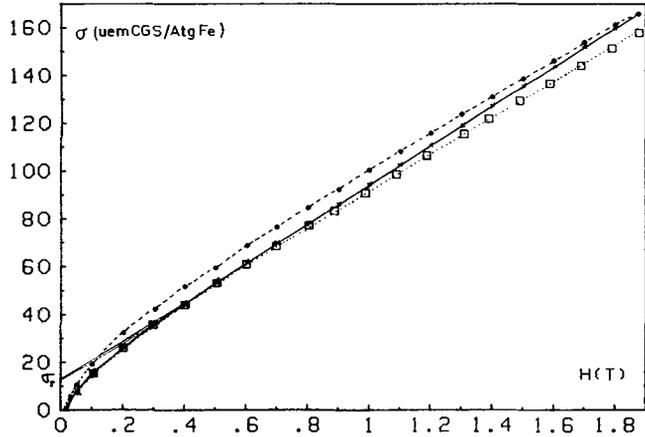
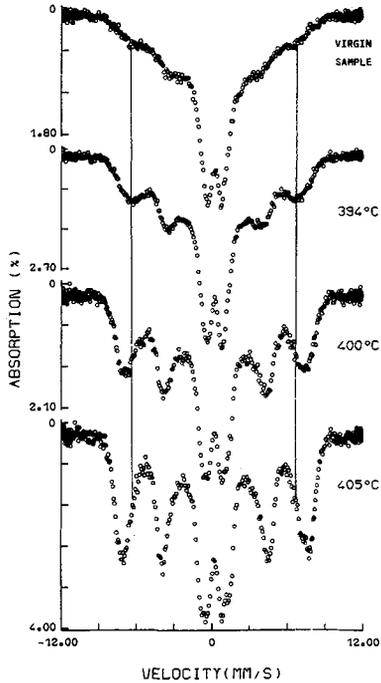


Fig.3.
4.2K Mössbauer spectra of amorphous samples annealed at different temperatures



Discussion:

The freezing temperatures increase with annealing temperature within the transition zone until the beginning of LiFe_5O_8 crystallization resulting in superparamagnetic behaviour. A growing of the cluster sizes or of the antiferromagnetic interactions could support this experimental finding. Cluster growth is evident if LiFe_5O_8 nucleation process starts in the transition zone. An increase of the antiferromagnetic interactions could result from a structural rearrangement increasing the superexchange coupling. The depressed paramagnetic susceptibility could also be explained by the raise of the antiferromagnetic interactions. The whole results can be related to a relaxation process (an evolution toward a more stable amorphous structure).

Prerenucleation or relaxation should occur when annealing in the transition zone. It may be questioned whether the two processes can be distinguished. For temperatures higher than T_g , the glass viscosity falls off rapidly so diffusion processes are accelerated and consequently nucleation and crystallization processes become faster.

5 - CONCLUSION.

This study demonstrates that the physical properties of these glasses are related to their composition variations or changes of valence states of minor components and not to structural changes of the amorphous state due to preparation process.

REFERENCES :

- 1) C. T. MOYNIHAN, A. J. EASTEAL, J. WILDER, J. Phys. Chem. 78 (1974) 2673
- 2) V. GOTTARDI, J. Non-cryst. Solids. 49 (1982) 461
- 3) G. S. GRETT, M. H. COHEN, Phys. Rev. B. 21 (1980) 4113
- 4) M. BARICCO, C. BATTEZZATI, F. MARINO, G. RIONTINO, Proceedings V int. conf. on Rapidly quenched Metals, Würzburg (1984)
- 5) C. CHAUMONT, J. C. BERNIER, J. Sol. State. Chem. 38 (1981) 246
- 6) J. P. SANCHEZ, J. M. FRIEDT, J. de Phys. 43 (1982) 1707
- 7) C. MARTIN, C. CHAUMONT, J. P. SANCHEZ, J. C. BERNIER, Rev. Chim. Min. 21 (1984) 657
- 8) J. F. SANCHEZ, J. M. FRIEDT, R. HORNE, A. J. VAN DUYNVELDT, J. Phys. C. Sol. State phys. 17 (1984) 127
- 9) A. SCHNELL, J. C. BERNIER, J. P. SANCHEZ, Mat. Res. Bull. 18 (1983) 251
- 10) C. MARAINE - GIROUX, R. BOUAZIZ, G. PEREZ, Rev. Chim. Min. 9 (1972) 779
- 11) P. W. Mc MILLAN, J. Non-cryst. Solids 52 (1982) 67