



HAL
open science

Europium retention onto clay minerals from 25 to 150 °C: Experimental measurements, spectroscopic features and sorption modelling

E. Tertre, G. Berger, E. Simoni, S. Castet, E. Giffaut, M. Loubet, H. Catalette

► **To cite this version:**

E. Tertre, G. Berger, E. Simoni, S. Castet, E. Giffaut, et al.. Europium retention onto clay minerals from 25 to 150 °C: Experimental measurements, spectroscopic features and sorption modelling. *Geochimica et Cosmochimica Acta*, 2006, 70, pp.4563-4578. 10.1016/j.gca.2006.06.1568 . in2p3-00105461

HAL Id: in2p3-00105461

<https://hal.in2p3.fr/in2p3-00105461>

Submitted on 11 Oct 2006

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

**EUROPIUM RETENTION ONTO CLAY MINERALS FROM 25 TO 150°C: EXPERIMENTAL
MEASUREMENTS, SPECTROSCOPIC FEATURES AND SORPTION MODELLING.**

Tertre E.^{1,2}, Berger G.¹, Simoni E.³, Castet S.¹, Giffaut E.², Loubet¹ M. and Catalette H.⁴

- ^{1.} LMTG (UMR CNRS/UPS/IRD n° 5563), 14 av. E. Belin, 31400 Toulouse, France
- ^{2.} ANDRA, Parc de la Croix Blanche, 92298 Châtenay-Malabry cedex, France
- ^{3.} Lab. Radiochimie, IN2P3, Université Paris 11, 91406 Orsay, France
- ^{4.} EDF, Site des Renardières, Ecuelles, 77818 Moret-sur-Loing cedex, France

Submitted December 2005 – Revised version in May 2006

Author to whom correspondence should be addressed:

Emmanuel TERTRE
USTL,
UMR PBDS/Bat SN5,
F-59655 Villeneuve d'Ascq, France

Email: Emmanuel.Tertre@univ-lille1.fr

Keywords: europium, sorption, clays, experimental, temperature, Time-Resolved Laser Fluorescence Spectroscopy (TRLFS)

Abstract

The sorption of Eu (III) onto kaolinite and montmorillonite was investigated up to 150°C.

The clays were purified samples, saturated with Na in the case of montmorillonite. Batch experiments were conducted at 25, 40, 80 and 150°C in 0.5 M NaClO₄ solutions to measure the distribution coefficients (K_d) of Eu as a trace element (<10⁻⁶ mol/L) between the solution and kaolinite. For the Na-montmorillonite, we used K_d results from a previous study (Tertre *et al.*, 2005) obtained under exactly the same conditions. The number and nature of the Eu species sorbed onto both clay minerals were investigated by Time Resolved Laser Fluorescence Spectroscopy (TRLFS) in specific experiments in the same temperature range. We identified a unique inner-sphere complex linked to the aluminol sites in both clays, assumed to be =AlOEu²⁺ at the edge of the particles, and a second exchangeable outer-sphere complex for montmorillonite, probably in an interlayer position.

The K_d values were used to adjust the parameters of a surface complexation model (DLM: Diffuse Layer Model) from 25 to 150°C. The number of Eu complexes and the stoichiometry of reactions were constrained by TRLFS. The acidity constants of the amphoteric aluminol sites were taken from another study (Tertre *et al.* submitted), which integrates the influence of the negative structural charge of clays on the acid-base properties of edge sites as a function of temperature and ionic strength. The results of the modelling show that the observed shift of the sorption edge towards low pH with increasing temperature results solely from the contribution of the =AlOEu²⁺ edge complexes.

Finally, we successfully tested the performance of our model by confronting the predictions with experimental K_d data. We used our own data obtained at lower ionic strength (previous study) or higher suspension density and higher starting concentration (TRLFS runs, this study), as well as published data from other experimental studies (Bradbury and Baeyens, 2002; Kowal-Fouchard, 2002).

Introduction

The radiative effects of nuclear waste are well known to drive an increase of temperature in the vicinity of the storage site. According to the scenarios developed for the French underground repository design, the temperature of the bentonite barrier would reach 100°C in the short term, thereafter decreasing with time (ANDRA, 2005). However, a greater rise in temperature, typically up to 150°C, remains to be considered as an option in evaluating performance. Such a thermal history raises a question about the retention of radionuclides on clay minerals at elevated temperatures. Indeed, sorption measurements of radionuclide analogues (lanthanides, Cs, divalent cations, etc.) have so far been mainly carried out at room temperature (Baeyens and Bradbury, 1997; Aja, 1998; Wang *et al.*, 2001; Coppin *et al.*, 2002; Kowal-Fouchard, 2002; Stumpf *et al.*, 2002; Stumpf *et al.*, 2004a; Rabung *et al.*, 2005). These studies point to the existence of two main types of sites, with exchange on interlayer sites and surface complexation on the edges of the particles. These conclusions are based on macroscopic properties (distribution coefficients (Kd) and acidity-basicity) as well as spectroscopic studies. Moreover, numerous studies have focused on evaluating the effect of an aqueous ligand (organic or not) on the sorption of trivalent cations, such as Eu(III) and Am(III), onto clay minerals. All of these studies were performed at room temperature. We can mention, for example, the studies of Pshinko *et al.* (2004) and Wang *et al.* (2001), which deal with the effect of fulvic acid on the sorption of Eu(III) onto clays. In addition, Sakuragi *et al.* (2002) studied the effect of both calcium ions and humic acid on the sorption of Am(III)/Eu(III) onto kaolinite. The main conclusions of these room-temperature studies is that humic/fulvic acid, as well as an aqueous ligand, can either increase or decrease the distribution coefficients of Eu(III)/Am(III), depending on pH and ionic strength, the nature of sorbent and the nature of the humic/fulvic substances or aqueous ligand. Therefore, to study the temperature dependence of lanthanide sorption onto clays, we performed our experiments in a simple electrolyte, such as NaClO₄, containing any competitor substances.

Some rare studies performed at higher temperatures with clay minerals (Angove *et al.*, 1998; Tertre *et al.*, 2005) show a weak but measurable thermal effect on the Kd of di- and trivalent cations. The authors found that the surface complexation increases slightly with increasing temperature, as generally observed for simple oxides, while there is only a negligible influence on the exchange processes at interlayer positions. On the other hand, Bauer *et al.* (2005) failed to detect any influence of temperature (up to 80°C) on the pH dependence of europium sorption onto montmorillonite. Even in the spectroscopic field, there are very few

studies characterizing sorbed lanthanides at elevated temperature. To our knowledge, only three research groups have developed specific experimentation for TRLFS measurements (time resolved laser fluorescence spectroscopy) at high temperature (see Kimura et al., 2002, Lindqvist-Reis et al., 2005, and Bauer et al., 2005). For experiments in aqueous solution, the first and second group show clearly that the change in hydration of trivalent cations with increasing temperature can be observed via the fluorescence properties of these elements. We develop these results in the section “TRLFS reference spectrum of aqueous Eu between 25 and 150°C”. On the other hand, using Eu sorption experiments onto smectite, Bauer et al. (2005) show that - at 60 and 80°C - europium forms complexes with aqueous organic matter. According to these authors, this organic material was released by the HDPE containers used in the experiments. Therefore, because of this artefact, they could not accurately attribute the temperature dependence of fluorescence properties of Eu(III).

Our previous experimental studies (Tertre *et al.*, 2005; Tertre *et al.*, this issue) provided Kd values up to 150°C for mono-, di- and trivalent metals, and proposed an original acidity-basicity model applicable up to 60°C. For the room-temperature results, the data were again interpreted in terms of complexation and exchange processes. We considered silanol and aluminol sites in the acidity-basicity model, rather than weak and strong sites.

The objective of this study was to develop a model for the sorption of lanthanides (acting as chemical analogues of trivalent actinides) onto two contrasted clay minerals, kaolinite and Na-montmorillonite. This model is based on the previously cited studies, while presenting some additional Kd measurements and specific spectroscopic analyses. We focus here on the retention of Eu (III), since this element is one of the lanthanides available for spectroscopic analysis.

In this study, we present a) new Kd measurements for sorption onto kaolinite up to 150°C, b) time-resolved-laser-fluorescence-spectroscopy (TRLFS) results for kaolinite, montmorillonite and single oxides at room temperature, as well as at 80° and 150°C, c) the sorption model inferred from these data and previously published measurements (Bradbury and Baeyens, 1997, 2002, and Kraepiel *et al.*, 1999, for trivalent and divalent cations; Liu *et al.*, 2003, for caesium). The advantage of our model is that it makes use of an original acidity-basicity model and considers spectroscopic constraints. Thus, it can even be applied at high temperature, since previous spectroscopic data are only available at room temperature: e.g. Stumpf *et al.*, 2001, for Cm (III) onto kaolinite and montmorillonite; Chisholm-Bause *et al.*, 2004, and Kowal-Fouchard *et al.*, 2004, for Eu (III) and uranyl onto montmorillonite.

Methods

Materials used for distribution coefficient measurements --- Suspensions of kaolinite and Na-montmorillonite were prepared for sorption experiments following the same procedure as presented in Tertre *et al.* (2005). The kaolinite sample originates from St Austell (UK), while the montmorillonite was isolated from the bentonite MX-80 and converted into the sodic form. The average grain size of particles is between 0.2-2 μm for both clay minerals, whereas the N_2 BET specific areas of the dried samples are 10 and 24 $\text{m}^2\cdot\text{g}^{-1}$ for kaolinite and Na-montmorillonite, respectively. As in the case of BET surface area, the cationic exchange capacity (CEC) is highly contrasted between the two clays. The exchange site density values range from 3.7 $\mu\text{mol}\cdot\text{m}^{-2}$ for kaolinite to 36 $\mu\text{mol}\cdot\text{m}^{-2}$ for Na-montmorillonite, according to the results reported by Coppin *et al.* (2002) and Baeyens and Bradbury (1997), respectively.

Distribution coefficient determination --- The distribution coefficients (K_d) of Eu (III) between kaolinite and 0.5M NaClO_4 solution were measured at 25, 40 80 and 150°C following the same protocol as presented in Tertre *et al.* (2005). Briefly, the protocol consists of:

- 1- Initial equilibration of the clay suspension (2.5g/L) at selected pH, ionic strength and temperature over a period of one week. After one week of reaction at a given temperature, the clay suspension remains stable in term of pH, silica and aluminium aqueous concentrations (see Tertre *et al.*, 2005).
- 2- Addition (without cooling) of 14 lanthanides, each at a concentration of 10^{-6} mol/L. This concentration is similar to that used in the previous studies of Coppin *et al.* (2002) and Tertre *et al.* (2005), and is assumed not to exceed the solubility of possible secondary solids, a hypothesis that is discussed in more detail below.
- 3- Three days of reaction at the selected temperature to achieve sorption equilibrium.
- 4- Separation (without cooling) of the aqueous phase from the clay fraction by centrifugation at 2000 g for 30 min for the runs at 25°C and 40°C, or by filtration at 0.2 μm for the runs at 80°C and 150°C.
- 5- Analysis of the filtrates by ICP-MS, followed by K_d determination using mass balance equations for each lanthanide (in particular Eu). Measurements of aluminium and silica contents were also performed to check that clay dissolution was negligible during the runs, assuming no concomitant aluminosilicate precipitation.

We did not perform specific new K_d measurements for montmorillonite, since the published data in Tertre *et al.* (2005) were considered sufficiently exhaustive. However, the runs carried out for the spectroscopic measurements in the present study (see next section) provide additional K_d data for montmorillonite as well as for kaolinite.

Evaluation of possible artefacts: As mentioned in Tertre *et al.* (2005), all batch experiments were performed in PTFE containers. Sorption onto the walls of the containers was investigated in the presence of clay mineral. The maximum europium retention was found under alkaline conditions (6 % at pH=8 and 25°C). These experiments were performed for a lanthanide initial concentration of 10^{-7} mol/L, which was 10 times lower than the concentration used in the sorption experiments (10^{-6} mol/L). Therefore, we assumed that the sorption onto the containers was negligible under acidic and neutral conditions, and that the protocol was valid under alkaline conditions with a 6% confidential interval. Moreover, to ensure that fluoride ions, a strong ligand for aqueous lanthanides, were not released by the PTFE containers during experiments, some specific “blank” measurements were carried out at 25 and 150°C using a pure 0.025 M NaClO₄ background electrolyte solution. After ten days (duration of the total batch experiments), the aqueous fluoride concentrations did not exceed 10^{-7} mol/L, corresponding to the detection limit of the ionic chromatographic method used.

Evaluation of dissolution effects --- Sorption measurements assume that the reactive sites are stable with respect to the solution during the duration of the process. This question becomes important at temperature as high as 150°C, as well as the possible modification of the clay surface by sorption or coating of secondary phases. Even if the clay material is not transformed into another more stable phase within 10 days at 100°C (Beaufort *et al.*, 2001), the dissolution of a fraction of the particles is almost certain during the experimental runs, at least at elevated temperature. In a previous study using the same experimental protocol, Tertre *et al.* (2005) reported that dissolution reactions consume 6% of the smectite mass at 150°C under alkaline conditions. On the other hand, Zysset and Schindler (1996) and Huertas *et al.* (1999) found, respectively, that smectite and kaolinite dissolution at 25°C is negligible (around 0.3% in mass for both minerals). Nevertheless, Carroll and Walther (1990) mentioned that this value increases significantly with temperature (9% for kaolinite at 80°C). Although dissolution is likely to occur at the same reactive site as sorption, its consequence is probably negligible during the development of fresh surfaces, because the dissolution rate is slower

than the sorption process. We can assume that the sole effect of dissolution is a limited modification of the V/M ratio, which is nevertheless included in the calculation of K_d uncertainties. However, the worst scenario is incongruent dissolution associated with the precipitation of secondary phases, which can then form a coating onto clay surfaces. In this case, the secondary phases can compete with the clay material for lanthanide retention. To avoid incongruent dissolution, and to minimize its extent, we shortcut the first stage of reaction far from equilibrium by using starting solutions previously equilibrated with respect to quartz. This method was successfully used in Berger *et al.* (2002) for approaching chemical equilibrium with the aluminosilicate K-feldspar up to 300°C, but without precipitation of secondary clay or hydroxides (which are common products when reactions are started very far from equilibrium). Then, we waited one week before adding the lanthanide-bearing solution and beginning the sorption process. According to methodological data reported in Tertre *et al.* (2005), most of the dissolution occurs during the first few days of equilibration before the sorption stage. Consequently, we assume that sorption sites are not significantly perturbed by dissolution during the sorption stage, even at 150°C.

Sample preparation for TRLFS measurements --- The TRLFS spectrum and fluorescence times were acquired on clay suspensions at 25, 80 and 150°C. The suspensions were prepared at various pH in 0.5M NaClO₄ solutions (see Table 1 for characteristics of each sample). The protocol used for the sample preparation was similar to that developed for the sorption experiments, but with some specific adaptations to enhance the intensity of the TRLFS signal. As in the “sorption experiments”, the suspensions were left to react over several days before the addition of europium to the system. The equilibration lasted three days in standard PTFE containers kept at constant temperature in ovens. Compared with the “sorption experiments”, the main differences in the sample preparation protocol were as follows; a solid/liquid ratio taken as equal to 25g/L (instead of 2.5 g/L), a starting europium concentration of 10⁻⁵mol/L (instead of 10⁻⁶ mol /L), and the use of standard PTFE containers at all temperatures. After completion of the sorption, to concentrate the solid fraction and consequently enhance the TRLFS signal, the containers were centrifuged for 5 minutes in preheated supports with thermal insulating properties (PTFE). The concentrated solutions were sampled and immediately introduced into preheated quartz tubes, which were sealed by a blowing technique and placed in a specific furnace within the optical device. The furnace maintained the sample at the desired temperature, while allowing the laser beam to traverse the quartz tube.

Because the PTFE containers were not equipped with either injectors or internal filters, we briefly reduced the temperature during the 150°C runs (for 5 min at 90°C), to allow the addition of Eu and sampling of the suspension. The centrifugation step also involved a short drop in temperature. Despite our efforts to minimize cooling during the 5 min of centrifugation, we measured a temperature drop to 80-90°C for the 150°C runs and down to 74°C for the 80°C-runs.

The TRLFS measurements were carried out with a tunable OPO Panther Continuous laser operating at 465 nm or 393 nm. The detection was performed by a Spectra-Pro-300 monochromator (Acton Research Corporation) coupled with a CCD camera (Princeton Instruments). The emission spectrum of Eu (III) was recorded in the range 550-650 nm using the software Winspec (Princeton Instruments). For measuring the emission decay, the delay time between laser pulse and camera gating was scanned with time intervals between 10 and 300 μ s. To determine the fluorescence times of the sorbed species, the decay profiles were then fitted to (multi)-exponential laws.

Table 1 summarizes the experimental parameters for clay samples analysed by the TRLFS technique. As previously mentioned, the initial concentration of Eu (III) in these samples, before sorption, was close to 10^{-5} mol/L. This is nearly 10 times higher than the value of 10^{-6} mol/L used for experiments performed without spectroscopic study. Consequently, these samples were duplicated in order to measure the distribution coefficients. The K_d values measured under such conditions (25g/L and $C_{Eu}=10^{-5}$ mol/L), as reported in Table 1, are similar to those obtained with lower initial concentration and solid/solution ratio (2.5g/L and $C_{Eu}=10^{-6}$ mol/L). The fact that the K_d values are independent of the aqueous concentration suggests that the sorption sites are not saturated (linear region on a Langmuir isotherm).

To achieve a better characterization of the surface sites involved in the sorption of europium onto clays, we compared the previous results with TRLFS measurements performed for europium experimentally sorbed onto single oxides: γ -alumina (γ -Al₂O₃), gibbsite (Al(OH)₃), quartz (SiO₂) and amorphous silica. The γ -alumina sample provided by Merck® (ref: 90) has grain size ranging from 63 to 200 μ m and a BET surface area of 135 m².g⁻¹ (Marmier, 1994). The gibbsite sample was previously used by Castet et al. (1993), and is a pure synthetic compound (α -Al(OH)₃) obtained from Riedel®-De Haën® AG SEELZE, Hannover-2504301). The quartz sample is the 50-125 μ m size fraction of the sample previously used by Gautier et al. (2001), with a BET surface of 0.061 m².g⁻¹. The amorphous silica is a silica powder

distributed by Baker[®]. The runs performed with single oxides were conducted at 25 and 80°C, but only at neutral pH and using suspensions prepared in 0.5 mol/L NaClO₄ solutions having a solid/solution ratio equal to 25 g/L. The rest of the protocol was identical to the clay experiments.

Experimental results and discussion

Sorption data --- Figures 1A and 1B show the K_d values for europium sorption onto kaolinite and Na-montmorillonite obtained in 0.5M NaClO₄ suspensions as a function of pH (2.5-9.5) and temperature (25, 40, 80 and 150°C). Data are similar for the other lanthanides (not shown). Previous studies using Eu alone (Coppin, 2002; Tertre, 2005) yield the same K_d values, suggesting that the results were unaffected by possible competition between the lanthanides at this concentration level ($<10^{-6}$ mol/L). For both minerals, we also report Eu data obtained at 25°C by Coppin *et al.* (2002), who used similar procedures and clay samples. All these data show that the Eu(III) sorption is clearly pH dependent, with a net increase of K_d with increasing pH. This behaviour is observed for both minerals and at all investigated temperatures. The strong influence of pH on K_d is well known and can be explained by the contribution of sorption on sites exhibiting amphoteric properties at the edges of particles, which occurs in addition to sorption on the exchange sites. Therefore, two different sorption mechanisms are proposed in the literature: (1) exchange reactions with little or no dependence on pH, which dominate sorption at low pH, and (2) strongly pH-dependent surface complexation reactions, which dominate sorption under neutral and alkaline conditions (e.g. Bradbury and Baeyens, 1997 and 2002; Kraepiel *et al.*, 1999). The temperature increase leads to a shift of the sorption curves towards lower pH, which is more pronounced for the Na-montmorillonite than for kaolinite. As previously mentioned for the Na-montmorillonite in Tertre *et al.* (2005), we observe that exchange reactions are practically temperature-independent, whereas surface complexation reactions are endothermic. Nevertheless, for both minerals, the K_d s obtained at 150°C are very high even at acidic pH ($\text{pH}>4.5$ for kaolinite and $\text{pH}>4$ for Na-montmorillonite). This strong K_d increase at 150°C, which contrasts with the behaviour observed at lower temperatures, raises questions about the nature of the retention at elevated temperature. It is debateable whether there is any diffusion of Eu inside the clay structure or precipitation of Eu-rich phases during the runs. The TRLFs spectroscopic analyses presented below, which allow a better characterization of the nature of Eu retention, are well suited to address this question.

TRLFs reference spectrum of aqueous Eu between 25 and 150°C--- For aqueous europium, the TRLFs method is usually performed at room temperature (Moulin *et al.*, 1999; Choppin and Peterman, 1998; Planque *et al.*, 2003). To our knowledge, only two published studies have addressed the acquisition and interpretation of fluorescence spectra of aqueous trivalent

cations (Eu^{3+} or Cm^{3+}) at high temperature: Kimura et al. (2002) up to 227 °C (500 K) and Lindqvist-Reis et al. (2005) up to 200 °C. We therefore began by recording fluorescence spectra and lifetimes of aqueous europium between 25 and 150°C in 0.5M NaClO_4 solutions. The experiments were conducted using the same procedure as for the sorption experiments. In addition, thermodynamic calculations were carried out to identify the speciation of aqueous europium under the conditions of TRLFS acquisition. These calculations were performed with the computer code CHESS[®] (Van der Lee, 2002) and using the reaction constants reported in Table 2. At this stage, we do not allow precipitation in the calculations. For the purposes of calculation, we assumed NaCl as background electrolyte because of the lack of thermodynamic data for NaClO_4 at high temperature. The dissociation constants related to H_2O , NaCl, NaOH and HCl are the same as those reported in Tertre et al. (2005), whereas the constants for hydrolysed and carbonated europium are reported from the studies of Turner et al. (1981), Wood (1990), Johnson et al. (1992) and Spahiu and Bruno (1995). Calculations were performed for 0.5M NaCl solutions containing 10^{-5} mol/L of total aqueous Eu(III) and with pH varying from 3 to 10. Although our experiments were run in hermetically sealed systems to prevent contamination by atmospheric CO_2 , the containers were opened for a few minutes to inject the europium and transfer the concentrated suspensions into the quartz tubes. Thus, due to possible minor contamination by CO_2 , we carried out speciation calculations assuming two boundary cases: absence of CO_2 contamination and full equilibrium with the atmosphere. In both cases, at 25°C and pH 3-7, the sole significant species of Eu(III) in solution is Eu^{3+} . At pH 7-8, the main species is EuCO_3^+ if the solution is equilibrated with the atmospheric CO_2 ($p_{\text{CO}_2}=10^{-3.5}$ atm), whereas the two species Eu^{3+} and EuOH^{2+} coexist in the absence of CO_2 contamination. In any case, whatever its absolute concentration, the proportion of hydrolysed species increases with temperature at constant pH. As an example, the proportion of $\text{Eu}(\text{OH})_{3\text{aq}}$ is nearly equal to 0% at 150°C and pH=3, but rises sharply with increasing pH: 16% at pH=5 and 80% at pH=7, while Eu^{3+} dominates at 25°C below pH 7. Moreover, we also note an increase in the proportion of carbonate species such as EuCO_3^+ with increasing temperature, since CO_2 is less soluble at high than at low temperatures. Consequently, even if Eu^{3+} is the main Eu(III) species at 25°C over the entire pH range investigated (pH 3-7), this is not necessarily the case at higher temperature since hydrolysed and carbonate species can become preponderant. Figure 2 reports the results of the speciation calculations performed at 25°C and 150°C as a function of pH and in the case of $p_{\text{CO}_2}=0$ atm. In comparison with Fig. 2A, Fig. 2B clearly show a shift of the percentages of hydrolysed species towards lower pH at 150°C.

Then, in a second stage of the speciation calculations, we include the precipitation of mineral phases. As expected, the thermodynamic calculations indicate an oversaturation of Eu hydroxides and carbonates (in the CO₂ system) under alkaline conditions (pH>7.5 at 25°C). For this reason, the subsequent TLRFS measurements were performed at lower pH, outside the stability range of these secondary phases.

Figure 3 presents fluorescence spectra of Eu (III) in 0.5M NaClO₄ solutions at pH=3, for temperatures of 25, 80 and 150°C. Spectra are characterized by two peaks, at 593 nm and 617 nm, corresponding to transitions from the excited state at 17374 cm⁻¹ (⁵D₀) to ground states at 374 cm⁻¹ and 1036 cm⁻¹ (⁷F₁ and ⁷F₂), respectively. The transition at 617 nm is hypersensitive and, as mentioned by Moulin *et al.* (1999) and Plancque *et al.* (2003), its intensity depends on the atomic environment of Eu(III). In our experiments at 25°C, the I₅₉₃/I₆₁₇ peak ratio is nearly equal to 1.5. This value is much lower than the range (4-8) reported in Plancque *et al.* (2003) at the same temperature and for similar conditions. Nevertheless, the same authors also mentioned that the I₅₉₃/I₆₁₇ peak ratio could vary widely (even if the main species is always free europium) due to changes in the second co-ordination sphere of europium. Such changes can be caused by interaction with the electrolyte used. For example, the recent study of Rabung *et al.* (2005) shows that the I₅₉₃/I₆₁₇ ratio is equal to 1.66 for free europium in CaClO₄, whereas this ratio falls to 1.2 when aqueous europium is complexed with nitrates at pH=3.5. Therefore, the low value of 1.5 obtained here for the I₅₉₃/I₆₁₇ ratio can be explained by the presence of small quantities of aqueous nitrates, since the lanthanide stock solutions were conditioned in HNO₃ 4%.

In our experiments, the I₅₉₃/I₆₁₇ peak ratio clearly decreases with increasing temperature (see figure 3). The same behaviour is observed at higher pH (not shown on the figure). Under acidic conditions (pH=3), the corresponding recorded lifetimes also decrease with temperature, from 110±5 μs at 25°C to 95±5 μs at 150°C. As mentioned by Kimura *et al.* (2002) for free aqueous europium, Lindqvist-Reis *et al.* (2005) for aqueous curium and Kirishima *et al.* (2004) for aqueous uranyl, both the fluorescence intensity and lifetime show a decrease with increasing temperature. Indeed, Kimura *et al.* (2002) found that the fluorescence time of free europium in 0.1M HClO₄ solutions decreases from 111 μs at 20°C to 100 μs at 130°C. On the other hand, Lindqvist-Reis *et al.* (2005), working on the hydration of Cm³⁺ in 0.1M HClO₄ solutions from 20 to 200°C, reported that the lifetime of free aqueous curium fell from 64 ± 3 μs at 25°C to 42 ± 4 μs at 200°C, which they interpreted by a

temperature dependant equilibrium between $[\text{Cm}(\text{H}_2\text{O})_9]^{3+}$ and $[\text{Cm}(\text{H}_2\text{O})_8]^{3+}$. Anyway, these authors (*op. cit.*) concluded that the decrease of the lifetime between 20 and 200°C, under acidic conditions, was not due to changes in the environment around free curium. Their interpretation is consistent with speciation calculations clearly demonstrating that Cm^{3+} is the main aqueous form of curium (III) at 20 and 200°C, when $\text{pH} < 3$. Consequently, following the conclusions of Lindqvist-Reis et al. (2005), the slight temperature effect observed here on the aqueous Eu (III) lifetime would not be due to a significant change of the aqueous speciation, and we assume that Eu^{3+} is the main species of Eu(III) in solution at $\text{pH} = 3$ whatever the temperature.

At $\text{pH} = 7.0$ and 25°C, the lifetimes and fluorescence spectra (not shown here) are both similar to those obtained under acidic conditions. We can conclude that Eu^{3+} is the main form of aqueous europium under these conditions, as concluded also for acidic medium. This is in agreement with results of thermodynamic calculations presented in the last section. On the other hand, at neutral pH and 150°C ($\text{pH} = 5.8$) the I_{593}/I_{617} ratio decreases compared to that measured for the spectrum recorded under acidic conditions ($I_{593}/I_{617} = 0.8$ at $\text{pH} = 3$; $I_{593}/I_{617} = 0.6$ at $\text{pH} = 5.8$). At 150°C and $\text{pH} = 5.8$, the corresponding fluorescence lifetime was equal to $50 \pm 10 \mu\text{s}$. In their study performed at room temperature, Plancque et al. (2003) showed that the fluorescence times of Eu^{3+} and hydrolyzed complexes are clearly different ($110 \pm 10 \mu\text{s}$ for Eu^{3+} ; $50 \pm 5 \mu\text{s}$ for EuOH^{2+} ; $40 \pm 5 \mu\text{s}$ for $\text{Eu}(\text{OH})_2^+$ and $\text{Eu}(\text{OH})_{3\text{aq}}$; $180 \pm 20 \mu\text{s}$ for EuCO_3^+). Therefore, our measured lifetime of $50 \pm 10 \mu\text{s}$ can be attributed to the presence of hydrolysed species: EuOH^{2+} , $\text{Eu}(\text{OH})_2^+$ and $\text{Eu}(\text{OH})_{3\text{aq}}$. Given the uncertainty of the measurements ($\pm 10 \mu\text{s}$), we cannot distinguish these three species. Nevertheless, this conclusion is in agreement with the speciation calculations showing that hydrolysed species are the main form of aqueous europium at neutral pH and 150°C.

Consequently, the results obtained by fluorescence measurements and thermodynamic calculations imply that Eu^{3+} is the dominant form of aqueous Eu (III) between pH 3 and 7 at 25°C, whereas the proportion of hydrolysed species increases not only with pH but also with temperature, becoming the dominant aqueous species above pH 4.3 at $T = 150^\circ\text{C}$. The relative abundance of EuCO_3^+ only becomes significant at low temperature and $\text{pH} > 7$ in solutions equilibrated with the atmosphere. However, these conditions are not representative of our experimental runs, suggesting that carbonate species can be ignored for our purposes.

TRLFS spectrum of sorbed Eu (III) --- Previous studies on lanthanide/clay systems using similar experimental protocols (Stumpf *et al.*, 2001; Stumpf *et al.*, 2002; Coppin *et al.*, 2002; Kowal Fouchard *et al.*, 2004; Rabung *et al.*, 2005) have shown that Ln(III)-bearing phases do not precipitate during sorption experiments conducted at 25°C. This was suggested by the consistence of the sorption data and further supported by TRLFS data (Stumpf *et al.*, 2002; Kowal Fouchard *et al.*, 2004). However, the Kd values obtained in the present study are much higher at 150°C and near neutral pH than otherwise at lower temperatures. The possible precipitation of a europium-bearing phase is backed up by the thermodynamic calculations in this study, which also provide information on the saturation state of the solutions. Even at 25°C, for pH higher than 7, the starting solutions (before Eu sorption is completed) are oversaturated with $\text{Eu}_2(\text{CO}_3)_3$ and $\text{Na}_3[\text{Eu}(\text{CO}_3)_3]6\text{H}_2\text{O}$ when equilibrated with the atmosphere, and with $\text{Eu}(\text{OH})_3$ in absence of CO_2 . To avoid any ambiguity in interpreting the Kd values, as well as improve characterization of the complexes sorbed onto the particles, we recorded fluorescence spectra and corresponding lifetimes for concentrated suspensions at 25 and 150°C, after sorption of 10^{-5} mol/L Eu. For both minerals, we analysed two samples for each temperature. One sample was prepared under acidic conditions, where sorption is independent of pH and mainly explained by an exchange process with the interlayer cations. A second sample was prepared at higher pH, where sorption is close to 100% and generally interpreted by surface complexation (Bradbury and Baeyens, 2002; Kowal-Fouchard, 2002). As mentioned in the methods section, the characteristics of these samples are reported on Table 1. Because most of the europium added to the system is sorbed onto the clays, or at least 50% at low pH, and given the increase of the clay fraction in the concentrated suspension, the monitored spectra are representative of the sorbed Eu, whatever the pH.

Figures 4A and 4B present examples of the emission spectra of europium sorbed onto Namontmorillonite at 25 and 150°C, respectively, under two different pH conditions. Irrespective of temperature or mineral, the spectra are similar and show the presence of the most intense broad band at 617 nm. These spectra are clearly different from the aqueous Eu(III) spectrum, suggesting significant changes in the atomic environment of europium. Such a behaviour have been already reported in the literature for the clay/europium system at 25°C under neutral and basic pH conditions, but is still discussed for the acidic conditions. Stumpf *et al.* (2002) found that the I_{593}/I_{617} peak ratio is higher than 1 (equal to 2 at pH=3.5), whereas Kowal Fouchard *et al.* (2004) found a lower value (equal to 0.6 at pH=3.1). We note that the two last authors recorded their fluorescence spectra with a similar montmorillonite

and in 0.025M NaClO₄ solutions. While the main species of Eu(III) in solution is Eu³⁺ in both cited studies, there is a difference in the results obtained by Stumpf et al. (2002) and Kowal Fouchard et al. (2004), who measured lifetimes of 110±5 μs and 75±10 μs, respectively. No rational interpretation has yet been advanced for this discrepancy. One possible explanation could be a change in the second coordination sphere around free europium, as mentioned by Plancque et al. (2003). In any case, for our own experiments at 25°C, the fluorescence spectra and lifetimes recorded under acid conditions for the clay/europium system are close to those observed by Kowal Fouchard et al. (2004). Indeed, we measured a fluorescence time of 85±10 μs for the montmorillonite/europium system (see figure 5) and 88±10 μs for the kaolinite/europium system, whereas Kowal Fouchard et al. (2004) measured a fluorescence time of 75 μs for a montmorillonite/europium system at similar pH and ionic strength.

Moreover, as shown on Figs. 4A and 4B, the I₅₉₃/I₆₁₇ peak ratio clearly depends on the pH of the suspension and decreases with increasing pH. This behaviour is in good agreement with the results of many studies (Stumpf et al., 2001; Stumpf et al., 2002; Kowal Fouchard et al., 2004; Rabung et al., 2005). We attribute this effect to a pH-related change in the atomic environment of the sorbed Eu(III), as suggested by the fluorescence time measurements. The decay curves obtained for all samples can be adequately fitted assuming two exponential decays corresponding to two different species. For the montmorillonite/europium system, one of the species has a fluorescence time of 85±10 μs at 25°C, which decreases slightly with increasing temperature (74±10 μs at 80°C and 66±10 μs at 150°C). The other species has a fluorescence time equal to 240±40 μs, which also slightly decreases with increasing temperature (200±40 μs at 80°C and 170±45 μs at 150°C). We should mention that the trends are similar in the kaolinite/europium system. At least at 25°C, these values are quite different from the lifetime obtained for solid Na₃[Eu(CO₃)₃]6H₂O, but are close to lifetimes for Eu₂(CO₃)_{3(solid)} and Eu(OH)_{3(solid)} (310 μs, 80 μs and 220 μs, respectively, according to Kowal-Fouchard, 2002). Even though europium carbonates and hydroxides can precipitate under neutral and basic conditions, lifetimes of 85 μs and 240 μs were also recorded at low pH. This suggests that Eu-bearing carbonates or hydroxides did not precipitate during our experiments. The species with a lifetime of 85 μs is assumed to be an outer-sphere complex since its lifetime is close to the value of 110±10 μs (at 25°C) for aqueous Eu³⁺ (Choppin and Peterman, 1998). The species with a lifetime of 240 μs is assumed to be an inner-sphere complex bound to the edges of the particles. Regarding the behaviour of Eu³⁺ in solution, we

assume that the effect of temperature on the lifetimes of the two sorbed complexes does not result from a significant change in the Eu (III) environment. On the contrary, it is more likely related rather to an intrinsic dependence of fluorescence time on temperature. To evaluate the relative abundance of each sorbed species as a function of pH, we calculated the proportion of outer-sphere complexes. Using the fit of the decay profiles, we estimated the proportion of outer-sphere complexes at each studied pH using the ratio between the number of sorbed complexes with a lifetime close to 85 μs (generally specified as an europium outer-sphere complex) and the number of the total surface complexes. For experiments performed at 25°C, the proportion of outer-sphere complex is always higher than 90% for samples sorbed under acidic conditions (pH=3.0 for montmorillonite; pH=4.5 for kaolinite) and progressively decreases to less than 10% for samples sorbed at higher pH (pH=7.1 for montmorillonite; pH=7.7 for kaolinite). Our calculations show that the proportion of outer-sphere complexes at 150°C is always less than 20%, even under acidic conditions (pH=2.6 for montmorillonite; pH=4 for kaolinite), and then decreases below 5% at higher pH.

At this stage, we can draw three conclusions:

- For both mineral surfaces, adsorption is the sole process of Eu retention at the studied temperatures.
- Two sorbed species are detected for all samples, whatever the temperature. An outer-sphere complex is dominant under acidic conditions, and is assumed to be exchanged with the interlayer cations of the clay mineral. The second species, an inner-sphere complex, is predominant at higher pH, and is probably bound onto the edges of the particles.
- The nature of the complexes does not change with temperature, although all of the recorded lifetimes decrease slightly with increasing temperature.

As mentioned above, the lifetime of 240 μs measured at 25°C (i.e. 170 μs at 150°C) corresponds to an inner-sphere complex and reflects the exclusion of water molecules in the first coordination sphere of Eu(III). According to the formula of Kimura et al. (1996), a lifetime of 240 μs at 25°C would correspond to approximately four water molecules in the first coordination sphere of Eu(III). To ascertain the nature of the surface sites involved (silanol or aluminol), we carried out comparisons with spectra acquired on simple oxides. Although we do not attribute an additive property to former oxides, we assume that the single oxides and former oxides of an aluminosilicate have lifetimes of the same order of magnitude.

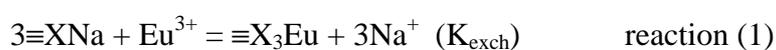
As described in the experimental section, we performed some experiments at 25°C and 80°C to study the sorption of Eu(III) onto simple oxides: alumina (γ -Al₂O₃), gibbsite (α -Al(OH)₃), quartz (SiO₂) and solid amorphous silica. The runs were conducted at 25 and 80°C at neutral pH (7.0 at 25°C; 6.3 at 80°C), where there is known to be almost complete sorption of trivalent metals (Kosmulski, 1997; Rabung *et al.*, 2000; Wang *et al.*, 2000; Stumpf *et al.*, 2001; Kowal-Fouchard *et al.*, 2004). Under such pH conditions, sorption is mainly attributed to surface complexation on aluminol sites (for alumina and gibbsite) or silanol sites (for quartz and amorphous silica). In addition, we tested two different oxides for each type of site (aluminol or silanol) to evaluate the influence of the sorbent crystallographic structure on the lifetime of the sorbed species. Table 3 reports the characteristics of these samples and the corresponding measured lifetimes. In the case of europium sorption onto gibbsite and alumina, we detected only one sorbed species for both minerals with a lifetime close to 220 μ s at 25°C. This lifetime is in good agreement with the values reported in Rabung *et al.* (2000) and Kowal-Fouchard (2002) for trivalent metals sorbed onto γ -alumina (235 μ s and 250 \pm 30 μ s, respectively). Fluorescence time measurements are similar for γ -alumina and α -gibbsite. This behaviour is not surprising because γ -alumina is unstable in aqueous solutions at high temperature (Laiti *et al.*, 1998), and therefore its surface consists of sites similar to those involved in the case of α -gibbsite. On the other hand, two lifetimes were measured at 25°C with quartz and amorphous silica (see Table 3). We observe a short fluorescence time close to 41 \pm 20 μ s at 25°C in the quartz/europium system. Such a short lifetime is very difficult to explain in terms of sorption and may be due to an inappropriate fitting of the lifetime data. Indeed, for this same system, Kowal-Fouchard (2004) measured two fluorescence times close to 130 and 350 μ s. Therefore, although we report the lifetime data obtained here for quartz/europium, the results are not used in the following discussion. On the other hand, the lifetimes measured in the amorphous silica/europium system suggest the existence of two distinct sorbed species. This behaviour is in good agreement with the results of Kowal-Fouchard *et al.* (2004). Indeed, these authors interpret the sorption of Eu onto amorphous silica (silica 60H Merck) at 25°C by the presence of two sorbed complexes. A first species appears at pH 6-7 with a lifetime of 130 \pm 20 μ s, whereas a second species appears at higher pH with a lifetime of 350 \pm 50 μ s. For all the oxide samples, the lifetimes decrease slightly with increasing temperature as observed previously with clay suspensions.

The similar lifetimes for europium sorbed onto Al₂O₃, Al(OH)₃ and both clay minerals (see figure 5) suggests that the inner-sphere complexes on clay surfaces are mainly bound to an

aluminol site. Such a process has already proposed by Stumpf *et al.* (2001) for the sorption of Cm(III) onto the same clay minerals studied here. On the other hand, Kowal-Fouchard (2002) provided experimental evidence that surface complexes were bound to both aluminol and silanol sites. As mentioned by these authors, the silanol sites are only involved at alkaline pH when the aluminol sites are saturated. In contrast to Kowal-Fouchard's experiments, saturation of the edge sites is prevented by the concentration of europium and the water/clay ratio used in our systems. This also rules out any significant contribution of silanol sites to europium retention, since aluminol sites are the major sorption sites for Eu. Thus, not all the possible silanol sites are taken into account in the subsequent modelling of surface complexation. The complex formed could be a monodentate species, such as $=\text{AlOHEu}^{2+}$ or $=\text{AlOHEu}^{3+}$, or a bidentate species such as $(=\text{AlOH})_2\text{Eu}^{3+}$ or $(=\text{AlO})_2\text{Eu}^+$. We currently have no objective arguments to support either of these hypotheses, so we consider the monodentate species for reasons of simplicity. This choice of stoichiometry is analogous to the approach adopted by Stumpf *et al.* (2001) for the sorption of Cm(III). In any case, the choice of complex species has no effect on the ability of the numerical modelling to reproduce the observed sorption results (the same fit can be obtained with different sets of values for the adjustable parameters).

Surface complexation modelling and concluding remarks.

We used the DLM (Diffuse Layer Model) formalism to interpret our sorption data. The calculations were performed with the CHESS[®] computer code. We selected the Davis equations option to calculate the aqueous species activity coefficients. Following our experimental results and spectroscopic analyses, we take into account two kinds of reactive sites implying two distinct mechanisms: an exchange reaction with the compensating Na^+ cations, described by reaction (1), and surface complexation on the amphoteric aluminol sites described by reaction (2). Note that, in order to satisfy the charge balance, the exchange reaction implies three exchange sites for each aqueous europium cation.



To propose a model for the sorption of europium onto clays, we require the following input: (1) values of the site densities (aluminol sites and structural negative sites), (2) values of the exchange constants involving negative structural sites over the entire range of ionic strength, (3) values of the acid/base constants for the aluminol sites and the exchange reactions at the studied temperatures. This step is not trivial, so we applied the model proposed in Tertre *et al.* (submitted) to the same clay minerals. Briefly, we quantify the influence of the permanent negative charge on the acid-base surface chemistry of kaolinite and Na-montmorillonite using potentiometric titrations at 25 and 60°C in 0.025 to 0.5M NaClO₄ solutions. The experimental results are then fitted to the model, with the parameters used in the acid/base model as reported in Table 4. The main conclusion of this study is that the zero net proton consumption value is temperature independent between 25 and 60°C, whereas the net proton consumption increases slightly with increasing temperature. The originality of our acid/base model is that it takes into account the effect of the negative potential, created by the negative structural charges of clay minerals, on the protonation/deprotonation reactions of the edge sites (i.e. aluminol and silanol sites). In this way, our acid/base model is useful for interpreting clay mineral potentiometric data obtained at different ionic strengths in the temperature range from 25 to 60°C.

Because exchange and acid/base constants are not available at 80°C and 150°C, and due to the weak effect of temperature on the surface charge properties of the clays - at least between 25 and 60°C - we initially ignore the temperature dependence of the exchange and acid base constants and use the values obtained at 25°C for the whole range of temperatures. Then, we perform a test to evaluate the sensibility of the sorption model to the temperature dependence of the acid/base constants. The test consists of increasing by one unit the pK_a corresponding to the protonation of the aluminol site, while, at the same time, decreasing by one unit the pK_a corresponding to the deprotonation of the aluminol site. These test conditions were chosen because the pK_a of silanol and aluminol sites varies by less than one unit, at least between 25 and 60°C (see Tertre *et al.*, submitted). The results of this test show that there was no significant change to the modelled sorption edge. Consequently, the modelled sorption curves are presented here at different temperatures (25, 40, 80 and 150°C) and calculated from the acid/base model proposed at 25°C. For the europium sorption reactions, we assume the exchange mechanism to be temperature-independent. In fact, the observed shift of the Eu sorption curves towards lower pH at higher temperature can be attributed solely to the presence of the $\equiv\text{AlO}^-\text{Eu}^{2+}$ complex.

Based on these assumptions, we can adjust the values of the intrinsic sorption constants by fitting the modelled sorption edge to the experimental data of the sorption runs (the TLRFs runs were not used for this fitting). We obtain satisfactory results for all the experimental conditions. Figures 1A and 1B compare the experimental data with the calculated curves obtained for sorption of Eu(III) onto kaolinite and Na-montmorillonite, respectively. The numerical values of the intrinsic sorption constants are reported in Tables 5 and 6 for Na-montmorillonite and kaolinite, respectively. These tables also give the acid/base parameters. For Na-montmorillonite, the $\log K_{\text{exch}}$ value for the Na/Eu exchange reaction is equal to 6.5 ± 0.8 , which is not far outside the range of 5.3 ± 0.3 to 6.1 ± 0.3 proposed by Kowal-Fouchard (2002). By contrast, the formation constant of $\equiv\text{AlOEu}^{2+}$ (K_{edge}), the only temperature-dependent intrinsic value in our model, increases with temperature by 8 log units from 25 to 150°C. However, these values cannot be compared with those used in the Kowal-Fouchard's model because of fundamental differences in surface complexation reactions between the two models for. In the present sorption model, the unique constant that varies with temperature is the constant related to the formation of the $\equiv\text{AlOEu}^{2+}$ complex. This apparently contradicts the temperature dependence of the dissociation constants of surface sites present at the surface of single oxides (for example, pKa for $\equiv\text{SiOH}$ for quartz; pKa for $\equiv\text{AlOH}$ for α -gibbsite). For single oxides, acid/base and electrokinetic measurements clearly show that the surface charges vary with temperature and these variations can be modelled by the temperature dependence of the surface site constants (Blesa et al., 1984; Brady, 1994; Mustafa et al., 1998; Wesolowski et al., 2000). Nevertheless, surface charge is always negative for clay minerals whatever the temperature, thus masking the variations of surface charge with temperature, at least between 25 and 150°C (Brady et al., 1996; Thomas et al., 1999). Therefore, the corresponding acid/base constants used in our sorption model were assumed temperature independent.

The temperature dependence of the complexation reaction of europium onto aluminol sites can be quantified by an enthalpic term that is related to the intrinsic constant K_{edge} through the Van't Hoff formalism. By integrating $\log K_{\text{edge}}$ with $1/T$, we obtain positive values of enthalpy: $150\text{kJ}\cdot\text{mol}^{-1}$ for Na-montmorillonite and $95\text{kJ}\cdot\text{mol}^{-1}$ for kaolinite. This latter value is close to the Eu enthalpy of sorption onto montmorillonite proposed by Tertre *et al.* (2005), as calculated from the measured Kd. However, the Kd values cannot be taken as a strictly thermodynamic constants, and the significance of this positive enthalpy remains speculative.

Among the elementary reactions involved in edge surface complexation, dehydration of the aqueous cation is probably the second step after diffusion toward the mineral surface. Thus, such a reaction is likely to be promoted with increasing temperature, as demonstrated by Berger *et al.* (1994) for Na^+ on silica surfaces. Hence, the higher the temperature, the stronger the chemical bonding between the metal and the aluminol site.

Figure 6 presents the distribution of the sorbed Eu species onto Na-montmorillonite as a function of pH, at two temperatures (25 and 150°C) and for $I=0.5\text{M}$. This is consistent with the results of the spectroscopic study, showing a predominance of X_3Eu under acidic conditions and a predominance of AlOEu^{2+} under neutral conditions (pH=7.0 at 25°C; pH=5.8 at 150°C). Finally, we test our model, and the set of previously fitted intrinsic data, over a wide range of conditions. We first calculate the apparent K_d for the samples used in the TLRFs study, assuming a solid/liquid ratio of 25g/L and an initial europium concentration of 10^{-5}mol/L . Table 1 compares the predicted and measured K_d values. We find a relatively good agreement at all the temperatures. Without changing the adjusted parameters, we also apply the model to sorption experiments reported in Tertre *et al.* (2005) that were carried out with montmorillonite in 0.025 M NaClO_4 solutions. As shown on Figure 7, our model also reproduces the observed Eu retention at low ionic strength. Finally, we compare our model with the published data of other experimental studies. The test is limited to studies carried out at 25°C (Bradbury and Baeyens, 2002; Kowal-Fouchard, 2002) because of the lack of data at higher temperatures. As in Table 1, Figure 8 compares the predicted values of K_d with experimental measurements reported in the literature. Even though we only consider two sorption sites, our model fits data obtained over a wide range of conditions: from 0.05M (Kowal-Fouchard, 2002) to 0.5M (this study), and from an initial europium concentration of 10^{-7}mol/L (Baeyens and Bradbury, 2002) to 10^{-4}mol/L (Kowal-Fouchard, 2002). Nevertheless, our model underestimates sorption when the initial europium concentration is higher than 10^{-3}mol/L (some data of Kowal-Fouchard, 2002, not reported here). Indeed, as mentioned above, sorption on silanol sites becomes significant under such conditions because all the aluminol sites are saturated. Future investigations will be carried out to improve the model over a wider range of chemical conditions.

Finally, the conclusions presented here for Eu^{3+} can likely be extrapolated to other Ln^{3+} . However, the accurate modelling of other Ln^{3+} would involve extracting the sorption constants from experimental measurements. The required data are available in Tertre (2005).

Acknowledgments

This research was supported by grants from ANDRA (Agence Nationale pour la gestion des Déchets Radioactifs) and EDF (Electricité De France). We would like to thank F. Javier Huertas and two anonymous reviewers for their substantial contribution to improving the paper. We are also grateful to R. Freydier and F. Candaudap for their assistance in the ICP-MS analyses at Toulouse, and G. Lagarde for his technical help in TRLFS measurements at Orsay. M. Carpenter post-edited the English style.

Editorial handling: Johnson R. Haas

References

- Aja S.U., 1998 The sorption of the rare earth element, Nd, onto kaolinite at 25°C. *Clays Clay Minerals* **46**, 103-109.
- ANDRA, 2005 La charge thermique d'un stockage – Site de Meuse/Haute-Marne. Rapport technique, n°C.NT.ASIT.03.118.
- Angove M.J., Johnson B.B. and Wells J.D., 1998 The influence of temperature on the adsorption of cadmium (II) and cobalt (II) on kaolinite. *J. Colloid Interf. Sc.* **204**, 93-103.
- Baeyens B. and Bradbury M.H., 1997 A mechanistic description of Ni and Zn sorption on Na-montmorillonite. Part I: Titration and sorption measurements. *J. Contam. Hydro.* **27**, 199-222.
- Bauer A., Rabung T., Claret F., Schäfer T, Buckau G. and Fanghänel T., 2005 Influence of temperature on sorption of europium onto smectite: The role of organic contaminants. *Applied Clay Science* **30**, 1-10.
- Beaufort D., Berger G., Lacharpagne J.C. and Meunier A., 2001 An experimental alteration of montmorillonite into a di + trioctahedral smectite assemblage at 100 and 200°C. *Clay Minerals* **36**, 211-225.
- Berger G., Cadoré E., Schott J. and Dove P.M., 1994 Dissolution rate of quartz in lead and sodium electrolyte solutions between 25 and 300°C. Effect of the nature of surface complexes and reaction affinity. *Geochim. Cosmochim. Acta* **58**, 541-551.
- Berger G., Beaufort D. and Lacharpagne J.C., 2002 Experimental dissolution of sanidine under hydrothermal conditions: mechanism and rate. *Amer J. Sci.* **302**, 663-685.
- Blesa M.A., Figliolia N.M., Maroto A.J.G. and Regazzoni A.E., 1984 The influence of temperature on the interface magnetite—aqueous electrolyte solution. *J. Colloid Interf. Sci.* **101**, 410-418.

Bradbury M.H. and Baeyens B., 2002 Sorption of Eu on Na and Ca-montmorillonite: Experimental investigations and modeling with cation exchange and surface complexation. *Geochim. Cosmochim. Acta* **66**, 2325-2334.

Brady P.V., 1994 Alumina surface chemistry at 25, 40, and 60°C. *Geochim. Cosmochim. Acta* **58**, 1213-1217.

Brady P.V., Cygan R.T. and Nagy K.L., 1996 Molecular controls on kaolinite surface charge. *J. Colloid Interf. Sci.* **183**, 356-364.

Carroll S. A. and Walther J. V., 1990 Kaolinite dissolution at 25°, 60° and 80°C. *Amer. J. Sci.* **290**, 797-810.

Castet S., Dandurand J.L., Schott J. and Gout R. 1993 Boehmite solubility and aqueous aluminium speciation in hydrothermal solutions (90-350°C): experimental study and modeling. *Geochim. Cosmochim. Acta* **57**, 4869-4884.

Chisholm-Brause C.J., Berg J.M., Little K.M., Matzner R.A. and Morris D.E., 2004 Uranyl sorption by smectites : spectroscopic assessment of thermodynamic modeling. *J. Colloid and Interf. Sci.* **277**, 366-382.

Choppin G. R. and Peterman D. R., 1998 Applications of lanthanide luminescence spectroscopy to solution studies of coordination chemistry. *Coordination Chemistry Reviews* **174**, 283-299.

Coppin F., Berger G., Bauer A., Castet S. and Loubet M., 2002 Sorption of lanthanides on smectite and kaolinite. *Chem. Geol.* **182**, 57-68.

Coppin F., 2002 Etude expérimentale de l'adsorption des lanthanides sur une kaolinite et une montmorillonite, a 25°C. Ph. D. Thesis, Université Paul Sabatier, Toulouse, France, 244p.

Gautier J.M., Oelkers E.H. and Schott J., 2001 Are quartz dissolution rates proportional to B.E.T. surface areas? *Geochim. Cosmochim. Acta* **65**, 1059-1070.

Helgeson H.C., Kirkham D.H. and Flowers G.C. 1981 Theoretical prediction of the thermodynamic behaviour of aqueous electrolytes at high pressures and temperature; IV. calculation of the activity coefficients, osmotic coefficients and apparent molal standard and relative partial molal properties to 600°C and 5 kb. *Am. J. Sc.* **281**, 1249-1516.

Huertas F. J., Chou L. and Wollast R., 1999 Mechanism of kaolinite dissolution at room temperature and pressure. Part II: Kinetic study. *Geochim. Cosmochim. Acta* **63**, 3261-3275.

Johnson J.W., Oelkers E.H. and Helgeson H.C. 1992 SUPCRT92: A software package for calculating the standard molars thermodynamic properties of minerals; gases, aqueous species and reactions from 1 to 5000 bars and 0 to 1000°C. *Computer Geosciences* **18**, 899-947.

Kimura, T., Choppin, G.R., Kato, Y. and Yoshida, Z., 1996 Determination of the Hydration Number of Cm(III) in Various Aqueous Solutions. *Radiochimica Acta* **72**, 61-64.

Kimura T., Nagaishi R., Arisaka M., Ozaki T. and Yoshida Z., 2002 Instrumental development for spectroscopic speciation of f-elements in hydrothermal solutions: luminescence properties of lanthanide (III) ions. *Radiochimica Acta* **90**, 715-719.

Kirishima A., Kimura T., Tochiyama O. and Yoshida Z., 2004 Speciation study on uranium (VI) hydrolysis at high temperatures and pressures. *Journal of Alloys and Compounds* **374**, 277-282.

Kosmulski, M., 1997 Adsorption of Trivalent Cations on Silica. *J. Colloid Interf. Sci.* **195**, 395-403.

Kowal-Fouchard A., 2002 Etude des mécanismes de rétention des ions U(IV) et Eu(III) sur les argiles : influence des silicates. Ph. D. Thesis, Université Paris Sud, France, 330p.

Kowal-Fouchard A., Drot R., Simoni E., Marmier N., Fromage F. and Ehrhardt J.J., 2004 Structural identification of europium(III) adsorption complexes on montmorillonite. *New. J. Chem.* **28**, 864-869.

Kraepiel A. M. L., Keller K. and Morel F. M. M., 1999 A model for metal adsorption on montmorillonite. *J. Colloid Interf. Sci.* **210**, 43-54.

Laiti E., Persson P. and Öhman L. O., 1998 Balance between Surface Complexation and Surface Phase Transformation at the Alumina/Water Interface. *Langmuir* **14**, 825-831.

Lindqvist-Reis P., Klenze R., Schubert G. and Fanghänel T., 2005 Hydration of Ce^{3+} in aqueous solution from 20 to 200°C. A Time-Resolved Laser Fluorescence Spectroscopy study. *J. Phys. Chem. B* **109**, 3077-3083.

Liu C., Zachara J.M., Qafoku O. and Smith S.C., 2003 Effect of temperature on Cs^+ sorption and desorption in subsurface sediments at the Hanford site, USA. *Environ. Sci. Technol.* **37**, 2640-2645.

Marmier N., 1994. Etude expérimentale et modélisation de la fixation d'éléments en trace sur des oxydes minéraux Contribution à l'étude des propriétés adsorbantes des solides naturels. Ph. D. Thesis, Université Reims Champagne-Ardenne, France.

Moulin C., Wei J., Van Iseghem P., Laszak I., planque G. and Moulin V., 1999 Europium complexes investigations in natural waters by time-resolved laser-induced fluorescence. *Analytica Chimica Acta* **396**, 253-261.

Mustapha S., Dilara B., Neelofer Z., Naeem A. and Tasleem S., 1998 Temperature effect on the surface charge properties of $\gamma\text{-Al}_2\text{O}_3$. *J. Colloid Inter. Sci.* **204**, 284-293.

Planque G., Moulin V., Toulhoat P. and Moulin C., 2003 Europium speciation by time-resolved laser-induced fluorescence. *Analytica Chimica Acta* **478**, 11-22.

Pshinko G., Spasnova L. and Kornilovich B., 2004 Complexation and sorption of europium (III) ions onto clay minerals in the presence of fulvic acids. *Adsorption Science and Technology* **22**, 669-678.

Rabung T., Stumpf T., Geckeis H., Klenze R. and Kim J. I., 2000 Sorption of Am(III) and Eu(III) onto γ -alumina: experiment and modelling. *Radiochimica Acta* **88**, 711-716.

Rabung T., Pierret M. C., Bauer A., Geckeis H., Bradbury M. H. and Baeyens B., 2005 Sorption of Eu(III)/Cm(III) on Ca-montmorillonite and Na-illite. Part 1: Batch sorption and time-resolved laser fluorescence spectroscopy experiments. *Geochim. Cosmochim. Acta* **69**, 5393-5402.

Ruaya J.R. and Seward T.M. 1987 The ion-pair constant and other thermodynamic properties of HCl up to 350°C. *Geochim. Cosmochim. Acta* **51**, 121-130.

Sakuragi T., Tokuyama A., Sato S., Kozaki T., Mitsugashira T., Hara M. and Suzuki Y., 2002 Effects of calcium ions on the sorption of Am(III) and Eu(III) onto kaolinite in the presence of humic acid. *Journal of Nuclear Science and Technology Suppl.* **3**, 520-523.

Spahiu K. and Bruno J., 1995 A selected thermodynamic database for REE to be used in HLNW performance assessment exercise. SKB Technical Report 95-35.

Stumpf, T., Rabung, T., Klenze, R., Geckeis, H. and Kim, J.-I., 2001 Spectroscopic study of Cm (III) Sorption onto Y-Alumina. *J. Colloid Interf. Sci.* **238**, 219-224.

Stumpf T., Bauer A., Coppin F. and Kim J. I., 2002 Time-Resolved Laser Fluorescence Spectroscopy Study of the Sorption of Cm(III) onto Smectite and Kaolinite. *Environ. Sci. Technol.* **35**, 3691-3694.

Stumpf T., Bauer A., Coppin F., Fanghanel T. and Kim J. I., 2002 Inner-sphere, outer-sphere and ternary surface complexes: a TRLFS study of the sorption process of Eu(III) onto smectite and kaolinite. *Radiochimica Acta* **90**, 345-349.

Stumpf T., Hennig C., Bauer A., Denecke M. A. and Fanghanel T., 2004a An EXAFS and TRLFS study of the sorption of trivalent actinides onto smectite and kaolinite. *Radiochimica Acta* **92**, 133-138.

Tanger J.C. and Helgeson H.C. 1997 Revised equation of state for the standard partial molal properties of ions and electrolytes. *Am. J. Sc.* **288**, 19-98.

Tertre E., Berger G., Castet S., Loubet M. and Giffaut E., 2005 Experimental study of adsorption of Ni²⁺, Cs⁺ and Ln³⁺ onto Na-montmorillonite up to 150°C. *Geochim. Cosmochim. Acta* **69**, 4937-4948.

Tertre E. 2005 Adsorption de Cs⁺, Ni²⁺, et des lanthanides sur une kaolinite et une smectite jusqu'à 150°C: étude expérimentale et modélisation. PhD thesis, Université Toulouse III.

Tertre E., Castet S., Berger G., Loubet M. and Giffaut E. Acid-base surface chemistry of kaolinite and Na-montmorillonite at 25 and 60°C: experimental study and modelling. *Geochim. Cosmochim. Acta*, Submitted.

Thomas F., Michot L.J., Vantelon D., Montarges E., Prélot B., Cruchaudet M. and Delon J.F., 1999 Layer charge and electrophoretic mobility of smectites. *Colloids Surf. A Physicochem. Eng. Aspects* **159**, 351-358.

Turner D.R., Whitfield M. and Dickson A.G., 1981 The equilibrium speciation of dissolved components in freshwater and seawater at 25°C and 1 atm pressure. *Geochim. Cosmochim. Acta* **45**, 855-881.

Van der Lee J. and De Windt L., 2002 CHESS Tutorial and Cookbook. Updated for version 3.0. Users Manuals Nr. LHM/RD/02/13. Ecole des Mines de Paris, Fontainebleau, France.

Wang X., Dong W., Dai X., Wang A., Du J. and Tao Z., 2000 Sorption and desorption of Eu and Yb on alumina: mechanisms and effect of fulvic acid. *Applied Radiation and Isotopes* **52**, 165-173.

Wang X., Dong W., Gong Y, Wang C. and Tao Z., 2001 Sorption characteristics of radioeuropium on bentonite and kaolinite. *Journal of Radioanalytical and Nuclear Chemistry* **250**, 267-270.

Wesolowski D.J., Machesky M.L., Palmer D.A. and Anovitz L.M., 2000 Magnetite surface charge studies to 290°C from in situ pH titrations. *Chem. Geol.* **167**, 193-229.

Wood S.A. 1990 The aqueous geochemistry of the rare-earth elements and yttrium. 2. Theoretical predictions of speciation in hydrothermal solutions to 350°C at saturation water vapor pressure. *Chem. Geol.* **88**, 99-125.

Zysset M. and Schindler P. W., 1996 The proton promoted dissolution kinetics of K-montmorillonite. *Geochim. Cosmochim. Acta* **60**, 921-931.

#	Sorbent	T (°C)	pH equi.	% sorbed	Measured log Kd	Predicted log Kd
25Kaol1	kaolinite	25	4.5	44	1.5	1.4
25Kaol2	“	25	7.7	96	2.6	2.5
80Kaol1	“	80	3.9	39	1.4	1.4
80Kaol2	"	80	6.1	95	2.8	2.5
150Kaol1	"	150	4	56	1.7	1.8
150Kaol2	"	150	5.1	100	3.5	3.0
25Mont1	Na-montmor.	25	3	50	1.6	1.7
25Mont2	"	25	7.1	100	3.6	3.5
80Mont1	"	80	2.7	71	2	1.9
80Mont2	"	80	4.2	100	4.7	4.6
150Mont1	"	150	2.6	94	2.8	2.6
150Mont2	"	150	3.6	100	4.6	4.4

Table 1 : Experimental conditions for samples of clay minerals analysed by TRLFS.

A

T(°C)	log K _{H2O} (0)	log K _{NaCl} (1)	log K _{NaOH} (2)	log K _{HCl} (3)
25	-13.995	0.925	0.455	0.67
50	-13.271	1.023	0.336	0.73
90	-12.437	0.986	0.145	0.64
100	-12.259	0.977	0.098	0.62
150	-11.637	0.733	-0.141	0.41

B

Reference	Reaction	log K 25°C	log K 150°C
(4)	$\text{Eu}^{3+} + \text{H}_2\text{O} = \text{Eu}(\text{OH})^{2+} + \text{H}^+$	-7.8	-3.7
..	$\text{Eu}^{3+} + 2\text{H}_2\text{O} = \text{Eu}(\text{OH})_2^+ + 2\text{H}^+$	-16.4	-8.7
..	$\text{Eu}^{3+} + 3\text{H}_2\text{O} = \text{Eu}(\text{OH})_{3(\text{aq})} + 3\text{H}^+$	-25.2	-13.7
(5)	$\text{Eu}^{3+} + 4\text{H}_2\text{O} = \text{Eu}(\text{OH})_4^- + 4\text{H}^+$	-35.3	-21.5
(6)	$\text{Eu}^{3+} + \text{CO}_3^{2-} = \text{Eu}(\text{CO}_3)^+$	-8.0	-
..	$\text{Eu}^{3+} + 2\text{CO}_3^{2-} = \text{Eu}(\text{CO}_3)_2^-$	-12.9	-
(7)	$\text{Eu}^{3+} + \text{HCO}_3^- + \text{H}_2\text{O} = \text{EuOHCO}_{3(\text{aq})} + 2\text{H}^+$	-8.5	-
..	$\text{Eu}^{3+} + 2\text{HCO}_3^- + \text{H}_2\text{O} = \text{EuOH}(\text{CO}_3)_2^{2-} + 3\text{H}^+$	-15.2	-
..	$\text{Eu}^{3+} + \text{HCO}_3^- + 2\text{H}_2\text{O} = \text{Eu}(\text{OH})_2\text{CO}_3^- + 3\text{H}^+$	-17.8	-
(7)	$\text{Eu}^{3+} + 3\text{H}_2\text{O} = \text{Eu}(\text{OH})_{3(\text{s})} + 3\text{H}^+$	-15.3	-9
..	$2\text{Eu}^{3+} + 3\text{HCO}_3^- = \text{Eu}_2(\text{CO}_3)_{3(\text{s})} + 3\text{H}^+$	-5.8	-
(4)	$\text{Eu}^{3+} + \text{CO}_3^{2-} + \text{H}_2\text{O} = \text{Eu}(\text{OH})(\text{CO}_3)_{(\text{s})} + \text{H}^+$	-7.8	-

(0): Tanger and Helgeson (1997).

(4): Spahiu and Bruno (1995).

(1): Helgeson et al. (1981).

(5): Turner et al. (1981).

(2): reported values of Castet et al. (1993).

(6): Wood, 1990.

(3): Ruaya and Seward (1987).

(7): Johnson et al., 1992.

Table 2 : Thermodynamic constants used for calculating the aqueous speciation of europium. (A) Dissociation constants for H₂O, NaCl, NaOH and HCl. (B) Constants for aqueous and solid europium species.

#	Sorbent	T (°C)	pH _{eq}	Number of sorbed species	Corresponding fitted lifetimes (μs)
25gibbs	Gibbsite	25	6.9	1	220±30
80gibbs	Gibbsite	80	5.9	1	195±20
25alum	γ-Alumina	25	6.9	1	232±30
80alum	γ-Alumina	80	6.1	1	194±40
25quar	Quartz	25	8.0	2	41±20 352±30
80quar	Quartz	80	6.5	2	30±20 302±40
25amorp-sil	Amorphous silica	25	7.2	2	137±20 335±40
80amorp-sil	Amorphous silica	80	6.1	2	113±20 290±40

Table 3: Experimental conditions for samples of simple oxides. For TRLFS analysis, the best fit of the decay profiles is obtained with one lifetime for the gibbsite and alumina system, whereas two lifetimes are needed for the quartz and amorphous silica systems.

	site	$\frac{\text{SiOH}}{\text{AlOH}}$	Density ($\mu\text{mol.m}^{-2}$)	Surface reactions	log K (25°C) ± 0.3
Na-montmor.	=AlOH (complexation)	2	1.70	$\text{AlOH} + \text{H}^+ = \text{AlOH}_2^+$	5.1
				$\text{AlOH} = \text{AlO}^- + \text{H}^+$	-8.5
	$\equiv\text{SiOH}$ (complexation)	3.40	$\text{SiOH} = \text{SiO}^- + \text{H}^+$	-7.9	
	X ⁻ (exchange)	-	36.3	$\text{X}^- + \text{H}^+ = \text{XH}$ $\text{X}^- + \text{Na}^+ = \text{XNa}$	-2.2 1.4
kaolinite	AlOH (complexation)	1	0.83	$\text{AlOH} + \text{H}^+ = \text{AlOH}_2^+$	4.8
				$\text{AlOH} = \text{AlO}^- + \text{H}^+$	-6.1
	SiOH (complexation)	0.83	$\text{SiOH} = \text{SiO}^- + \text{H}^+$	-7.7	
	X ⁻ (exchange)	-	3.7	$\text{X}^- + \text{H}^+ = \text{XH}$ $\text{X}^- + \text{Na}^+ = \text{XNa}$	-2.2 5.1

Table 4: Na-montmorillonite and kaolinite site densities and their acid/base constants, from Tertre *et al.* (submitted).

Na-montmorillonite	$\log K_{\text{exch}}$	$\log K_{\text{edge}}$			
Sorption reactions	25°C to 150°C :	25°C	40°C	80°C	150°C
$3 \text{XNa} + \text{Eu}^{3+} = \text{X}_3\text{Na} + 3 \text{Na}^+$	5.4±0.4		-		
$\text{AlOH} + \text{Eu}^{3+} = \text{AlOEu}^{2+} + \text{H}^+$	-	-1.0±0.2	2.5±0.2	6.5±0.3	7.5±0.4
Reactions of the acid/base model					
$\text{X}^- + \text{H}^+ = \text{XH}$	-2.2		-		
$\text{X}^- + \text{Na}^+ = \text{XNa}$	1.4		-		
$\text{AlOH} + \text{H}^+ = \text{AlOH}_2^+$	-		5.1		
$\text{AlOH} = \text{AlO}^- + \text{H}^+$	-		-8.5		
$\text{SiOH} = \text{SiO}^- + \text{H}^+$	-		-7.9		

Total Site densities (in $\mu\text{mol}/\text{m}^2$):

- AlOH = 1.70
- SiOH = 3.40
- X⁻ = 36.3

Table 5: Sorption reactions and corresponding values of constants used for the modelling of europium sorption onto Na-montmorillonite.

kaolinite	log K_{exch}	log K_{edge}			
		25°C	40°C	80°C	150°C
Sorption reactions	25°C to 150°C :				
$3 \text{XNa} + \text{Eu}^{3+} = \text{X}_3\text{Na} + 3 \text{Na}^+$	11.0±0.4		-		
$\text{AlOH} + \text{Eu}^{3+} = \text{AlOEu}^{2+} + \text{H}^+$	-	-2.3±0.3	-1.4±0.2	0.2±0.2	2.7±0.4
Reactions of the acid/base model					
$\text{X}^- + \text{H}^+ = \text{XH}$	-2.2		-		
$\text{X}^- + \text{Na}^+ = \text{XNa}$	5.1		-		
$\text{AlOH} + \text{H}^+ = \text{AlOH}_2^+$	-		4.8		
$\text{AlOH} = \text{AlO}^- + \text{H}^+$	-		-6.1		
$\text{SiOH} = \text{SiO}^- + \text{H}^+$	-		-7.7		

Total Site densities (in $\mu\text{mol}/\text{m}^2$):

- AlOH = 0.83
- SiOH = 0.83
- X^- = 3.7

Table 6: Sorption reactions and corresponding values of constants used for the modelling of europium sorption onto kaolinite.

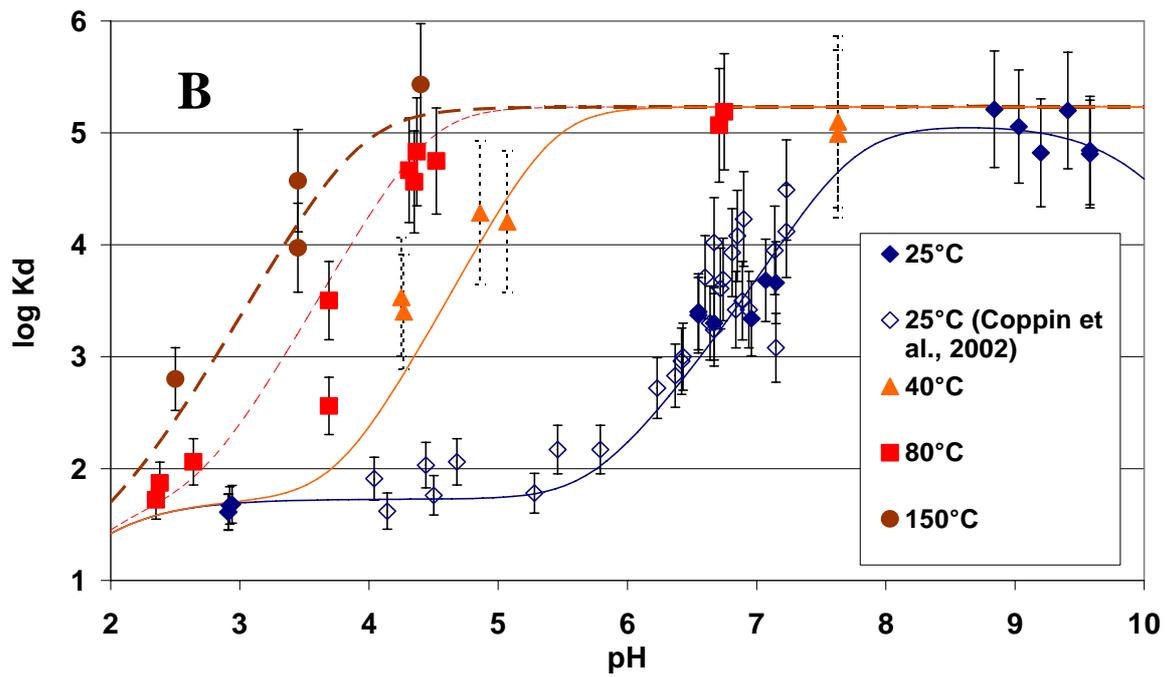
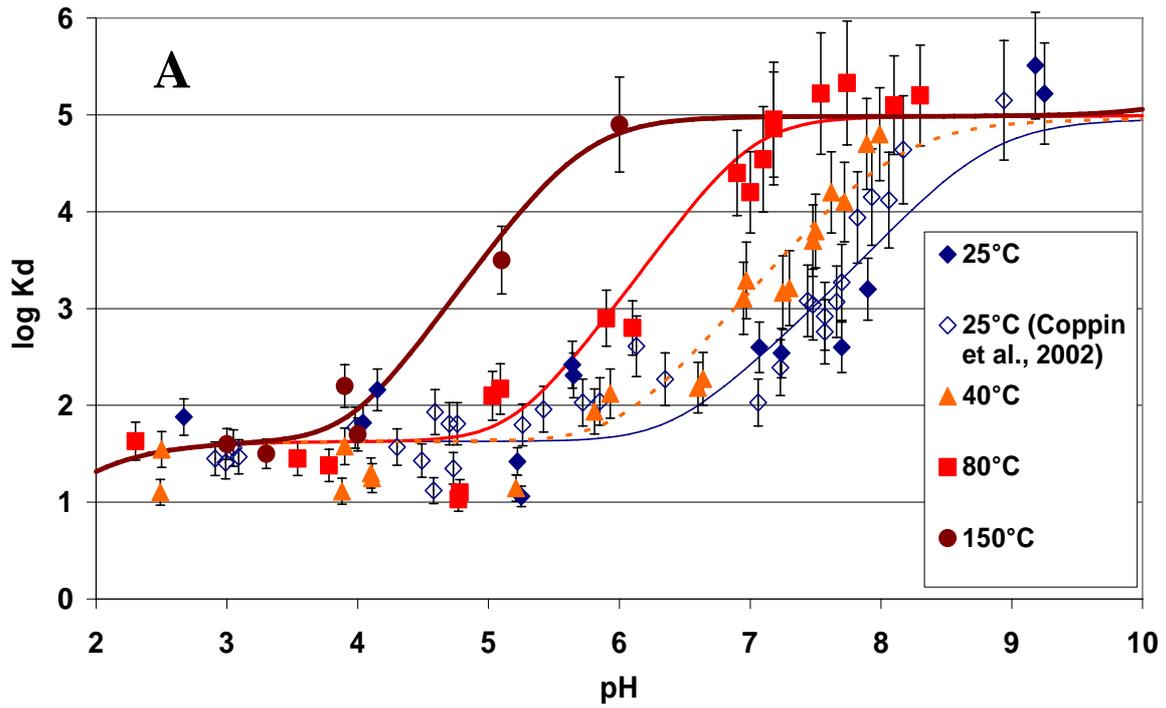


Figure 1: Logarithm of the distribution coefficients of Eu between kaolinite (A) and 0.5M NaClO₄ solutions at various temperatures and pH, compared with previous data on Namontmorillonite (B) under the same conditions. Symbols refer to experimental measurements whereas curves indicate model results.

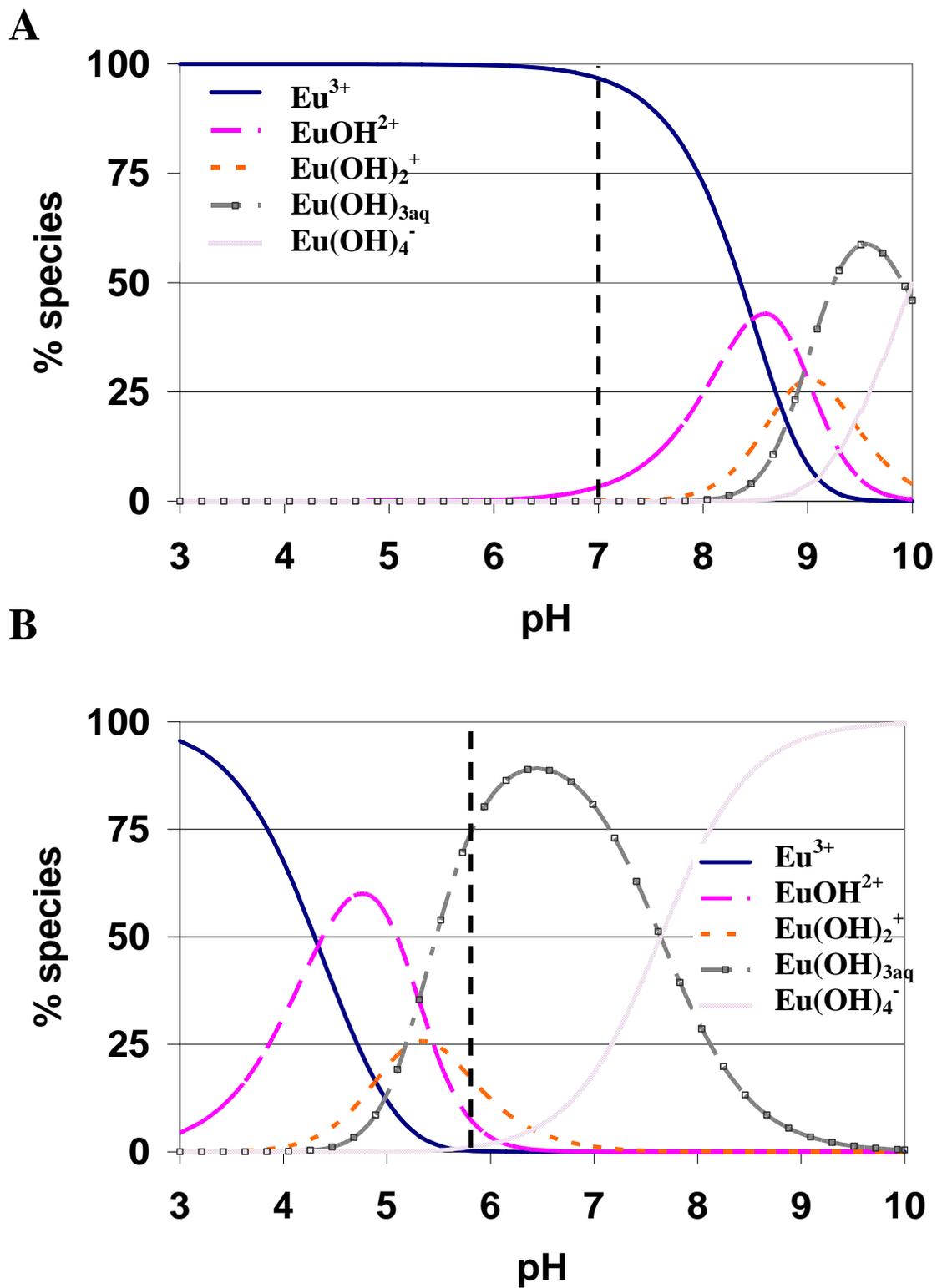


Figure 2: Speciation of 10^{-5} mol/L Eu(III) in 0.5M NaClO_4 , CO_2 -free solutions at 25°C (A) and 150°C (B). The vertical dotted line shows the value of neutral pH (5.8 at 150°C, cf. 7.0 at 25°C).

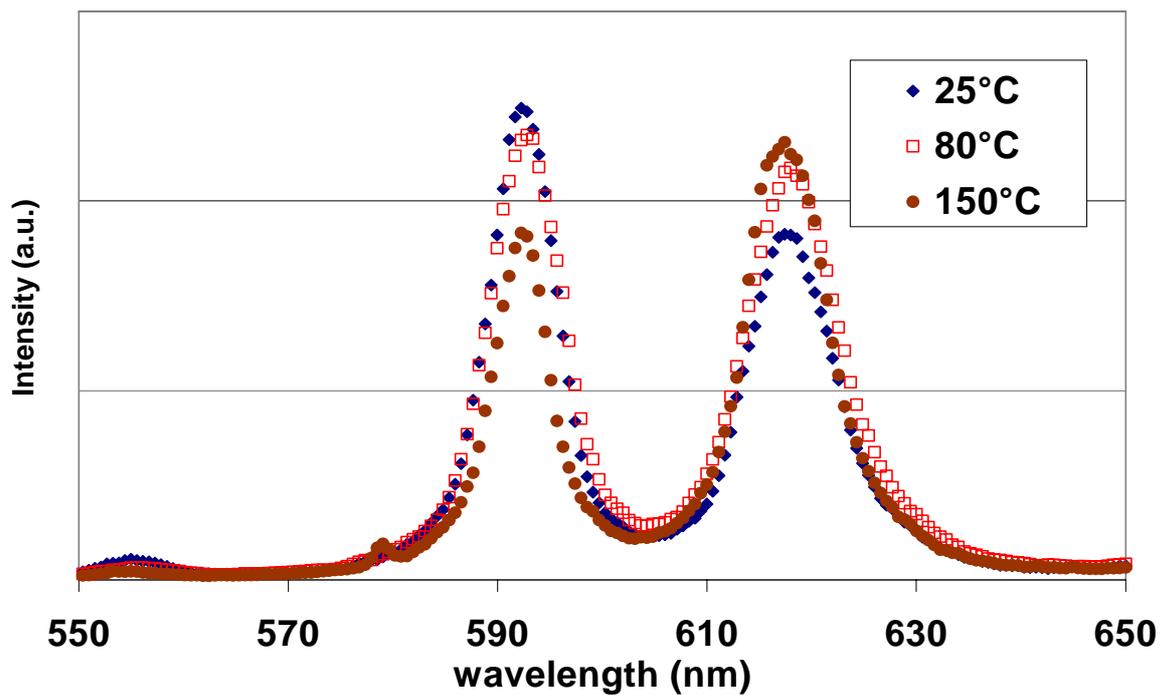


Figure 3: Fluorescence spectra of aqueous Eu(III) in NaClO₄ solutions at pH=3 at 25, 80 and 150°C (similar spectra were recorded at neutral pH 7.0 at 25°C, 6.5 at 80°C and 5.8 at 150°C).

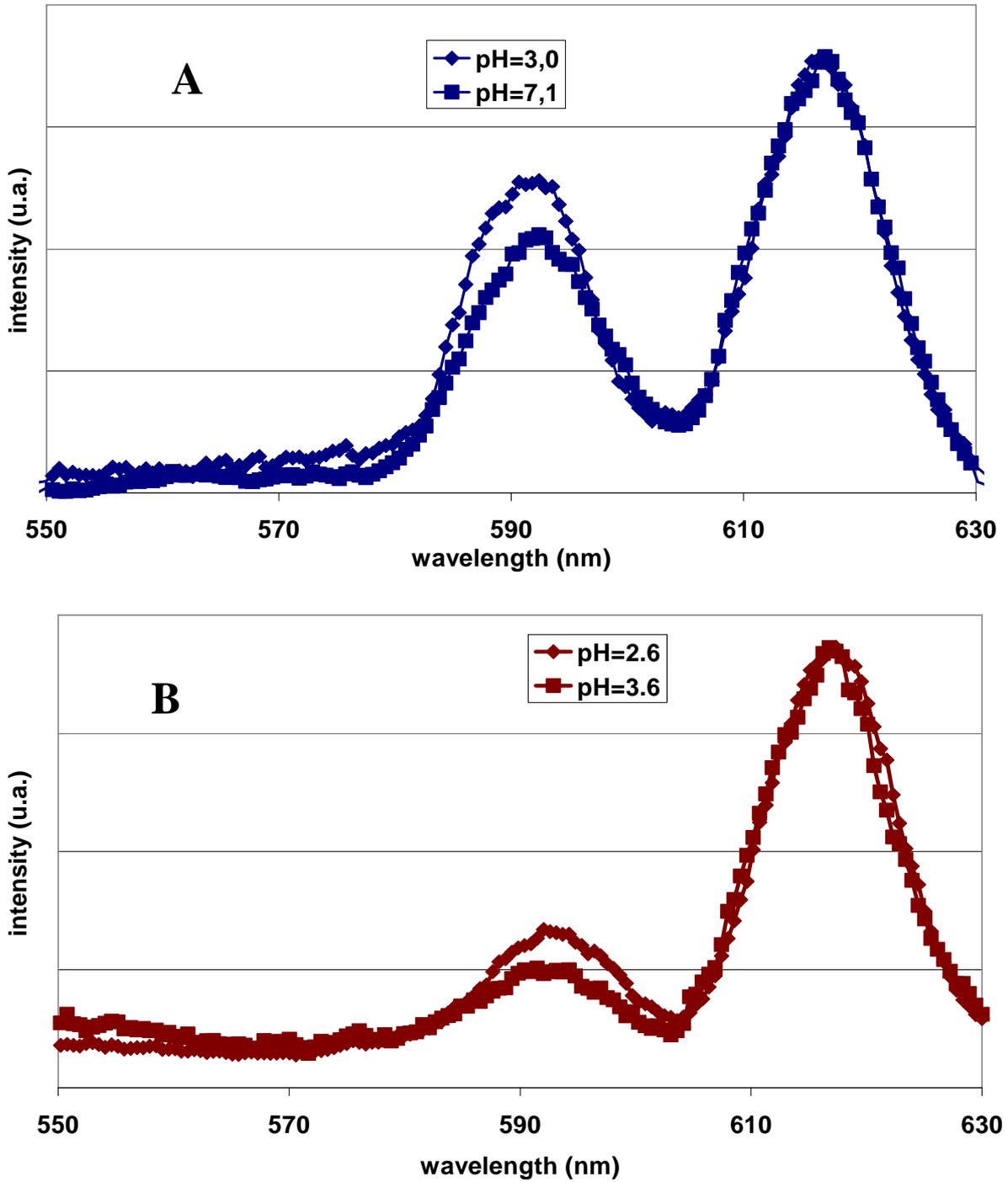


Figure 4: Fluorescence spectra of Eu(III) sorbed onto Na-montmorillonite at 25°C (A) and 150°C (B).

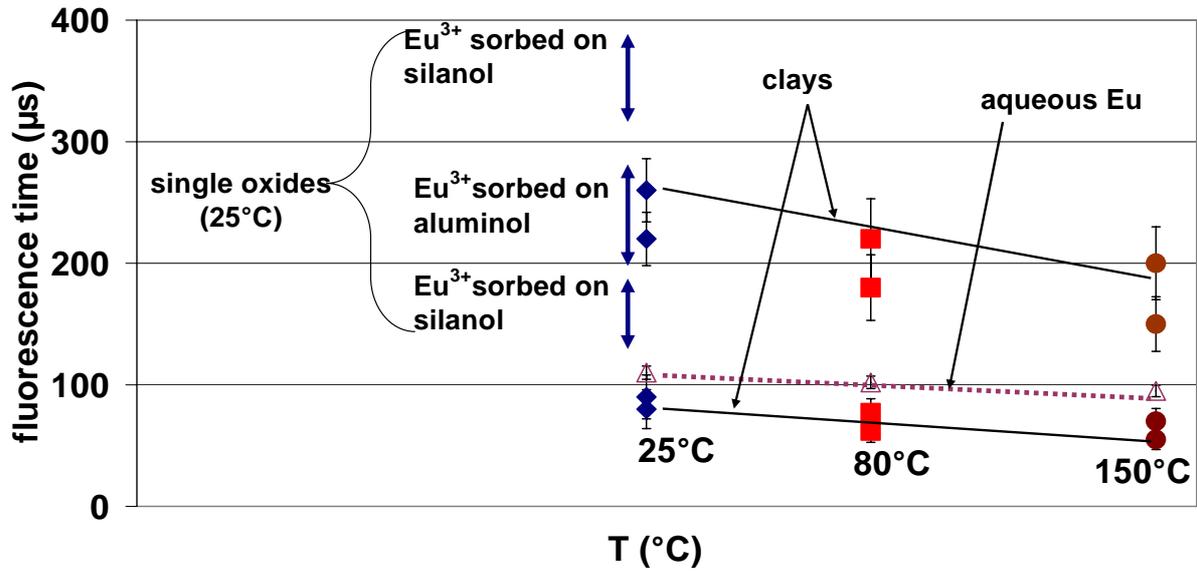


Figure 5: Fluorescence times of Eu(III) sorbed onto kaolinite (similar results for Na-montmorillonite) as a function of temperature. Lifetimes of aqueous europium obtained in NaClO₄ solutions at pH=3 are indicated on the figure. Blue arrows indicate lifetimes of europium sorbed onto silanol sites of amorphous silica and aluminol sites of gibbsite or alumina at 25°C.

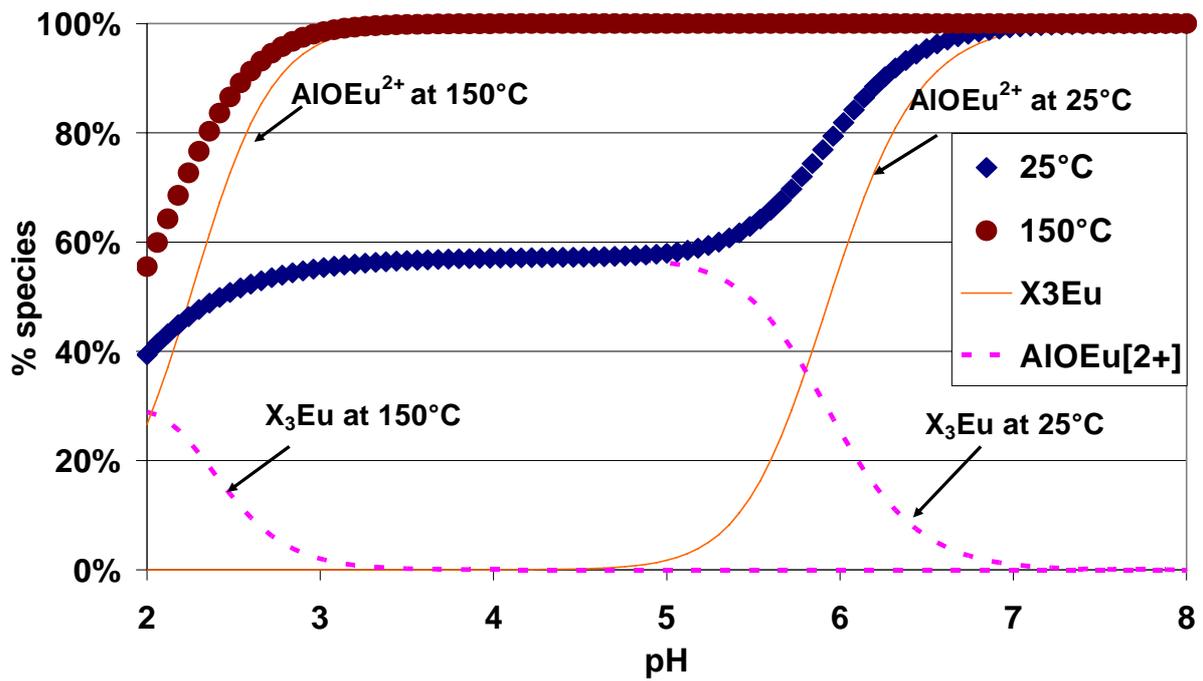


Figure 6: Distribution of europium sorbed onto Na-montmorillonite as a function of pH in 0.5M NaClO₄ solutions at 25 and 150°C. All data plotted on the graph refer to the results of the sorption model. Symbols represent total sorbed europium whereas curves represent the proportion of europium species sorbed onto the exchange sites and the aluminol sites.

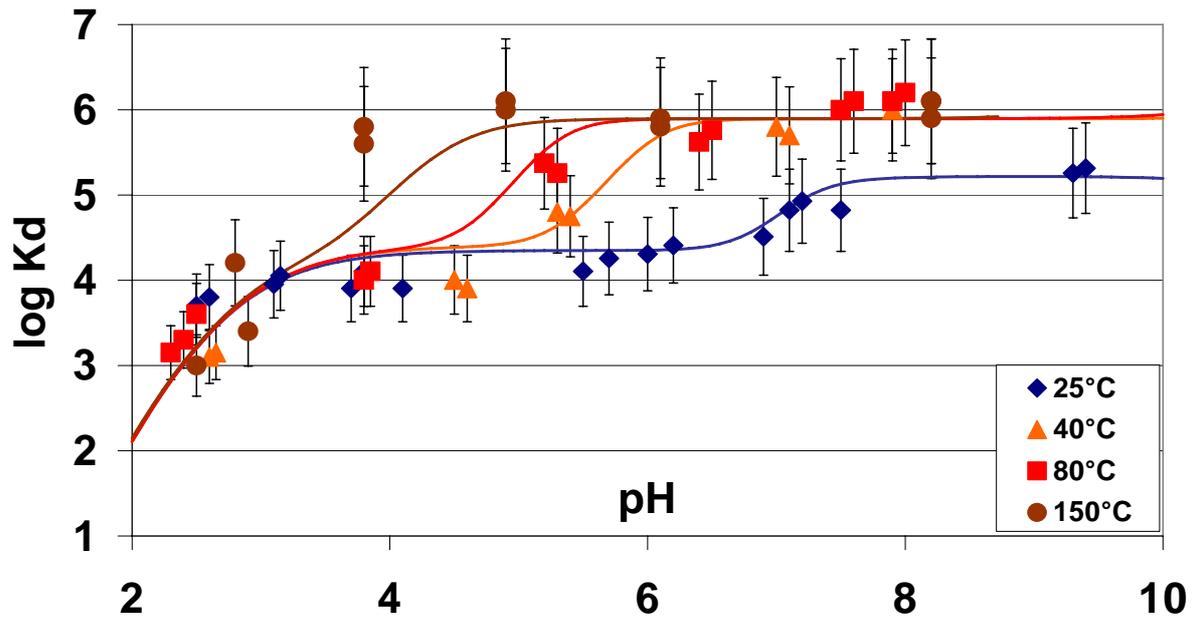


Figure 7: Distribution coefficients of Eu(III) sorbed onto Na-montmorillonite in 0.025 M NaClO₄ (Tertre *et al.*, 2005; data points with symbols) compared with predicted values according to our model (curves).

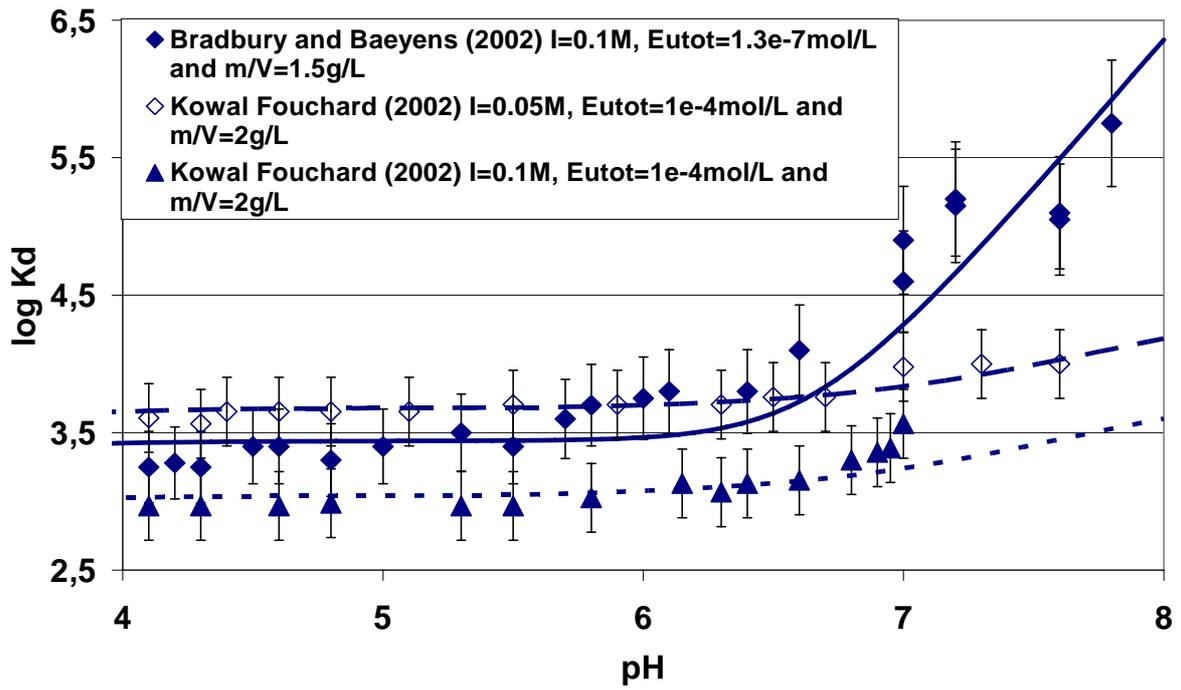


Figure 8: Other published distribution coefficients of Eu(III) sorbed onto Na-montmorillonite (dots) compared with predicted values according to our model (curves).